



Designation: **E508—13 E508 – 21**

## Standard Test Method for Determination of Calcium and Magnesium in Iron Ores by Flame Atomic Absorption Spectrometry<sup>1</sup>

This standard is issued under the fixed designation E508; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reappraisal. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reappraisal.

### 1. Scope

1.1 This test method covers the determination of calcium and magnesium in iron ores, concentrates, and agglomerates in the mass fraction (%) range from 0.05 % to 5 % of calcium and 0.05 % to 3 % of magnesium.

1.2 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.4 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

#### 2.1 ASTM Standards:<sup>2</sup>

[D1193 Specification for Reagent Water](#)

[E29 Practice for Using Significant Digits in Test Data to Determine Conformance with Specifications](#)

[E50 Practices for Apparatus, Reagents, and Safety Considerations for Chemical Analysis of Metals, Ores, and Related Materials](#)

[E135 Terminology Relating to Analytical Chemistry for Metals, Ores, and Related Materials](#)

[E276 Test Method for Particle Size or Screen Analysis at No. 4 \(4.75-mm\) Sieve and Finer for Metal-Bearing Ores and Related Materials](#)

[E663 Practice for Flame Atomic Absorption Analysis \(Withdrawn 1997\)<sup>3</sup>](#)

[E882 Guide for Accountability and Quality Control in the Chemical Analysis Laboratory](#)

### 3. Terminology

3.1 *Definitions*—For definitions of terms used in this test method, refer to Terminology [E135](#).

### 4. Summary of Test Method

4.1 The sample is dissolved in HCl with the addition of a small amount of HNO<sub>3</sub>. After evaporation and dehydration, any insoluble

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee [E01](#) on Analytical Chemistry for Metals, Ores, and Related Materials and is the direct responsibility of Subcommittee [E01.02](#) on Ores, Concentrates, and Related Metallurgical Materials.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](#), or contact ASTM Customer Service at [service@astm.org](#). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

<sup>3</sup> The last approved version of this historical standard is referenced on [www.astm.org](#).

residue is filtered, treated for the recovery of calcium and magnesium, and added to the main solution. The solution is aspirated into the flame of the atomic absorption spectrometer. For calcium, a nitrous oxide-acetylene burner is used, for magnesium, an air-acetylene burner is preferred (**Note 1**). The absorbed energy at the resonance line of the spectrum of calcium or magnesium, or both, is measured by atomic absorption spectrometry and compared with that of the calibration solutions of the two elements.

**NOTE 1**—If a nitrous oxide burner is used for the determination of magnesium, the sensitivity is decreased by a factor of about three.

If an air-acetylene burner is used for the determination of calcium, several elements, if present in a mass fraction (%) of 1.0 or greater, interfere with the calcium determination. Also, lanthanum chloride (LaCl<sub>3</sub>) must be employed and the sensitivity is decreased by a factor of about two.

## 5. Significance and Use

5.1 This test method is intended as a referee method for compliance with compositional specifications for impurity content. It is assumed that all who use this procedure will be trained analysts capable of performing common laboratory practices skillfully and safely. It is expected that work will be performed in a properly equipped laboratory and that proper waste disposal procedures will be followed. Follow appropriate quality control practices such as those described in Guide **E882**.

## 6. Interferences

6.1 None of the elements normally found in iron ores interfere with this test method (**Note 1**).

## 7. Apparatus

7.1 *Atomic Absorption Spectrometer*, meeting the following criteria:

7.1.1 *Minimum Sensitivity*—The absorbance of the highest calibration solution (see **8.4**) must be at least 0.3.

7.1.2 *Curve-Linearity*—The difference between the readings of the highest calibration solutions must be more than 1.4 times the difference between the readings for the zero solutions and the lowest calibration solution.

7.1.3 *Minimum Stability*—The coefficient of variation of a number of measurements of the highest calibration solution and of the zero calibration solution must be less than 1.5 % and 0.5 % respectively.

**NOTE 2**—Use of digital readout devices or printers, or both, is advisable. **E508-21**

**NOTE 3**—Instrument parameters will vary with each instrument. The following parameters were successfully used in several laboratories and they can be used as guidelines. In the case of calcium, solutions were aspirated into a nitrous oxide-acetylene flame of a premix burner. For magnesium an air-acetylene burner was used.

	Ca	Mg
Hollow cathode lamp, mA	15	15
Wavelength, nm	422.7	285.2
Nitrous oxide flow rate, L/min	13.8	...
Air flow rate, L/min	...	22
Acetylene flow rate, L/min	6.6	4.2

## 8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society where such specifications are available.<sup>4</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Type I or Type II of Specification **D1193**. Type III or Type IV may be used if they effect no measurable change in the blank or sample.

<sup>4</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC—DC, [www.chemistry.org](http://www.chemistry.org). For suggestions on the testing of reagents not listed by the American Chemical Society, see the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD—MD, <http://www.usp>.

8.3 *Calcium, Standard Solution* (1 mL = 25 µg Ca)—Dissolve 1.2487 g of anhydrous calcium carbonate (CaCO<sub>3</sub>) in 100 mL of dilute HCl (1 + 3). When solution is complete, cool and dilute to 1 L in a volumetric flask. Transfer 10 mL of this solution to a ~~200-mL~~200 mL volumetric flask. Dilute to volume and mix.

8.4 *Calcium and Magnesium Calibration Solutions*—Using pipets, transfer, (5.0, 10.0, 20.0, 40.0, and ~~50.0~~50.0) mL portions of the standard calcium and magnesium solutions to 200-mL volumetric flasks (**Note 4**). Dilute to 90 mL and add 6 mL of HCl, 60 mL of background solution, and 40 mL of LaCl<sub>3</sub> solution. Dilute to volume and mix.

8.5 *Iron, Background Solution* (10 g/L)—Dissolve 10 g of high-purity iron in 50 mL of HCl, and oxidize by the dropwise addition of HNO<sub>3</sub>. Evaporate to a syrupy consistency. Add 20 mL of HCl and dilute to 200 mL with water. Add 20 g of sodium chloride dissolved in 200 mL water and dilute to 1 L with water.

8.6 *Lanthanum Chloride Solution* (50 g/L)—Dissolve 50 g of LaCl<sub>3</sub> in 50 mL of HCl and 300 mL of hot water. Cool and dilute to 1 L.

8.7 *Magnesium, Standard Solution* (1 mL = 17.5 µg Mg)—Dissolve 0.3500 g of high-purity magnesium by slowly adding 75 mL of HCl (1 + 3). When dissolved, dilute with water to 1 L in a volumetric flask. Transfer 10 mL of this solution to a ~~200-mL~~200 mL volumetric flask, dilute to volume, and mix.

8.8 *Zero Calibration Solution*—Transfer 60 mL of the background solution to a ~~200-mL~~200 mL volumetric flask, add 6 mL of HCl, 40 mL of LaCl<sub>3</sub> solution, dilute to volume, and mix.

NOTE 4—The range of calcium or magnesium that can be covered may vary from instrument to instrument. Attention should be paid to the minimum criteria given in 7.1.

## 9. Hazards

9.1 For precautions to be observed in this method, refer to Practices E50.

## 10. Sample Preparation

10.1 The analytical sample shall be pulverized to pass a ~~150-µm~~150 µm sieve (see Test Method E276).

NOTE 5—To facilitate decomposition, some ores, such as specular hematite, require grinding to pass a ~~75-µm~~75 µm sieve.

## 11. Procedure

11.1 Transfer approximately 1.0 g of the sample to a small weighing bottle previously dried at about 150 °C. Dry the bottle and contents for 1 h at 105 °C to 110 °C. Cap the bottle and cool to room temperature in a desiccator. Momentarily release the cap to equalize the pressure and weigh the capped bottle and sample to the nearest 0.1 mg. Repeat the drying and weighing until there is no further loss of mass. Transfer the sample to a ~~250-mL~~250 mL beaker and reweigh the capped bottle to the nearest 0.1 mg. The difference between the two masses is the mass of the sample taken for analysis.

11.2 Carry a reagent blank through all steps of the procedure starting with 11.3.

11.3 *Decomposition of Sample*—Moisten the sample with a few milliliters of water, add 25 mL of HCl, cover, and digest below the boiling point until no further attack is apparent. Add 2 mL of HNO<sub>3</sub> and digest for several minutes. Remove the cover and evaporate the solution to dryness. Heat the salts on a hot plate at 105 °C to 110 °C for 15 min. Add 5 mL of HCl, cover the beaker, and warm for several minutes. Add 50 mL of water and warm until soluble salts are in solution. Filter on a fine-textured paper, and collect the filtrate in a ~~250-mL~~250 mL beaker. Transfer the residue quantitatively to the filter paper. Wash alternately three times with hot HCl (1 + 9) and hot water until the yellow color of ferric chloride is no longer apparent. Transfer the paper and residue to a platinum crucible. Reserve the filtrate.

11.4 *Treatment of Residue*—Ignite the paper and residue in the platinum crucible. Cool, moisten with several drops of water, add