



Designation: **E321–96 (Reapproved 2012) E321 – 20**

## Standard Test Method for Atom Percent Fission in Uranium and Plutonium Fuel (Neodymium-148 Method)<sup>1</sup>

This standard is issued under the fixed designation E321; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

### 1. Scope

1.1 This test method covers the determination of stable fission product <sup>148</sup>Nd in irradiated uranium (U) fuel (with initial plutonium (Pu) content from 0 to 50 %) as a measure of fuel burnup **(1-3)**.<sup>2</sup>

1.2 It is possible to obtain additional information about the uranium and plutonium concentrations and isotopic abundances on the same sample taken for burnup analysis. If this additional information is desired, it can be obtained by precisely measuring the spike and sample volumes and following the instructions in Test Method **E267**.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.5 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

#### 2.1 ASTM Standards:<sup>3</sup>

[C859 Terminology Relating to Nuclear Materials](#)

[C1168 Practice for Preparation and Dissolution of Plutonium Materials for Analysis](#)

[C1267 Test Method for Uranium by Iron \(II\) Reduction in Phosphoric Acid Followed by Chromium \(VI\) Titration in the Presence of Vanadium](#)

[C1347 Practice for Preparation and Dissolution of Uranium Materials for Analysis](#)

[C1411 Practice for The Ion Exchange Separation of Uranium and Plutonium Prior to Isotopic Analysis](#)

[C1415 Test Method for <sup>238</sup>Pu Isotopic Abundance By Alpha Spectrometry](#)

[C1625 Test Method for Uranium and Plutonium Concentrations and Isotopic Abundances by Thermal Ionization Mass Spectrometry](#)

[C1672 Test Method for Determination of Uranium or Plutonium Isotopic Composition or Concentration by the Total Evaporation Method Using a Thermal Ionization Mass Spectrometer](#)

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee C26 on Nuclear Fuel Cycle and is the direct responsibility of Subcommittee C26.05 on Methods of Test. Current edition approved June 1, 2012; Dec. 1, 2020. Published June 2012; January 2021. Originally approved in 1967-1967. Last previous edition approved in 2005-2012 as E321 – 96 (2012), (2005). DOI: 10.1520/E0321-96R12.10.1520/E0321-20.

<sup>2</sup> The boldface numbers in parentheses refer to the list of references appended to the end of this test method.

<sup>3</sup> For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

- [C1832 Test Method for Determination of Uranium Isotopic Composition by the Modified Total Evaporation \(MTE\) Method Using a Thermal Ionization Mass Spectrometer](#)
- [C1845 Practice for The Separation of Lanthanide Elements from Uranium Matrices Using High Pressure Ion Chromatography \(HPIC\) for Isotopic Analyses by Inductively Coupled Plasma Mass Spectrometry \(ICP-MS\)](#)
- [D1193 Specification for Reagent Water](#)
- [E180 Practice for Determining the Precision of ASTM Methods for Analysis and Testing of Industrial and Specialty Chemicals \(Withdrawn 2009\)<sup>4</sup>](#)
- [E244 Test Method for Atom Percent Fission in Uranium and Plutonium Fuel \(Mass Spectrometric Method\) \(Withdrawn 2001\)<sup>4</sup>](#)
- [E267 Test Method for Uranium and Plutonium Concentrations and Isotopic Abundances](#)

### 3. Terminology

#### 3.1 Definitions:

3.1.1 For definitions of terms used in this test method but not defined herein, refer to Terminology [C859](#).

### 4. Summary of Test Method

4.1 Fission product neodymium (Nd) is chemically separated from irradiated fuel and determined by isotopic dilution mass spectrometry. Enriched <sup>150</sup>Nd is selected as the Nd isotope diluent, and the mass-142 position is used to monitor for natural Nd contamination. The two rare earths immediately adjacent to Nd do not interfere. Interference from other rare earths, such as natural or fission product <sup>142</sup>Ce or natural <sup>148</sup>Sm and <sup>150</sup>Sm is avoided by removing them in the chemical purification (4 and 5).

4.2 After addition of a blended <sup>150</sup>Nd, <sup>233</sup>U, and <sup>242</sup>Pu spike to the sample, the Nd, U, and Pu fractions are separated from each other by ion exchange. Each fraction is further purified for mass analysis. Two alternative separation procedures are provided.

4.3 The gross alpha, beta, and gamma decontamination factors are in excess of 10<sup>3</sup> and are normally limited to that value by traces of <sup>242</sup>Cm, <sup>147</sup>Pm, and <sup>241</sup>Am, respectively (and sometimes <sup>106</sup>Ru), none of which interferes in the analysis. The 70 ng <sup>148</sup>Nd minimum sample size recommended in the procedure is large enough to exceed by 100-fold a typical natural Nd blank of 0.7 ± 0.7 ng <sup>148</sup>Nd (for which a correction is made) without exceeding radiation dose rates of 20 μ Sv/h (20(20 mR mR/h)/h) at 1 m. Since a constant amount of fission products is taken for each analysis, the radiation dose from each sample is similar for all burnup values and depends principally upon cooling time. Gamma dose rates vary from 200 μ Sv/h (20(20 mR mR/h)/h) at 1 m for 60-day cooled fuel to 20 μ Sv/h (2 mR/h) at 1 m for 1-year cooled fuel. Beta dose rates are an order of magnitude greater, but can be shielded out with a 1/2-in. (12.7 mm) (12.7 mm) thick plastic sheet. By use of such simple local shielding, dilute solutions of irradiated nuclear fuel dissolver solutions can be analyzed for burnup without an elaborate shielded analytical facility. The decontaminated Nd fraction is mounted on a rhenium (Re) filament for mass analysis. Samples from 20 ng to 20 μg run well in the mass spectrometer with both NdO<sup>+</sup> and Nd<sup>+</sup> ion beams present. The metal ion is enhanced by deposition of carbonaceous material on the filament as oxygen getter. (Double and triple filament designs do not require an oxygen getter.)

### 5. Significance and Use

5.1 The burnup of an irradiated nuclear fuel can be determined from the amount of a fission product formed during irradiation. Among the fission products, <sup>148</sup>Nd has the following properties to recommend it as an ideal burnup indicator: (1) It is not volatile, does not migrate in solid fuels below their recrystallization temperature, and has no volatile precursors. (2) It is nonradioactive and requires no decay corrections. (3) It has a low destruction cross section and formation from adjacent mass chains can be corrected for. (4) It has good emission characteristics for mass analysis. (5) Its fission yield is nearly the same for <sup>235</sup>U and <sup>239</sup>Pu and is essentially independent of neutron energy. (6) It has a shielded isotope, <sup>142</sup>Nd, which can be used for correcting natural Nd contamination. (7) It is not a normal constituent of unirradiated fuel.

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<sup>4</sup> The last approved version of this historical standard is referenced on [www.astm.org](http://www.astm.org).

**TABLE 1 K Factors to Correct <sup>148</sup>Nd for <sup>147</sup>Nd Thermal Neutron Capture<sup>A</sup>**

Total Neutron Flux, $\phi$ (neutrons/cm <sup>2</sup> /s)	Total Neutron Exposure, $\phi t$ (neutrons/cm <sup>2</sup> )				
	$1 \times 10^{20}$	$3 \times 10^{20}$	$1 \times 10^{21}$	$2 \times 10^{21}$	$3 \times 10^{21}$
$3 \times 10^{12}$	0.9985	0.9985	0.9985	0.9985	0.9985
$1 \times 10^{13}$	0.9956	0.9952	0.9950	0.9950	0.9950
$3 \times 10^{13}$	0.9906	0.9870	0.9856	0.9853	0.9852
$1 \times 10^{14}$	0.9858	0.9716	0.9598	0.9569	0.9559
$3 \times 10^{14}$	0.9835	0.9592	0.9187	0.9008	0.8941
$1 \times 10^{15}$	0.9826	0.9526	0.8816	0.8284	0.8006

<sup>A</sup> Assuming continuous reactor operation and a  $274 \pm 91$  barn <sup>147</sup>Nd effective neutron absorption cross section for a thermal neutron power reactor. This cross section was obtained by adjusting the  $440 \pm 150$  barn <sup>147</sup>Nd cross section (7) measured at  $20 \pm 20$  °C to a Maxwellian spectrum at a neutron temperature of  $300 \pm 300$  °C.

5.1.5 Its fission yield is nearly the same for <sup>235</sup>U and <sup>239</sup>Pu and is essentially independent of neutron energy (6).

5.1.6 It has a shielded isotope, <sup>142</sup>Nd, which can be used for correcting natural Nd contamination.

5.1.7 It is not a normal constituent of unirradiated fuel.

5.2 The analysis of <sup>148</sup>Nd in irradiated fuel does not depend on the availability of preirradiation sample data or irradiation history. Atom percent fission is directly proportional to the <sup>148</sup>Nd-to-fuel ratio in irradiated fuel. However, the production of <sup>148</sup>Nd from <sup>147</sup>Nd by neutron capture will introduce a systematic error whose contribution must be corrected for. In power reactor fuels, this correction is relatively small. In test reactor irradiations where fluxes can be very high, this correction can be substantial (see Table 1).

5.3 The test method can be applied directly to U fuel containing less than 0.5 % initial Pu with 1 to 100 GW days/metric ton burnup. For fuel containing 5 to 50 % initial Pu, increase the Pu content by a factor of 10 to 100, respectively in both reagents 5.36.3 and 5.46.4.

## 6. Reagents and Materials

NOTE 1—For chemical processing (dissolution and separations) use of the appropriate referenced practices in Section 2 is acceptable: Practices C1168, C1347, C1411, and C1845.

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.<sup>5</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water as defined in Specification D1193.

6.3 *Blended <sup>148</sup>Nd, <sup>239</sup>Pu, and <sup>238</sup>U Calibration Standard*—Prepare a solution containing about 0.0400 mg <sup>148</sup>Nd/litre, 50 mg <sup>238</sup>U/litre, and 2.5 mg <sup>239</sup>Pu/litre, in nitric acid (HNO<sub>3</sub>, 1 + 1) with 0.01 M hydrofluoric acid (HF) as follows. With a new calibrated, clean, Kirk-type micropipet, add 0.500 mL of <sup>239</sup>Pu known solution (see 5.16.11) to a calibrated 1-litre volumetric flask. Rinse the micropipet into the flask three times with HNO<sub>3</sub> (1 + 1). In a similar manner, add 0.500 mL of <sup>238</sup>U known solution (see 5.126.12) and 1.000 mL of <sup>148</sup>Nd known solution (see 5.96.9). Add 10 ten drops of concentrated HF and dilute exactly to the 1-litre mark with HNO<sub>3</sub> (1 + 1) and mix thoroughly.

6.3.1 From  $K_{148}$  (see 5.96.9), calculate the atoms of <sup>148</sup>Nd/mL of calibration standard,  $C_{148}$ , as follows:

$$C_{148} = \frac{\text{mL } ^{148}\text{Nd known solution}}{1000 \text{ mL calibration standard}} \times K_{148} \quad (1)$$

<sup>5</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

6.3.2 From  $K_{238}$  (see [5.46.12](#)), calculate the atoms of  $^{238}\text{U}/\text{mL}$  of calibration standard,  $C_{238}$ , as follows:

$$C_{238} = \frac{\text{mL } ^{238}\text{U known solution}}{1000 \text{ mL calibration standard}} \times K_{238} \quad (2)$$

6.3.3 From  $K_{239}$  (see [5.46.11](#)), calculate the atoms of  $^{239}\text{Pu}/\text{mL}$  of calibration standard,  $C_{239}$ , as follows:

$$C_{239} = \frac{\text{mL } ^{239}\text{Pu known solution}}{1000 \text{ mL calibration standard}} \times K_{239} \quad (3)$$

6.3.4 Flame seal 3 to ~~5-mL~~ 5 mL portions in glass ampoules to prevent evaporation after preparation until time of use. For use, break off the tip of an ampoule, pipet promptly the amount required, and discard any unused solution. If more convenient, calibration solution can be flame-sealed in pre-measured ~~1000- $\mu\text{L}$~~  1000  $\mu\text{L}$  portions for quantitative transfer when needed.

6.4 *Blended  $^{150}\text{Nd}$ ,  $^{233}\text{U}$ , and  $^{242}\text{Pu}$  Spike Solution*—Prepare a solution containing about 0.4 mg  $^{150}\text{Nd}$ /litre, 50 mg  $^{233}\text{U}$ /litre, and 2.5 mg  $^{242}\text{Pu}$ /litre in  $\text{HNO}_3$  (1 + 1) with 0.01 M HF. These isotopes are obtained in greater than 95, 99, and 99 % isotopic purity, respectively, from the ~~Isotopes Sales Department of DOE Isotope Business Office located at Oak Ridge National Laboratory-Laboratory.~~ Standardize the spike solution as follows:

6.4.1 In a ~~5-mL~~ 5 mL beaker, place about 0.1 mL of ferrous solution, exactly 500  $\mu\text{L}$  of calibration standard (see [5.36.3](#)) and exactly 500  $\mu\text{L}$  of spike solution (see [5.46.4](#)). In a second beaker, place about 0.1 mL of ferrous solution and 1 mL of calibration standard without any spike. In a third beaker, place about 0.1 mL of ferrous solution and 1 mL of spike solution without standard. Mix well and allow to stand for 5 min to reduce Pu (VI) to Pu (III) or Pu (IV).

6.4.2 Follow the procedure described in [7.2.49.2.4 – 7.5.89.5.8](#) or [7.6.29.6.2 – 7.7.119.7.11](#). Measure the Pu, U, and Nd isotopes by surface ionization mass spectrometry following the procedure described in [7.8.19.8.1 – 7.8.3.29.8.3.2](#). On the Pu fractions, record the atom ratios of  $^{242}\text{Pu}$  to  $^{239}\text{Pu}$  in the calibration standard,  $C_{2/9}$ ; in the spike solution,  $S_{2/9}$ ; and in the standard-plus-spike mixture,  $M_{2/9}$ . On the U fractions record the corresponding  $^{233}\text{U}$ -to- $^{238}\text{U}$  ratios,  $C_{3/8}$ ,  $S_{3/8}$ , and  $M_{3/8}$ . On the Nd fractions, record the corresponding  $^{150}\text{Nd}$ -to- $^{148}\text{Nd}$  ratios,  $C_{50/48}$ ,  $S_{50/48}$ , and  $M_{50/48}$ . Correct all average measured ratios for mass discrimination bias (see [6.28.2](#)).

6.4.3 Calculate the number of atoms of  $^{150}\text{Nd}/\text{mL}$  of Spike,  $A_{50}$ , as follows:

$$A_{50} = C_{148}[(M_{50/48} - C_{50/48})/(1 - M_{50/48}/S_{50/48})] \quad (4)$$

6.4.4 Calculate the number of atoms of  $^{233}\text{U}/\text{mL}$  of spike,  $A_{33}$ , as follows:

$$A_{33} = C_{238}[(M_{3/8} - C_{3/8})/(1 - M_{3/8}/S_{3/8})] \quad (5)$$

6.4.5 Calculate the number of atoms of  $^{242}\text{Pu}/\text{mL}$  spike,  $A_{42}$ , as follows:

$$A_{42} = C_{239}[(M_{2/9} - C_{2/9})/(1 - M_{2/9}/S_{2/9})] \quad (6)$$

6.4.6 Store in the same manner as the calibration standard (see [5.36.3](#)), that is, flame seal 3 to ~~5-mL~~ 5 mL portions in glass ampoules. For use, break off the tip of an ampoule, pipet promptly the amount required, and discard any unused solution. If more convenient, spike solution can be flame sealed in a premeasured ~~1000- $\mu\text{L}$~~  1000  $\mu\text{L}$  portions for quantitative transfer to individual samples.

6.5 *Ferrous Solution (0.001 M)*—Add 40 mg of reagent grade ferrous ammonium sulfate ( $\text{Fe}(\text{NH}_4)_2(\text{SO}_4)_2 \cdot 6\text{H}_2\text{O}$ ) and 1 drop of concentrated  $\text{H}_2\text{SO}_4$  to 5 mL of redistilled water. Dilute to 100 mL with water and mix. This solution does not keep well. Prepare fresh daily.

6.6 *Filament Mounting Solution*—Dissolve 70 mg of sucrose in 100 mL of water (single filament only).

6.7 *Hydrofluoric Acid*—Reagent grade concentrated HF (28 M).

6.8 *Methanol*, absolute.

6.9 <sup>148</sup>Nd *Known Solution*—Heat natural Nd<sub>2</sub>O<sub>3</sub> (>99.9 % pure) in an open crucible at ~~900°C~~900 °C for 1 h to destroy any carbonates present and cool in a desiccator. Weigh 0.4071 g of Nd<sub>2</sub>O<sub>3</sub> and place it in a calibrated ~~500-mL~~500 mL volumetric flask. Dissolve the oxide in HNO<sub>3</sub> (1 + 1) and dilute to the ~~500-mL~~500 mL mark with HNO<sub>3</sub> (1 + 1) and mix thoroughly. By using the weight of Nd<sub>2</sub>O<sub>3</sub> in grams, and the purity, calculate the atoms of <sup>148</sup>Nd/mL of known solution,  $K_{148}$ , as follows:

$$K_{148} = \text{g Nd}_2\text{O}_3 / 500 \text{ mL} \times \% \text{ purity} / 100 \quad (7)$$

$$\times 50.38 \text{ mg } ^{148}\text{Nd} / 1 \text{ g Nd}_2\text{O}_3 \times (6.025$$

$$\times 10^{20} \text{ atoms}) / 147.92 \text{ molecular weight}$$

6.10 *Perchloric Acid*—70 % HClO<sub>4</sub>.

6.11 <sup>239</sup>Pu *Known Solution*—Add 10 mL of HCl (1 + 1) to a clean calibrated ~~100-mL~~100 mL flask. Cool the flask in an ice water bath. Allow time for the acid to reach approximately 0°C and place the flask in a glove box. Displace the air in the flask with inert gas (Ar, He, or N<sub>2</sub>). Within the glove box, open the U.S. National Institute of Standards and Technology Plutonium Metal Standard Sample 949, containing about 0.5 g of Pu (actual weight individually certified), and add the metal to the cooled HCl. After dissolution of the metal is complete, add 1 drop of concentrated HF and 40 mL of HNO<sub>3</sub> (1 + 1) and swirl. Place the flask in a stainless-steel beaker for protection and invert a ~~50-mL~~50 mL beaker over the top and let it stand for at least 8 days to allow any gaseous oxidation products to escape. Dilute to the mark with HNO<sub>3</sub> (1 + 1) and mix thoroughly. By using the individual weight of Pu in grams, the purity, and the molecular weight of the Pu given on the NIST certificate, with the atom fraction,  $A_9$ , determined as in [8-810.8](#), calculate the atoms of <sup>239</sup>Pu/mL of <sup>239</sup>Pu known solution,  $K_{239}$ , as follows:

$$K_{239} = [(\text{mg Pu} / 100 \text{ mL solution}) \times \% \text{ purity} / 100] \quad (8)$$

$$\times (6.025 \times 10^{20} \text{ atoms/Pu molecular weight}) \times A_9]$$

6.12 <sup>238</sup>U *Known Solution*—Heat U<sub>3</sub>O<sub>8</sub> from the National Institute of Standards and Technology Natural Uranium Oxide Standard Sample 950 in an open crucible at ~~900°C~~900 °C for 1 h and cool in a desiccator in accordance with the certificate accompanying the standard sample. Weigh about 12.0 g of U<sub>3</sub>O<sub>8</sub> accurately to 0.1 mg and place it in a calibrated ~~100-mL~~100 mL volumetric flask. Dissolve the oxide in HNO<sub>3</sub> (1 + 1). Dilute to the ~~100-mL~~100 mL mark with HNO<sub>3</sub> (1 + 1) and mix thoroughly. By using the measured weight of U<sub>3</sub>O<sub>8</sub> in grams, the purity given on the NIST certificate, and the atom fraction <sup>238</sup>U,  $A_8$ , determined as in [8-510.5](#), calculate the atoms <sup>238</sup>U/mL of <sup>238</sup>U solution,  $K_{238}$ , as follows:

$$K_{238} = [( \text{g U}_3\text{O}_8 / 100 \text{ mL solution}) \times (\% \text{ purity} / 100) \quad (9)$$

$$\times 848.0 \text{ mg U} / 1 \text{ g U}_3\text{O}_8) \times (6.025$$

$$\times 10^{20} \text{ atoms} / 238.03 \text{ molecular weight}) \times A_8$$

6.13 *Reagents and Materials for Procedure A:*

6.13.1 *Dowex AGMP-1 Resin*—Convert Dowex AGMP-1 (200 to 400 mesh) chloride form resin to nitrate form by washing 200 mL of resin in a suitable column (for example, a ~~250-mL~~250 mL buret) with HNO<sub>3</sub> (1 + 1) until a drop of effluent falling into an AgNO<sub>3</sub> solution remains clear. Finally, rinse with water, and dry overnight in a vacuum desiccator. Store the resin in an airtight container. Since the elution characteristics of ion exchange resins depend upon their actual percentage cross linkage and particle size (surface-to-volume ratio), which may vary from one lot to the next, it is most convenient to set aside a bottle of resin to be used solely for this procedure. Before use on actual samples, a small amount of tracer <sup>147</sup>Nd should be taken through the procedure. Collect each consecutive 80 mm fraction of eluant and count for  $\gamma$  radioactivity. If over 80 % of the <sup>147</sup>Nd appears in the Nd fraction, the resin can be used as directed; if not, small adjustments can be made in the elution volumes collected.

6.13.2 *Hydrochloric Acid*—Prepare reagent low in U and dissolved solids by saturating redistilled water in a polyethylene container to 12 M with HCl gas which has passed through a quartz-wool filter. Dilute 1 + 1 and 1 + 24 with redistilled water. Store each solution in a polyethylene container. One drop of acid, when evaporated on a clean microscope slide cover glass, must leave no visible residue. Test monthly. Commercial HCl (cp grade) in glass containers has been found to contain excessive residue (dissolved glass) which inhibits ion emission.

6.13.3 *Dowex 1 Resin*—Dowex 1-X4 (200 to 400 mesh) chloride form resin.

6.13.4 *Ion Exchange Column (Type I)*—Type I ion exchange columns are used whenever Dowex AG 1-X4 columns are specified in the procedure. These columns are prepared from ~~230-mm~~ 230 mm disposable glass capillary (Pasteur) pipets that have a glass wool plug inserted to contain the resin beads. Filling this column to the top is considered a ~~2-mL~~ 2 mL addition of reagent solution.

6.13.5 *Ion Exchange Column (Type II)*—Type II ion exchange columns are used whenever AGMP-1 columns are specified in the procedure. These columns are prepared from ~~4-mm~~ 4 mm (inside diameter) glass tubing that has been heated and drawn, forming a long, fine tip. A coating of paraffin wax melted on the long tip keeps the methanol from climbing the outside surface. A small plug of glass wool is inserted to contain the resin beads. The length of the column above the glass wool plug should be a little more than 22 cm. The columns are carefully marked every 4 cm above the top of the resin bed (4 cm = 0.5 mL of solution).

6.13.6 *Methanolic HNO<sub>3</sub> Eluant*—Pipet 10 mL of HNO<sub>3</sub> (1 + 500) into a ~~100-mL~~ 100 mL volumetric flask and dilute to the mark with absolute methanol. Protect this reagent against preferential evaporation of methanol by keeping it in a polyethylene wash bottle. Prepare fresh daily.

6.13.7 *Methanolic HNO<sub>3</sub> Loading Solution*—Pipet 1 mL of HNO<sub>3</sub> (1 + 1) into a ~~10-mL~~ 10 mL volumetric flask and dilute to the mark with absolute methanol. Store as ~~5.13.66.13.6~~. Prepare fresh daily. High nitrate loading solution is used to ensure absorption of Nd in a tight band and to overcome interference from sulfate and fluoride ions.

6.13.8 *Methanolic HNO<sub>3</sub> Wash Solution*—Pipet 10 mL of HNO<sub>3</sub> (1 + 100) into a ~~100-mL~~ 100 mL volumetric flask and dilute to the mark with absolute methanol. Store as ~~5.13.66.13.6~~. Prepare fresh daily.

6.13.9 *Nitric Acid (8 M, 4 M, 3 M)<sup>8</sup>*—Prepare by diluting Ultrapure concentrated HNO<sub>3</sub> (15.6 M) with deionized water.

6.13.10 *Sodium Nitrite Stock Solution(2M)*—Dissolve 3 g of reagent grade sodium nitrite (NaNO<sub>2</sub>) in 20 mL of 0.1 M NaOH.

6.13.11 *Sodium Nitrite Working Solution*—Dilute 100 µL of stock solution from ~~5.13.106.13.10~~ to 10 mL with 8 M HNO<sub>3</sub>. Prepare fresh daily.

#### 6.14 *Reagents and Materials for Alternative Procedure B:*

6.14.1 *Eluting Solution (0.094 M HNO<sub>3</sub> in 80 % CH<sub>3</sub>OH)*—Prepare 100 mL of 0.47 M HNO<sub>3</sub> by diluting 3.00 mL of ~~15.6-15.6 M~~ HNO<sub>3</sub> to 100 mL with water in a volumetric flask. Prepare the eluting solution just before use by pipetting ~~20.0-mL~~ 20.0 mL of the 0.47 M HNO<sub>3</sub> into a ~~100-mL~~ 100 mL volumetric flask and diluting to the mark with anhydrous methanol. The methanol must be free of aldehydes. Absence of a characteristic aldehyde odor is an adequate criterion.

6.14.2 *First Column Resin*—Transfer a water slurry of analytical grade macroporous anion resin (~~AGMP-1~~)(AGMP-1); 50 to 100 mesh, chloride-form resin to a column until the settled height is just below the reservoir. Pass 4 mL of water through, then ~~6-mL~~ 6 mL of 12 M HCl. Keep the resin wet with 12 M HCl until use.

6.14.3 *Hydrochloric Acid (12 M, 0.1 M)*—Using plastic apparatus and an ice bath, bubble filtered HCl gas through quartz-distilled acid until it is saturated. Verify 12 M concentration by titration with standard base. Prepare the 0.1 M by dilution with quartz-distilled water.

6.14.4 *Hydrofluoric Acid (1 M)*—Dilute 1 mL of concentrated analytical reagent grade HF to 30 mL with quartz-distilled water.

6.14.5 *Hydroiodic Acid-Hydrochloric Acid Mixture (0.1 M HI-12 M HCl)*—Dilute 1 mL of distilled 57 % HI to 74 mL with 12 M HCl. Prepare fresh for each use. Store distilled HI in flame-sealed bottles to prevent air oxidation.

6.14.6 *Hydrogen Peroxide (30 %)*—Refrigerate when not in use.

6.14.7 *Ion Exchange Column*—Use commercial disposable polyethylene droppers, 6 mm inside diameter and 60 mm long, with a ~~2-mL~~ 2 mL reservoir. Cut off the top of the dropper to form a reservoir and place a glass wool plug in the tip to support the resin bed. The reservoir of the second column can be made cylindrical to accommodate the feeder by inserting as a mold a 1-dram glass vial and heating with a hot air gun. Cool and remove the glass vial mold.

6.14.8 *Loading Solution (1.56 M HNO<sub>3</sub> in 80 % CH<sub>3</sub>OH)*—Prepare 100 mL of 7.8 M HNO<sub>3</sub> by diluting 50 mL of quartz-distilled HNO<sub>3</sub> to 100 mL with water. Prepare the loading solution by diluting 20 mL of 7.8 M HNO<sub>3</sub> to 100 mL with anhydrous methanol. The methanol must be free of aldehydes. The absence of a characteristic aldehyde odor is an adequate criterion.

6.14.9 *Nitric Acid (15.6 M, 2 M, 1 M)*<sup>8</sup>—Dilute quartz-distilled 15.6 M HNO<sub>3</sub> with distilled water to prepare the 2 M HNO<sub>3</sub> and 1 M HNO<sub>3</sub>.

6.14.10 *Perchloric Acid (6 M)*—Dilute 12 M HClO<sub>4</sub> with water.

6.14.11 *Second Column Feeder*—Use polyethylene dispensing bottles (coaxial tip) of about ~~30-mL~~30 mL capacity. Cut off the delivery tip to a length of about 15 mm.

6.14.12 *Second Column Resin*—Convert ~~AGMP-1~~AGMP-1<sub>2-7</sub><sup>8</sup>, 200 to 400 mesh, chloride-form resin to nitrate form. One satisfactory method is to fill a ~~30-mm~~30 mm diameter by ~~120-mm~~120 mm high glass column with a water slurry of the resin, then pass 160 mL of HNO<sub>3</sub> (1 + 1) and 160 mL of loading solution (5.2.8) through the column. Verify the absence of chloride by AgNO<sub>3</sub> test on the final effluent. Store the resin in loading solution in a closed container until ready for use. When a sample is ready, transfer the resin to a column to a settled height just below the reservoir and keep wet with loading solution until use. Nitrate-form resin is not as chemically stable as chloride-form resin. For this reason it is best not to store nitrate-form resin for longer than a few months.

## 7. Hazards

7.1 Hydrofluoric acid is a highly corrosive and toxic acid that can severely burn skin, eyes, and mucous membranes. Hydrofluoric acid differs from other acids because the fluoride ion readily penetrates the skin, causing destruction of deep tissue layers. Unlike other acids that are rapidly neutralized, hydrofluoric acid reactions with tissue may continue for days if left untreated. Familiarization and compliance with the Safety Data Sheet is essential.”

## 8. Instrument Calibration

NOTE 2— For the determination of isotopic distributions and elemental assay use of the appropriate referenced test methods in Section 2 is acceptable: Test Methods C1267, C1415, C1625, C1672, and C1832.

8.1 In the calibration of the mass spectrometer for the analysis of Nd, U, and Pu, the measurement and correction of mass discrimination bias is an important factor in obtaining accurate and consistent results. The mass discrimination bias can be readily measured on natural Nd where the <sup>142</sup>Nd-to-<sup>150</sup>Nd ratio spans over a 5 % spread in mass. The mass discrimination bias factor, *B*, is constant for Nd, U, and Pu analysis for a given method of scanning (for example, by varying either acceleration voltage or magnetic field) and for a given method of detection (for example, by pulse counting or current integration) on a given detector (for example, electron multiplier, scintillation detector, or d-c collector plate). Calculate *B* as follows:

$$B = (1/c) \left[ \left( \bar{R}_{ij}/R_s \right) - 1 \right] \quad (10)$$

where:

where:

$R_{ij}^-$  = average measured atom ratio of isotope *i* to isotope *j*. For the most accurate determination of *B*, let  $R_{ij}^-$  be the average measured atom ratio of <sup>142</sup>Nd to <sup>150</sup>Nd,

$R_s$  = known value of the measured atom ratio. For the ratio of <sup>142</sup>Nd to <sup>150</sup>Nd in natural neodymium,  $R_s = 4.824$ , and

*e* =  $\Delta$  mass/mass. The value of *e* for various ratios and ion species include

*c* =  $\Delta$  mass/mass.

Ratio	Nd <sup>+</sup> , U <sup>+</sup> , or Pu <sup>+</sup>	NdO <sup>+</sup> , UO <sub>2</sub> <sup>+</sup> , or PuO <sub>2</sub> <sup>+</sup>
<sup>148</sup> Nd/ <sup>150</sup> Nd	+2/150	+2/166
<sup>150</sup> Nd/ <sup>148</sup> Nd	-2/148	-2/164
<sup>142</sup> Nd/ <sup>150</sup> Nd	+8/150	+8/166
<sup>234</sup> U/ <sup>238</sup> U	+4/238	+4/270
<sup>235</sup> U/ <sup>238</sup> U	+3/238	+3/270
<sup>236</sup> U/ <sup>238</sup> U	+2/238	+2/270
<sup>238</sup> U/ <sup>233</sup> U	-5/233	-5/265
<sup>233</sup> U/ <sup>238</sup> U	+5/238	+5/270
<sup>240</sup> Pu/ <sup>239</sup> Pu	-1/239	-1/271