

SLOVENSKI STANDARD

SIST EN 15077:2013

01-julij-2013

Nadomešča:
SIST EN 15077:2006

Kemikalije, ki se uporabljajo za pripravo bazenske vode - Natrijev hipoklorit

Chemicals used for treatment of swimming pool water - Sodium hypochlorite

Produkte zur Aufbereitung von Schwimm- und Badebeckenwasser - Natriumhypochlorit

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Produits chimiques utilisés pour le traitement de l'eau des piscines - Hypochlorite de sodium
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Ta slovenski standard je istoveten z: [EN 15077:2013](#) 3d6-4f2b-9a91-
9b6c47095438/sist-en-15077-2013

ICS:

- | | | |
|-----------|------------------------------|-------------------------------------|
| 13.060.25 | Voda za industrijsko uporabo | Water for industrial use |
| 71.100.80 | Kemikalije za čiščenje vode | Chemicals for purification of water |

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EUROPEAN STANDARD
NORME EUROPÉENNE
EUROPÄISCHE NORM

EN 15077

May 2013

ICS 71.100.80

Supersedes EN 15077:2006

English Version

**Chemicals used for treatment of swimming pool water - Sodium
hypochlorite**

Produits chimiques utilisés pour le traitement de l'eau des
piscines - Hypochlorite de sodium

Produkte zur Aufbereitung von Schwimm- und
Badebeckenwasser - Natriumhypochlorit

This European Standard was approved by CEN on 14 March 2013.

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This European Standard exists in three official versions (English, French, German). A version in any other language made by translation under the responsibility of a CEN member into its own language and notified to the CEN-CENELEC Management Centre has the same status as the official versions.

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EUROPEAN COMMITTEE FOR STANDARDIZATION
 COMITÉ EUROPÉEN DE NORMALISATION
 EUROPÄISCHES KOMITEE FÜR NORMUNG

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EN 15077:2013 (E)**Foreword**

This document (EN 15077:2013) has been prepared by Technical Committee CEN/TC 164 "Water supply", the secretariat of which is held by AFNOR.

This European Standard shall be given the status of a national standard, either by publication of an identical text or by endorsement, at the latest by November 2013, and conflicting national standards shall be withdrawn at the latest by November 2013.

Attention is drawn to the possibility that some of the elements of this document may be the subject of patent rights. CEN [and/or CENELEC] shall not be held responsible for identifying any or all such patent rights.

This document supersedes EN 15077:2006.

Significant technical differences between this edition and EN 15077:2006 are as follows:

- Replacement of warning and safety precautions notes by labelling according to Regulation (EC) No 1272/2008 [3].

According to the CEN-CENELEC Internal Regulations, the national standards organisations of the following countries are bound to implement this European Standard: Austria, Belgium, Bulgaria, Croatia, Cyprus, Czech Republic, Denmark, Estonia, Finland, Former Yugoslav Republic of Macedonia, France, Germany, Greece, Hungary, Iceland, Ireland, Italy, Latvia, Lithuania, Luxembourg, Malta, Netherlands, Norway, Poland, Portugal, Romania, Slovakia, Slovenia, Spain, Sweden, Switzerland, Turkey and the United Kingdom.

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Introduction

In respect of potential adverse effects on the quality of water for swimming pools caused by the product covered by this European Standard, the following statements apply:

- This European Standard provides no information as to whether the product may be used without restriction in any of the Member States of the EU or EFTA.
- It should be noted that, while awaiting the adoption of verifiable European criteria, existing national regulations concerning the use and/or the characteristics of this product remain in force.

NOTE 1 Conformity with this European standard does not confer or imply acceptance or approval of the product in any of the Member States of the EU or EFTA. Use of the product covered by this European Standard is subject to regulation or control by National Authorities.

NOTE 2 This product is a biocide and needs to comply with the relevant legislation in force. In the European Union, at the time of publication, this legislation is Directive 1998/8/EC [1].

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1 Scope

This European Standard is applicable to sodium hypochlorite used directly or for the production of formulations for treating swimming pool water. It describes the characteristics of sodium hypochlorite and specifies the requirements and the corresponding test methods for sodium hypochlorite. It gives information on its use for treating swimming pool water and determines the rules relating to safe handling and use of sodium hypochlorite (see Annex B).

2 Normative references

The following documents, in whole or in part, are normatively referenced in this document and are indispensable for its application. For dated references, only the edition cited applies. For undated references, the latest edition of the referenced document (including any amendments) applies.

EN 901, *Chemicals used for treatment of water intended for human consumption — Sodium hypochlorite*

3 Description

3.1 Identification

3.1.1 Chemical name

Sodium hypochlorite.

3.1.2 Synonym or common names **iTech STANDARD PREVIEW (standards.itech.ai)**

3.1.3 Relative molecular mass

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74,44.
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3.1.4 Empirical formula

NaClO.

3.1.5 Chemical formula

NaClO.

3.1.6 CAS Registry Number¹⁾

7681-52-9.

3.1.7 EINECS reference²⁾

231-668-3.

3.2 Commercial form

The product is supplied as an aqueous solution with an available (active) chlorine concentration up to a mass fraction of 18 %.

1) Chemical Abstracts Service Registry Number.

2) European Inventory of Existing Commercial Chemical Substances.

3.3 Physical properties

3.3.1 Appearance and odour

The product is a clear yellowish-green solution with a faint chlorinous odour.

3.3.2 Density

The density of the product varies between 1,13 g/ml and 1,30 g/ml at 20 °C.

3.3.3 Solubility in water

The product is capable of being mixed with water in any proportion.

3.3.4 Vapour pressure

Approximately 2,5 kPa at 20 °C.

3.3.5 Boiling point at 100 kPa³⁾

Not applicable.

3.3.6 Crystallisation and freezing point

At approximately -10 °C, crystallisation of NaOCl·6H₂O starts.

Freezing of the concentrated product takes place between - 20 °C and - 30 °C.
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3.3.7 Specific heat

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3.3.8 Viscosity (dynamic)

2.6 mPa s at 20 °C

3.3.9 Critical temperature

Not applicable

3.3.10 Critical pressure

Not applicable.

3.3.11 Physical hardness

Not applicable.

3.4 Chemical properties

The product is an alkaline solution with a pH value greater than 11 at 20 °C.

It reacts with acids and acidic salts to form chlorine

$$3) \quad 100 \text{ kPa} = 1 \text{ bar.}$$