



Designation: ~~E1584 – 17~~ E1584 – 23

## Standard Test Method for Assay of Nitric Acid<sup>1</sup>

This standard is issued under the fixed designation E1584; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

### 1. Scope\*

1.1 This test method covers determination of the assay of nitric acid by total acidity.

1.2 ~~This test method covers determination of the assay of nitric acid by total acidity.~~ This test method is suitable for concentrations between approximately ~~50~~ 50 % and 70 %, calculated as nitric acid.

1.3 The following applies for the purposes of determining the conformance of the test results using this test method to applicable specifications, results shall be rounded off in accordance with the rounding-off method of Practice ~~E29~~.

1.4 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

~~1.3 Review the current Safety Data Sheets (SDS) for detailed information concerning toxicity, first aid procedures, and safety precautions.~~

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use. Specific precautionary statements are given in* For specific hazard statements, see Section 8.

1.6 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

#### 2.1 ASTM Standards:<sup>2</sup>

[E29 Practice for Using Significant Digits in Test Data to Determine Conformance with Specifications](#)

[D1193 Specification for Reagent Water](#)

[D6809 Guide for Quality Control and Quality Assurance Procedures for Aromatic Hydrocarbons and Related Materials](#)

[E180 Practice for Determining the Precision of ASTM Methods for Analysis and Testing of Industrial and Specialty Chemicals](#)

(Withdrawn 2009)<sup>3</sup>

[E200 Practice for Preparation, Standardization, and Storage of Standard and Reagent Solutions for Chemical Analysis](#)

[E300 Practice for Sampling Industrial Chemicals](#)

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee D16 on Aromatic, Industrial, Specialty and Related Chemicals and is the direct responsibility of Subcommittee D16.10 on Acids.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

<sup>3</sup> The last approved version of this historical standard is referenced on www.astm.org.

\*A Summary of Changes section appears at the end of this standard

2.2 Other Documents:

OSHA Regulations, 29 CFR, paragraphs 1910.1000 and 1910.1200<sup>4</sup>

### 3. Summary of Test Method

3.1 A weighed sample of acid is diluted in water and titrated with 1.0 *N* sodium hydroxide solution, using phenolphthalein as the end-point indicator.

### 4. Significance and Use

4.1 This test method provides a means for assaying nitric acid, based on total acidity. The concentration of nitric acid is important in many of the uses of nitric acid, including specification compliance and manufacturing control.

### 5. Interferences

5.1 Acids other than nitric, and compounds that consume sodium hydroxide, will yield erroneously high results.

### 6. Apparatus

6.1 *Erlenmeyer Flask*, 250 mL, glass stoppered.

6.2 *Buret*, 50 mL, Class A.

NOTE 1—A digital buret or automated titrator capable of measuring volumes to the nearest 0.01 mL may be used in place of a conventional buret.

### 7. Reagents

7.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society where such specifications are available.<sup>5</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

7.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water as defined by Type **HI** or Type **HHI** of Specification **D1193**.

7.3 *Phenolphthalein Indicator Solution* (10 g/L)—Dissolve 1 g of phenolphthalein in 100 mL of ethanol (95 %), methanol, or isopropanol.<sup>6</sup>

7.4 *Sodium Hydroxide, Standard Solution* (1.0*N*)—Prepare and standardize in accordance with Practice **E200**, remembering to record the temperature of the NaOH solution during the standardization to the nearest °C.

NOTE 2—Because of atmospheric CO<sub>2</sub> absorption by NaOH, 1.0 *N* NaOH solutions will need frequent standardization.

### 8. Safety Precautions-Hazards

8.1 Consult current OSHA regulations, suppliers' Safety Data Sheets, and local regulations for all materials used in this test method.

8.2 Nitric acid is dangerous if handled improperly. It is a strong corrosive acid and a strong oxidizer. Contact with other material may cause fire. Avoid inhalation of fumes or mist and any contact with skin or eyes. Use of safety goggles and gloves highly recommended.

<sup>4</sup> Available from Superintendent of Documents, U.S. Government Printing Office, 732 N. Capitol St. NW, Mail Stop: SDE, Washington, DC 20401.

<sup>5</sup> *Reagent Chemicals, American Chemical Society Specifications, ACS Reagent Chemicals, Specifications and Procedures for Reagents and Standard-Grade Reference Materials*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

<sup>6</sup> This reagent is also described in Practice **E200**.

**TABLE 1 Sample Size for Assay of Nitric Acid<sup>A</sup>**

HNO <sub>3</sub> , %	Sample Size, mL	Nominal Weight, g
70	2.5 ± 0.2	3.6
65	2.8 ± 0.2	3.9
60	3.1 ± 0.2	4.3
55	3.4 ± 0.2	4.6
50	3.8 ± 0.2	5.0

<sup>A</sup> In some cases, where the concentration range of the sample is unknown, a preliminary test may be required to determine the proper size test sample to use.

8.3 Clean up all spills immediately by covering the spill with vermiculite or some other inert absorbent material and sweeping it into a pan. Dispose of the vermiculite by flooding with water and discarding in a suitable container. Flush the area with water.

## 9. Sampling

9.1 Sampling of nitric acid is not within the scope of this test method. General guidelines for sampling may be found in Practice E300.

9.2 The sample to be analyzed shall be considered to be that sample in a single bottle submitted to the analytical laboratory.

## 10. Procedure

10.1 Tare to the nearest 0.1 mg a glass stoppered 250-mL Erlenmeyer flask containing approximately 15 mL of water.

10.2 Remove the stopper. Using a suitable pipet, quickly add an appropriate volume of sample (depending on the acid strength as given in Table 1), while keeping the tip of the pipet just above the surface of the liquid. Immediately replace the stopper, and reweigh the flask to the nearest 0.1 mg to obtain the weight of the sample.

NOTE 3—A disposable dropping pipet is convenient for the sample transfer.

10.3 Remove the stopper, add 50 mL of water and 4 drops of phenolphthalein indicator solution to the flask, and swirl to mix the contents. Record the temperature of the 1.0 N NaOH solution, and then titrate the sample to a pale pink end point. Record the volume of the titrant to the nearest 0.01 mL.

## 11. Calculation

11.1 If necessary, correct the buret reading for calibration errors, and record the volume of the titrant as  $V$  and its temperature during the titration as  $t$ .

11.2 Correct the normality of the sodium hydroxide standard solution for any difference in temperature between time of standardization and time of use in accordance with the following equation:

$$N = N_s + 0.00035 (s - t) \quad (1)$$

where:

- $N$  = normality of NaOH solution at time of titration,
- $N_s$  = normality of NaOH solution at time of standardization,
- $s$  = temperature of NaOH during standardization, °C, and
- $t$  = temperature of NaOH during titration °C.

11.3 Calculate the assay as percent nitric acid as follows:

$$\text{nitric acid, \% by weight} = \frac{VN \times 6.301}{W} \quad (2)$$