



Designation: ~~E2036 – 15~~ E2036 – 24

## Standard Test Method for Nitrogen Trichloride in Liquid Chlorine by High Performance Liquid Chromatography (HPLC)<sup>1</sup>

This standard is issued under the fixed designation E2036; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reappraisal. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reappraisal.

### 1. ~~Scope~~ Scope\*

1.1 This test method uses high performance liquid chromatography (HPLC) to determine nitrogen trichloride levels in liquid chlorine at the  $\theta$ -10.1  $\mu\text{g/g}$  (ppm) to 600  $\mu\text{g/g}$  (ppm) range. Solvent samples from chlorine production facilities containing very high concentrations of nitrogen trichloride may be analyzed by dilution with methanol.

1.2 Review the current safety data sheet (SDS) for detailed information concerning toxicity, first aid procedures, and safety precautions.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 The following applies for the purposes of determining the conformance of the test results using this test method to applicable specifications, results shall be rounded off in accordance with the rounding-off method of Practice E29.

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use. Specific hazard statements are given in Section 8.*

1.6 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

#### 2.1 *ASTM Standards:*<sup>2</sup>

[D6809 Guide for Quality Control and Quality Assurance Procedures for Aromatic Hydrocarbons and Related Materials](#)

[E29 Practice for Using Significant Digits in Test Data to Determine Conformance with Specifications](#)

[E180 Practice for Determining the Precision of ASTM Methods for Analysis and Testing of Industrial and Specialty Chemicals](#) (Withdrawn 2009)<sup>3</sup>

[E691 Practice for Conducting an Interlaboratory Study to Determine the Precision of a Test Method](#)

[E806 Test Method for Carbon Tetrachloride and Chloroform in Liquid Chlorine by Direct Injection \(Gas Chromatographic Procedure\)](#)

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee D16 on Aromatic, Industrial, Specialty and Related Chemicals and is the direct responsibility of Subcommittee D16.13 on Chlorine.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [service@astm.org](mailto:service@astm.org). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

<sup>3</sup> The last approved version of this historical standard is referenced on [www.astm.org](http://www.astm.org).

\*A Summary of Changes section appears at the end of this standard

## 2.2 Federal Standards:<sup>4</sup>

CFR 173 Title 49 Transportation; Shippers' General Requirements for Shipments and Packaging, including Sections:  
 173.304 Charging of Cylinders with Liquefied Compressed Gas  
 173.314 Requirements for Compressed Gases in Tank Cars  
 173.315 Compressed Gases in Cargo Tanks and Portable Tank Containers

## 2.3 Other Document:<sup>5</sup>

Chlorine Institute Pamphlet No. 152 Safe Handling of Chlorine Containing Nitrogen Trichloride

### 3. Summary of Test Method

3.1 Weighed samples of chlorine are delivered into a cooled graduated centrifuge tube. One mL of cooled HPLC eluent is added before the chlorine is allowed to evaporate. After the chlorine has evaporated the remaining eluent is analyzed directly on the HPLC for nitrogen trichloride concentration.

3.2 In-process solvent samples from chlorine production facility may be analyzed by dilution in methanol followed by direct HPLC analysis of the diluted solution.

### 4. Significance and Use

4.1 This test method was developed for the determination of nitrogen trichloride in samples of carbon tetrachloride liquid taken from the compressor suction chiller bottoms of chlorine production plants and other places in the plants that may collect and concentrate nitrogen trichloride to levels that could explode. The test method was then modified to measure the lower levels of nitrogen trichloride observed in product liquid chlorine. This test method is sensitive enough to measure the levels of nitrogen trichloride observed in the normal production of liquid chlorine.

4.2 This test method for nitrogen trichloride will require the dilution (100:1) of highly concentrated in-process samples to bring them within the working range of the analysis.

### 5. Interferences

5.1 This test method is selective for nitrogen trichloride. At this time there is no known interference in the materials used in this test method.

<https://standards.iteh.ai/catalog/standards/astm/ebc4461f-6776-4ba8-af24-313eae6f8ce0/astm-e2036-24>

5.2 Contact with any metal surfaces should be avoided due to the corrosive nature of the sample.

5.3 Nitrogen trichloride is decomposed by UV light, heat or strong acid. The presence of metal ions will increase the rate of decomposition. Special care should be taken to avoid exposure of the samples to direct sunlight. Samples and standards should be analyzed immediately upon preparation. Samples not immediately analyzed must be stored in dry ice until just before analysis.

### 6. Apparatus

6.1 A high performance liquid chromatograph (HPLC) composed of the following:

6.1.1 *HPLC Pump*, capable of 1 mL/min flow,

6.1.2 *HPLC UV Detector*, capable of operating at 221 nm with a 1-cm light path,

6.1.3 *HPLC Injection Valve*, 20 µL sample loop, all nonmetal, and

6.1.4 *HPLC Column*, C18 reverse phase, 25 cm by 4.6 mm.

6.2 *Plastic Syringes*. 1, 2, 5, 10, 20, 1 mL, 2 mL, 5 mL, 10 mL, 20 mL, and 60 mL.

<sup>4</sup> Available from U.S. Government Publishing Office, 732 N. Capitol St., NW, Washington, DC 20401-0001, <http://www.gpo.gov>.

<sup>5</sup> Available from The Chlorine Institute (CI), 1300 Wilson Blvd., Ste. 525, Arlington, VA 22209, <http://www.chlorineinstitute.org>.

6.3 *Nonmetallic Syringe Needles.*

6.4 *Top Loader Balance*, capable of 0.01 g resolution and 1 kg capacity.

6.5 *TFE-fluorocarbon Tubing*, 1.59 mm outside diameter.

6.6 *Stainless Steel Sample Cylinder*, with a needle valve on one end.

6.7 *Graduated Centrifuge Tube*, 15 mL.

6.8 *Reducing Tube Fitting*, 6.35 mm to 1.59 mm tubing.

## 7. Reagents

7.1 *Purity of Reagents*—Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.<sup>6</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

7.2 *Water*, HPLC grade.

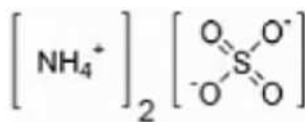
7.3 *Methanol*, HPLC grade. CAS# 67-56-1.

7.4 *Sodium Acetate*, reagent grade. CAS# 127-09-3.

7.5 *Glacial Acetic Acid*, reagent grade. CAS# 64-19-7.

7.6 *Dry Ice*. CAS# 124-38-9.

7.7 *Ammonium Sulfate*, reagent grade. CAS# 7783-20-2.



**FIG. 1 Ammonium Sulfate, reagent grade: CAS# 7783-20-2**

7.8 *Chlorine*, reagent grade. CAS# 7782-50-5.

7.9 *Helium*.

## 8. Hazards

8.1 Consult current OSHA regulations, suppliers' Safety Data Sheets, and local regulations for all materials used in this test method.

<sup>6</sup> *Reagent Chemicals, American Chemical Society Specifications, ACS Reagent Chemicals, Specifications and Procedures for Reagents and Standard-Grade Reference Materials*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

8.2 *Safety and Health Precautions*—Exposure to all solvents used in this test method should be avoided.

8.3 Concentrations of nitrogen chloride over 2 % pose the risk of explosion and should be considered dangerous. Never allow any solutions containing nitrogen chloride to evaporate, concentrating the nitrogen chloride in the remaining residue. The concentrations of nitrogen trichloride synthesized in this test method are small and should not pose any significant risk of explosion. All of the solutions prepared in this test method can be disposed of by flushing them down any laboratory sink using a reasonable amount of water.

8.4 Chlorine is a corrosive and toxic material. A well-ventilated fume hood should be used to house all sample handling and to vent the test equipment when this product is analyzed in the laboratory.

8.5 This analysis should only be performed by persons who are thoroughly familiar with the handling of chlorine, and even an experienced person should not work alone. The operator must be provided with adequate eye protection and a respirator. Splashes of liquid chlorine destroy clothing and, if such clothing is next to the skin, will produce irritations and burns.

8.6 If liquid samples are to be taken in cylinders, do not allow the sample cylinder to become liquid full. Test Method E806, 49 CFR 173.314, 173.315, and 173.304 advise that the weight of the chlorine in the cylinder should not be more than 125 % of the weight of the water that the cylinder could contain.

8.7 When sampling and working with chlorine out-of-doors, people downwind from such an operation should be warned of the possible release of chlorine vapors.

8.8 In the event chlorine is inhaled, first aid should be summoned immediately and oxygen administered without delay.

8.9 Store pressurized samples where involuntary release would not cause excessive risk to people or property.

8.10 It is recommended that means be available for disposal of excess chlorine in an environmentally safe and acceptable manner. If chlorine cannot be disposed of in a chlorine consuming process, a chlorine absorption system should be provided. When the analysis and sampling regimen requires an initial purging of chlorine from a container, the purged chlorine should be similarly handled. Purging to the atmosphere should be avoided.

## 9. Typical Instrument Parameters

9.1 Adjust the chromatograph in accordance with the following parameters and allow the instrument to equilibrate until a steady baseline is obtained:

9.1.1 *Column*—C18 reverse phase ODS (C18) 25 cm by 4.6 mm, 10  $\mu$ m,

9.1.2 *Eluent*—60 % by volume methanol, 40 % by volume 0.1 mol/L (M) acetate buffer, pH 4.5, helium sparged,

9.1.3 *Flow Rate*—1 mL/min,

9.1.4 *Injection Volume*—~~20  $\mu$ l~~ 20  $\mu$ l sample loop, and

9.1.5 *Detector Wavelength*—UV at 221 nm.

## 10. Preparation of Buffer Solution

10.1 *Sodium Acetate Buffer Stock Solution (1 mol/L (M))*—Dissolve 136 g sodium acetate ( $\text{NaOOCCH}_3 \times 3\text{H}_2\text{O}$ ) and 60 g glacial acetic acid ( $\text{HOOCCH}_3$ ) in water and dilute to 1 L.

10.2 *Sodium Acetate Buffer (0.1 mol/L (M))*—Transfer 100 mL of the stock buffer solution into a ~~1~~ 1 L volumetric flask and dilute to volume with water.

## 11. Preparation of Eluent

11.1 Add 600 mL methanol to 400 mL 0.1 M sodium acetate buffer solution and mix well. Before use purge the solution with helium for 20 min to remove dissolved oxygen.

## 12. Preparation of the Sample Cylinder

12.1 Clamp the sample cylinder in a vertical position with the needle valve in the downward position. Insert the 6.35 mm end of the reducing tube fitting into the needle valve and set the 6.35 mm nut and ferrule of the fitting. Insert the 1.59 mm fluoropolymer tubing in the reducing tube fitting and tighten the 1.59 mm nut and ferrule. See Fig. 42. It may be helpful to cut a 6.35 mm circle of fluoropolymer frit material and place it into the reducing fitting prior to assembly to prevent plugging of the 1.59 mm tubing. Cut the length of the 1.59 mm tubing so that only 38.10 mm protrudes out of the fitting.

## 13. Standardization of the HPLC

13.1 Prepare a ~~1000 µg/g~~ 1000 µg/g (ppm) stock solution of ammonium sulfate in water. Make a series of standards of ammonium sulfate by serial dilution covering the range between ~~1 µg/g~~ 1 µg/g and 650 µg/g (ppm) nitrogen trichloride. Calculate the nitrogen trichloride concentration of each standard from the ammonium sulfate concentration by multiplying by the ratio of the molecular weights, as follows:

$$[2\text{NCl}_3/(\text{NH}_4)_2\text{SO}_4] = (2 \times 120.5/132) = 1.826 \quad (1)$$

13.2 Withdraw 0.5 mL of an ammonium sulfate standard into a ~~1-mL~~ 1 mL plastic syringe and then pull the plunger back to the ~~1-mL~~ 1 mL mark to fill the remainder of the syringe with air.

13.3 Clamp a sample cylinder containing reagent chlorine vertically in a ring stand and attach a 152.4 mm piece of 1.59 mm outside diameter fluoropolymer tubing to the valve. Take a second ~~1-mL~~ 1 mL plastic syringe and pull the plunger back to the ~~1-mL~~ 1 mL mark. Insert the 1.59 mm fluoropolymer tubing into the second syringe and gently purge chlorine through the syringe filling this second syringe with reagent chlorine gas. Remove the fluoropolymer tubing from the syringe and attach a nonmetal needle to the luer tip.

13.4 Holding the first syringe pointed upward, use the second syringe equipped with the nonmetal needle to bubble the reagent chlorine gas through the ammonium sulfate standard solution, converting the ammonium ions to nitrogen trichloride. Use of a ~~1-mL~~ 1 mL syringe should avoid overchlorination. See Fig. 23.

13.5 Use the entire 0.5 mL standard to flush the sample valve and fill the sample loop. Immediately inject this standard into the HPLC for analysis. The nitrogen trichloride peak will elute between ~~11 min~~ 11 min and 12 min in the chromatogram. See Fig. 34.

13.6 Measure and record the peak height of the nitrogen trichloride peak in the analysis of each of the standard solutions. Plot the peak heights of each standard versus the concentrations of the nitrogen trichloride in each standard analyzed. The slope and intercept of this line will be used in the calculation of the sample analysis values. See Figs. 45 and 56.

NOTE 1—Standards over 200 µg/g (ppm) require 2 mL of chlorine gas for complete conversion of the ammonium ions to nitrogen trichloride. If the calibration curve is extended beyond 650 µg/g (ppm), more than 2 mL of reagent chlorine gas may be needed to obtain complete conversion of the ammonium ions in the standards. Although the amount of chlorine used for the chlorination of the standards is extremely small when compared to that of the sample, it is always a good idea to prepare the first standard in any calibration curve without the addition of the ammonium sulfate to serve as a reagent blank.

NOTE 2—Two calibrations, one completed in December 1995 and one completed in February 1996, were not significantly different. This would indicate

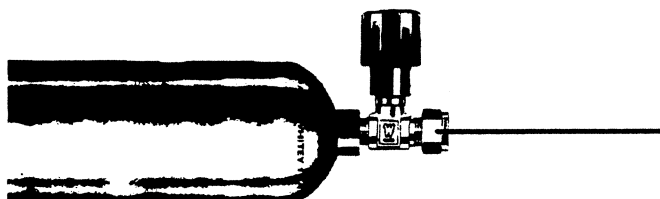
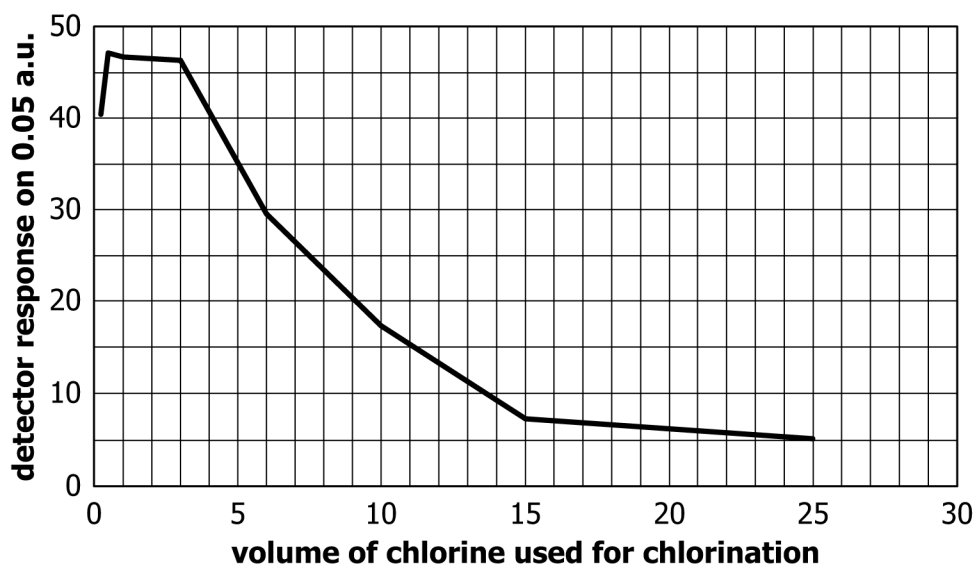


FIG. 42 Sample Cylinder



26 ppm ammonia standard

FIG. 23 Standard Preparation By the Chlorination of Ammonia

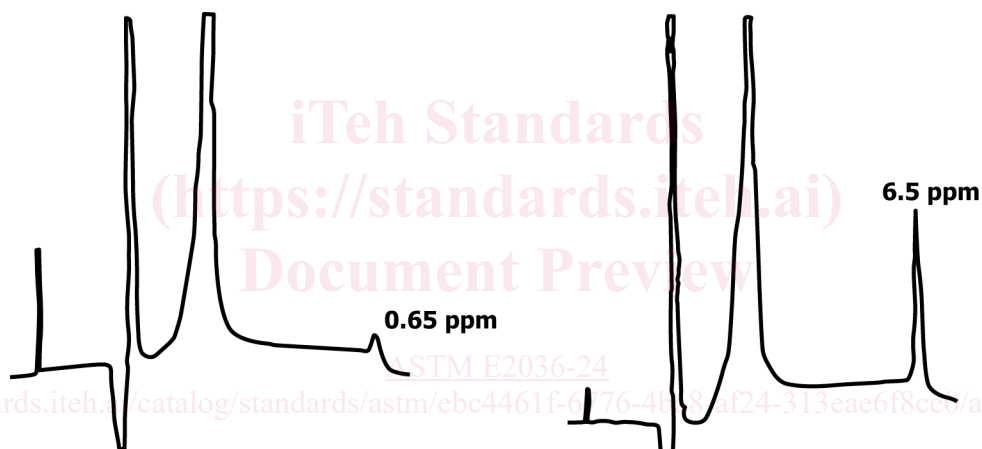


FIG. 34 Typical Chromatograms

considerable stability of the UV detector. Considering the ease with which this test method can be calibrated, a one-point check of the calibration curve is recommended on a weekly basis to ensure that the system is working properly.

#### 14. Procedure for Liquid Chlorine

14.1 Tare a metal sample cylinder equipped with a needle valve and cool the cylinder in dry ice making sure that the needle valve is closed. Sample the liquid chlorine source using metal tubing pipe connections. Do not use fluoropolymer or other type of plastic tubing that could transmit UV light. Nitrogen trichloride is very rapidly decomposed in the presence of UV light. Consult The Chlorine Institute Pamphlet No. 152 for additional information.

14.2 Re-weigh the sample cylinder to determine the weight of chlorine collected in the cylinder before the cylinder is allowed to warm. If the cylinder is more than 60 % full, drain some of the liquid chlorine out of the bottom of the cylinder and reweigh until the weight of the chlorine is within acceptable limits for that cylinder.

NOTE 3—For example, a 150-mL cylinder should not contain more than 90 mL of chlorine at a density of 1.467 g/mL (that is, 132 g of chlorine).

14.3 Attach a 38.10 mm piece of 1.59 mm TFE-fluorocarbon tubing and reducing fitting to the needle valve on the chlorine cylinder, weigh, and record the sample cylinder weight to the nearest 0.1 g. Clamp the cylinder in a ring stand with the needle valve and fluoropolymer tubing pointed down.