



Designation: **D1976—19 D1976 – 20**

Standard Test Method for Elements in Water by Inductively-Coupled Plasma Atomic Emission Spectroscopy¹

This standard is issued under the fixed designation D1976; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope*

1.1 This test method covers the determination of dissolved, total-recoverable, or total elements in drinking water, ground water, surface water, domestic, commercial or industrial wastewaters,^{2,3} within the following concentration ranges of **Table 1**.

1.2 It is the user's responsibility to ensure the validity of the test method for waters of untested matrices.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.* For specific hazard statements, see **Note 2** and **Section 9**.

1.5 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

2. Referenced Documents

2.1 ASTM Standards:⁴

[D1066 Practice for Sampling Steam](#)

[D1129 Terminology Relating to Water](#)

[D1193 Specification for Reagent Water](#)

[D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water](#)

[D3370 Practices for Sampling Water from Flowing Process Streams](#)

[D4841 Practice for Estimation of Holding Time for Water Samples Containing Organic and Inorganic Constituents](#)

[D5673 Test Method for Elements in Water by Inductively Coupled Plasma—Mass Spectrometry](#)

[D5744 Test Method for Laboratory Weathering of Solid Materials Using a Humidity Cell](#)

[D5810 Guide for Spiking into Aqueous Samples](#)

[D5847 Practice for Writing Quality Control Specifications for Standard Test Methods for Water Analysis](#)

[D6234 Test Method for Shake Extraction of Mining Waste by the Synthetic Precipitation Leaching Procedure](#)

[D8006 Guide for Sampling and Analysis of Residential and Commercial Water Supply Wells in Areas of Exploration and Production \(E&P\) Operations](#)

[E1915 Test Methods for Analysis of Metal Bearing Ores and Related Materials for Carbon, Sulfur, and Acid-Base Characteristics](#)

[E2242 Test Method for Column Percolation Extraction of Mine Rock by the Meteoric Water Mobility Procedure](#)

2.2 USEPA Standards:²

[Method 200.7 Determination of Metals and Trace Elements in Water and Wastes by Inductively Coupled Plasma-Atomic Emission Spectrometry](#)

¹ This test method is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.05 on Inorganic Constituents in Water. Current edition approved Nov. 1, 2019/May 1, 2020. Published December 2019/June 2020. Originally approved in 1991. Last previous edition approved in 2018/2019 as ~~D1976—18~~D1976 – 19. DOI: ~~10.1520/D1976-19~~10.1520/D1976-20.

² The detailed report of EPA Method Study 27, Method 200.7 is available from the National Technical Information Service, 5285 Port Royal Road, Springfield, VA. A summary of the project is available from the U.S. Environmental Protection Agency, Environmental Monitoring and Support Laboratory, Cincinnati, OH, <http://www.epa.gov>.

³ Fishman, M. J. and Friedman, L., "Methods for Determination of Inorganic Substances in Water and Fluvial Sediments," *U.S. Geological Survey Techniques of Water-Resources Investigations*, Book 5, Chapter D1066, Open File Report 85-495, 1985, p. 659–671.

⁴ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

*A Summary of Changes section appears at the end of this standard

TABLE 1 Summary of Tested Concentration Ranges

Element	From	To	Unit
Aluminum	0.083	1.43	µg/mL
Antimony	0.411	1.41	µg/mL
Arsenic	0.083	0.943	µg/mL
Barium	0.030	250	µg/mL
Beryllium	0.017	0.076	µg/mL
Boron	0.330	1.18	µg/mL
Cadmium	0.018	0.776	µg/mL
Calcium	0.400	1100	µg/mL
Chromium	0.025	0.47	µg/mL
Cobalt	0.058	0.843	µg/mL
Copper	0.017	0.189	µg/mL
Iron	0.074	2.34	µg/mL
Lead	0.085	0.943	µg/mL
Lithium	0.800	450	µg/mL
Magnesium	0.073	4.62	µg/mL
Manganese	0.017	0.94	µg/mL
Molybdenum	0.073	1.09	µg/mL
Nickel	0.043	0.943	µg/mL
Phosphorus	10.0	310	µg/mL
Potassium	8.00	5200	µg/mL
Selenium	0.083	0.755	µg/mL
Silica	1.00	3000	µg/mL
Silver	0.017	0.189	µg/mL
Sodium	5.00	3500	µg/mL
Strontium	0.500	500	µg/mL
Sulfur	2.00	600	µg/mL
Thallium	0.126	0.953	µg/mL
Vanadium	0.041	1.877	µg/mL
Zinc	0.068	0.759	µg/mL

3. Terminology

3.1 Definitions:

3.1.1 For definitions of terms used in this standard, refer to Terminology [D1129](#).

3.2 Definitions of Terms Specific to This Standard:

3.2.1 *calibration blank, n*—a volume of water containing the same acid matrix as the calibration standards (see [11.1](#)).

3.2.2 *calibration standards, n*—a series of known standard solutions used by the analyst for calibration of the instrument (preparation of the analytical curve) (see [8.9](#)).

3.2.3 *instrumental detection limit, n*—the concentration equivalent to a signal, due to the analyte, that is equal to three times the standard deviation of a series of ten replicate measures of a reagent-blank signal at the same wavelength.

3.2.4 *laboratory control sample, n*—a solution with the certified concentration(s) of the analytes.

3.2.5 *method blank, n*—a volume of water carried through the entire sample preparation, preservation, and analytical procedure.

3.2.6 *reagent blank, n*—a volume of water containing the same matrix as the calibration standards, carried through the entire analytical procedure.

3.2.7 *total, n*—the concentration determined on an unfiltered sample following vigorous digestion (see [12.3](#)).

3.2.8 *total-recoverable, adj*—determinable by the digestion method that is included in this procedure (see [12.2](#)).

4. Summary of Test Method

4.1 Elements are determined, either sequentially or simultaneously, by inductively-coupled plasma atomic emission spectroscopy.

4.2 A background correction technique may be used to compensate for variable background contribution from high concentrations of major and trace elements.

5. Significance and Use

5.1 This test method is useful for the determination of element concentrations in many natural waters and wastewaters. It has the capability for the simultaneous determination of up to 29 elements. High sensitivity analysis and larger dynamic range can be achieved for some elements that are difficult to determine by other techniques such as Flame Atomic Absorption.

5.2 The test method is useful for multi-element analysis of domestic and commercial well produced drinking water for metals and nonmetals for use in baseline analysis and monitoring during exploration, hydraulic fracturing, production, closure and reclamation activities related to oil and gas operations (see Guide [D8006](#)).

5.2.1 Minimum analyses include arsenic, barium, iron, magnesium, sodium, calcium, manganese, and lead.

5.2.2 Boron, potassium, chromium, selenium, cadmium, and strontium may be required on a site specific basis.

5.2.3 The most abundant elements in oil and gas produced water are sodium, potassium, lithium, magnesium, calcium, strontium, iron, silica, phosphorus, and sulfur.

5.3 The test method is useful for multi-element analysis of acid rock drainage and other major and some trace elements in mining influenced water.

5.4 Where low quantitation limits are required, Test Method **D5673** may be applicable.

5.5 The test method is also useful for testing leachates and effluents for ore and mining and metallurgical waste characterization tests including Test Methods **D6234**, **E2242**, **D5744**, and solutions from the Biological Acid Production Potential and Peroxide Acid Generation Methods in the Appendix of Test Methods **E1915**.

6. Interferences

6.1 Several types of interference effects may contribute to inaccuracies in the determination of trace elements. These interferences can be summarized as follows:

6.1.1 Spectral interferences can be categorized as (1) overlap of a spectral line from another element; (2) unresolved overlap of molecular band spectra; (3) background contribution from continuous or recombination phenomena; and (4) background contribution from stray light from line emission of high concentration elements.

6.1.1.1 The effects described in **6.1.1** can be compensated for by utilizing a computer correction of the raw data, requiring the monitoring and measurement of the interfering element. The second effect may require selection of an alternate wavelength. The third and fourth effects can usually be compensated for by a background correction adjacent to the analyte line.

6.1.1.2 **Table 2** lists some interference effects for the recommended wavelengths given in **Table 2**. The data in **Table 2** are intended for use only as a rudimentary guide for the indication of potential spectral interferences. For this purpose, linear relations between concentration and intensity for the analytes and the interferents can be assumed.

6.1.1.3 Only those interferents listed in **Table 2** were investigated for the analytes in **Table 3**. The blank spaces in **Table 2** indicate that measurable interferences were not observed for the interferent concentrations listed in **Table 4**. Generally, interferences were considered as discernible if the interferent produced interference peaks or background shifts that corresponded to 2 to 5 % of the peaks generated by the analyte concentrations listed in **Table 3**.

6.1.2 Physical interferences are generally considered to be effects associated with the sample nebulization and transport processes. Such properties as change in viscosity and surface tension can cause significant inaccuracies, especially in samples that may contain high dissolved solids or acid concentrations, or both. The use of a peristaltic pump may lessen these interferences. If these types of interferences are operative, they must be reduced by dilution of these samples or utilization of standard addition techniques, or both.

6.1.2.1 Salt buildup at the tip of the nebulizer is another problem that can occur from high dissolved solids. This salt buildup affects aerosol flow rate that can cause instrumental drift. To control this problem, wet the argon prior to nebulization, use a tip washer, or dilute the sample.

TABLE 2 Analyte Concentration Equivalents, mg/L, Arising from Interferents at the 100 mg/L Level^A

Analyte	Wavelength, nm	Interferent										
		Al	Ca	Cr	Cu	Fe	Mg	Mn	Ni	Ti	V	
Aluminum	308.215	0.21	1.4	
Antimony	206.833	0.47	...	2.9	...	0.08	0.25	0.45	
Arsenic	193.696	1.3	...	0.44	1.1	
Barium	455.403	
Beryllium	313.042	0.04	0.05	
Boron	249.773	0.04	0.32	
Cadmium	226.502	0.03	0.02	
Calcium	317.933	0.08	...	0.01	0.01	0.04	...	0.03	0.03	
Chromium	267.716	0.003	...	0.04	0.04	
Cobalt	228.616	0.03	...	0.005	0.03	0.15	...	
Copper	324.754	0.003	0.05	0.02	
Iron	259.940	0.12	0.12	
Lead	220.353	0.17	
Magnesium	279.079	...	0.02	0.11	...	0.13	0.002	0.25	...	0.07	0.12	
Manganese	257.610	0.005	...	0.01	...	0.002	
Molybdenum	202.030	0.05	0.03	
Nickel	231.604	
Selenium	196.026	0.23	0.09	
Silicon	288.158	0.07	0.01	
Sodium	588.995	0.08	...	
Thallium	190.864	0.30	
Vanadium	292.402	0.05	...	0.005	0.02	...	
Zinc	213.856	0.14	0.29	

^A See **Table 4** for concentrations used.

TABLE 3 Analyte Elemental Concentrations Tested for Interferents^A

Analytes	mg/L
Al	10
As	10
B	10
Ba	1
Be	1
Ca	1
Cd	10
Co	1
Cr	1
Cu	1
Fe	1
Mg	1
Mn	1
Na	10
Ni	10
Pb	10
Sb	10
Se	10
Si	1
Tl	10
V	1
Zn	10

^A This table indicates concentrations used for interference measurements in [Table 2](#).

TABLE 4 Interferent Elemental Concentrations for Analytes Tested^A

Interferents	mg/L
Al	1000
Ca	1000
Cr	200
Cu	200
Fe	1000
Mg	1000
Mn	200
Ni	200
Ti	200
V	200

^A This table indicates concentrations used for interference measurements in [Table 2](#).

NOTE 1—Periodic inspection and cleaning of the nebulizer and torch components are highly recommended.

6.1.2.2 Reports indicate that better control of the argon flow rate improves instrument performance. This control of the argon flow rate can be accomplished with the use of mass flow controllers.

6.1.3 Chemical interferences are characterized by molecular compound formation, ionization effects, and solute vaporization effects. Normally these effects are not pronounced with the inductively coupled plasma (ICP) technique; however, if observed, they can be minimized by careful selection of operating conditions (incident power, plasma observation position, and so forth), by buffering the sample, by matrix matching, and by standard addition procedures. These types of interferences can be highly dependent on matrix type and the specific analyte.

6.2 Analysis for silica precludes the use of borosilicate glassware due to potential contamination.

7. Apparatus

7.1 See the manufacturer's instruction manual for installation and operation of inductively-coupled argon plasma spectrometers. [Table 5](#) lists elements for which this test method applies, with recommended wavelengths and typical estimated instrumental detection limits using conventional pneumatic nebulization. Actual working detection limits are sample dependent and as the sample matrix varies, these detection limits may also vary. In time, other elements may be added as more information becomes available and as required.

7.1.1 Use of a vacuum or purged path is necessary for determination of sulfur.

7.1.2 Use of glass in the sample path may not be acceptable for silica, use of an inert material is recommended to avoid silica contamination.

TABLE 5 Suggested Wavelengths and Estimated Detection Limits^A

Element	Wavelength, nm ^B	Estimated detection limit, µg/L ^C
Aluminum	308.215	45
Arsenic	193.696	53
Antimony	206.833	32
Barium	455.403	2
Beryllium	313.042	0.3
Boron	249.773	5
Cadmium	226.502	4
Calcium	317.933	10
Chromium	267.716	7
Cobalt	228.616	7
Copper	324.754	6
Iron	259.940	7
Lead	220.353	42
Lithium	670.784	4
Magnesium	279.079	30
Manganese	257.610	2
Molybdenum	202.030	8
Nickel	231.604	15
Phosphorous	214.914	76
Potassium	766.491	700
Selenium	196.026	75
Silica	288.158	27
Silver	328.068	7
Sodium	588.995	29
Strontium	421.552	0.77
Sulfur	182.037	3
Thallium	190.864	40
Vanadium	292.402	8
Zinc	213.856	2

^A Winge, R. K., Fassel, V. A., Peterson, V. J., and Floyd, M. A., "Inductively Coupled Plasma-Atomic Emission Spectroscopy," An Atlas of Spectral Information, Elsevier Science Publishing Co., Inc., New York, NY, 1985.

^B The wavelengths listed are recommended because of their sensitivity and overall acceptance. Other wavelengths may be substituted if they can provide the needed sensitivity and are treated with the same corrective techniques for spectral interference (see 6.1.1).

^C The estimated detection limits as shown are taken from Winge et al.,^A USEPA Method 200.7, or task group data. They are given as a guide for approximate detection limits for the listed wavelengths. The actual test method instrumental detection limits are sample-dependent and may vary as the sample matrix varies (see 3.2.3).

<https://standards.iteh.ai/catalog/standards/sist/c8cfe88f-9d83-4bf0-9cde-08a171675e7f/astm-d1976-20>

8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society.⁵ The high sensitivity of inductively-coupled argon plasma atomic emission spectrometry may require reagents of higher purity. Stock standard solutions are prepared from high purity metals, oxides, or nonhygroscopic reagent grade salts using Types I, II, and III reagent water, and ultrapure acids. Other grades may be used, provided it is first ascertained that the reagent is of sufficient purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Type I, II, or III of Specification **D1193**. It is the analyst's responsibility to assure that water is free of interferences. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the precision and bias of the test method. Type II water was specified at the time of round robin testing of this test method.

8.3 *Aqua Regia*—Mix three parts hydrochloric acid (sp gr 1.19) and one part concentrated nitric acid (sp gr 1.42) just before use.

NOTE 2—Exercise caution when mixing this reagent, use of a fume hood is recommended.

8.4 *Argon*—Welding grade equivalent or better.

8.5 *Hydrochloric Acid (1 + 1)*—Add 1 volume of hydrochloric acid (sp gr 1.19) ultrapure or equivalent to 1 volume of water.

⁵ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

8.6 *Nitric Acid* (1 + 1)—Add 1 volume of nitric acid (sp gr 1.42) ultrapure or equivalent to 1 volume of water.

8.7 *Nitric Acid* (1 + 499)—Add 1 volume of nitric acid (sp gr 1.42) ultrapure or equivalent to 499 volumes of water.

8.8 *Stock Solutions*—Preparation of example stock solutions for each element is listed in **Table 6**. Use of commercially prepared certified stock solutions is recommended.

8.9 *Mixed Calibration Standard Solutions*—Prepare mixed calibration standard solutions by combining appropriate volumes of the stock solutions in volumetric flasks (see **Note 3**). Prior to preparing mixed standards, each stock solution should be analyzed separately to determine possible spectral interference or the presence of impurities. Care should be taken when preparing the mixed standards to ensure the elements are compatible and stable. It is common practice to have all or nearly all elements in one mixed calibration standard.

NOTE 3—Mixed calibration standards will vary depending on the number of elements being determined. An example of mixed calibration standards for the simultaneous determination of 20 elements is as follows:

Mixed Standard Solution I—manganese, beryllium, cadmium, lead, and zinc
 Mixed Standard Solution II—copper, vanadium, iron, and cobalt
 Mixed Standard Solution III—molybdenum, arsenic, and selenium
 Mixed Standard Solution IV—aluminum, chromium, and nickel
 Mixed Standard Solution V—antimony, boron, magnesium, silver, and thallium

9. Hazards

9.1 The toxicity or carcinogenicity of each reagent used in this test method has not been precisely defined; however, each chemical should be treated as a potential health hazard. Adequate precautions should be taken to minimize personnel exposure to chemicals used in this procedure.

TABLE 6 Preparation of Example Element Stock Solutions^{A,B}

Element (Compound)	Weight, g	Solvent
Al	0.1000	HCl (1 + 1)
Sb	0.1000	Aqua regia
As ₂ O ₃ ^C	0.1320	Water + 0.4 g NaOH
BaCl ₂ ^E	0.1516	HCl (1 + 1)
Be	0.1000	Aqua regia
H ₃ BO ₃	0.5716	Water
Cd	0.1000	HNO ₃ (sp gr 1.42)
CaCO ₃ ^F	0.2498	Water + HCl (1 + 1)
Cr	0.1000	HCl (1 + 1)
Co	0.1000	HNO ₃ (1 + 1)
Cu	0.1000	HNO ₃ (1 + 1)
Fe	0.1000	HNO ₃ (sp gr 1.42)
Pb	0.1000	HNO ₃ (sp gr 1.42)
Li ₂ CO ₃	0.5323	HNO ₃ (1 + 1)
Mg	0.1000	HNO ₃ (1 + 1)
Mn	0.1000	HNO ₃ (1 + 1)
Ni	0.1000	HNO ₃ (sp gr 1.42)
NH ₄ H ₂ PO ₄	0.3745	HCl (1 + 9)
KCl	0.1907	Water
(NH ₄) ₂ MoO ₄	0.2043	Water
Na ₂ SeO ₄ ^D	0.2393	Water
Na ₂ SiO ₃ ·5H ₂ O	0.3531	Water
Ag	0.1000	HNO ₃ (sp gr 1.42)
NaCl	0.2542	Water
SrCO ₃	0.1685	HCl (1 + 9)
Na ₂ SO ₄	0.4431	Water
TiNO ₃	0.1303	Water
NH ₄ VO ₃	0.2297	HNO ₃ (1 + 1)
Zn	0.1000	HNO ₃ (1 + 1)

^A Example element stock solutions, 1.00 mL = 100 µg of metal. Dissolve the listed weights of each compound or metal in 20 mL of specified solvent and dilute to 1 L. The metals may require heat to increase rate of dissolution.

^B Where water is used as the solvent, acidify with 10 mL of HNO₃ (sp gr 1.42) and dilute to 1 L. See Section 8 for concentration of acids. Commercially available standards may be used. Alternative salts or oxides may also be used.

^C Add 2 mL of HNO₃ (sp gr 1.42) and dilute to 1 L.

^D Add 1 mL of HNO₃ (sp gr 1.42) and dilute to 1 L.

^E Dry for 1 h at 180°C.

^F Dry for 1 h at 180°C. Add to approximately 600 mL of water and dissolve cautiously with a minimum of dilute HCl. Dilute to 1 L with water.

10. Sampling

10.1 Collect the samples in accordance with Practices **D1066** or Practices **D3370** as applicable.

10.1.1 Analysis for silica precludes the use of borosilicate glassware due to potential contamination.

10.2 Preserve the samples by immediately adding nitric acid to adjust the pH to 2 at the time of collection. Normally, 2 mL of HNO₃ is required per L of sample. If only dissolved elements are to be determined, filter the sample through a 0.45- μ m membrane filter before acidification (see **Note 4**). The holding time for the sample may be calculated in accordance with Practice **D4841**.

NOTE 4—Depending on the manufacturer, some filters have been found to be contaminated to various degrees with heavy metals. Care should be exercised in selecting a source for these filters. It is good practice to wash the filters with dilute nitric acid and a small portion of the sample before filtering.

11. Calibration and Standardization

11.1 Calibrate the instrument over a suitable concentration range for the elements chosen by atomizing the calibration blank and mixed standard solutions and recording their concentrations and signal intensities.

11.1.1 Multiple-point calibration standards may be used, and it is the user's responsibility to ensure the validity of the test method. Regardless of the calibration procedure used, appropriate quality control (QC) is required to verify the calibration curve at the anticipated concentration range(s) before proceeding to the sample analysis. It is recommended that the calibration blank and standard(s) be matrix matched with the same acid concentration contained in the samples.

12. Procedure

12.1 To determine dissolved elements, proceed with **12.4**.

12.2 When determining total-recoverable elements for a two-point calibration, choose a volume of a well mixed, acid-preserved sample appropriate for the expected level of elements.

12.2.1 Transfer the sample to a beaker (use tetrafluoroethylene or equivalent for silica analysis) and add 2 mL of HNO₃ (1 + 1) and 10 mL of HCl (1 + 1) and heat on a steam bath or hot plate until the volume has been reduced to near 25 mL, making certain the sample does not boil. Cool the sample, and if necessary filter or let insoluble material settle to avoid clogging of the nebulizer. Adjust to the original sample volume. To determine total-recoverable elements, proceed with **12.4**.

12.3 When determining total elements (hard digestion), choose a volume of well mixed, acid-preserved sample appropriate for the expected level of elements.

NOTE 5—Addition of HF acid may be required in order to effect complete dissolution of all siliceous material, so this digestion method will result in low results for samples containing insoluble silica.

12.3.1 Transfer the sample to a beaker (use tetrafluoroethylene or equivalent for silica analysis). Add 3 mL of HNO₃ (sp gr 1.42). Place the beaker on a hot plate and cautiously evaporate to near dryness, making certain that the sample does not boil and that no area of the bottom of the beaker is allowed to go dry. Cool the beaker and add 5 mL of HNO₃ (sp gr 1.42). Cover the beaker with a watch glass (use tetrafluoroethylene or equivalent for silica analysis) and return it to the hot plate. Increase the temperature of the hot plate so a gentle reflux action occurs. Continue heating, adding additional acid as necessary, until the digestion is complete (generally indicated when the digestate is light in color or does not change in appearance with continued refluxing). Again, evaporate to near dryness and cool the beaker. Add 10 mL of HCl (1 + 1) and 15 mL of water per 100 mL of final solution and warm the beaker gently for 15 min to dissolve any precipitate or residue resulting from evaporation. Allow the sample to cool, wash the beaker walls and watch glass with water, and if necessary, filter or let insoluble material settle to avoid clogging the nebulizer. Adjust to the original sample volume. To determine total elements, proceed with **12.4**.

NOTE 6—Many laboratories have found block digestion systems a useful way to digest samples for trace metals analysis. Systems typically consist of either a metal or graphite block with wells to hold digestion tubes. The block temperature controller must be able to maintain uniformity of temperature across all positions of the block. For trace metals analysis, the digestion tubes should be constructed of polypropylene and have a volume accuracy of at least 0.5 %. All lots of tubes should come with a certificate of analysis to demonstrate suitability for their intended purpose.

12.4 Atomize each solution to record its emission intensity or concentration. A sample rinse of HNO₃ (1 + 499) is recommended between samples.

13. Calculation

13.1 Include the blank in the calibration as the zero point.

NOTE 7—The original interlaboratory study subtracted reagent blanks from all samples. This subtraction was particularly important for digested samples requiring large quantities of acids to complete the digestion.

13.2 If dilutions are required, apply the appropriate dilution factor to sample values.

13.3 Report results in the calibration concentration units.