
**Iron ores — Determination of sodium
content — Flame atomic absorption
spectrometric method**

*Minerais de fer — Dosage du sodium — Méthode par spectrométrie
d'absorption atomique dans la flamme*

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ISO 13313:1997

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Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

Draft International Standards adopted by the technical committees are circulated to the member bodies for voting. Publication as an International Standard requires approval by at least 75 % of the member bodies casting a vote.

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International Standard ISO 13313 was prepared by Technical Committee ISO/TC 102, *Iron ores*, Subcommittee SC 2, *Chemical analysis*.
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This first edition, together with ISO 13312, cancels and replaces the first edition of ISO 6831 (ISO 6831:1986), of which it constitutes a partial revision.

Annex A forms an integral part of this International Standard. Annexes B and C are for information only.

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Iron ores — Determination of sodium content — Flame atomic absorption spectrometric method

WARNING — This International Standard may involve hazardous materials, operations and equipment. This International Standard does not purport to address all of the safety problems associated with its use. It is the responsibility of the user of this International Standard to establish appropriate health and safety practices and determine the applicability of regulatory limitations prior to use.

1 Scope

This International Standard specifies a flame atomic absorption spectrometric method for the determination of the sodium content of iron ores.

This method is applicable to sodium contents between 0,002 % (*m/m*) and 1,0 % (*m/m*)¹⁾ in natural iron ores, iron ore concentrates and agglomerates, including sinter products.

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2 Normative references

The following standards contain provisions which, through reference in this text, constitute provisions of this International Standard. At the time of publication, the editions indicated were valid. All standards are subject to revision, and parties to agreements based on this International Standard are encouraged to investigate the possibility of applying the most recent editions of standards indicated below. Members of IEC and ISO maintain registers of currently valid International Standards.

ISO 648:1977, *Laboratory glassware — One-mark pipettes.*

ISO 1042:1983, *Laboratory glassware — One-mark volumetric flasks.*

ISO 3081:1986, *Iron ores — Increment sampling — Manual method.*

ISO 3082:1987, *Iron ores — Increment sampling and sample preparation — Mechanical method.*

ISO 3083:1986, *Iron ores — Preparation of samples — Manual method.*

ISO 3696:1987, *Water for analytical laboratory use — Specification and test methods.*

1) The method has not been tested on ores containing more than 0,50 % (*m/m*) of sodium (see annex B). However, it is possible to extend its application to 1,0 % (*m/m*) using the calibration Standards prepared.

ISO 7764:1985, *Iron ores — Preparation of predried test samples for chemical analysis.*

ISO 11323:1996, *Iron ores — Vocabulary.*

3 Definitions

For the purposes of this International Standard, the definitions given in ISO 11323 apply.

4 Principle

Decomposition of the test portion by treatment with hydrochloric acid and hydrofluoric acid. Evaporation to dryness. Wetting and repetition of the evaporation with a new portion of hydrochloric acid. Dissolution with hydrochloric acid and appropriate dilution. Aspiration into the air-acetylene flame of the atomic absorption apparatus.

Measurement of the absorbance value obtained for sodium in comparison with those obtained from calibration solutions.

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5 Reagents

During the analysis, use only reagents of recognized analytical grade, and only water that complies with grade 2 of ISO 3696.

NOTE 1 Reagents are to be selected or purified for the lowest possible blank value.

5.1 Hydrochloric acid, ρ 1,16 g/ml to 1,19 g/ml.

5.2 Hydrofluoric acid, ρ 1,13 g/ml, 40 % (m/m), or ρ 1,19 g/ml, 48 % (m/m).

5.3 Hydrochloric acid, ρ 1,16 g/ml to 1,19 g/ml, diluted 1 + 2.

5.4 Background solution

Dissolve 43 g of high purity iron oxide powder²⁾ in 500 ml of hydrochloric acid (5.1). Allow to cool and dilute with water to 1 000 ml.

2) Instead of iron oxide, the use of metallic iron with a suitable oxidant is permitted. (The alkali content of the oxidant shall be low.)

5.5 Sodium, standard solution, 10 µg Na/ml.

Pulverize about 4 g of high purity sodium chloride in an agate mortar, dry in an oven at 105 °C to 110 °C for two hours and allow to cool to room temperature in a desiccator. Dissolve 2,542 g in water, dilute with water to 1 000 ml in a volumetric flask and mix.

Transfer³⁾ 10,0 ml of this solution to a 1 000 ml volumetric flask, dilute with water to volume and mix.

Store this standard solution in a plastic bottle.

1 ml of this standard solution contains 10 µg of sodium.

6 Apparatus

Ordinary laboratory equipment, including one-mark pipettes and one-mark volumetric flasks complying with the specifications of ISO 648 and ISO 1042 respectively, and

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6.1 Polytetrafluoroethylene (PTFE) beakers, of capacity 100 ml, provided with PTFE covers.

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6.2 PTFE-coated magnetic stirring bars.

6.3 PTFE digestion bomb.

6.4 Plastic pipettes.

6.5 Plastic volumetric flasks and storage bottles.

6.6 Magnetic stirring hotplates.

3) Glass equipment may be used.

NOTES

- 2 Platinum vessels may be used instead of PTFE beakers.
- 3 Except where stated, glass equipment should be avoided, as it could contaminate the solutions.
- 4 The caps of the plastic bottles should be of a type that does not contain a separate wad insert. Such inserts usually contain a sodium compound that will contaminate the solution.
- 5 To obtain reliable values, the equipment should be cleaned and checked as follows:
 - a) Rinse all volumetric ware, including the pipettes used for preparing the calibration solutions, with hydrochloric acid solution (5.3) before use. Check calibration regularly or as needed.
 - b) Clean PTFE vessels and stirring bars by stirring with 50 ml of hydrochloric acid solution (5.3) and heating for 15 min. Discard the rinsings and conduct a blank test in each vessel in turn exactly as specified in 8.3. If any absorbance value is above the limit specified in 8.3, the cleaning procedure should be repeated or acid reagents of a higher purity should be used. At no stage should the stirring bars be handled with the fingers.
 - c) Platinum vessels, exclusively used for sodium analysis according to this International Standard, can be cleaned by the same method as the PTFE vessels [see b)]. Otherwise, they should be pre-cleaned by fusion with lithium tetraborate or lithium borate, until the absorbance readings fall to those for the lithium salt alone.
 - d) Rinse storage bottles with hydrochloric acid solution (5.3) before use.

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6.7 Atomic absorption spectrometer

WARNING — Follow the manufacturer's instructions for igniting and extinguishing the air-acetylene flame to avoid possible explosion hazards. Wear tinted safety glasses whenever the burner is in operation.

The atomic absorption spectrometer shall meet the following criteria.

- a) *Minimum sensitivity* — the absorbance of the most concentrated calibration solution (see 8.4.3) shall be at least 0,25.
- b) *Graph linearity* — the slope of the calibration graph covering the top 20 % of the concentration range (expressed as a change in absorbance) shall not be less than 0,7 of the value of the slope for the bottom 20 % of the concentration range determined in the same way.
- c) *Minimum stability* — the standard deviation of the absorbance of the most concentrated calibration solution and that of the zero calibration solution, each being calculated from a sufficient number of repetitive measurements, shall be less than 1,5 % and 0,5 % respectively of the mean value of the absorbance of the most concentrated calibration solution.

NOTES

- 6 The use of a strip-chart recorder and/or digital readout device is recommended to evaluate criteria a), b) and c) and for all subsequent measurements.
- 7 Instrument parameters will vary with each instrument. The following parameters were successfully used in several laboratories and they can be used as guidelines. Solutions were aspirated into an air-acetylene flame of premix burner.

— hollow cathode lamp, mA	10
— wavelength, nm	589,0
— air flow-rate, l/min	10
— acetylene flow-rate, l/min	2

In systems where the values shown for gas flow-rates do not apply, the ratio of the gas flow-rates may still be a useful guideline.

7 Sampling and samples

7.1 Laboratory sample

For analysis, use a laboratory sample of minus 100 µm particle size which has been taken in accordance with ISO 3081 or ISO 3082 and prepared in accordance with ISO 3082 or ISO 3083. In the case of ores having significant contents of combined water or oxidizable compounds, use a particle size of minus 160 µm.

NOTE 8 A guideline on significant contents of combined water and oxidizable compounds is incorporated in ISO 7764.

7.2 Preparation of predried test samples

Thoroughly mix the laboratory sample and, taking multiple increments, extract a test sample in such a manner that it is representative of the whole contents of the container. Dry the test sample at $105\text{ °C} \pm 2\text{ °C}$ as specified in ISO 7764. (This is the predried test sample.)

8 Procedure

8.1 Number of determinations

Carry out the analysis at least in duplicate in accordance with annex A, independently, on one predried test sample.

NOTE 9 The expression "independently" means that the second and any subsequent result is not affected by the previous result(s). For the particular analytical method, this condition implies that the repetition of the procedure is carried out either by the same operator at a different time or by a different operator including, in either case, appropriate recalibration.

8.2 Test portion

Taking several increments, weigh, to nearest 0,000 2 g, 0,2 g to 0,5 g (depending on the sodium concentration) of the predried test sample obtained in accordance with 7.2.

NOTE 10 The test portion should be taken and weighed quickly to avoid reabsorption of moisture.

8.3 Blank test and check test

Before proceeding to the treatment of test portions, ensure that the cleaning procedures conducted in note 5 together with the quality of the reagents being used have produced a blank test value for the sodium determination not greater than the equivalent of 0,002 % (*m/m*) sodium in the ore.

In each run, one blank test and one analysis of a certified reference material of the same type of ore shall be carried out in parallel with the analysis of the ore sample(s) under the same conditions. A predried test sample of the certified reference material shall be prepared as specified in 7.2.

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When the analysis is carried out on several samples at the same time, the blank value may be represented by one test, provided that the procedure is the same and the reagents are from the same reagent bottles.

When the analysis is carried out on several samples of the same type of ore at the same time, the analytical value of one certified reference material may be used.

NOTE 11 The certified reference material should be of the same type as the sample to be analysed and the properties of the two materials should be sufficiently similar to ensure that in either case no significant changes in the analytical procedure will become necessary. Where a certified reference material is not available, a reference material may be used (see 9.2.4).

8.4 Determination

To prevent contamination during analysis, the following precautions shall be taken:

- a) finger contact with sample, solutions and stirring bars shall be avoided;
- b) mouth suction of pipettes shall not be permitted.

8.4.1 Decomposition of the test portion

Transfer the test portion (8.2) to a 100 ml PTFE beaker (6.1)⁴⁾. Moisten it with a few drops of water, then add 10 ml of hydrochloric acid (5.1) and 10 ml of hydrofluoric acid (5.2). Add a PTFE-coated magnetic stirring bar (6.2) and cover with a PTFE cover. Adjust the temperature of the stirring hotplate (6.6) so that a temperature of about 98 °C will be maintained in water in a covered PTFE beaker. Heat, with stirring, for 45 min or until no further dissolution of the test portion occurs. Remove the cover, stop the stirrer, leaving the bar in the solution, and evaporate to dryness. Add 5 ml of hydrochloric acid (5.1) and evaporate again to dryness. Dissolve the salts in 5 ml of hydrochloric acid (5.1) and 40 ml of water and transfer to a 100 ml one-mark plastic volumetric flask (6.5). Dilute to volume with water and mix.

NOTE 12 If any significant amount of residue remains, conduct the digestion process in a stirred PTFE digestion bomb (6.3) for 45 min at 160 °C.

8.4.2 Treatment of the solution

If the concentration of sodium is too high, it is necessary to dilute the test solution. Transfer by plastic pipette (6.4) y ml of the test solution to a 100 ml one-mark plastic volumetric flask, add $0,1 \times (100 - y)$ ml of the background solution (5.4), dilute with water to volume and mix (see table 1).

A diluted test solution shall be measured together with a diluted blank test solution, containing the same amount of background solution as the test solution. Prepare the diluted blank test solution as follows. Pipette y ml of the blank test solution into a 100 ml one-mark plastic volumetric flask, add $0,1 \times (100 - y)$ ml of the background solution, dilute to volume with water and mix.

4) See note 2.

Table 1 — Dilution guide for test solution

Sodium concentration range %	Aliquot from 100 ml <i>y</i>
0,002 to 0,030	—
0,030 to 0,10	30,0
0,10 to 0,30	10,00
0,30 to 0,60	5,00
0,60 to 1,00	3,0 ¹⁾

1) Take 30 ml and dilute to 100 ml, aliquot 10 ml.

8.4.3 Preparation of the set of calibration solutions

From the sodium standard solution (5.5) prepare calibration solutions as follows.

Using plastic pipettes, transfer 0 ml; 2,0 ml; 5,0 ml; 10,0 ml; 15,0 ml of standard solution (5.5) to 100 ml one-mark plastic volumetric flasks. Add by plastic pipette 10 ml of the background solution (5.4) to each, dilute with water to volume and mix. These calibration solutions cover the concentration range 0 µg Na/ml to 1,5 µg Na/ml and contain 3 000 µg Fe/ml.

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Store the calibration solutions in plastic bottles.

8.4.4 Adjustment of atomic absorption spectrometer

Optimize the response of the instrument as specified in 6.7. Set the wavelength of sodium (589,0 nm) to obtain minimum absorbance. After 10 min preheating of the burner, adjust fuel flow and burner position to obtain maximum absorbance while aspirating the most concentrated calibration solution (see 8.4.3). Aspirate water and the calibration solution to establish that the absorbance reading is not drifting, and then set the reading for water to zero absorbance.

8.4.5 Atomic absorption measurements

Aspirate the calibration and test solutions or diluted test solutions in order of increasing absorption, starting with the blank test solution, or diluted blank test solution, and the zero calibration solution. When a stable response is obtained for each solution, record the readings. Aspirate the test solutions or diluted test solutions at the proper points in the

calibration series and record their readings. Aspirate water between each calibration and test solution. Repeat the measurements at least twice.

Obtain the net absorbance of each calibration solution by subtracting the average absorbance of the zero calibration solution. In a similar manner, obtain the net absorbance of the test solution or diluted test solution by subtracting the absorbance of the corresponding blank test solution.

Prepare calibration graphs by plotting the net absorbance values of the calibration solutions against the concentration, in micrograms of sodium per millilitre (the test solution or, if diluted, the diluted test solution is the final test solution).

Convert the net absorbance of the final test solution to micrograms of sodium per millilitre by means of the calibration graph.

NOTE 13 With concentration readings the calculation should be made from absorbances to permit checking of the graph's linearity and the blank test value.

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9 Expression of results

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9.1 Calculation of sodium content

The sodium content, as a percentage by mass, is calculated to five decimal places for contents higher than 0,01 % (*m/m*) and to six decimal places for contents lower than 0,01 % (*m/m*), using the equation

$$X = \frac{\rho_M}{m_1 \times 100} \quad \dots (1)$$

where

X is the sodium content, as a percentage by mass, of the test sample;

ρ_M is the concentration, in micrograms per millilitre, of sodium in the final test solution;

m_1 is the mass, in grams, of test sample represented in 100 ml of the final test solution (8.4.5), calculated from the equation