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Standard Terminology Relating to Thermophysical Properties¹

This standard is issued under the fixed designation E1142; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This is a compilation of terms and corresponding definitions commonly used in the study of thermophysical properties. Terms that are generally understood or defined adequately in other readily available sources are either not included or their sources identified.

1.2 A definition is a single sentence with additional information included in a *Discussion*.

1.3 Definitions of terms specific to a particular field (such as *dynamic mechanical measurements*) are identified with an italicized introductory phrase.

1.4 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.5 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

2. Referenced Documents

2.1 *ASTM Standards:*²

[D4092 Terminology for Plastics: Dynamic Mechanical Properties](#)

[E7 Terminology Relating to Metallography](#)

[E344 Terminology Relating to Thermometry and Hydrometry](#)

[E2744 Test Method for Pressure Calibration of Thermal Analyzers](#)

3. Terminology

3.1 *Definitions:*

absolute pressure, n —pressure measured relative to zero pressure corresponding to empty space.

DISCUSSION—

Absolute pressure is atmospheric pressure plus gage pressure.

activation energy (E), n —in chemical kinetics, the energy that must be overcome in order for a chemical reaction to occur.

¹ This terminology is under the jurisdiction of ASTM Committee E37 on Thermal Measurements and are the direct responsibility of Subcommittee E37.03 on Nomenclature and Definitions.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

DISCUSSION—

The term *activation energy* was introduced in 1889 by Svante Arrhenius as a mathematical term in the eponymous, empirical relationship between temperature and reaction rate constant.

admittance, Y, n —the reciprocal of impedance.

alpha (α) loss peak, n —in *dynamic mechanical measurement*, first peak in the damping curve below the melt, in order of decreasing temperature or increasing frequency. **E7**

amorphicity, n —a relative measure of amorphous material content, expressed as a percent of the total material content.

angular frequency, ω, n —the number of radians per second traversed by a rotating vector that represents any periodically varying quantity.

DISCUSSION—

Angular frequency, ω , is equal to two π times the frequency, f .

anisotropic, *adj*—having different values for a property in different directions.

anti-thixotropy, n —an increase of the apparent viscosity under constant shear stress or shear rate followed by a gradual recovery when the stress or shear rate is reduced to zero.

arrhenius equation, n —a mathematical relationship between the specific reaction rate constant and the temperature given as:

$$k = Ae^{-E/RT} \quad (1)$$

where:

- k = the reaction rate constant,
- A = the pre-exponential factor,
- E = the energy of activation,
- R = the gas constant, and
- T = the absolute temperature.

atmospheric pressure, n —the pressure due to the weight of the atmosphere. **E2744**

DISCUSSION—

Atmospheric pressure varies with elevation above sea level, acceleration due to gravity, and weather conditions.

autocatalytic reaction, n —a chemical reaction in which a product or reaction intermediate functions as a catalyst.

barometer, n —an instrument for measuring atmospheric pressure.

beta (β) loss peak, n —in *dynamic mechanical measurement*, second discrete peak in damping curve below the melt, in order of decreasing temperature or increasing frequency. **D4092**

boiling pressure, n —at a specific temperature, the value of the vapor pressure of the liquid at which it is equal to the external pressure.

boiling temperature, n —at a specific pressure, the temperature at which the vapor pressure of the liquid is equal to the external pressure.

capacitance, n —that property of a system of conductors and dielectrics that permits the storage of electrical charge when a potential difference exists between the conductors.

DISCUSSION—

Capacitance is the ratio of a quantity of electric charge, Q , to a potential difference, V . A capacitance value is always positive. The unit of capacitance is the farad, F, which is equivalent to one coulomb per volt.

catalyst, n —a substance that increases the rate of a chemical reaction but is not consumed or changed by that reaction.

Celsius, n —designation of the degree on the International Practical Temperature Scale; also used for the name of the scale, as “Celsius Temperature Scale.” Formerly (prior to 1948) called “Centigrade.” The Celsius temperature scale is related to the International Kelvin Temperature Scale by the equation: $T_c = T - 273.16$ K.

Centigrade, n —see **Celsius**.

coefficient of expansion, n —see **coefficient of linear thermal expansion**.

coefficient of linear thermal expansion, α, n —change in length, relative to the length of the specimen, accompanying a unit change of temperature, at a specified temperature.

coefficient of viscosity, n —the ratio between an infinitesimally small increase in stress and the corresponding increase in strain rate.

coefficient of volume thermal expansion α_v, n —for a solid or liquid, the change in volume, relative to the volume of the specimen, accompanying a change of temperature at a specified temperature.

color temperature, n —temperature in degrees Kelvin (K) at which a black body must be operated to give a color equal to that of the source in question.

complex modulus, E^*, G^* , or K^*, n —ratio of the stress to strain where each is a factor that may be represented by a complex number as follows: $E^* = E' + iE''$, $G^* = G' + iG''$, and $K^* = K' + iK''$.

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where:

E^* = complex modulus, measured in tension or flexure,

E' = storage modulus, measured in tension or flexure,

E'' = loss modulus, measured in tension or flexure,

G^* = complex modulus, measured in shear,

G' = storage modulus, measured in shear,

G'' = loss modulus, measured in shear,

K^* = complex modulus, measured in compression,

K' = storage modulus, measured in compression

K'' = loss modulus, measured in compression, and

$i = \sqrt{-1}$, measured in compression.

The complex modulus may be measured in tension or flexure, (E^*), compression, (K^*), or in shear, (G^*).

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complex shear compliance, J^*, n —reciprocal of complex shear modulus, where $J^* = 1/G^*$.

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complex tensile compliance, D^*, n —reciprocal of complex tensile modulus, where $D^* = 1/E^*$.

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complex viscosity, η^*, n —the complex modulus divided by the imposed frequency in rad/s.

compliance, J, n —the strain divided by the corresponding stress.

DISCUSSION—

Compliance is the reciprocal of modulus.

composition, *n*—quantity of the components of a mixture; usually expressed in terms of the weight percentage, or the atomic percentage of each of the components in the mixture. **E7**

conductivity, electrical (volume), σ , *n*—the ratio of the current density ($A\cdot cm^{-2}$) through a specimen to the potential gradient (V/cm) in the same direction as the current.

DISCUSSION—

Conductivity is normally expressed in units $(ohm\cdot cm)^{-1}$, but the correct SI units are Siemen-m.

congruent phases, *n*—those states of matter of unique composition that co-exist at equilibrium at a single point in temperature and pressure; for example, the two coexisting phases of a two-phase equilibrium. **E7**

congruent transformation, *n*—an isothermal, or isobaric, phase change in which both of the phases concerned have the same composition throughout the process; the order of a system becomes unary at a composition of congruency. **E7**

constitutional diagram, *n*—graphical representation of the compositions, temperatures, pressures, or combinations thereof at which the heterogeneous equilibria of a system occur.

cooling curve, *n*—graphical representation of specimen temperature or temperature change as a function of time or decreasing environment temperature.

cooling rate, *n*—average slope of the time-temperature curve taken over a specific time and temperature interval as the temperature is decreased.

critical curve, *n*—in a binary, or higher order, phase diagram, a locus of points along which two or more phases exist in stable thermodynamic equilibrium.

critical point, *n*—in a binary phase diagram, that specific value of composition, temperature, pressure, or combinations thereof at which the phases of a heterogeneous equilibrium become identical. ^a

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critical pressure, *n*—that pressure at the critical point.

critical surface, *n*—in a ternary or higher order phase diagram, the area upon which the phases in equilibrium become identical. **E7**

critical temperature, *n*—that temperature at the critical point.

crystal, *n*—solid composed of atoms, ions, or molecules, arranged in a pattern which is periodic in three dimensions. **E7**

crystallinity, *n*—regular arrangement of the atoms of a solid in space.

DISCUSSION—

In most materials, this state is usually imperfectly achieved. The crystalline regions (ordered regions) are submicroscopic volumes in which there is more or less regularity of arrangement of the component molecules.

crystallite, *n*—crystalline grain not bounded by habit planes. **E7**

crystallization, *n*—arrangement of previously disordered material segments of repeating patterns into geometric symmetry.

crystallization temperature, *n*—that temperature at which a specimen undergoes crystallization upon cooling.

Curie point, n—see **Curie temperature**.

Curie temperature, n—temperature above which a ferromagnetic or ferroelectric material becomes paramagnetic, or paraelectric, respectively.

DISCUSSION—

There may be more than one if there are multiple materials.

damping, n—loss in energy, dissipated as heat, that results when a material or material system is subjected to an oscillatory load or displacement. **D4092**

devitrification, n—crystallization of an amorphous substance. **E7**

dielectric constant, n—see **permittivity, relative**.

dielectric dissipation factor, D, n—the ratio of the loss factor, ϵ'' , to the absolute permittivity, ϵ' , or:

$$D = \epsilon''/\epsilon' \quad (2)$$

DISCUSSION—

The dielectric dissipation factor is numerically equal to the tangent of the dielectric loss angle and may be referred to as the loss tangent, $\tan \delta$, or the cotangent of the phase angle, θ .

dielectric loss angle, n—the angle whose tangent is the dissipation factor or $\arctan \epsilon''/\epsilon'$.

DISCUSSION—

It is also the difference between 90 degrees and the phase angle.

differential thermocouple, n—see **differential thermopile**.

differential thermopile, n—a number of temperature sensors connected in series-opposing and arranged so that there is an increase in output signal for a given temperature difference between alternate junctions maintained at a reference temperature and the measured temperature.

dilatancy, n—the increase in volume caused by shear.

dipole relaxation time, γ, n —the exponential decay time required for the electric polarization of any point of a suitably charged dielectric to fall from its original value to $1/e$ of that value, due to the loss of dipole orientation.

DISCUSSION—

Under conditions of an alternating applied field and in systems with a single dipole relaxation time, it is equal to $1/\omega$ at the loss factor peak in cases where the peak is caused by a dipole mechanism.

dissipation factor, n—see **tangent delta**.

dissociation, n—as applied to heterogeneous equilibria, the transformation of one phase into two or more new phases, all of different composition. **E7**

dynamic modulus, n—see **complex modulus**.

elasticity, n—that property of materials that causes them to return to their original form or condition after the applied force is removed. **D4092**

elastic limit, n—the greatest stress that can be applied to a material without permanent deformation.

elastic modulus, n —the ratio of stress to corresponding strain within the elastic limit of the stress-strain curve.

DISCUSSION—

The elastic modulus may also be measured in tension (E'), compression (K'), flexure (E''), or shear (G'). (See also **complex modulus**.)

enthalpy, n —a thermodynamic *function* defined by the equation $H = U + PV$ where H is the enthalpy, U is the *internal* energy, P is the pressure, and V the volume of the system.

DISCUSSION—

At constant pressure the change in enthalpy measures the quantity of heat exchanged by the system and its surrounding.

equilibrium diagram, n —see **constitutional diagram**.

eutectic point, n —see **eutectic**.

eutectic, *adj*—mixture of two or more substances which solidifies as a whole when cooled from the liquid state, without change in composition.

DISCUSSION—

The temperature at which the eutectic mixture solidifies is called the eutectic point. This temperature is constant for a given composition, and represents the lowest melting point of the system.

expansivity, n —the change in dimension resulting from an infinitesimal change in an independent variable (such as temperature or humidity).

failure, n —the point beyond which a material ceases to be functionally capable of its intended use.

failure criterion, n —specification of the chemical, physical, mechanical, electrical, or other condition under which a material ceases to be functionally capable of its intended use.

failure temperature (T_F), n —the temperature at which a material fails.

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Fahrenheit, n —designation of a degree on the Fahrenheit temperature scale that is related to the International Practical Temperature Scale by means of the equation:

$$T_F = 1.8 T_C + 32 \quad (3)$$

where:

T_F = the temperature in degree Fahrenheit, and

T_C = the temperature in degrees Celsius.

freezing temperature, n —see **crystallization temperature**.

frequency, f, n —the number of cycles per unit time of periodic process.

DISCUSSION—

The unit is Hertz (Hz) which is equal to 1 cycle per/s.

frequency profile, n —in *dynamic mechanical measurement*, plot of the dynamic properties of a material, at a constant temperature, as a function of test frequency. **D4092**

gage pressure, n —pressure measured relative to atmospheric pressure.

DISCUSSION—

Gage pressure is the difference between absolute pressure and atmospheric pressure.

gamma (γ) loss peak, *n*—in *dynamic mechanical measurement*, third peak in the damping curve below the melt, in the order of decreasing temperature or increasing frequency. **D4092**

Gibbs Phase Rule, *n*—maximum number of phases (*P*) that may coexist at equilibrium is equal to two, plus the number of components (*C*) in the mixture, minus the number of degrees of freedom (*F*): $P + F = C + 2$. **E7**

glass transition, *n*—reversible change in an amorphous material or in amorphous regions of a partially crystalline material, from (or to) a viscous or rubbery condition to (or from) a hard and relatively brittle one.

DISCUSSION—

The glass transition generally occurs over a relatively narrow temperature region and is similar to the solidification of a liquid to a glassy state. Not only do hardness and brittleness undergo rapid changes in this temperature region, but other properties, such as coefficient of thermal expansion and specific heat capacity, also change rapidly. This phenomenon sometimes is referred to as a second order transition, rubber transition, or rubbery transition. When more than one amorphous transition occurs in a material, the one associated with segmental motions of the backbone molecular chain, or accompanied by the largest change in properties is usually considered to be the glass transition.

glass transition temperature, *n*—a temperature chosen to represent the temperature range over which the glass transition takes place.

DISCUSSION—

The glass transition temperature can be determined readily by observing the temperature region at which a significant change takes place in some specific electrical, mechanical, thermal, or other physical property. Moreover, the observed temperature can vary significantly depending on the property chosen for observation and on details of the experimental technique (for example, heating rate, frequency of test). Therefore, the observed T_g should be considered valid only for that particular technique and set of test conditions.

heat capacity, *n*—quantity of heat necessary to change the temperature of an entity, substance or system by one Kelvin of temperature.

DISCUSSION—

The SI units of measurement are J/K .

impedance, *Z*, *n*—the ratio of the time dependent voltage, $v(t)$, across a circuit, a circuit element, or material to the time dependent current, $i(t)$, through it; that is:

$$Z = v(t)/i(t) \quad (4)$$

DISCUSSION—

The impedance of a circuit, circuit element, or material is a measure of its ability to oppose the transmission of an alternating current. It is expressed in ohms. Its value depends on the angular frequency, ω , of the measurement.

invariant equilibrium, *n*—stable state among a number of phases exceeding by two the number of components in the system and in which more of the external variables (pressure, temperatures, or concentrations) may be varied without causing a decrease in the number of phases present. **E7**

invariant point, *n*—point defined by the unique values of temperature, pressure, and concentrations in a system with the maximum number of phases that can coexist in equilibrium. **E7**

isohume, *n*—constant relative humidity.

isotropic, *adj*—having the same magnitude of a property in all directions.

Kelvin, *n*—designation of the thermodynamic temperature scale and the degree on this scale. **E344**

DISCUSSION—

This Kelvin scale was defined by the 10th General Conference on Weights and Measure in 1954 by assigning the temperature of 273.16 K to the triple point of water. Also, the degree on the International Practical Kelvin Temperature Scale.

kinetics (chemical), *n*—the study of rates of chemical reactions.