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**INTERNATIONAL STANDARD**



**424**

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**Photographic grade sodium carbonate, anhydrous —  
Specification**

*Carbonate de sodium anhydre de qualité photographique — Spécifications*

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## FOREWORD

ISO (the International Organization for Standardization) is a worldwide federation of national standards institutes (ISO Member Bodies). The work of developing International Standards is carried out through ISO Technical Committees. Every Member Body interested in a subject for which a Technical Committee has been set up has the right to be represented on that Committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work.

Draft International Standards adopted by the Technical Committees are circulated to the Member Bodies for approval before their acceptance as International Standards by the ISO Council.

Prior to 1972, the results of the work of the Technical Committees were published as ISO Recommendations; these documents are now in the process of being transformed into International Standards. As part of this process, Technical Committee ISO/TC 42 has reviewed ISO Recommendation R 424 and found it technically suitable for transformation. International Standard ISO 424 therefore replaces ISO Recommendation R 424-1965.

[ISO 424:1976](https://standards.iteh.ai/catalog/standards/sist/c7b49521-2b67-4cf2-95a4-48b876d16115/iso-424-1976)

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ISO Recommendation R 424 was approved by the Member Bodies of the following countries :

Belgium	Italy	Sweden
Brazil	Japan	Switzerland
Canada	Netherlands	United Kingdom
Chile	New Zealand	U.S.A.
Germany	Romania	U.S.S.R.

No Member Body expressed disapproval of the Recommendation.

No Member Body disapproved the transformation of ISO/R 424 into an International Standard.

# Photographic grade sodium carbonate, anhydrous — Specification

## 0 INTRODUCTION

This International Standard is one of a series of specifications for photographic grade chemicals which are commonly used in the processing of sensitized photographic materials. These specifications have been prepared to establish criteria of purity which will provide a practical and economical grade, and prevent possible faulty processing which might be caused by chemicals of inferior quality, and to furnish manufacturers, suppliers, and processors with reliable and readily available specifications for photographic chemicals of satisfactory quality.

Photographic grade chemicals are those which meet the requirements specified in the appropriate International Standards. These specifications set out purity standards and state the limiting concentrations and test methods for certain inert or photographically harmful impurities that may be present.

Originally these specifications were based on known requirements for black-and-white photographic processing, but increased attention has been paid to the requirements of colour processing. Experience to date indicates that chemicals meeting these specifications are satisfactory for colour processes in general use.

### 0.1 Specification requirements

These specifications set out chemical and physical requirements. While it is recognized that the ultimate criterion of the quality of a photographic chemical is its successful performance in a photographic test, present knowledge indicates that, from a practical standpoint, chemical and physical methods of testing are generally adequate. The photographic industry has accumulated a comprehensive collection of such chemical tests for impurities. These tests, which correlate with objectionable photographic effects, have been drawn upon in the formulation of these specifications. Chemical tests are generally more sensitive, less variable, and less costly than photographic tests.

Purity requirements have been set as low as possible, consistent with the objectives mentioned. If, however, the purity of a commonly available grade of chemical exceeds photographic processing requirements, and if there is no economic penalty in its use, the purity requirements have been set to take advantage of the higher quality materials.

Every effort has been made to keep the number of requirements in each specification to a minimum. The requirements generally include only those photographically harmful impurities which, through experience, are likely to be present. Inert impurities are limited to amounts which will not unduly reduce the assay.

Assay procedures have been included in all cases where a satisfactory method is available. An effective assay requirement serves not only as a safeguard of chemical purity, but also as a valuable complement to the identity test. All assays are intended to be made on undried samples in view of the fact that photographic processing chemicals are normally used "as received".

Identity tests have been included in the specifications wherever a possibility exists that another chemical or a mixture of chemicals could pass the other tests.

All requirements listed in clause 3 of each specification are mandatory. The physical appearance of the material and any footnotes are for general information only and are not part of the requirements.

### 0.2 Selection of test methods

Efforts have been made to employ tests which are capable of being run in any normally equipped laboratory and, wherever possible, to avoid tests which require highly specialized equipment or techniques. Instrumental methods have been specified only as alternative methods or alone in those cases where no other satisfactory method is available.

While the test methods set out in the specifications are recommended, the use of other equally reliable methods is allowed. In case of disagreement in results, the method called for in the specification shall prevail. Where a requirement states "to pass test", however, alternative methods shall not be used.

### 0.3 Reagents

An effort has been made to minimize the number of reagents employed in this series of specifications. The methods of preparation and of standardization have been included in all cases where these are not common, or where a preferred method is desirable.

Details of reagent preparation and standardization are included in each specification in which the reagent is called for so that each specification shall be self-sufficient.

## 1 SCOPE AND FIELD OF APPLICATION

This International Standard specifies the purity requirements of, and test methods for, photographic grade sodium carbonate, anhydrous.

## 2 CHARACTERISTICS

Sodium carbonate, anhydrous, is in the form of white granules or powder, of chemical formula  $\text{Na}_2\text{CO}_3$  and relative molar mass 106,0.

## 3 REQUIREMENTS

### 3.1 Assay

The assay shall be not less than 98,0 % (*m/m*), expressed as  $\text{Na}_2\text{CO}_3$ , when determined by the method described in 4.1.

### 3.2 Appearance of solution

An aqueous solution shall be clear and free from sediment, other than a slight flocculence, when examined by the method described in 4.2.

### 3.3 Free alkali content

The free alkali content, expressed as NaOH, shall be not greater than 0,2 % (*m/m*).

Conformity with this requirement shall be determined by the limit test described in 4.3, when not more than 1,0 ml of 0,1 M hydrochloric acid solution shall be required to discharge the colour from the test solution.

### 3.4 Halide content

The halide content, expressed as NaCl, shall be not greater than 0,6 % (*m/m*).

Conformity with this requirement shall be determined by the limit test described in 4.4, when the turbidity produced in the test solution shall be not greater than that produced in the control solution.

### 3.5 Bicarbonate content

The bicarbonate content, expressed as  $\text{NaHCO}_3$ , shall be not greater than 0,7 % (*m/m*).

Conformity with this requirement shall be determined by the limit test described in 4.5, when not more than 1,7 ml of 0,1 M sodium hydroxide solution shall be required to produce a faint pink colour, persisting for 30 s, in the test solution.

### 3.6 Heavy metals content

The heavy metals content, expressed as lead (Pb), shall be not greater than 10 mg/kg.

Conformity with this requirement shall be determined by the limit test described in 4.6, when the colour produced

in the test solution shall be not greater than that produced in the control solution.

### 3.7 Iron content

The iron content, expressed as iron (Fe), shall be not greater than 20 mg/kg.

Conformity with this requirement shall be determined by the limit test described in 4.7, when the colour produced in the test solution shall be not greater than that produced in the control solution.

### 3.8 Reaction to ammoniacal silver nitrate

The colour or turbidity produced in the test solution by ammoniacal silver nitrate solution shall be not greater than that produced in the control solution by ammonia solution, when examined by the method described in 4.8.

## 4 TEST METHODS

Reagents used in the tests shall be recognized reagent grade chemicals normally used for careful analytical work. In all the directions the acids and ammonia solution referred to shall be of full strength unless dilution is specified. Dilution is specified in terms of molar concentration (molarity)<sup>1)</sup> when standardization of the reagent is required. When dilution is indicated as (1 + x), it means that 1 volume of the reagent or strong solution is added to x volumes of distilled water.

Distilled water, or water otherwise produced of at least equal purity, shall be used whenever water is required.

### 4.1 Assay

#### 4.1.1 Reagents

**4.1.1.1 Hydrochloric acid**, 1 M standard volumetric solution.

**4.1.1.2 Methyl orange** indicator solution, 0,4 g/l.

#### 4.1.2 Apparatus

Ordinary laboratory apparatus and

**4.1.2.1 Burette**, 50 ml capacity, conforming to class A of ISO/R 385.

#### 4.1.3 Procedure

Weigh, to the nearest 0,001 g, a test portion of about 2,0 g of the laboratory sample and dissolve in 50 ml of water. Add 2 drops of the methyl orange indicator solution (4.1.1.2) and titrate with the standard volumetric hydrochloric acid solution (4.1.1.1) to the colour change from yellow to red.

1) 1 mol/l = 1 kmol/m<sup>3</sup> = 1 mol/dm<sup>3</sup> = 1 M

#### 4.1.4 Calculation

The assay, expressed as a percentage by mass of sodium carbonate, anhydrous ( $\text{Na}_2\text{CO}_3$ ), is given by the formula

$$\frac{5,30 VT}{m}$$

where

$V$  is the volume, in millilitres, of the hydrochloric acid solution (4.1.1.1) used for the titration;

$T$  is the exact molarity of the hydrochloric acid solution (4.1.1.1);

$m$  is the mass, in grams, of the test portion.

#### 4.2 Appearance of solution test

Dissolve 200 g of the laboratory sample in 1 000 ml of water, and examine for clarity and sediment.

#### 4.3 Limit test for free alkali content

##### 4.3.1 Reagents

##### 4.3.1.1 Barium chloride, neutral solution.

Dissolve 243 g of barium chloride dihydrate ( $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ ) in 1 000 ml of carbon dioxide-free water. Check that it is neutral to the phenolphthalein indicator solution (4.3.1.4) and, if not, adjust with a few drops of the sodium hydroxide solution (4.3.1.3).

**4.3.1.2 Hydrochloric acid**, 0,1 M standard volumetric solution.

**4.3.1.3 Sodium hydroxide**, 0,1 M standard volumetric solution.

**4.3.1.4 Phenolphthalein** indicator, ethanol/water solution, 5 g/l.

Dissolve 5 g of phenolphthalein in 500 ml of ethanol and add 500 ml of water, with constant stirring. Filter if necessary.

##### 4.3.2 Apparatus

Ordinary laboratory apparatus and

**4.3.2.1 Pipette**, 25 ml capacity, conforming to class A of ISO/R 648.

**4.3.2.2 Burette**, 50 ml capacity, conforming to class A of ISO/R 385.

##### 4.3.3 Procedure

Weigh, to the nearest 0,1 g, a test portion of about 2 g of the laboratory sample and dissolve in 30 ml of freshly boiled water in a 125 ml stoppered conical flask. Add 25 ml of the neutral barium chloride solution (4.3.1.1) and

10 drops of the phenolphthalein indicator solution (4.3.1.4) and mix well. If the solution has a pink colour, titrate with the standard volumetric hydrochloric acid solution (4.3.1.2) until the colour is discharged. (If no pink colour is produced, proceed to 4.5, using the same solution.)

#### 4.4 Limit test for halides

##### 4.4.1 Reagents

**4.4.1.1 Nitric acid** solution, dilute (1 + 9).

**4.4.1.2 Silver nitrate** solution, 100 g/l.

**4.4.1.3 Chloride**, standard solution.

Dissolve 20 mg of potassium chloride in 1 000 ml of water.

##### 4.4.2 Apparatus

Ordinary laboratory apparatus and

**4.4.2.1 One-mark volumetric flask**, 500 ml capacity, conforming to class A of ISO 1042.

**4.4.2.2 Pipette**, 2 ml capacity, conforming to class A of ISO/R 648.

**4.4.2.3 Two matched Nessler cylinders**, 50 ml capacity.

##### 4.4.3 Procedure

Weigh, to the nearest 0,01 g, a test portion of about 1,3 g of the laboratory sample, dissolve in about 80 ml of water, filter, transfer the filtrate to the one-mark volumetric flask (4.4.2.1) and dilute to the mark. Using the pipette (4.4.2.2), transfer 2 ml aliquots of this solution and of the standard chloride solution (4.4.1.3) to separate Nessler cylinders (4.4.2.3). Then add 10 ml of the nitric acid solution (4.4.1.1) and 1 ml of the silver nitrate solution (4.4.1.2) to each. Dilute to 50 ml and mix well.

Compare, in the Nessler cylinders, the turbidities produced in the test and control solutions.

#### 4.5 Limit test for bicarbonate

##### 4.5.1 Reagents

As specified in 4.3.1.

##### 4.5.2 Apparatus

As specified in 4.3.2.

##### 4.5.3 Procedure

If the solution prepared under 4.3.3 is not pink, titrate with the standard volumetric sodium hydroxide solution (4.3.1.3) until a faint pink colour, which persists for 30 s, is obtained.

**4.6 Limit test for heavy metals**

**4.6.1 Reagents**

**4.6.1.1 Hydrochloric acid** solution,  $\rho$  approximately 1,18 g/ml.

**4.6.1.2 Hydrochloric acid** solution, dilute (1 + 99).

**4.6.1.3 Ammonia** solution, dilute (1 + 9).

**4.6.1.4 Heavy metals**, standard solution.

Dissolve a soluble lead salt in water to give a solution containing 10 mg of lead per 1 000 ml.

**4.6.1.5 Water, saturated at room temperature with hydrogen sulphide.**

**4.6.1.6 *p*-Nitrophenol** indicator, solution, 2,5 g/l.  $\rho$ , 2,5 g/l.

**4.6.2 Apparatus**

Ordinary laboratory apparatus and

**4.6.2.1 Two matched Nessler cylinders**, 50 ml capacity.

**4.6.3 Procedure**

Weigh, to the nearest 0,1 g, a test portion of about 5 g of the laboratory sample and dissolve in 15 ml of water. Also take 5 ml of the standard heavy metals solution (4.6.1.4) and treat this and the test solution in the following manner. Add 10 ml of the hydrochloric acid solution (4.6.1.1) and evaporate to dryness on a steam-bath. Take up the residues in 5 ml of the hydrochloric acid solution (4.6.1.2) and 25 ml of water. To each, add 1 drop of the *p*-nitrophenol indicator solution (4.6.1.6) followed by the ammonia solution (4.6.1.3), drop by drop, until the solutions turn yellow. Add the hydrochloric acid solution (4.6.1.2), drop by drop, until the solutions become colourless and then add 0,5 ml in excess. Dilute each to 50 ml with water.

Treat 20 ml aliquots of each solution separately in the Nessler cylinders (4.6.2.1), retaining the balance of the test solution for the iron test under 4.7.3. Add 5 ml of the hydrogen sulphide water (4.6.1.5) to each, dilute to 50 ml and mix well.

Compare, in the Nessler cylinders, the colours produced in the test and control solutions.

**4.7 Limit test for iron**

**4.7.1 Reagents**

As specified under 4.6.1 and

**4.7.1.1 Acetate** buffer solution, pH 5,0.

Dissolve 23 g of anhydrous sodium acetate in 58 ml of 2 M acetic acid and dilute to 1 000 ml with water. Adjust the final pH to  $5,0 \pm 0,1$  with glacial acetic acid or 100 g/l sodium hydroxide solution.

**4.7.1.2 Iron**, standard solution.

Dissolve a soluble iron(III) salt in water to give a solution containing 10 mg of iron(III) per 1 000 ml.

**4.7.1.3 1,10-Phenanthroline** reagent solution.

Thoroughly mix equal volumes of a 1 g/l aqueous solution of 1,10-phenanthroline, a 100 g/l aqueous solution of hydroxylammonium chloride and the acetate buffer solution (4.7.1.1).

**4.7.2 Apparatus**

Ordinary laboratory apparatus and

**4.7.2.1 Two matched Nessler cylinders**, 50 ml capacity.

**4.7.3 Procedure**

Take 10 ml of the standard iron solution (4.7.1.2) and treat in the same manner as the 5 ml of the standard heavy metals solution (4.6.1.4) under 4.6.3, as far as the dilution to 50 ml.

Transfer a 20 ml aliquot of this treated standard iron solution to one of the Nessler cylinders (4.7.2.1) and 20 ml of the balance of the treated test solution from 4.6.3 to the other Nessler cylinder. Add 5 ml of the 1,10-phenanthroline reagent solution (4.7.1.3) to each, mix and allow to stand for 10 min. Dilute each to 50 ml and mix well.

Compare, in the Nessler cylinders, the colours produced in the test and control solutions.

**4.8 Reaction to ammoniacal silver nitrate test**

**4.8.1 Reagent**

**4.8.1.1 Silver nitrate**, ammoniacal solution.

Immediately before use, mix equal volumes of ammonia solution,  $\rho$  approximately 0,910 g/ml, and 100 g/l silver nitrate solution.

**4.8.2 Apparatus**

Ordinary laboratory apparatus and

**4.8.2.1 Two matched Nessler cylinders**, 50 ml capacity.

**4.8.3 Procedure**

Weigh, to the nearest 0,1 g, a test portion of about 2 g of the laboratory sample and dissolve in 40 ml of water. Divide this solution equally between two Nessler cylinders (4.8.2.1). To one, the test solution, add 10 ml of the freshly prepared ammoniacal silver nitrate solution (4.8.1.1) and mix well. To the other, the control solution, add 5 ml of ammonia solution,  $\rho$  approximately 0,910 g/ml, and 5 ml of water and mix well. Allow each to stand for 2 min.

Compare, in the Nessler cylinders, the colours and turbidities of the test and control solutions.

**CAUTION — Dispose of all test solutions and rinse apparatus used immediately. Explosive compounds may be formed on standing.**

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