



Designation: D820 – 93 (Reapproved 2003)

Standard Test Methods for Chemical Analysis of Soaps Containing Synthetic Detergents¹

This standard is issued under the fixed designation D820; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 These test methods cover procedures for the chemical analysis of soaps containing synthetic detergents.

1.2 The analytical procedures appear in the following order:

	Sections
Moisture and Other Matter Volatile at 105°C (Oven Method)	5 and 6
Free Alkali or Free Acid	7 and 8
Anhydrous, Salt-Free, Soda Soap	9-12
Alcohol-Soluble Matter	13 and 14
Matter Insoluble in Water	15
Total Alkalinity of Matter Insoluble in Alcohol (Alkaline Salts)	16 and 17
Sodium Silicate	18-20
Phosphates	21-28
Phosphate (Colorimetric Method Using Molybdenum Blue)	29-34
Unsaponified and Unsaponifiable Matter	35-39
Free Fatty Matter	40
Chlorides in Alcohol-Soluble Matter	41-43
Rosin (McNicol Method)	44-47
Synthetic Detergent (by Difference)	48
Neutral, Inorganic Salts	49

1.3 *This standard does not purport to address all of the safety problems, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*

[D216 Method for Distillation of Natural Gasoline](#)^{2 3}

[D459 Terminology Relating to Soaps and Other Detergents](#)

[D875 Method for Calculation of Olefins and Aromatics in Petroleum Distillates from Bromine Number and Acid Absorption](#)^{4 5}

¹ These test methods are under the jurisdiction of ASTM Committee D12 on Soaps and Other Detergents and are the direct responsibility of Subcommittee D12.12 on Analysis of Soaps and Synthetic Detergents.

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² Discontinued, see 1987 *Annual Book of ASTM Standards*, Vol 05.01.

³ *Annual Book of ASTM Standards*, Vol 15.04.

⁴ Discontinued, see 1984 *Annual Book of ASTM Standards*, Vol 05.01.

⁵ *Annual Book of ASTM Standards*, Vol 11.01.

[D1193 Specification for Reagent Water](#)³

[D1768 Test Method for Sodium Alkylbenzene Sulfonate in Synthetic Detergents by Ultraviolet Absorption](#)³

[D2357 for Qualitative Classification of Surfactants by Infrared Absorption](#)³

[D2358 Test Method for Separation of Active Ingredient from Surfactant and Syndet Compositions](#)³

[D3049 Test Method for Synthetic Anionic Ingredient by Cationic Titration](#)

3. Terminology

3.1 *Definitions:*

3.2 The term *synthetic detergent* in these test methods is defined in accordance with Terminology D459, as follows:

3.3 *synthetic detergent*—a detergent produced by chemical synthesis and comprising an organic composition other than soap.

3.4 For definitions of other terms used in these test methods, refer to Terminology D459.

4. Purity of Reagents and Materials

4.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁶ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

4.2 Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Specification D1193.

⁶ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

MOISTURE AND OTHER MATTER VOLATILE AT 105°C (OVEN METHOD)

5. Apparatus

5.1 *Dish*—A porcelain or glass dish about 6 to 8 cm in diameter and about 2 to 4 cm in depth will be required.

6. Procedure

6.1 Weigh 5 ± 0.01 g of the sample in the dish, and dry to constant weight in an air oven at a temperature of $105 \pm 2^\circ\text{C}$. Constant weight is attained when heating for successive 1-h periods shows a loss (or gain) of not more than 0.1 %.

NOTE 1—Because of its established use in the trade, the term *weight* is employed in these methods in place of the technically correct term *mass*.

FREE ALKALI OR FREE ACID

7. Reagents

7.1 *Ethyl Alcohol, Neutral (95 %)*—Freshly boiled, reagent grade, ethyl alcohol, 95 % or higher, neutral to phenolphthalein, and containing only volatile denaturants, 95 mL,⁷ plus 5 mL of water.

7.2 *Phenolphthalein Indicator Solution (10 g/litre)*—Dissolve 1 g of phenolphthalein in 50 mL of neutral ethyl alcohol (95 %) and mix with 50 mL of water (see 7.1).

8. Procedure

8.1 Weigh 5 to 10 ± 0.01 g of the sample into a 300-mL Erlenmeyer flask. Add 200 mL of neutral ethyl alcohol (95 %). Equip the flask with an air-cooled reflux condenser, and digest the sample on a steam bath until the soap is dissolved (see Note 2). Remove the condenser, add 0.5 mL of the phenolphthalein indicator solution, and titrate immediately with standard acid or alkali. Calculate as NaOH, if alkaline, or as oleic acid, if acid.

NOTE 2—In the analysis of soaps known to contain little or no alkaline salts, it is unnecessary to filter the hot alcoholic soap solution. However, the filtration should be carried out in all cases where alkaline salts such as silicates, phosphates, borates, and similar salts are present, since these are known to affect the free alkali determination. Free alkali figures in soap or surfactant mixtures containing borax are unreliable, due to solubility of borax in hot alcohol.

ANHYDROUS, SALT-FREE, SODA SOAP

9. Apparatus

9.1 *Extraction Cylinder*, 250-mL, graduated, glass-stoppered, about 39 mm (1½ in.) in diameter and about 35.5 cm (14 in.) in length.

9.2 *Stokes Flask*, 100-mL, round-bottom (with the bottom blown out), sealed onto a 150-mL Erlenmeyer flask. A diagram of the Stokes flask is shown in Fig. 1.

9.3 *Siphon*, consisting of a two-hole rubber stopper fitted with small-diameter glass tubing as shown in Fig. 2.

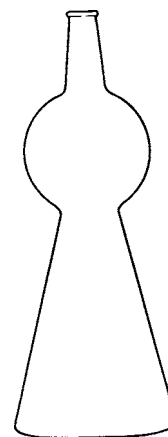


FIG. 1 Stokes Flask

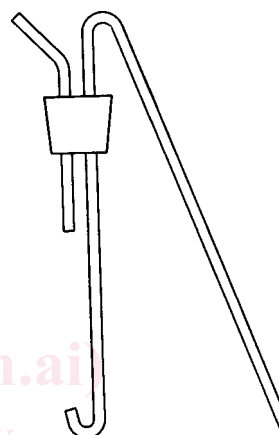


FIG. 2 Siphon

10. Reagents

10.1 *Ethyl Alcohol, Neutral (95 %)*—Freshly boiled, reagent grade, ethyl alcohol, 95 % or higher, neutral to phenolphthalein, and containing only volatile denaturants, 95 mL,⁷ plus 5 mL of water.

10.2 *Methyl Orange Indicator Solution (1 g/litre)*—Dissolve 0.1 g of methyl orange in water and dilute to 100 mL.

10.3 *Petroleum Ether*—The solvent used shall be of the pentane type, containing a minimum amount of isopentane, isohexane, and hexane, and boiling in the range 35 to 60°C.⁸

Distillation test: ^A	
Initial boiling point	35 to 38°C
Dry flask end point	52 to 60°C
Distilling under 54°C, min	95 %
Distilling under 40°C, max	60 %
Specific gravity at 15.5/15.5°C (60/60°F)	0.630 to 0.660
Color	water-white
Doctor test	sweet
Evaporation residue, 100 mL, max	0.0011 g
Copper-strip corrosion test ^B	noncorrosive
Unsaturated compounds ^C	trace only permitted
Residue in distilling flask	neutral to methyl orange
Blotter-strip odor test ^D	odorless within 12 min
Aromatic compounds ^E	no nitrobenzene odor

⁷ Fischer Scientific A962, or its equivalent, is suitable for this purpose.

⁸ J. T. Baker Analyzed Reagent 9268, or its equivalent, is suitable for this purpose.

Saponification value less than 1.0 mg KOH/100 mL

^A The distillation test shall be made in accordance with Method D216. As a check on the evaporation residue, 250 mL of the petroleum ether and 0.25 g of stearin or other hard fat (previously brought to constant weight by heating) when dried as in the actual determination (10.4) shall not show an increase in weight exceeding 0.003 g.

^B The copper-strip corrosion test shall be made by inserting a small polished copper strip into the petroleum ether in the distilling flask. There should be no appreciable darkening of the copper.

^C Unsaturated compounds shall be determined by the method for determining olefins described in Method D875.

^D Odor test: Immerse 1 in. of a strip of white unglazed blotting paper, approximately 1 by 4 by 0.166 in. in size, in the petroleum ether for 30 s, remove the strip, and allow to dry at room temperature in still air for 12 min.

^E Aromatic compounds: Add 5 drops of petroleum ether to 40 drops of sulfuric acid (H₂SO₄, sp gr 1.84) and 10 drops of nitric acid (HNO₃, sp gr 1.42) in a test tube, warm for 10 min, allow to cool for 30 min, transfer to a shallow dish, and dilute with water.

10.4 *Phenolphthalein Indicator Solution (10 g/litre)*—Dissolve 1 g of phenolphthalein in 50 mL of neutral ethyl alcohol (95 %) and then mix with 50 mL of water (see 10.1).

10.5 *Sodium Hydroxide, Standard Solution (0.1 N)*—Prepare and standardize a 0.1 N sodium hydroxide (NaOH) solution.

10.6 *Sodium Sulfate* (Na₂SO₄), anhydrous.

10.7 *Sulfuric Acid, Standard (0.5 N)*—Prepare and standardize a 0.5 N sulfuric acid (H₂SO₄) solution.

10.8 *Sulfuric Acid (sp gr 1.84)*—Concentrated sulfuric acid (H₂SO₄).

10.9 *Sulfur Acid (1+1)*—Gradually pour 10 g of concentrated sulfuric acid (H₂SO₄) (sp gr 1.84) onto 10 g of cracked ice made from distilled water, gently swirling the mixing vessel; or gradually pour the acid down the sides of the mixing vessel into an equal weight of water, swirling gently, while submersing the vessel in an ice bath.

11. Procedure

11.1 Weigh 2 ± 0.001 g of the sample into a tared 250-mL beaker. Add 25 mL of water and 25 mL of neutral ethyl alcohol (95 %), and warm on the steam bath until solution is complete. Cool, add 5 drops of methyl orange indicator solution, and titrate with 0.5 N H₂SO₄ to a pink color. Add 5 mL of H₂SO₄ (1+1) in excess.

11.2 Transfer the contents of the beaker to a 250-mL extraction cylinder or a Stokes flask, equipped with a siphon. Wash the beaker alternately with equal parts of hot water and hot ethyl alcohol (95 %), adding the washings to the extraction cylinder or Stokes flask. Keep the total volume for extraction under 160 mL in the extraction cylinder, or within the constricted portion of the Stokes flask. Wash the beaker with a small amount of petroleum ether to remove any traces of fatty acids and fatty matter and add to the extraction cylinder or Stokes flask. Cool the cylinder or flask under tap water to a temperature not to exceed 25°C. Add 50 mL of petroleum ether and allow to stand for ½ h without shaking. Remove the greater part of the fatty acids by drawing off the petroleum ether layer as closely as possible, by means of a glass siphon, into a 500-mL separatory funnel. Repeat the extractions five more times with petroleum ether, using 50-mL portions, and shaking the cylinder thoroughly each time.

NOTE 3—If an emulsion appears at this point, it may be broken by the

addition of 10 g of anhydrous Na₂SO₄.

11.3 Combine the petroleum ether extracts and wash with small portions of distilled water until the water washings are no longer acid to methyl orange indicator solution. Dry the combined, washed, petroleum ether extracts with anhydrous Na₂SO₄, and filter through paper into the original tared 250-mL beaker. Wash the separatory funnel with two small portions of petroleum ether, filtering and adding the washings to the beaker.

11.4 Evaporate the petroleum ether extract on the steam bath until about 1 mL remains. Then swirl manually until the last trace of solvent evaporates and the odor of petroleum ether is no longer perceptible. Cool in a desiccator and weigh as total fatty matter, which is defined as fatty and rosin acids plus unsaponified and unsaponifiable fatty matter.

11.5 Dissolve the total fatty matter in 50 mL of neutral ethyl alcohol (95 % v) with warming. Add phenolphthalein indicator and titrate with 0.1 NaOH solution to a pink end point.

12. Calculations

12.1 Calculate the percentage of anhydrous, salt-free, soda soap as follows:

$$A = G - F \quad (1)$$

$$G = [(VN \times 0.022) + E/W] \times 100$$

where:

- A = weight percent of anhydrous, salt-free, soda soap,
- G = weight percent of soda soap plus unsaponified and unsaponifiable fatty matter,
- F = weight percent of unsaponified and unsaponifiable fatty matter (Section 39),
- V = millilitres of NaOH solution used in titration (11.5),
- N = normality of NaOH solution,
- E = grams of extract (11.4), and
- W = grams of sample (11.1), and
- 0.022 = net gain in milliequivalent weight from the conversion of the fatty acid to the sodium salt by replacement of a proton with a sodium ion.

ALCOHOL-SOLUBLE MATTER

13. Reagents

13.1 *Ethyl Alcohol, Neutral (95 %)*—Freshly boiled, reagent grade, ethyl alcohol, 95 % or higher, neutral to phenolphthalein, and containing only volatile denaturants, 95 mL⁷ plus 5 mL of water.

13.2 *Ethyl Alcohol, Neutral (absolute)*—Freshly boiled absolute ethyl alcohol, neutral to phenolphthalein.⁷

14. Procedure

14.1 Weigh 2 ± 0.001 g of the sample into a 250-mL beaker. Add 100 mL of neutral ethyl alcohol (95 %), cover the beaker, and heat on the steam bath with frequent stirring and maceration of the sample until completely disintegrated. Let settle and filter the supernatant liquid through a tared Gooch crucible with a glass wool pad, with suction into a tared 300-mL Erlenmeyer flask, retaining as much of the residue as possible in the beaker. Repeat this extraction three times with

25-mL portions of hot neutral ethyl alcohol (95 %), each time retaining as much of the residue as possible in the beaker. Finally, evaporate any remaining alcohol and dissolve the residue in the smallest possible quantity of hot water (5 mL if sufficient). Reprecipitate the alcohol-insoluble matter by slowly adding, while stirring vigorously, 50 mL of neutral ethyl alcohol (absolute).

NOTE 4—Solution and reprecipitation of alcohol-insoluble matter is necessary for complete separation from alcohol-soluble matter.

14.2 Heat the solution to boiling on the steam bath, filter, and transfer the precipitate quantitatively to the Gooch crucible, washing several times with neutral ethyl alcohol (95 %). Evaporate the combined filtrate and washings in the Erlenmeyer flask on the steam bath, and then dry to constant weight at $105 \pm 2^\circ\text{C}$. Calculate total alcohol-soluble matter.

14.3 Reserve the flask and contents for the determination of free fatty matter (Section 37). Reserve the Gooch crucible and contents, without drying, for the determination of matter insoluble in water (Section 15) and total alkalinity of matter insoluble in alcohol (alkaline salts) (Section 17).

MATTER INSOLUBLE IN WATER

15. Procedure

15.1 Wash the alcohol-insoluble matter retained in the Gooch crucible (Section 14) thoroughly with hot water until the washings are no longer alkaline to phenolphthalein. Reserve the filtrate for determination of total alkalinity of matter insoluble in alcohol (Section 17). Dry the crucible and residue to constant weight at $105 \pm 2^\circ\text{C}$ and calculate the percentage of matter insoluble in water.

TOTAL ALKALINITY OF MATTER INSOLUBLE IN ALCOHOL (ALKALINE SALTS)

16. Reagents

16.1 *Hydrochloric Acid*—Prepare and standardize a 1 N hydrochloric acid solution.

16.2 *Methyl Orange Indicator Solution (1 g/litre)*—Dissolve 0.1 g of methyl orange in water and dilute to 100 mL.

17. Procedure

17.1 Titrate the water solution obtained in the determination of matter insoluble in water (Section 15) with the standard 1 N hydrochloric acid solution of 16.1, using methyl orange as indicator.

17.2 Calculate the alkalinity as sodium carbonate (Na_2CO_3) as follows:

$$\text{Na}_2\text{CO}_3, \% \text{ weight} = 106 \text{ VN}/20 \text{ W} \quad (2)$$

where:

V = millilitres of standard acid used,

N = normality of standard acid, and

W = weight of sample (14.1), g

SODIUM SILICATE

18. Reagents

18.1 *Hydrochloric Acid (sp gr 1.19)*—Concentrated hydrochloric acid (HCl).

18.2 *Hydrofluoric Acid (sp gr 1.15)*—Prepare a solution of hydrofluoric acid (HF) having a specific gravity of 1.15.

18.3 *Sulfuric Acid (sp gr 1.84)*—Concentrated sulfuric acid (H_2SO_4).

19. Procedure

19.1 When the material contains no mineral matter that is insoluble in water, ignite a portion of the sample containing not to exceed 0.2 g of silica (SiO_2) in a platinum dish (Note 5) at a low temperature. When charred, extract the soluble salts with water, return the paper and charred residue to the dish, and complete the ignition. Unite the residue in the dish and the water extract, carefully acidify with HCl, and finally add the equivalent of from 5 to 10 mL of HCl in excess. The dish or casserole containing the solution should be covered with a watch glass while adding acid, so as to avoid loss by spray.

NOTE 5—When phosphates are present, a platinum dish should not be used.

19.2 When the material contains mineral matter insoluble in water, or a determination of highest accuracy is not necessary, take a portion of the solution after titrating the matter insoluble in alcohol (Section 17) containing not more than 0.2 g of SiO_2 and add 5 to 10 mL of HCl.

19.3 Evaporate the acidified solution obtained in accordance with 19.1 or 19.2 (washing off and removing the cover glass if used) to dryness on a steam bath or hot plate at a temperature not exceeding 120°C . Cool, moisten with HCl and let stand 5 to 10 min, breaking up all lumps with a stirring rod. Add about 25 mL of hot water. Heat a few minutes and filter through a small ashless paper. Wash thoroughly with hot water.

19.4 Evaporate the filtrate to dryness and proceed as described in 19.3, filtering on a second paper. Carefully ignite the two papers and contents in a weighed platinum crucible, first at a low temperature until the paper is consumed, then over a blast lamp. Cool in a desiccator, weigh, and repeat until constant weight is obtained.

19.5 If extreme accuracy is desired, moisten the weighed contents of the crucible with water, add 10 mL of HF and 4 drops of concentrated H_2SO_4 , and evaporate to dryness over a low flame. Ignite at the temperature of the blast lamp for about 2 min, cool in a desiccator, and weigh.

20. Calculation

20.1 Calculate the sodium silicate having the ratio 1 Na_2O :3.25 SiO_2 as follows:

$$\text{Sodium silicate, \% w} = (A - B) \times 1.308 \quad (3)$$

where:

A = grams of ignited residue before treatment with HF (19.4), and

B = grams of ignited residue after treatment with HF (19.5).

PHOSPHATES

21. Application

21.1 This test method is applicable to any species of alkali metal phosphates free of interfering ions. This test method can

be used for the analysis of soap and synthetic detergent builders if the sample is properly prepared (see Section 26). This test method does not apply when the level of phosphate present is equivalent to or less than 2 % P₂O₅.

Phenolphthalein solution	32 mL
Thymol blue solution	8 mL
Methylene blue solution	4 mL
Alcohol	24 mL

22. Summary of Test Method

22.1 All of the phosphate present is converted, by acid hydrolysis, to the ortho form and titrated between pH 4.3 and 8.8 with NaOH solution.

23. Interferences

23.1 Heavy metals such as iron, aluminum, calcium, magnesium, etc., that will precipitate, either as insoluble phosphates or hydroxides, before the upper end point is reached, will interfere. Interference also occurs if borates, sulfites, carbonates, or other buffering materials are present. The last two compounds and some of the borate will be expelled during the acid hydrolysis boil. Borate is removed by conversion to methyl borate and subsequent volatilization. Ammonia or other weak bases also will interfere. The most common interference is from silicic acid. Experiment and experience in analysis of spray-dried synthetics have shown that unless the ratio of the percentage of SiO₂ to the percentage of P₂O₅ approaches or exceeds 0.2, the interference by silicates will be so slight that it may be neglected. Larger amounts must be dehydrated as directed, but need not be removed by filtration during preparation of the sample.

24. Apparatus

24.1 *Electrometric Titration Apparatus*, equipped with glass and calomel electrodes. Any standard pH meter, capable of performing titrations accurate to ±0.1 pH and accurately standardized at pH 4.0 and 8.0 is suitable.

24.2 *Gas Burners*, suitable for heating the sample to approximately 550°C.

24.3 *Muffle Furnace*, with suitable pyrometer and controls for maintaining temperatures up to 550°C.

24.4 *Motor Stirrer*, air or electric.

25. Reagents

25.1 *Hydrochloric Acid (sp gr 1.19)*—Concentrated hydrochloric acid (HCl).

25.2 *Mixed Indicator (optional)*—Prepare the following solutions:

25.2.1 *Methyl Orange Solution (0.5 g/litre)*—Dissolve 0.05 g of methyl orange in water and dilute to 100 mL.

25.2.2 *Phenolphthalein, Alcohol Solution (5.0 g/litre)*—Dissolve 0.50 g of phenolphthalein in alcohol (50 %) and dilute to 100 mL with alcohol.

25.2.3 *Thymol Blue Solution (0.4 g/litre)*—Dissolve 0.04 g of thymol blue in water and dilute to 100 mL.

25.2.4 *Methylene Blue Solution (1.0 g/litre)*—Dissolve 0.10 g of methylene blue in water and dilute to 100 mL.

25.2.5 *Alcohol (90 %)*—Alcohol (90 %) prepared from alcohol conforming to Formula No. 3A or No. 30 of the U.S. Bureau of Internal Revenue. Mix the solutions in the following proportions:

Methyl orange solution	32 mL
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The individual components are stable indefinitely. The mixed indicator should be prepared at least weekly. In practice, 3 mL of this mixed indicator are used in a final volume of approximately 250 mL of solution to be titrated. The lower end point is taken as the first change from gray to a definite green; the upper end point is the change from pink to a bright purple.

25.3 *Sodium Hydroxide, Standard Solution (0.5 or 1.0 N)*—Prepare a 0.5 or 1.0 N carbonate-free solution of sodium hydroxide (NaOH) and standardize accurately.

25.4 *Sodium Hydroxide Solution (1+1)*—Dissolve sodium hydroxide (NaOH) in an equal weight of water. When using, decant the solution from the settled carbonate. A more dilute solution may be used. NaOH solutions must be protected from carbon dioxide (CO₂) contamination.

26. Preparation of Sample Solutions

26.1 Commercial sodium or potassium phosphates need no special preparation except solution in water. Weigh a portion of the well-mixed sample to the nearest 0.001 g, transfer directly to a 400-mL beaker, and dissolve in about 100 mL of water. Neutralize to litmus paper with HCl (sp gr 1.19) and add 10 mL excess. The optimum size of sample is given by the formula:

$$\text{Grams of sample} = (N \times 280) / P \quad (4)$$

where:

N = normality of the NaOH solution to be used in the titration, and

P = percentage of P₂O₅ expected in the sample.

26.2 Soap products may be analyzed by using the filtrate from the silicon dioxide (SiO₂) determination. Use care not to exceed the sample weight prescribed in 26.1. Alternatively the sample may be prepared as described in 26.3.

26.3 Built synthetic detergent samples are treated by ashing. Weigh a sample of suitable size (but not to exceed 10 g) to the nearest 0.001 g. When the expected percentage of P₂O₅ in the sample is known, the equation in 26.1 may be used to calculate a suitable sample weight. Place the sample in a 400- mL high-silica glass beaker, or a new, well-glazed porcelain or silica evaporating dish or large crucible and ignite gently over a low gas burner until most of the volatile combustible matter is burned off. Use care so that the sample is not heated to above 550°C with the gas burner. Transfer to a muffle, operated at not over 550°C, for 10 to 15 min. The ignited residue need not be free from carbon and usually is of a grayish color. Cool and add cautiously 10 mL of HCl. Evaporate to dryness on open steam. If the ratio of the percentage of SiO₂ to the percentage of P₂O₅ approaches or exceeds 0.2 or is unknown, dehydrate the silicates completely by cooling the sample and repeating the HCl addition and evaporation two additional times. After the third evaporation, continue to heat the residue for an additional 15 to 20 min after dryness is attained to ensure complete dehydration of SiO₂. After the sample appears dry, last traces of water or HCl may be removed in a 110 to 120°C oven, if

desired, in this step. Cool the sample and transfer into a 400-mL beaker using distilled water and proceed as in 26.3.1 or 26.3.2.

26.3.1 If the sample contains perborate or borate, evaporate to dryness on a steam bath, add about 200 mL of methanol, 10 mL of HCl and two or three hollow glass beads. Partially cover the beaker with a watch glass and boil down to a volume of about 20 mL. (The boiling time must be *at least* 30 min.) Evaporate down to less than 10 mL on a steam bath under a stream of nitrogen or clean, dry air. Proceed as described in Section 27.

26.3.2 If the sample is known to be free from perborate and borate, add distilled water to make a total of about 90 mL. Add 10 mL of HCl and proceed as in Section 27.

27. Procedure

27.1 Each solution in a 400-mL beaker, prepared as described in 26.1, should have a volume of about 100 mL and contain an excess of at least 10 mL of HCl (sp gr 1.19). Cover with a watch glass and boil gently for a minimum of 30 min. Up to 60 min may be necessary for phosphates of the glass type. All phosphate must be in the ortho form. Cool to room temperature (20 to 30°C).

27.2 Dilute to 200 mL, place on an electrometric titration stand (Note 6), and neutralize to a pH of 4.3. Most of the neutralization may be made with NaOH solution (1+1), but final adjustment should be made with the standard NaOH solution (0.5 or 1.0 *N*) to be used in the titration. Cool again, if necessary, to maintain the temperature below 30°C. Titrate carefully to the upper end point (pH 8.8) recording the titration between end points (*T*).

NOTE 6—The mixed indicator may be used for this titration but with some small sacrifice of accuracy. If the samples have been prepared by the ignition method, they must be filtered and the paper washed thoroughly, after the acid hydrolysis, as particles of carbon obscure the visual end point. The color changes can be checked by comparison with pH meter readings to acquire familiarity with the exact shade required. For greatest accuracy, titration with a pH meter is recommended.

28. Calculation

28.1 Calculate the percentage of total P₂O₅ as follows:

$$\text{Total P}_2\text{O}_5, \% \text{ weight} = (TN \times 7.098)/W \quad (5)$$

where:

T = millilitres of NaOH solution required for titration of the sample,

N = normality of the NaOH solution, and

W = grams of sample in the sample solution.

PHOSPHORUS (COLORIMETRIC METHOD USING MOLYBDENUM BLUE)

29. Summary of Test Method

29.1 The sample is dissolved in water and the phosphates hydrolyzed to orthophosphate with mineral acid. A diluted aliquot of the hydrolyzed sample is reacted with a single-solution molybdate reagent to form the molybdenum blue complex with the ortho-phosphate present. After extraction into isobutanol to remove any interference from product

colorants or turbidity, the molybdenum blue complex is measured colorimetrically.

30. Scope

30.1 This procedure must be used to determine the level of phosphate present in detergents when the level present is equal to or less than that equivalent to 2 % P₂O₅. Results are reported as percent weight P.

31. Reagents

31.1 The single-solution molybdate reagent is prepared from solutions of sulfuric acid, ammonium molybdate, potassium antimonyltartrate, and ascorbic acid as follows:

31.1.1 *Solution A, Sulfuric Acid*—Cautiously add, with cooling, 140 mL of concentrated sulfuric acid (H₂SO₄, sp gr 1.82) to 900 mL of water. Cool to room temperature.

31.1.2 *Solution B, Ammonium Molybdate*—Dissolve 15.0 g of ammonium molybdate in 500 mL of water. Store in the dark.

31.1.3 *Solution C, Ascorbic Acid*—Dissolve 13.5 g of ascorbic acid in 250 mL of water. Store at 4°C. Discard after 1 week.

31.1.4 *Solution D, Potassium Antimonyltartrate*—Dissolve 0.35 g of potassium antimonyltartrate (C₄H₄KO₇Sb) in 500 mL of distilled water. Store at 4°C in the dark.

31.2 Mix thoroughly 125 mL of Solution A, 50 mL of Solution B, 50 mL of Solution C, and 25 mL of Solution D. The resulting single-solution molybdate reagent should be yellow in color. Allow to reach room temperature prior to use. A greenish or blue color indicates phosphate contamination, in which case the solution should be discarded. The single-solution molybdate reagent must be prepared fresh daily. The individual solutions are stable as stated above. *Add solutions in the order specified* “A” 1st, “B” 2nd, “C” 3rd, and “D” 4th.

31.3 *Potassium dihydrogen phosphate (anhydrous)*, reagent grade.

32. Calibration

32.1 A phosphate stock solution for calibration may be prepared by dissolving in a 1000-mL volumetric flask 0.4394 ± 0.0002 g of potassium dihydrogen phosphate (anhydrous) (KH₂PO₄) which has been previously dried at 105°C for 1 h. This stock solution contains ortho-phosphate equivalent to 100 µg P/mL.

32.2 Make a dilute stock solution containing 1.00 µg P/ mL by pipeting 10 mL of the original stock solution into a 1-L volumetric flask and diluting to volume with water.

32.3 Transfer by pipet 10, 20, 30, 40, and 50-mL aliquots of the dilute stock solution to separate 250-mL separatory funnels containing about 50 mL of water. Add additional water to each separatory funnel to bring the total volume to 100 mL. Use as a blank 100 mL of water added to an additional separatory funnel.

32.4 To each calibration point and the blank add 20 mL of mixed reagent from a dispensing flask or graduated cylinder; stopper and shake the separatory funnel vigorously to mix well. Let stand for 10 min but no longer than 15 min. (The formation of the molybdenum blue complex is not instantaneous; however color formation is essentially complete in the first 2 to 3 min.) Add 40 mL of isobutanol using a graduated cylinder and