



Standard Test Methods for Instrumental Determination of Carbon, Hydrogen, and Nitrogen in Petroleum Products and Lubricants¹

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1. Scope

1.1 These test methods cover the instrumental determination of carbon, hydrogen, and nitrogen in laboratory samples of petroleum products and lubricants. Values obtained represent the total carbon, the total hydrogen, and the total nitrogen.

1.2 These test methods are applicable to samples such as crude oils, fuel oils, additives, and residues for carbon and hydrogen and nitrogen analysis. These test methods were tested in the concentration range of at least 75 to 87 mass % for carbon, at least 9 to 16 mass % for hydrogen, and < 0.1 to 2 mass % for nitrogen.

1.3 The nitrogen test method is not applicable to light materials or those containing < 0.75 mass % nitrogen, or both, such as gasoline, jet fuel, naphtha, diesel fuel, or chemical solvents.

1.4 These test methods are not recommended for the analysis of volatile materials such as gasoline, gasoline-oxygenate blends, or gasoline type aviation turbine fuels.

1.5 The results of these tests can be expressed as mass % carbon, hydrogen or nitrogen.

1.6 The values stated in SI units are to be regarded as the standard.

1.7 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 ASTM Standards:

D 4057 Practice for Manual Sampling of Petroleum and Petroleum Products²

D 4177 Practice for Automatic Sampling of Petroleum and Petroleum Products²

3. Summary of Test Methods

3.1 In these test methods, carbon, hydrogen, and nitrogen are determined concurrently in a single instrumental procedure.

With some systems, the procedure consists of simply weighing a portion of the sample, placing the portion in the instrument, and initiating the (subsequently automatic) analytical process. In other systems, the analytical process, to some degree, is manually controlled.

3.2 The actual process can vary substantially from instrument to instrument, since a variety of means can be utilized to effect the primary requirements of the test methods. All satisfactory processes provide for the following:

3.2.1 The conversion of the subject materials (in their entirety) to carbon dioxide, water vapor, and elemental nitrogen, respectively, and

3.2.2 The subsequent, quantitative determination of these gases in an appropriate gas stream.

3.3 The conversion of the subject materials to their corresponding gases takes place largely during combustion of the sample at an elevated temperature in an atmosphere of purified oxygen. Here, a variety of gaseous materials are produced, including the following:

3.3.1 Carbon dioxide from the oxidation of organic and elemental carbon,

3.3.2 Hydrogen halides from organic halides (and organic hydrogen, as required),

3.3.3 Water vapor from the oxidation of (the remaining) organic hydrogen and the liberation of moisture,

3.3.4 Nitrogen and nitrogen oxides from the oxidation of organic nitrogen, and

3.3.5 Sulfur oxides from the oxidation of organic sulfur. In some systems, sulfurous and sulfuric acids can also be obtained from a combination of the sulfur oxides and the water vapor.

3.4 There are several accepted ways of isolating the desired gaseous products and quantitatively determining them. These are as follows:

3.4.1 *Test Method A*³—From the combustion product gas stream, oxides of sulfur are removed with calcium oxide in the secondary combustion zone. A portion of the remaining mixed gases is carried by helium gas over a hot copper train to remove oxygen, and reduce NO_x to N₂, over NaOH to remove CO₂, and over magnesium perchlorate to remove H₂O. The remaining elemental nitrogen is measured by the thermal conductivity

¹ These test methods are under the jurisdiction of ASTM Committee D-2 on Petroleum Products and Lubricants and are the direct responsibility of Subcommittee D02.03.0B on Spectrometric Methods.

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² *Annual Book of ASTM Standards*, Vol 05.02.

³ Leco CHN-600 instrument has been found satisfactory for the analyses in Test Method A, and is available from Leco Corporation, 3000 Lakeview Ave., St. Joseph, MI 49085.

cell. Simultaneously, but separately from the nitrogen measurement, the carbon and hydrogen selective infrared cells measure the CO₂ and H₂O levels.

3.4.2 *Test Method B*⁴—From the combustion product gas stream (which is cleaned from sulfur oxides, excess oxygen, etc. as in 3.4.1), the remaining CO₂, water vapor, and N₂ are flushed into a mixing chamber and are thoroughly homogenized at a precise volume, temperature, and pressure. After homogenization, the chamber is depressurized to allow the gases to pass through a heated column, where the gases separate as a function of selective retention times. The separation occurs in a stepwise steady-state manner for nitrogen, carbon dioxide, and water.

3.4.3 *Test Method C*⁵—The combustion product gas stream, after full oxidation of component gases, is passed over heated copper to remove excess oxygen and reduce NO_x to N₂ gas. The gases are then passed through a heated chromatographic column to separate and elute N₂, CO₂, and H₂O in that order. The individual eluted gases are measured by a thermal conductivity detector.

3.5 In all cases, the concentrations of carbon, hydrogen and nitrogen are calculated as functions of the following:

3.5.1 The measured instrumental responses,

3.5.2 The values for response per unit mass for the elements (established via instrument calibration), and

3.5.3 The mass of the sample.

3.6 A capability for performing these computations automatically can be included in the instrumentation utilized for these test methods.

4. Significance and Use

4.1 This is the first ASTM standard covering the simultaneous determination of carbon, hydrogen, and nitrogen in petroleum products and lubricants.

4.2 Carbon, hydrogen, and particularly nitrogen analyses are useful in determining the complex nature of sample types covered by this test method. The CHN results can be used to estimate the processing and refining potentials and yields in the petrochemical industry.

4.3 The concentration of nitrogen is a measure of the presence of nitrogen containing additives. Knowledge of its concentration can be used to predict performance. Some petroleum products also contain naturally occurring nitrogen. Knowledge of hydrogen content in samples is helpful in addressing their performance characteristics. Hydrogen to carbon ratio is useful to assess the performance of upgrading processes.

5. Apparatus

5.1 Since a variety of instrumental components and configurations can be satisfactorily utilized for these test methods, no specifications are given here with respect to overall system design.

⁴ Perkin Elmer 240C, 2400 series and CEC 240XA and 440 instruments have been found satisfactory for the analyses in Test Method B and are available from Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

⁵ Carlo Erba 1106, 1108, and 1500 instruments have been found satisfactory for the analyses in Test Method C and are available from Carlo Erba Strumentazione, Strada Rivoltana, 20090 Rodano, Milan, Italy.

5.2 Functionally, however, the following are specified for all instruments:

5.2.1 The conditions for combustion of the sample must be such that (for the full range of applicable samples) the subject components are completely converted to carbon dioxide, water vapor (except for hydrogen associated with volatile halides and sulfur oxides), and nitrogen or nitrogen oxides. Generally, instrumental conditions that affect complete combustion include availability of the oxidant, temperature, and time.

5.2.2 Representative aliquots of the combustion gases must then be treated:

5.2.2.1 To liberate (as water vapor) hydrogen present as hydrogen halides and sulfur oxyacids, and

5.2.2.2 To reduce (to the element) nitrogen present as nitrogen oxides.

5.2.3 The water vapor and nitrogen so obtained must be included with the materials originally present in these aliquots.

5.2.4 Additional treatment of the aliquots (prior to detection) depends on the detection scheme utilized for the instrument (see Note 1).

NOTE 1—These additional treatments can be provided by the instrumental components utilized to satisfy 5.2.2.

5.2.5 The detection system (in its full scope) must determine the analytical gases individually and without interference. Additionally, for each analyte, either:

5.2.5.1 The detectors must provide linear responses with respect to concentration over the full range of possible concentrations from the applicable samples, or

5.2.5.2 The system must include provisions for appropriately evaluating non-linear responses so that they can be accurately correlated with these concentrations.

5.2.6 Such provisions can be integral to the instrumentation, or they can be provided by (auxiliary) computation schemes.

5.2.7 Lastly, except for those systems where the concentration data are output directly, the instrument must include an appropriate readout device for the detector responses.

5.3 Additionally consumables needed for the analyses include:

5.3.1 *Tin Capsules*, large and small,

5.3.2 *Ceramic Crucibles*,

5.3.3 *Copper Capsules*,

5.3.4 *Tin Plugs*,

5.3.5 *Tin Boats*,

5.3.6 *Copper Plugs*,

5.3.7 *Aluminum Capsules*,

5.3.8 *Combustion Tubes*,

5.3.9 *Adsorption Tubes*,

5.3.10 *Nickel Capsules*, and

5.3.11 *Reduction Tubes*.

5.4 *Analytical Balance*, capable of weighing to the nearest 0.00001 g.

5.5 *Syringes or Pipettes*, to transfer the test specimens to capsules.

6. Reagents

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that

all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁶ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Calibration Standards*—Table 1 lists the pure organic compounds most commonly used to calibrate the instruments operated according to 3.4.1-3.4.3; other suitable pure compounds can also be used.

6.3 *Carrier and Combustion Gases*:

6.3.1 *Oxygen*, high purity (99.998 %),

6.3.2 *Helium*, high purity (99.995 %),

6.3.3 *Compressed Air, Nitrogen, or Argon*, for operating pneumatic valves, if needed, and

6.3.4 *Carbon Dioxide*.

6.4 *Additional Reagents (as Specified by the Instrument Manufacturer)*—This specification covers the reagents utilized to provide for the functional requirements cited in 5.2.2 and 5.2.3. These reagents can vary substantially for different instruments. Consequently, these reagents shall be those recommended by the manufacturer. Specifically, these reagents will be for:

6.4.1 *Test Method A*³:

6.4.1.1 *Sodium Hydroxide Coated Silica*,

6.4.1.2 *Quartz Wool*,

6.4.1.3 *Magnesium Perchlorate*,

6.4.1.4 *Copper Turnings*,

6.4.1.5 *Coated Calcium Oxide (Furnace Reagent)*,

6.4.1.6 *Nitrogen Catalyst*, and

6.4.1.7 *Magnesium Oxide*⁷, for liquids.

6.4.2 *Test Method B*⁴:

6.4.2.1 *EA 1000 Reagent*⁸,

6.4.2.2 *Silver Tungstate on MgO*,

6.4.2.3 *Silver Vanadate*,

6.4.2.4 *Quartz Wool*,

6.4.2.5 *Silver Gauze*,

6.4.2.6 *Copper Oxide*⁹,

6.4.2.7 *Tungstic Oxide*,

6.4.2.8 *Cobalt Oxide*,

6.4.2.9 *Copper Powder*,

6.4.2.10 *Sodium Hydroxide Coated Silica*,

6.4.2.11 *Alumina*,

6.4.2.12 *Magnesium Perchlorate*, and

6.4.2.13 *Platinum Gauze*.

6.4.3 *Test Method C*⁵:

6.4.3.1 *Quartz Wool*,

6.4.3.2 *Chromic Oxide* (oxidation catalyst),

6.4.3.3 *Silver Coated Cobalt Oxide*,

6.4.3.4 *Reduced Copper* (reduction catalyst),

6.4.3.5 *Magnesium Perchlorate*,

6.4.3.6 *Molecular Sieve*, 3A 1/16 in. (1.6 mm),

6.4.3.7 *Sodium Hydroxide Coated Silica*,

6.4.3.8 *Chromosorb*, (Absorber,¹⁰ for liquid samples; calcined silica), and

6.4.3.9 *Copper Grains*.

7. Sampling, Test Specimens, and Test Units

7.1 *Laboratory Sample*—Take a representative sample as specified in Practices D 4057 or D 4177.

7.2 *Test Specimen*—Take an aliquot from the laboratory sample for analysis as follows:

7.2.1 *Preparation*—Warm viscous samples until they are fluid, and shake for 5 s.

7.2.2 *Transfer*—Use any convenient, clean syringe or pipet to transfer test specimens to the capsules as described in Section 9.

8. Preparation of Apparatus

8.1 Prepare the instrumental system (in its entirety) in strict accordance to the manufacturer's instructions.

8.2 Calibrate the system using acetanilide or other suitable calibration standard mentioned in Table 1, using the standardization procedure specified by the manufacturer. This procedure, with respect to the actual analytical process involved, must not differ from that specified for the samples.

8.3 *Tubes/Columns Preparation* (see Note 2)—Clean all quartz and glass parts prior to use with soap and water followed

⁷ Com-aid, a registered trademark of Leco, has been found suitable for this purpose and is available from Leco Corporation, 3000 Lakeview Ave., St Joseph, MI 49085.

⁸ EA 1000 Reagent is a registered trademark of Perkin Elmer and is available from Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

⁹ Cuprox, a registered trademark of Perkin Elmer, has been found suitable for this purpose and is available from Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

¹⁰ Chromosorb, a registered trademark of Carlo Erba Strumentazione, has been found suitable for this purpose and is available from Carlo Erba Strumentazione, Strada Rivoltana, 20090 Rodano, Milan, Italy.

⁶ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

TABLE 1 Calibration Standards for CHN Instrumental Analysis^{AB}

Compound	Molecular Formula	Carbon, Mass%	Hydrogen, Mass %	Nitrogen, Mass %
Acetanilide	C ₈ H ₉ NO	71.09	6.71	10.36
Atropine	C ₁₇ H ₂₃ NO ₃	70.56	8.01	4.84
Benzoic acid	C ₇ H ₆ O ₂	68.84	4.95	...
Cyclohexanone-2,4-dinitrophenylhydrazone	C ₁₂ H ₁₄ N ₄ O ₄	51.79	5.07	20.14
Cystine	C ₆ H ₁₂ N ₂ O ₄ S ₂	29.99	5.03	11.66
Diphenyl	C ₁₂ H ₁₀	93.46	6.54	...
EDTA	C ₁₀ H ₁₆ N ₂ O ₈	41.10	5.52	9.59
Imidazol	C ₃ H ₄ N ₂	52.92	5.92	41.15
Nicotinic acid	C ₆ H ₅ NO ₂	58.53	4.09	11.38
Stearic acid	C ₁₈ H ₃₆ O ₂	75.99	12.76	...
Succinamide	C ₄ H ₈ N ₂ O ₂	41.37	6.94	24.13
Sucrose	C ₁₂ H ₂₂ O ₁₁	42.10	6.48	...
Sulphanilamide	C ₆ H ₈ N ₂ O ₂ S	41.84	4.68	16.27
Triethanol amine	C ₆ H ₁₅ NO ₃	48.30	10.13	9.39

^AThe Merck Index, 10th Edition, Merck and Company, Inc., Rahway, New Jersey, 1983.

^BMany of these compounds can be obtained from the National Institute of Standards and Technology as well as other commercial chemical manufacturers, such as Aldrich, Alfa, Kodak and others. See 6.1 for the purity of these reagents.