



Designation: C1411 – 20

Standard Practice for The Ion Exchange Separation of Uranium and Plutonium Prior to Isotopic Analysis¹

This standard is issued under the fixed designation C1411; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This practice is for the ion exchange separation of uranium and plutonium from each other and from other impurities for subsequent isotopic analysis by thermal ionization mass spectrometry. Plutonium-238 and uranium-238, and plutonium-241 and americium-241, will appear as the same mass peak and must be chemically separated prior to analysis. Only high purity solutions can be analyzed reliably using thermal ionization mass spectrometry.

1.2 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.4 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

2. Referenced Documents

2.1 ASTM Standards:²

- C698 Test Methods for Chemical, Mass Spectrometric, and Spectrochemical Analysis of Nuclear-Grade Mixed Oxides ((U, Pu)O₂)
- C833 Specification for Sintered (Uranium-Plutonium) Dioxide Pellets for Light Water Reactors
- C859 Terminology Relating to Nuclear Materials

¹ This practice is under the jurisdiction of ASTM Committee C26 on Nuclear Fuel Cycle and is the direct responsibility of Subcommittee C26.05 on Methods of Test.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

- C1008 Specification for Sintered (Uranium-Plutonium) Dioxide Pellets—Fast Reactor Fuel (Withdrawn 2014)³
- C1168 Practice for Preparation and Dissolution of Plutonium Materials for Analysis
- C1347 Practice for Preparation and Dissolution of Uranium Materials for Analysis
- C1625 Test Method for Uranium and Plutonium Concentrations and Isotopic Abundances by Thermal Ionization Mass Spectrometry
- D1193 Specification for Reagent Water

3. Terminology

3.1 Definitions:

3.1.1 For definitions of terms used in this practice, refer to C859.

4. Summary of Practice

4.1 Solid samples are dissolved according to Practices C1168, C1347, or other appropriate methods. The resulting solution is processed by this practice to prepare separate solutions of plutonium and uranium for mass spectrometric isotopic analysis using Test Method C698 or C1625. Appropriate aliquants are taken to provide up to 1 mg of plutonium on the ion exchange column to be separated from 10 mg or less of uranium. Valence adjustment is obtained by using one of two procedures as described in 4.1.1 and 4.1.2 or by an alternative method demonstrated by the user to perform the equivalent reduction/oxidation procedure.

NOTE 1—Reduction of all higher plutonium oxidation states to plutonium (III) by the addition of hydroxylamine or NH₂ClO₄, followed by oxidation to plutonium (IV) by sodium nitrite and subsequent boiling to eliminate the nitrous fumes has been found to be acceptable. This method avoids the addition of Fe, which could interfere with electrodeposition of prior to mass spectrometry analysis samples.

4.1.1 For any sample type, especially those containing large amounts of impurities, ferrous sulfate may be used for reduction. The aliquant is dissolved in 3 M HNO₃. Ferrous sulfate is added to reduce all plutonium (VI) to plutonium (III), then 16 M HNO₃ is added to oxidize plutonium (III) to plutonium (IV), and to adjust the final acid concentration to 8 M HNO₃.

³ The last approved version of this historical standard is referenced on www.astm.org.

4.1.2 A hydrogen peroxide reduction may be used for relatively pure samples which do not contain excessive amounts of oxidizing impurities. The aliquant is dissolved in 8 M HNO₃. Hydrogen peroxide is added to the aliquant prior to fuming to reduce plutonium (VI) to the lower oxidation states. The solution is warmed on a hot plate to destroy excess hydrogen peroxide and stabilize plutonium (IV) in solution.^{4,5}

4.2 After valence adjustment, the resulting solution is passed through an anion exchange column in the nitrate form which retains the plutonium; uranium and americium are not absorbed. The adsorbed plutonium is washed with additional 8 M nitric acid (HNO₃) to remove impurities and then stripped from the column with 0.36 M hydrochloric acid (HCl) and 0.01 M hydrofluoric acid (HF). The effluent containing the uranium and americium is converted to a HCl medium, and this solution is passed through an anion exchange column in the chloride form which retains the uranium. The adsorbed uranium is washed with additional 9 M HCl to remove the impurities and then stripped from the column with 0.1 M HCl.

5. Significance and Use

5.1 Uranium and plutonium are used in nuclear reactor fuel and must be analyzed to insure that they meet certain criteria for isotopic composition as described in Specifications C833 and C1008. This practice is used to chemically separate the same mass peak interferences from uranium and plutonium and from other impurities prior to isotopic abundance determination by thermal ionization mass spectrometry.

5.2 In those facilities where perchloric acid use is tolerated, the separation in Test Method C698 may be used prior to isotopic abundance determination. Uranium and plutonium concentrations as well as isotopic abundances using thermal ionization mass spectrometry can be determined using this separation and following Test Method C1625.

6. Interferences

6.1 The separated heavy element fractions placed on mass spectrometric filaments must be very pure. The quantity required depends upon the sensitivity of the instrument detection system. Chemical purity of the sample becomes more important as the sample size decreases, because the ion emission of the sample is repressed by impurities.

6.2 Organics from ion exchange resin degradation products, if present, could affect the response of the mass spectrometer during the plutonium and uranium isotopic abundance measurements. Evaporation of the samples with concentrated nitric acid after the ion exchange separation has been found to destroy any resin degradation products. Organics from extraction resins may not be destroyed by this process, and will require a perchloric acid treatment, heating in a furnace at 500 °C, or both.

6.3 The use of hydrogen peroxide for valence adjustment, when possible, avoids the addition of iron, an element which is not cleanly removed from uranium by HCl ion exchange.

6.4 Elemental impurities, especially alkali elements, tend to produce unstable ion emission and alter observed ratios in an unpredictable manner.

6.5 Isobaric impurities or contaminants will alter the observed isotopic ratios; most notable of these for plutonium are americium-241 and uranium-238; the most notable isobaric impurity for uranium is plutonium-238.

6.6 Extreme care must be taken to avoid contamination of the sample by environmental uranium. The level of uranium contamination should be measured by analyzing an aliquant of 8 M HNO₃ reagent as a blank taken through the same chemical processing as the sample and computing the amount of uranium it contains.

7. Apparatus

7.1 *Ion Exchange Columns*—Disposable, 0.9 cm id × 3 cm with a 15 mL reservoir.

7.2 *Beakers or Alternate Acceptable Containers*—Pretreated, 20 to 30 mL, borosilicate glass. To avoid cross contamination, use only new borosilicate glass containers pretreated by heating in 4 M HNO₃ to leach uranium, rinsed in deionized water, and air or oven dried prior to use.

7.3 *Infrared Heating Lamps or Hot Plate*, with adjustable low and high heat settings.

7.4 *Transfer Pipets*—Disposable.

8. Reagents

8.1 Reagent grade or better chemicals should be used. Unless otherwise indicated, it is intended that all reagents conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society⁶ where such specifications are available. Other grades of reagents may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of measurements made on the prepared materials. Store solutions in appropriate polyethylene or glass bottles except as noted.

8.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean laboratory accepted demineralized or deionized water conforming to Specification D1193.

8.3 *Nitric Acid (sp gr 1.42), 16 M*—Concentrated nitric acid (HNO₃).

8.4 *Nitric Acid, 8 M*—Add 500 mL of HNO₃ (sp gr 1.42) to about 400 mL of water and dilute to 1 L.

8.5 *Nitric Acid, 4 M*—Add 250 mL of HNO₃ (sp gr 1.42) to about 700 mL of water and dilute to 1 L with water.

8.6 *Nitric Acid, 3 M*—Add 187 mL of HNO₃ (sp gr 1.42) to about 750 mL of water and dilute to 1 L with water.

⁶ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

⁴ Kressin, I.V. and Waterbury, G.R., *Anal. Chem.* Vol 34, No. 12, p. 1598, 1962.

⁵ Pietri, C.E., Freeman, B.P., and Weiss, J.R., *DOE/NBL-298*, Sept. 1981.

8.7 *Hydrochloric Acid (sp gr 1.19), 12 M*—Concentrated hydrochloric acid (HCl).

8.8 *Hydrochloric Acid, 9M*—Add 750 mL of HCl (sp gr 1.19) to about 250 mL of water and dilute to 1 L.

8.9 *Hydrofluoric Acid (sp gr 1.18), 29 M*—Concentrated hydrofluoric acid (HF).

8.10 *Hydrochloric Acid, 0.1 M*—Add 8 mL of HCl (sp gr 1.19) to about 900 mL of water and dilute to 1 L with water.

8.11 *Stripping solution (0.36 M HCl, 0.01 M HF)*—Add 30 mL of HCl (sp gr 1.19) and 0.4 mL HF (sp gr 1.18) to about 900 mL of water and dilute to 1 L with water.

8.12 Anion exchange resin, 50-100 mesh, wet, chloride form for uranium separation.

8.13 Anion exchange resin, nitrate form, 50-100 mesh, wet, for plutonium separation.

8.13.1 **Warning**—The dry and wet mesh size of the resins differ; for the chloride form of the resin, a 100-200 mesh, dry resin is purchased to provide 50-100 mesh, wet resin. The 100-200 mesh dry chloride resin may be used to prepare the 50-100 mesh, wet nitrate form of resin.

NOTE 2—The exchange capacity of the resin should be between 0.6 and 0.7 milliequivalents/gram of dry resin for optimum separation. A strong base anion exchange resin (quarternary amine functional groups on styrene divinyl benzene polymeric beads) is most often used, with DVB cross linkages of 8 %. The resin is available in the chloride form. Others may be used but the volume of resin may need to be adjusted to provide the desired degree of separation.

8.14 *Sulfuric Acid (sp gr 1.84), 18 M*—Concentrated sulfuric acid (H₂SO₄).

8.15 *Ferrous Sulfate Solution (0.1 M)*—Add 1.5 g of ferrous sulfate heptahydrate (FeSO₄ • 7H₂O) to approximately 40 mL of water; add 0.3 mL (7 drops) concentrated sulfuric acid and dilute to 50 mL with water.

8.16 *Hydrogen Peroxide (H₂O₂, 30 %)*, stabilized.

9. Precautions

9.1 Strong acids are used during this analysis. Safety glasses and gloves must be worn when handling these solutions. Extreme care should be exercised in using hydrofluoric acid and other hot concentrated acids. Acid solutions are evaporated during this analysis. These operations must be conducted in a fume hood.

9.2 Hydrofluoric acid is a highly corrosive acid that can severely burn skin, eyes, and mucous membranes. Hydrofluoric acid differs from other acids because the fluoride ion readily penetrates the skin, causing destruction of deep tissue layers. Unlike other acids that are rapidly neutralized, hydrofluoric acid reactions with tissue may continue for days if left untreated. Familiarization and compliance with the Safety Data Sheet is essential.

10. Procedure

10.1 *Plutonium Anion Exchange Separation:*

10.1.1 *Sample Preparation:*

10.1.1.1 Dissolve solid samples according to Practices C1168, C1347, or other appropriate methods. Aliquants of the

solution containing the approximate desired quantity of element are taken; the desired quantity of element will depend upon whether or not the solution is diluted prior to filament loading. (**Warning**—No initial aliquant should contain more than 10 mg of uranium to prevent inadequate rinsing of the ion exchange resin by the volumes given, and hence, inadequate separation of uranium and plutonium.)

NOTE 3—If uranium or plutonium concentration is desired, tracers should be added and equilibrated with the sample during dissolution; concentration may then be determined by isotope dilution TIMS. If concentration is to be determined by other means, total dissolution of the sample and quantitative transfers of the sample is essential. These precautions may not be necessary if the only measurement to be made is the isotope ratios of the uranium or plutonium, or both, in the sample.

10.1.1.2 An aliquant containing up to 1 mg of plutonium is transferred to a new, acid leached 20 to 30 mL container for ion exchange preparation.

NOTE 4—From this separation, a maximum of 200 ng of plutonium or 2 µg of uranium will be loaded on a filament for thermal ionization mass spectrometric isotopic abundance determination.

10.1.1.3 Fume the aliquant just to dryness on a hot plate or under an infrared heat lamp. If the sample contains halides, add 8 M HNO₃ and dry two more times. (**Warning**—Overheating may cause polymerization of plutonium and difficulty in subsequent dissolution.)

For most samples, proceed with 10.1.2 for ferrous sulfate reduction. For relatively pure samples, free of transition elements, proceed with 10.1.3 for hydrogen peroxide reduction. (See 4.1 for a further discussion of the valence adjustment.)

10.1.2 *Ferrous Sulfate Valence Adjustment:*

10.1.2.1 Dissolve the residue from 10.1.1.3 in 8 mL of 3 M HNO₃.

10.1.2.2 Add 1 drop of 0.1 M FeSO₄ solution, prepared the day of the ion exchange, from a disposable pipet. Swirl to mix well. This will reduce all plutonium in higher oxidation states to plutonium (III).

10.1.2.3 Add 5 mL of 16 M HNO₃. Swirl to mix well. This will oxidize plutonium (III) to plutonium (IV) and adjust the solution concentration to 8 M HNO₃.

10.1.2.4 Proceed to 10.1.4.

10.1.3 *Hydrogen Peroxide Valence Adjustment:*

10.1.3.1 Dissolve the residue from 10.1.1.3 in 5 mL of 8 M HNO₃.

10.1.3.2 Add 4 drops of 30 % H₂O₂ from a disposable pipet. Place the beaker on a hot plate at low heat (80 °C) for a minimum of 1.5 h while the beaker is covered with a watch glass. Gentle effervescence should occur; continue heating until the effervescence has stopped. Reduction of all plutonium (VI) to plutonium (IV) is critical to a successful ion exchange separation.

10.1.3.3 Cool solution to room temperature.

10.1.4 *Anion Exchange:*

10.1.4.1 Fill the ion exchange column with a water slurry of the nitrate resin to a settled height of 1 to 2 cm. Pass 10 mL of 8 M HNO₃ through the resin, and then add 5 mL of 8 M HNO₃ and drain just before the sample is added. The column should