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Surface active agents — Sulfated ethoxylated alcohols and alkylphenols — Estimation of the mean relative molecular mass

*Agents de surface — Sulfates d'alcools et d'alkylphénols éthoxylés — Évaluation de la masse
moléculaire relative moyenne*

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Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

Draft International Standards adopted by the technical committees are circulated to the member bodies for approval before their acceptance as International Standards by the ISO Council. They are approved in accordance with ISO procedures requiring at least 75 % approval by the member bodies voting.

International Standard ISO 6843 was prepared by Technical Committee ISO/TC 91, *Surface active agents*.

Annex A forms an integral part of this International Standard.

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Surface active agents — Sulfated ethoxylated alcohols and alkylphenols — Estimation of the mean relative molecular mass

1 Scope

This International Standard specifies a method for the estimation of the mean relative molecular mass of the anionic-active matter present in ordinary commercial neutralized products of sulfation of ethoxylated alcohols or alkylphenols (alkyl oxyethylene sulfates (ethoxylated alcohol sulfates) or alkylphenol oxyethylene sulfates (ethoxylated alkylphenol sulfates)) containing an average of not more than 20 oxyethylene groups per molecule.

It also sets out, in annex A, a general scheme of analysis.

2 Normative references

The following standards contain provisions which, through reference in this text, constitute provisions of this International Standard. At the time of publication, the editions indicated were valid. All standards are subject to revision, and parties to agreements based on this International Standard are encouraged to investigate the possibility of applying the most recent editions of the standards listed below. Members of IEC and ISO maintain registers of currently valid International Standards.

ISO 607 : 1980, *Surface active agents and detergents — Methods of sample division*.

ISO 1042 : 1983, *Laboratory glassware — One-mark volumetric flasks*.

ISO 2271 : 1972, *Surface active agents — Detergents — Determination of anionic-active matter (Direct two-phase titration procedure)*.

ISO 4800 : 1977, *Laboratory glassware — Separating funnels and dropping funnels*.

ISO 6842 : 1983, *Surface active agents — Polyethoxylated alcohol and alkylphenol sulfates — Determination of total active matter*.

ISO 8799 : 1988, *Sulfated ethoxylated alcohols and alkylphenols — Determination of content of unsulfated matter*.

3 Principle

From a solution of the test portion saturated with sodium chloride, extraction of alkylether sulfate with an ethyl acetate/butan-1-ol mixture, then evaporation of the aqueous phase containing the polyglycol, the polyglycol sulfate and possibly traces of ether sulfates; then removal of salts from the residue by treatment with methanol and filtration.

Evaporation of an aliquot portion of the filtrate and weighing of the residue, then redissolution in water and determination of the sodium chloride and anionic-active matter contents.

Determination of the polyglycol content by passing the remaining fraction of filtrate through an ion-exchange resin.

Determination of the polyglycol sulfate content by difference between the above determinations.

From the content of anionic surface active agent (alkylether sulfate) and the anionic surface active matter content determined by two-phase titration, estimation of the mean relative molecular mass.

NOTE — The content of anionic surface active agent (alkylether sulfate) is obtained by the difference between the total active matter content and the unsulfated matter and polyglycol sulfate contents.

4 Reagents and materials

During the analysis, use only reagents of recognized analytical reagent grade and only distilled water or water of equivalent purity.

4.1 Methanol.

4.2 Sodium chloride.

4.3 Ethyl acetate/butan-1-ol, mixture (9 + 1) by volume.

4.4 Sodium chloride, 59 g/l solution.

4.5 Hydrochloric acid, 73 g/l solution.

4.6 Sodium hydroxide, 80 g/l solution.

4.7 Silver nitrate, standard volumetric solution, $c(\text{AgNO}_3) = 0,100 \text{ mol/l}$.

4.8 Cation-exchange resin, polystyrene sulfonic acid type, 2 % crosslinked, 150 to 330 μm , hydrogen form.

4.9 Anion-exchange resin, polystyrene quaternary ammonium type, 2 % crosslinked, 150 to 330 μm , chloride form.

5 Apparatus

Ordinary laboratory apparatus and

5.1 Separating funnel, of capacity 250 ml, in accordance with ISO 4800.

5.2 Separating funnel, jacketed, of capacity 250 ml.

5.3 One-mark volumetric flask, of capacity 100 ml, in accordance with ISO 1042.

5.4 Water bath, capable of being controlled at 50 °C.

5.5 Reflux condenser.

5.6 Rotary evaporator, with round-bottom flasks of capacity 250 ml.

5.7 Ion-exchange column: Glass tube of internal diameter 25 mm and length 200 mm, restricted at the bottom and provided with a glass stopcock.

The ion-exchange resin is supported at the bottom by a 10 to 20 mm layer of glass wool or by a sintered glass filter.

6 Sampling

The laboratory sample of surface active agent shall be prepared and stored in accordance with the instructions given in ISO 607.

7 Procedure

7.1 Test portion

From the laboratory sample, if necessary homogenized by introducing a definite and appropriate quantity of water, weigh, to the nearest 0,1 g, into a 100 ml beaker, a quantity corresponding to about 25 mmol of anionic-active matter.

7.2 Preparation of the ion-exchange resins

If preferred, the resins may be prepared in smaller amounts than 1 kg with proportionately smaller volumes of reagents.

7.2.1 Anion-exchange resin

Take 1 kg of the anion-exchange resin (4.9) and allow to distend in water for 48 h. Transfer the resin to a suitable column and pass 5 litres of the sodium hydroxide solution (4.6) through the column, followed by 2 to 3 litres of water. Then pass 4 litres of the hydrochloric acid solution (4.5) through the column and again wash with 2 to 3 litres of water. The treated resin may be stored in water.

7.2.2 Cation-exchange resin

Take 1 kg of the cation-exchange resin (4.8) and allow to distend in water for 48 h. Transfer the resin to a suitable column and pass 5 litres of the hydrochloric acid solution (4.5) through the column, then wash with water until the washings are neutral. The treated resin may be stored in water.

7.3 Final preparation of resins

Take the required amount of anion-exchange resin, prepared as specified in 7.2.1, namely 25 ml per determination, and transfer to a suitable column. Pass through the column five times its volume of the sodium hydroxide solution (4.6), wash with water until neutral, then wash with 1 to 2 volumes of the methanol (4.1).

Take the required amount (25 ml) of cation-exchange resin, prepared as specified in 7.2.2, and, after introduction into a suitable column, wash with twice its volume of methanol (4.1).

7.4 Arrangement of the mixed-bed exchange column

Mix 25 ml of the cation-exchange resin and 25 ml of the anion-exchange resin, prepared as specified in 7.3, in a beaker. Fill the column (5.7) with the mixed resin in small portions and compress the mixed resin with a glass rod to a volume between 50 and 60 ml, and wash with 500 ml of the methanol (4.1).

7.5 Separation of polyglycol sulfate

7.5.1 Dissolve the test portion (7.1) in 50 ml of the sodium chloride solution (4.4) and 50 ml of the ethyl acetate/butan-1-ol mixture (4.3).

Transfer the solution to a separating funnel (5.1) and rinse the beaker with a few millilitres of the sodium chloride solution (4.4) and of the ethyl acetate/butan-1-ol mixture, then add the rinsing solution to the separating funnel.

Shake the separating funnel vigorously and allow the phases to separate.

Draw off the aqueous phase into the jacketed separating funnel (5.2), thermoregulated at 60 °C, add 5 g of the sodium chloride (4.2) and shake to complete dissolution. Add 50 ml of the ethyl acetate/butan-1-ol mixture and shake.

After separation of the phases, draw off the aqueous phase into a 250 ml round-bottom flask (5.6) and evaporate to dryness by means of the rotary evaporator (5.6) on a water bath (5.4), controlled at 50 °C, under vacuum, using a water-jet air pump.