



Designation: C696 – 19

# Standard Test Methods for Chemical, Mass Spectrometric, and Spectrochemical Analysis of Nuclear-Grade Uranium Dioxide Powders and Pellets<sup>1</sup>

This standard is issued under the fixed designation C696; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

## 1. Scope

1.1 These test methods cover procedures for the chemical, mass spectrometric, and spectrochemical analysis of nuclear-grade uranium dioxide powders and pellets to determine compliance with specifications.

1.2 *Units*—The values stated in SI units are to be regarded as standard. The values given in parentheses are for information only.

1.3 The analytical procedures appear in the following order:

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<sup>1</sup> These test methods are under the jurisdiction of ASTM Committee C26 on Nuclear Fuel Cycle and are the direct responsibility of Subcommittee C26.05 on Methods of Test.

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<sup>2</sup> Discontinued January 1999. See C696–80.

<sup>3</sup> Discontinued September 2011.

<sup>4</sup> Discontinued as of May 30, 1980.

<sup>5</sup> Discontinued as of Nov. 1, 2019

1.4 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

## 2. Referenced Documents

### 2.1 ASTM Standards:<sup>6</sup>

- C753 Specification for Nuclear-Grade, Sinterable Uranium Dioxide Powder
- C761 Test Methods for Chemical, Mass Spectrometric, Spectrochemical, Nuclear, and Radiochemical Analysis of Uranium Hexafluoride
- C776 Specification for Sintered Uranium Dioxide Pellets for Light Water Reactors
- C859 Terminology Relating to Nuclear Materials
- C1267 Test Method for Uranium by Iron (II) Reduction in Phosphoric Acid Followed by Chromium (VI) Titration in the Presence of Vanadium
- C1287 Test Method for Determination of Impurities in Nuclear Grade Uranium Compounds by Inductively Coupled Plasma Mass Spectrometry
- C1347 Practice for Preparation and Dissolution of Uranium Materials for Analysis
- C1408 Test Method for Carbon (Total) in Uranium Oxide Powders and Pellets By Direct Combustion-Infrared Detection Method
- C1413 Test Method for Isotopic Analysis of Hydrolyzed Uranium Hexafluoride and Uranyl Nitrate Solutions by Thermal Ionization Mass Spectrometry
- C1453 Test Method for the Determination of Uranium by Ignition and the Oxygen to Uranium (O/U) Atomic Ratio of Nuclear Grade Uranium Dioxide Powders and Pellets
- C1457 Test Method for Determination of Total Hydrogen Content of Uranium Oxide Powders and Pellets by Carrier Gas Extraction

<sup>6</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [service@astm.org](mailto:service@astm.org). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

**C1502** Test Method for Determination of Total Chlorine and Fluorine in Uranium Dioxide and Gadolinium Oxide

**C1517** Test Method for Determination of Metallic Impurities in Uranium Metal or Compounds by DC-Arc Emission Spectroscopy

**D1193** Specification for Reagent Water

### 3. Terminology

3.1 Except as otherwise defined herein, definitions of terms are as given in Terminology **C859**.

### 4. Significance and Use

4.1 Uranium dioxide is used as a nuclear-reactor fuel. In order to be suitable for this purpose, the material must meet certain criteria for uranium content, stoichiometry, isotopic composition, and impurity content. These test methods are designed to show whether or not a given material meets the specifications for these items as described in Specifications **C753** and **C776**.

4.1.1 An assay is performed to determine whether the material has the minimum uranium content specified on a dry weight basis.

4.1.2 The stoichiometry of the oxide powder is useful for predicting its sintering behavior in the pellet production process.

4.1.3 Determination of the isotopic content of the uranium in the uranium dioxide powder is made to establish whether the effective fissile content is in compliance with the purchaser's specifications.

4.1.4 Impurity content is determined to ensure that the maximum concentration limit of certain impurity elements is not exceeded. Determination of impurities is also required for calculation of the equivalent boron content (EBC).

4.1.5 Determination of the oxygen-to-uranium ratio is performed on the completed pellets to determine whether they have the appropriate stoichiometry for optimal performance during irradiation.

### 5. Reagents

5.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.<sup>7</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

5.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Specification **D1193**.

<sup>7</sup> *ACS Reagent Chemicals, Specifications and Procedures for Reagents and Standard-Grade Reference Materials*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

### 6. Safety Precautions

6.1 Proper precautions should be taken to prevent inhalation, or ingestion of uranium dioxide powders or dust during grinding or handling operations.

### 7. Sampling

7.1 Criteria for sampling this material are given in Specification **C753** and Specification **C776**.

7.2 Samples can be dissolved using the appropriate dissolution techniques described in Practice **C1347**, but final determination of applicability must be made by the user.

#### URANIUM BY FERROUS SULFATE REDUCTION IN PHOSPHORIC ACID AND DICHROMATE TITRATION METHOD

**This test method was withdrawn in January 1999 and replaced by Test method **C1267**.**

#### URANIUM AND OXYGEN TO URANIUM ATOMIC RATIO BY THE IGNITION (GRAVIMETRIC) IMPURITY CORRECTION METHOD

**This test method was withdrawn in September 2011 and replaced by Test Method **C1453**.**

#### TOTAL CHLORINE AND FLUORINE BY PYROHYDROLYSIS ION-SELECTIVE ELECTRODE METHOD

**This test method was withdrawn in September 2011 and replaced by Test Method **C1502**.**

#### CARBON (TOTAL) BY DIRECT COMBUSTION-THERMAL CONDUCTIVITY METHOD

**This test method was withdrawn in January 1999 and replaced by Test Method **C1408**.**

#### MOISTURE BY THE COULOMETRIC ELECTROLYTIC MOISTURE ANALYZER METHOD

### 8. Scope

8.1 This test method covers the determination of moisture in uranium dioxide samples. Detection limits are as low as 10 µg.

### 9. Summary of Test Method

9.1 The sample is heated in an oven (up to 400 °C) to drive off any water. The moisture is carried from the oven into the

electrolytic cell by a flowing stream of *dry* nitrogen. Two parallel platinum wires wound in a helix are attached to the inner surface of the tube, the wall of which is evenly coated with phosphorus pentoxide ( $P_2O_5$ ) (a strong desiccant that becomes electrically conductive when wet). A potential applied to the wires produces a measurable electrolysis current when moisture wets the desiccant. Electrolysis of the water continuously regenerates the cell enabling it to accept additional water.

9.2 Precautions must be taken to prevent interference from the following sources. Hydrogen fluoride will cause permanent damage to the cell and sample system and should not be run under any conditions. Corrosive acidic gases, such as chlorine and hydrogen chloride, will corrode the instrument. Entrained liquids and solids can cause cell failure and should be prevented from entering the gas stream. Ammonia and other basic materials react with the acidic cell coating and renders the cell unresponsive. Hydrogen, and to a lesser extent, oxygen or air, may cause a high reading due to recombination, in the cell, or in the case of hydrogen, due to reaction with oxide coating of the sample boat to produce water. Alcohols and glycols, particularly the more volatile ones, respond like water and therefore must not be present.

## 10. Apparatus

10.1 *Moisture Analyzer*, for solids, with quartz glass oven capable of being heated from ambient temperatures to 1000 °C. The assembly includes electrolytic cell, flow meter, range 30 to 140 cm<sup>3</sup>/min air, and a dryer assembly.<sup>8</sup>

10.2 *Balance*,<sup>9</sup> for weighing samples in the range from 1 to 100 mg.

10.3 *Nitrogen Gas Cylinder*, with a pressure regulator, a flow meter and a drying tower.

## 11. Reagents

11.1 *Barium Chloride Dihydrate* ( $BaCl_2 \cdot 2 H_2O$ ).

## 12. Operation

12.1 Turn the main power switch ON.

12.2 Adjust nitrogen gas pressure to 41.4 kPa (6 psi) and the flow rate to 50 mL/min measured at the exit of the apparatus.

12.3 Weigh the sample into a small, dry, aluminum boat (**Note 1**) and insert it into the instrument oven as follows:

**NOTE 1**—For samples that have been reduced in a hydrogen atmosphere and thus contain excess hydrogen, the use of a platinum boat in place of the aluminum tube and nickel boat will minimize any interference due to the hydrogen.

<sup>8</sup> A CEC Solids Moisture Analyzer, of Type 26-321A-MA is available from DuPont Instruments Inc., S. Shamrock Ave., Monrovia, CA 91016. If you are aware of alternative suppliers, please provide this information to ASTM International Headquarters. Your comments will receive careful consideration at a meeting of the responsible technical committee,<sup>1</sup> which you may attend.

<sup>9</sup> A Cahn Electrobalance, or equivalent, available from Cahn Division, Ventrum Instrument Corp., Paramount, CA has been found satisfactory. If you are aware of alternative suppliers, please provide this information to ASTM International Headquarters. Your comments will receive careful consideration at a meeting of the responsible technical committee,<sup>1</sup> which you may attend.

12.3.1 Open the top of the analyzer and remove the TFE-fluorocarbon plug. Do not touch with gloves.

12.3.2 With forceps pull the nickel boat one third of the way out of the tube and place the aluminum boat and the sample inside the nickel boat, then reposition the nickel boat near the center of the heating coils.

12.3.3 Replace the TFE-fluorocarbon plug and close the lid of the analyzer.

12.4 Reset the counter to 0 µg.

12.5 Set the timer at 1 h.

12.6 Set the temperature at 400 °C. This will activate the analyzer and start the heating cycle.

12.7 When the preset temperature has been reached and the counter ceases counting, record the reading, *S*.

## 13. Standardization

13.1 Determine the blank by processing dry, empty, aluminum boats according to steps 12.3 – 12.7 until constant values are obtained.

13.2 Weigh and analyze replicate 5-mg samples of  $BaCl_2 \cdot 2 H_2O$  until consistent results are obtained. Sodium tungstate dihydrate ( $Na_2WO_4 \cdot 2 H_2O$ ) may also be used for calibration.

## 14. Calculation

14.1 Calculate the moisture recovery, *Z*, for the standard as follows:

$$Z = (A - B)147.2Y \quad (1)$$

where:

*A* = micrograms of moisture on counter when standard is tested,

*B* = micrograms of moisture on counter from blank, and

*Y* = milligrams of  $BaCl_2 \cdot 2 H_2O$ . Each milligram of  $BaCl_2 \cdot H_2O$  contains 147.2 µg of water.

14.2 Calculate the percent moisture in the sample as follows:

$$\text{Moisture, \%} = [(S - B)/1000 WZ] \times 100 = (S - B)/10 WZ \quad (2)$$

where:

*S* = micrograms of moisture on counter when sample is tested,

*B* = micrograms of moisture on counter from blank,

*W* = milligrams of sample, and

*Z* = recovery of moisture from standard.

## 15. Precision

15.1 The relative standard deviation for moisture in a concentration range of 100 µg/g is approximately 2 % but increases to 10 % at the 20 µg/g level.

15.2 *Bias*—Since there is no accepted reference material for determining bias in this test method, no statement of bias is being made.

## NITROGEN BY THE KJELDAHL METHOD

### 16. Scope

16.1 This test method covers the determination of nitrogen in uranium dioxide in the range from 10 to 250 µg.

## 17. Summary of Test Method

17.1 The sample is decomposed with acid, the resulting solution is made strongly alkaline with sodium hydroxide solution, and the nitrogen is separated as ammonia by steam distillation. The distillate is collected in boric acid solution and the ammonia present is titrated with 0.01 *N* standard acid using a mixed indicator.

NOTE 2—Although a simple acid digestion is usually adequate for dissolution of uranium samples, some uranium nitrides do not yield to such treatment. The use of potassium dichromate in phosphoric acid (1)<sup>10</sup> has proved to be successful with nitrides that are difficult to decompose. Therefore, this medium has been recommended although, in most cases, a mixture of phosphoric and sulfuric acids would be adequate.

## 18. Interferences

18.1 There should be no interferences in nuclear-grade uranium dioxide.

## 19. Apparatus

19.1 *Nitrogen Distillation Apparatus*, micro.<sup>11</sup>

19.2 *Heater*, 750-W electric, full-control.

19.3 *Burner*, bunsen-type.

19.4 *Buret*, micro, class A, 5- or 10-mL capacity, graduated in 0.02-mL divisions.

## 20. Reagents

20.1 *Ammonia-Free Water*—Prepare by distillation or from an ion-exchange column.

20.2 *Boric Acid-Indicator Solution*—Dissolve 20 g of boric acid ( $H_3BO_3$ ) in 800 mL of hot ammonia-free water, cool the solution, add 4 mL of mixed indicator solution (52.3), and dilute to 1 litre.

20.3 *Mixed Indicator Solution*—Mix 100 mL of a 1 % alcoholic solution of bromocresol green and 20 mL of a 1 % alcoholic solution of methyl red.

20.4 *Phosphoric Acid ( $H_3PO_4$ , 85 %)*—Heat acid to 190 °C to remove excess water.

NOTE 3—Some lots of  $H_3PO_4$  give high blanks and cannot be used.

20.5 *Potassium Dichromate Solution (65 g/litre)*—Dissolve 65 g of potassium dichromate ( $K_2Cr_2O_7$ ) in ammonia-free water and dilute to 1 litre. If necessary to reduce the blanks prepare the dichromate by recrystallization of  $K_2CrO_4$  from alkaline solution (1).

20.6 *Sodium Hydroxide Solution*—Dissolve 500 g of sodium hydroxide (NaOH) in 1 litre of ammonia-free water.

20.7 *Sulfuric Acid, Standard*—( $H_2SO_4$ , 0.01 *N*)—Standardize against a standard sodium hydroxide solution that has been standardized against potassium hydrogen phthalate.

<sup>10</sup> The boldface numbers in parentheses refer to the list of references at the end of these methods.

<sup>11</sup> Kemmerer-Hallett Type, Fisher Scientific Co., has been found satisfactory. If you are aware of alternative suppliers, please provide this information to ASTM International Headquarters. Your comments will receive careful consideration at a meeting of the responsible technical committee,<sup>1</sup> which you may attend.

NOTE 4—Hydrochloric acid (HCl, 0.01 *N*) may be used instead of  $H_2SO_4$ .

## 21. Procedure

### 21.1 *Blank Determinations:*

21.1.1 Fill the boiler of the distillation apparatus with ammonia-free water and distill for at least 30 min with a digestion flask in place in order to purge the apparatus of any traces of ammonia present.

21.1.2 Place 10 mL of  $H_3PO_4$  and 15 mL of potassium dichromate solution (65 g/litre) in a digestion flask and attach to the apparatus. Add 50 mL of NaOH solution and start passing the steam from the boiler through the digestion flask.

21.1.3 Place a 125-mL Erlenmeyer flask containing 5 mL of the boric acid-indicator solution over the tip of the condenser and collect 25 mL of distillate. Lower the flask so that the tip of the condenser is above the level of the distillate and continue the distillation for an additional 30 s to rinse down the inside of the tube.

21.1.4 Titrate the distillate with the 0.01 *N*  $H_2SO_4$  from a microburet until the solution turns to a pink color.

21.1.5 Repeat the blank determination, steps 21.1.2 – 21.1.4, until the blanks are constant. If the blank exceeds 0.03 to 0.04 mL, look for a source of contamination.

### 21.2 *Analysis of the Sample:*

21.2.1 Transfer up to 2 g of a weighed, powdered sample (Note 5) to the digestion flask.

NOTE 5—Samples in pellet form must be crushed in a diamond mortar to – 100 mesh powder and sampled by riffing or quartering to obtain a representative sample.

21.2.2 Add 10 mL of  $H_3PO_4$  and heat the flask gently with a small burner until a clear green solution is obtained. Inspect the solution carefully to ensure that no undissolved uranium nitrides remain.

21.2.3 Cool the flask, then add 15 mL of  $K_2Cr_2O_7$  solution (65 g/litre) slowly with mixing. Warm at low heat for 3 to 4 min.

21.2.4 Attach the digestion flask to the distillation apparatus and add 50 mL of NaOH solution.

21.2.5 Place the receiving flask containing 5 mL of the boric acid-indicator solution over the condenser tip and distill and titrate following the procedure used to determine the blank.

## 22. Calculation

22.1 Calculate the nitrogen content as follows:

$$N, \mu\text{g/g on } UO_2 \text{ basis} = (A - B) 14.01 N \times 10^3 / W \quad (3)$$

where:

- A = millilitres of standard acid to titrate sample,
- B = millilitres of standard acid to titrate blank,
- N = normality of standard acid solution, and
- W = grams of  $UO_2$  sample.

## 23. Precision

23.1 This test method will determine nitrogen to within 7  $\mu\text{g}$  of the amount present.