



Designation: D3561 – 16 (Reapproved 2021)^{e1}

Standard Test Method for Lithium, Potassium, and Sodium Ions in Brackish Water, Seawater, and Brines by Atomic Absorption Spectrophotometry¹

This standard is issued under the fixed designation D3561; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

^{e1} NOTE—The WTO caveat was editorially added in November 2021.

1. Scope

1.1 This test method covers the determination of soluble lithium, potassium, and sodium ions in brackish water, seawater, and brines by atomic absorption spectrophotometry.²

1.2 Samples containing from 0.1 to 70 000 mg/L of lithium, potassium, and sodium may be analyzed by this test method.

1.3 This test method has been used successfully with artificial brine samples. It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

1.4 The values stated in SI units are to be regarded as standard. The values given in parentheses are mathematical conversion to inch-pound units that are provided for information only and are not considered standard.

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.6 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

¹ This test method is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.05 on Inorganic Constituents in Water.

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² Fletcher, G. F. and Collins, A. G., *Atomic Absorption Methods of Analysis of Oilfield Brines: Barium, Calcium, Copper, Iron, Lead, Lithium, Magnesium, Manganese, Potassium, Sodium, Strontium, and Zinc*. U.S. Bureau of Mines, Report of Investigations 7861, 1974, 14 pp., Collins, A. G. *Geochemistry of Oilfield Waters*, Elsevier, New York, NY, 1975.

2. Referenced Documents

2.1 *ASTM Standards*:³

D1129 Terminology Relating to Water

D1193 Specification for Reagent Water

D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water

D3370 Practices for Sampling Water from Flowing Process Streams

D5810 Guide for Spiking into Aqueous Samples

D5847 Practice for Writing Quality Control Specifications for Standard Test Methods for Water Analysis

3. Terminology

3.1 *Definitions*:

3.1.1 For definitions of terms used in this standard, refer to Terminology D1129.

4. Summary of Test Method

4.1 This test method is dependent on the fact that metallic elements, in the ground state, will absorb light of the same wavelength they emit when excited. When radiation from a given excited element is passed through a flame containing ground state atoms of that element, the intensity of the transmitted radiation will decrease in proportion to the amount of ground state element in the flame. A hollow cathode lamp whose cathode is made of the element to be determined provides the radiation. The metal atoms to be measured^{4,5} are placed in the beam of radiation by aspirating the specimen into

³ For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

⁴ Angino, E. E., and Billings, G. K., *Atomic Absorption Spectrophotometry in Geology*, Elsevier Publishing Co., New York, NY, 1967.

⁵ Dean, J. A., and Rains, T. C., Editors, *Flame Emission and Atomic Absorption Spectrometry*, Volume 1, Theory, Marcel Dekker, New York, NY, 1969.

an oxidant fuel flame. A monochromator isolates the characteristic radiation from the hollow cathode lamp, and a photosensitive device measures the attenuated transmitted radiation, which may be read as absorbance units or directly as concentration on some instruments.

4.2 Since the variable and sometimes high concentrations of matrix materials in the waters and brines affect absorption differently, it is difficult to prepare standards sufficiently similar to the waters and brines. To overcome this difficulty, the method of additions is used in which three identical samples are prepared and varying amounts of a standard added to two of them. The three samples are then aspirated, the concentration readings recorded, and the original sample concentration calculated.

5. Significance and Use

5.1 Identification of a brackish water, seawater, or brine is determined by comparison of the concentrations of their dissolved constituents. The results are used to evaluate the water as a possible pollutant, or as a commercial source of a valuable constituent such as lithium.

6. Interferences

6.1 Ionization interference is controlled by adding large excesses of an easily ionized element. Sodium ion is added in the potassium and lithium determinations, and potassium ion is added in the sodium determinations.

7. Apparatus

7.1 *Atomic Absorption Spectrophotometer*—The instrument shall consist of an atomizer and burner, suitable pressure-regulating devices capable of maintaining constant oxidant and fuel pressure for the duration of the test, a hollow cathode lamp for each metal to be tested, an optical system capable of isolating the desired line of radiation, an adjustable slit, a photomultiplier tube or other photosensitive device as a light measuring and amplifying device, and a readout mechanism for indicating the amount of absorbed radiation.

7.1.1 *Multielement Hollow-Cathode Lamps.*

7.2 *Pressure-Reducing Valves*—The supplies of fuel and oxidant shall be maintained at pressures somewhat higher than the controlled operating pressure of the instrument by suitable valves.

8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specification of the Committee on Analytical Reagents of the American Chemical Society,⁶ where such specifications are available. Other grades may be used, provided it is first ascertained that the reagent is of

sufficiently high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification **D1193**, Type I. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the precision and bias of the test method. Type III water was specified at the time of round robin testing of this test method.

8.3 *Lithium Solution, Standard* (1 mL = 1 mg Li)—Dissolve 5.324 g of lithium carbonate (Li_2CO_3) in a minimum volume of HCl (1 + 1). Dilute to 1 L with water. One millilitre of this solution contains 1 mg of lithium. A purchased stock solution of adequate purity is also acceptable.

8.4 *Potassium Solution, Stock* (1 mL = 100 mg K)—Dissolve 190.7 g of potassium chloride (KCl) in water and dilute to 1 L with water. A purchased stock solution of appropriate known purity is also acceptable.

8.5 *Potassium Solution, Standard* (1 mL = 1 mg K)—Dissolve 1.907 g of potassium chloride (KCl) in water and dilute to 1 L with water. One millilitre of this solution contains 1 mg of potassium. A purchased stock solution of appropriate known purity is also acceptable.

8.6 *Sodium Solution, Stock* (1 mL = 100 mg Na)—Dissolve 254.2 g of sodium chloride (NaCl) in water and dilute to 1 L with water. A purchased stock solution of appropriate known purity is also acceptable.

8.7 *Sodium Solution, Standard* (1 mL = 10 mg Na)—Dissolve 25.42 g of sodium chloride (NaCl) in water and dilute to 1 L with water. One millilitre of this solution contains 1 mg of sodium. A purchased stock solution of appropriate known purity is also acceptable.

8.8 *Oxidant:*

8.8.1 Air that has been cleaned and dried through a suitable filter to remove oil, water, and other foreign substances, is the usual oxidant.

8.9 *Fuel:*

8.9.1 *Acetylene*—Standard, commercially available acetylene is the usual fuel. Acetone, always present in acetylene cylinders, can be prevented from entering and damaging the burner head by replacing a cylinder that has only 689.4 kPa (100 psi) of acetylene remaining.

8.10 *Filter Paper*—Purchase suitable filter paper. Typically the filter papers have a pore size of 0.45- μm membrane. Material such as fine-textured, acid-washed, ashless paper, or glass fiber paper are acceptable. The user must first ascertain that the filter paper is of sufficient purity to use without adversely affecting the bias and precision of the test method.

9. Sampling

9.1 Collect the sample in accordance with the applicable ASTM standard (see Practices **D3370**).

⁶ *ACS Reagent Chemicals, Specifications and Procedures for Reagents and Standard-Grade Reference Materials*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

10. Procedure

10.1 Potassium is determined at the 766.5-nm wavelength, lithium at the 670.8-nm wavelength, and sodium at the 330.2 to 330.3-nm wavelength with an air-acetylene flame. For much greater sensitivity, sodium is determined at the 589.0 to 589.6-nm wavelength.

10.2 *Preliminary Calibration*—Using micropipets prepare lithium standards containing 1 to 5 mg/L of lithium, potassium standards containing 1 to 5 mg/L of potassium, and sodium standards containing 100 to 500 mg/L of sodium using the standard lithium (8.3), potassium (8.5), and sodium (8.7) solutions to 50-mL volumetric flasks. Before making up to volume, add 0.5 mL of the sodium stock (8.6) solution to the potassium and lithium standards, and to a blank. Before

making up to volume, add 0.5 mL of the potassium stock (8.4) solution to the sodium standards and to a blank. Aspirate these standards and the appropriate blank (for background setting) and adjust the curvature controls, if necessary, to obtain a linear relationship between absorbance and the actual concentration of the standards.

10.3 Transfer an aliquot of water or brine (previously filtered through a 0.45- μ m filter [8.10]) to a 50-mL volumetric flask. The specific gravity of the water or brine can be used to estimate the lithium, potassium, or sodium content of the sample and, thereby, serve as a basis for selecting the aliquot sizes that will contain about 0.05 mg of lithium, 0.05 mg of potassium, or 5 mg of sodium. Fig. 1 shows the relationship between sodium concentration and specific gravity for some oil

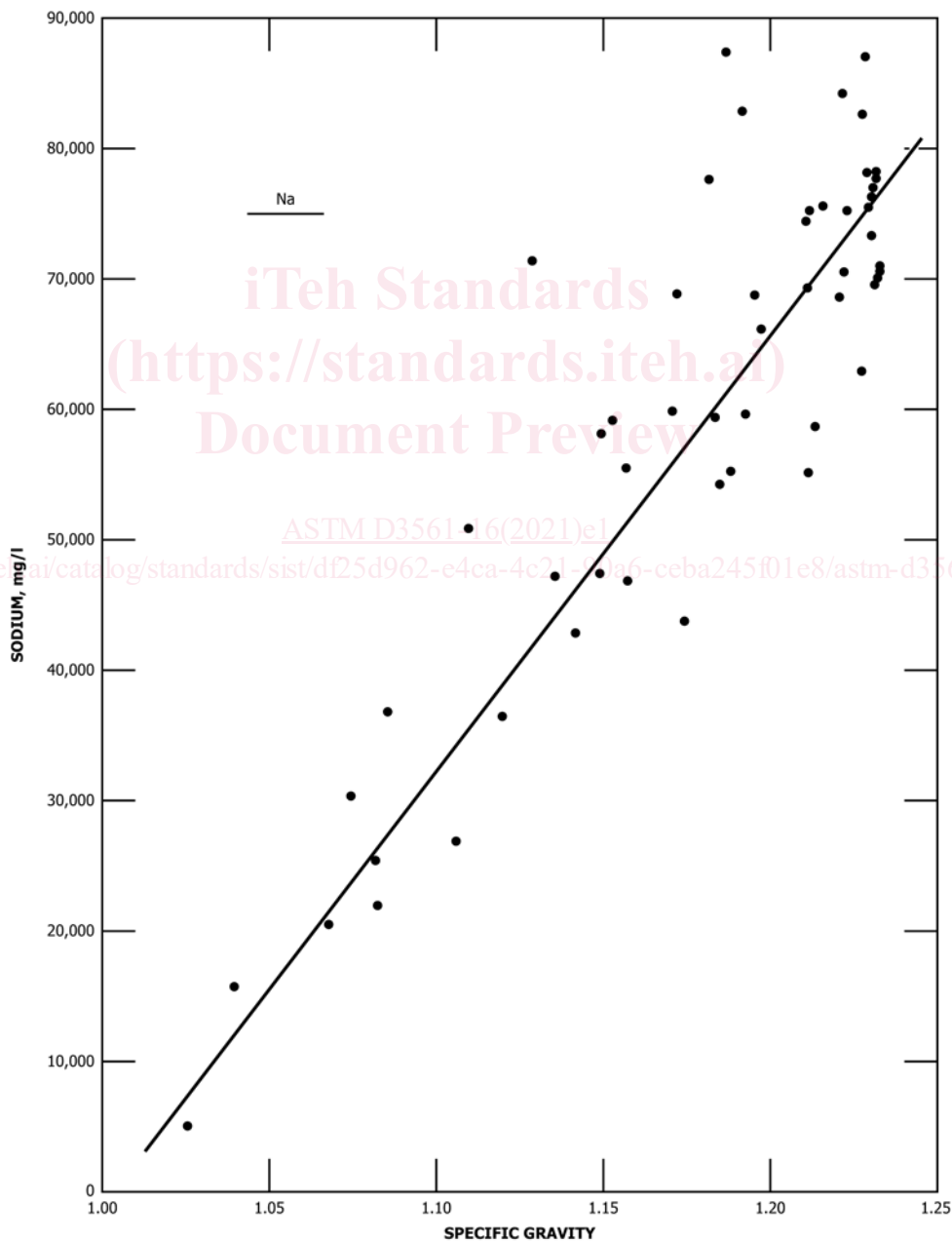


FIG. 1 Relationship of the Concentration of Sodium in Some Oilfield Brines to Specific Gravity

field brines from the Smackover formation. The concentrations of sodium and also of lithium and potassium will not necessarily correlate with the concentrations found in other formations. Therefore, the user of this test method may find it necessary to draw similar curves for brine samples taken from other formations. Add 0.5 mL of the sodium stock (8.6) solution to the lithium and potassium samples and 0.5 mL of the potassium stock (8.4) solution to the sodium samples, dilute to volume, and aspirate. Calculate the approximate sample concentration from the preliminary calibration readings, and determine the aliquot sizes that will contain about 0.05 mg of lithium, 0.05 mg of potassium, or 5 mg of sodium.

10.4 Transfer equal aliquots containing about 0.05 mg of potassium or lithium, or 5 mg of sodium to three 50-mL volumetric flasks. Add no potassium or lithium standard to the first flask, using a micropipet add 0.05 mg to the second, and 0.1 mg to the third. For sodium, add no standard to the first flask, 5 mg to the second, and 10 mg to the third.

10.5 Add 0.5 mL of the sodium stock (8.6) solution to the potassium and lithium samples and 0.5 mL of the potassium stock (8.4) solution to the sodium samples, dilute to volume, aspirate, and record the absorbance readings for each sample.

11. Calculation

11.1 Calculate the concentration of potassium, lithium, or sodium ion in the original sample in milligrams per litre as follows:

11.2

$$\text{Concentration, mg/L} = \frac{V_1(A_s \times C_{std})}{V_2(A_{std} - A_s)} \quad (1)$$

where:

- V_1 = volume of the dilute sample, mL,
- V_2 = volume of the original sample, mL,
- A_s = absorbance of dilute sample,
- A_{std} = absorbance of one of the standard additions, and
- C_{std} = concentration of the same standard addition as A_{std} ² in mg/L.

Since there are two standard additions, calculate for each and average the two results.

12. Precision and Bias⁷

12.1 The precision of this test method within its designated range may be expressed as follows:

Lithium,

$$S_t = 0.0677X + 3.127$$

$$S_o = 0.0486X + 1.936$$

Potassium,

$$S_t = 0.1443X - 2.317$$

$$S_o = 0.0847X - 61.15$$

Sodium,

$$S_t = 0.08905X + 729$$

$$S_o = 0.0295X + 195$$

where:

- S_t = overall precision,
- S_o = single-operator precision, and
- X = concentration of lithium, potassium, or sodium determined, mg/L.

12.2 The bias of this test method determined from recoveries of known amounts of lithium, potassium, and sodium in a series of prepared standards were as follows:

Lithium, Amount Added, mg/L	Recovery, % Relative
21.0	102.0
52.3	101.1
74.1	100.5
164	95.0
Potassium, Amount Added, mg/L	Recovery, % Relative
591	111.0
1650	110.9
1670	113.2
1921	125.2
Sodium, Amount Added, mg/L	Recovery, % Relative
9 140	105.7
29 000	103.9
62 500	105.4
66 200	108.3

NOTE 1—The preceding precision and bias estimates are based on an interlaboratory study of lithium, potassium, and sodium and interfering ions as shown in Table 1. Two analysts in each of four laboratories and one analyst in each of two laboratories performed duplicate determinations on each of two days. Practice D2777 was used in developing these precision and bias estimates.

12.3 It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

12.4 Precision and bias for this test method conforms to Practice D2777 – 77, which was in place at the time of collaborative testing. Under the allowances made in 1.4 of Practice D2777 – 13, these precision and bias data do meet existing requirements for interlaboratory studies of Committee D19 test methods.

13. Quality Control

13.1 In order to be certain that analytical values obtained using these test methods are valid and accurate within the

TABLE 1 Compositions of Artificial Brine Samples

Sample No.	g/L			
	1	2	3	4
Na ^A	9.14	62.5	29.0	66.2
K ^B	0.591	1.670	1.650	1.921
Li ^C	0.0210	0.0523	0.0741	0.164
CaCl ₂	1.47	2.86	1.93	4.67
MgCl ₂ ·6H ₂ O	9.40	10.19	4.12	1.99
BaCl ₂	0.05	0.95	0.48	0.47

^A Added to the aqueous solution as NaCl.

^B Added to the aqueous solution as KCl, KBr, and KI: Sample No. 1 contained 0.470 g/L of KCl, 0.979 g/L of KBr, and 0.098 g/L of KI; Sample No. 2 contained 1.774 g/L of KCl, 1.909 g/L of KBr, and 0.477 g/L of KI; Sample No. 3 contained 1.721 g/L of KCl, 1.928 g/L of KBr, and 0.482 g/L of KI; and Sample No. 4 contained 2.072 g/L of KCl, 1.870 g/L of KBr, and 0.934 g/L of KI.

^C Added to the aqueous solution as LiCl.

⁷ Supporting data have been filed at ASTM International Headquarters and may be obtained by requesting Research Report RR:D19-1029. Contact ASTM Customer Service at service@astm.org.