



Designation: ~~E3007--15~~ E3007 – 21

Standard Practice for ~~Kinetic Values Used to Evaluate Selection and Use of~~ Kinetic Reference Values in the Study of Decomposition Reactions by Thermogravimetry¹

This standard is issued under the fixed designation E3007; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. ~~Scope~~ Scope*

1.1 It is the purpose of this practice to provide kinetic parameters for reference material(s) to be used for evaluation of thermogravimetry methods, apparatus, and software where decomposition mass loss and associated temperature are measured. This practice addresses n th order reactions.

1.2 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.3 There is no International Organization of Standards (ISO) equivalent to this standard.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate ~~safety~~ safety, health, and health environmental practices and determine the applicability of regulatory limitations prior to use.*

1.4 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

2. Referenced Documents

2.1 ASTM Standards:²

[E473 Terminology Relating to Thermal Analysis and Rheology](#)

[E1142 Terminology Relating to Thermophysical Properties](#)

[E1641 Test Method for Decomposition Kinetics by Thermogravimetry Using the Ozawa/Flynn/Wall Method](#)

[E2958 Test Methods for Kinetic Parameters by Factor Jump/Modulated Thermogravimetry](#)

NOTE 1—Test Methods [E1641](#) and [E2958](#) determine the pre-exponential factor A in units of min^{-1} . These results shall be converted to units of s^{-1} for comparison to values quoted in this practice using:

$$\ln[A, \text{s}^{-1}] = \ln[A, \text{min}^{-1}] + \ln[60 \text{ s/min}] = \ln[A, \text{min}^{-1}] + 4.094$$

¹ This practice is under the jurisdiction of ASTM Committee [E37](#) on Thermal Measurements and is the direct responsibility of Subcommittee [E37.02](#) on Standard Reference Materials.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the ~~standard's~~ Document Summary page on the ASTM website.

*A Summary of Changes section appears at the end of this standard

3. Terminology

3.1 *Definitions*—Specific technical terms used in this practice are defined in Terminologies E473 and E1142, including Arrhenius equation, derivative, reaction, and thermogravimetric analysis.

3.1.1 *decomposition, n*—the process by which a substance is broken up into constituent parts, elements or simpler compounds.

3.1.2 *kinetics, chemical, n*—the study of the dependence of the chemical reaction rate on temperature and time.

4. Summary of Practice

4.1 Kinetics is the measurement and study of the rate of a chemical reaction to the independent parameters of time and temperature. This relationship is often described using the Arrhenius expression, where:

$$d\alpha/dt = A f(\alpha) \exp[-E/RT] \quad (1)$$

and:

- α = fraction reacted (dimensionless),
- $f(\alpha)$ = ~~some function of α ,~~
- $f(\alpha)$ \equiv some function of α ,
- E = activation energy (J mol⁻¹),
- R = ~~gas constant (= 8.314 J mol⁻¹ K⁻¹),~~
- R \equiv gas constant (=8.314 J mol⁻¹ K⁻¹),
- T = absolute temperature (K),
- t = time (s),
- exp = natural logarithm base,
- da/dt = ~~fraction reaction rate (s⁻¹),~~
- da/dt \equiv fraction reaction rate (s⁻¹),
- A = ~~pre-exponential factor (s⁻¹), and~~
- n = reaction order.

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4.2 The function $f(\alpha)$ is commonly in the form:

$$f(\alpha) = (1 - \alpha)^n \quad (2)$$

and is known as an n th order reaction model.

4.3 Eq 1 may be evaluated in either its exponential or logarithmic form:

$$\ln[da/dt] = \ln[A] + \ln[f(\alpha)] - E/RT \quad (3)$$

4.4 The study of kinetics involves the determination of values of E , A , and n for a given chemical reaction.

NOTE 2—Activation energy and pre-exponential factor are not independent parameters.

NOTE 3—The descriptions provided in Eq 1-3 are only mathematical models. That is, they represent the fitting of mathematical equations to often “noisy” experimental data. In practice, no such model will faithfully describe the complete reaction(s) under all conditions for the materials described in this practice.

4.5 Values for the kinetic parameters are typically in the ranges of:

$$\begin{aligned} 20 < E < 350 \text{ kJ/mol,} \\ 5 < \ln[A] < 50 \text{ with } A \text{ in s}^{-1}, \text{ and} \\ 0.0 < n < 2.0 \end{aligned}$$

4.6 Some, thermally reactive materials may change with time. For this reason, certified reference materials may not be available for use in the evaluation of kinetic parameters. For any purchased reference material(s) of high purity from a commercial laboratory supply house or national metrology institute, the source, storage conditions and storage time of the material(s) shall be reported.

4.7 The recommended kinetic values for the thermally reactive materials identified in **Table 1** are derived from a meta-analysis of literature values.³ These reference values may be used for the evaluation of methods, apparatus or software where decomposition mass loss and temperature are measured.>

NOTE 4—Kinetic parameters are known to be affected by specimen particle size. Large particle size is thought to retard escaping decomposition produced thereby affecting the reaction kinetics. While not specifically addressed by this practice, the user is advised to cryogenically grind samples for 1 minute prior to use.

NOTE 5—Kinetic parameters may be affected by specimen atmosphere. Kinetic results obtained in an inert atmosphere shall be used for comparison purposes (see **Table 1**).

5. Significance and Use

5.1 The kinetic parameters provided in this practice may be used to evaluate the performance of a standard, apparatus, technique or software for the determined mean parameters (such as Test Methods **E1641** and **E2958**). The results obtained by these approaches may be compared to the values provided by this practice.

6. Hazards

6.1 Thermally reactive materials evolve decomposition products as part of the indicated reaction. These products may be toxic or corrosive. Operators shall use as small an amount of material as is practical for the measurement. Handling of these materials shall be performed by trained workers who are knowledgeable with their Safety Data Sheets (SDS) and adopt appropriate safety precautions according to local regulations.

7. Procedure

7.1 Experimentally determined mean kinetic parameters for a reference material are compared to the values described in this practice as their quotient, expressed as percent. The closeness of the conformance quotient to 100 % is an indication of accuracy. Values less than 100% indicate that the determined value is less than the reference value while those greater than unity or 100 % indicate that the determined value is greater than the reference value.

NOTE 6—The kinetic values for polytetrafluoroethylene are considered the standard. Other values presented in **Table 1** are provided for user information.

8. Calculation

8.1 $\text{Conformance} = (\text{Observed Value} \times 100 \%) / (\text{Referenced Value})$

NOTE 7—Observed values include $E, \ln[A, \text{s}^{-1}]$ and n .

8.2 *Example Calculation*—If the observed experimental value for activation energy of polytetrafluoroethylene decomposition is $E_o = 320 \text{ kJ/mol}$ and the reference value of E_r taken from **Table 1** is 319 kJ/mol , then:

$$\text{Conformance} = (320 \text{ kJ/mol} \times 100 \%) / 319 \text{ kJ/mol} = 100 \%$$

TABLE 1 Kinetic Parameters for Kinetic Reference Materials (see Note 1 in 2.1 and Note 6 in 7.1)

NOTE 1—Where: E = activation energy, A = pre-exponential factor, and n is reaction order.

NOTE 2—Derivation of these values is described in supporting data filed at ASTM International Headquarters and may be obtained by requesting Research Report RR:E37-1044. Contact ASTM Customer Service at service@astm.org.

| Material | CAS | E , kJ/mol | $\ln[A, \text{s}^{-1}]$ | n | Description |
|--------------------------------|------------------|-------------------------|---------------------------|--------------------|------------------|
| Calcium Carbonate | 471-34-1 | 194 ± 17 (8.7 %) | 17.3 ± 1.5 (8.7 %) | 0.51 ± 0.42 | nitrogen |
| Calcium Oxalate Monohydrate | 24804-31-70 | 97 ± 16 (16 %) | 19.3 ± 3.1 (16 %) | 0.83 ± 0.18 | nitrogen and air |
| Polytetrafluoroethylene | 9002-84-0 | 319 ± 24 (7.5 %) | 45.1 ± 3.4 (7.5 %) | 1.00 ± 0.01 | nitrogen |
| Polyethylene | 9002-88-4 | 255 ± 41 (16 %) | 37.5 ± 6.0 (16 %) | 0.96 ± 0.12 | nitrogen |
| Polypropylene | 9003-07-0 | 232 ± 27 (12 %) | 34.3 ± 4.1 (12 %) | 0.96 ± 0.12 | nitrogen |
| Polystyrene | 9003-53-6 | 211 ± 60 (28 %) | 34.0 ± 9.5 (28 %) | 0.96 ± 0.15 | nitrogen |

³ Supporting data have been filed at ASTM International Headquarters and may be obtained by requesting Research Report RR:E37-1044. Contact ASTM Customer Service at service@astm.org.