



Designation: D 4692 – 01

Standard Practice for Calculation and Adjustment of Sulfate Scaling Salts (CaSO_4 , SrSO_4 , and BaSO_4) for Reverse Osmosis and Nanofiltration¹

This standard is issued under the fixed designation D 4692; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This practice covers the calculation and adjustment of calcium, strontium, and barium sulfates for the concentrate stream of a reverse osmosis or nanofiltration system. The calculations are used to determine the need for scale control in the operation and design of reverse osmosis and nanofiltration installations. This practice is applicable for all types of reverse osmosis devices (tubular, spiral wound, and hollow fiber) and nanofiltration devices.

1.2 This practice is applicable to both brackish waters and seawaters.

2. Referenced Documents

2.1 ASTM Standards:

D 511 Test Methods for Calcium and Magnesium in Water²

D 516 Test Method for Sulfate Ion in Water²

D 1129 Terminology Relating to Water²

D 3352 Test Method for Strontium Ion in Brackish Water, Seawater, and Brines³

D 4194 Test Methods for Operating Characteristics of Reverse Osmosis Devices³

D 4195 Guide for Water Analysis for Reverse Osmosis Application³

D 4382 Test Method for Barium in Water, Atomic Absorption Spectrophotometry, Graphite Furnace²

3. Terminology

3.1 *Definitions*—For definitions of terms used in this practice, refer to Terminology D 1129.

¹ This practice is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.08 on Membranes and Ion Exchange Materials.

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² Annual Book of ASTM Standards, Vol 11.01.

³ Annual Book of ASTM Standards, Vol 11.02.

3.2 Definitions of Terms Specific to This Standard:

3.2.1 For definitions of terms relating to reverse osmosis, refer to Test Methods D 4194.

4. Summary of Practice

4.1 This practice consists of calculating the potential for scaling by CaSO_4 , SrSO_4 , and BaSO_4 in a reverse osmosis or nanofiltration concentrate stream from the concentration of Ca^{+} , Sr^{++} , Ba^{++} , and SO_4^{-} in the feed solution and the recovery of the reverse osmosis or nanofiltration system.

4.2 This practice also presents techniques to eliminate scaling by decreasing the recovery, by decreasing the Ca^{+} , Sr^{++} , and Ba^{++} concentrations in the feed water, and by addition of scale inhibitors.

5. Significance and Use

5.1 In the design and operation of reverse osmosis and nanofiltration installations, it is important to predict the CaSO_4 , SrSO_4 , and BaSO_4 scaling properties of the concentrate stream. Because of the increase in total dissolved solids and the increase in concentration of the scaling salts, the scaling properties of the concentrate stream will be quite different from those of the feed solution. This practice permits the calculation of the scaling potential for the concentrate stream from the feed water analyses and the reverse osmosis or nanofiltration operating parameters.

5.2 Scaling by CaSO_4 , SrSO_4 , and BaSO_4 will adversely affect the reverse osmosis or nanofiltration performance. This practice gives various procedures for the prevention of scaling.

6. Procedure

6.1 Determine the concentrations of Ca^{+} , Sr^{++} , Ba^{++} , and SO_4^{-} in the feed stream in accordance with Test Methods D 511, D 3352, D 4382, and D 516, respectively.

NOTE 1—If H_2SO_4 is used for control of CaCO_3 scale, measure the SO_4^{-} after acid addition.

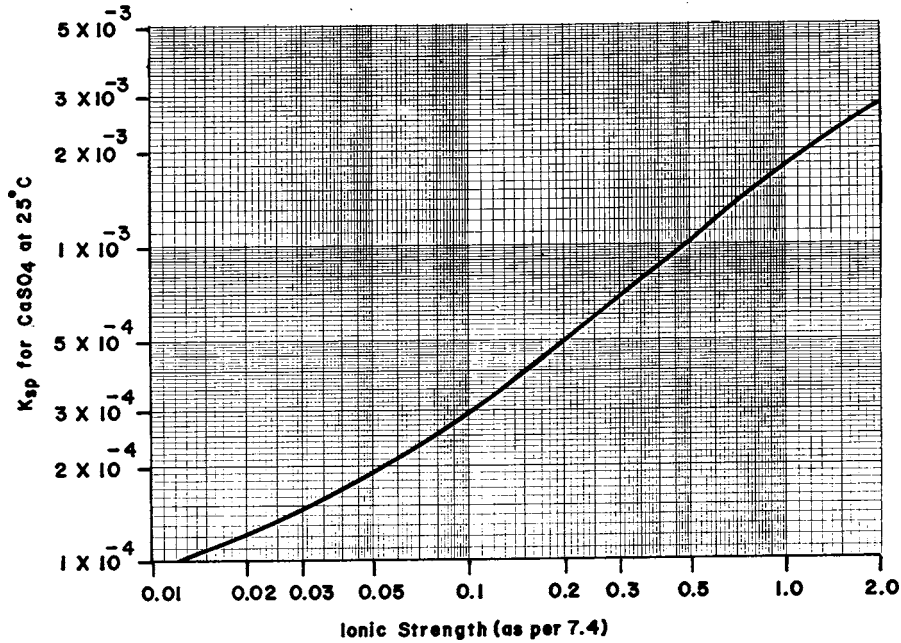


FIG. 1 K_{sp} for CaSO_4 versus Ionic Strength

6.2 Determine the concentration of all major ions using the appropriate methods given in Guide D 4195. At a minimum, the concentrations of Mg^{++} , Na^+ , HCO_3^- , and Cl^- must be determined.

7. Calculation

7.1 Calculate the calcium concentration in the concentrate stream from the calcium concentration in the feed solution, from the recovery of the reverse osmosis or nanofiltration system, and from the calcium ion passage as follows:

$$Ca_c = Ca_f \times \frac{1 - Y(SP_{Ca})}{1 - Y}$$

where:

- Ca_c = calcium ion concentration in concentrate, mg/L,
- Ca_f = calcium ion concentration in feed, mg/L,
- Y = recovery of the reverse osmosis system, expressed as a decimal, and
- SP_{Ca} = calcium ion passage, expressed as a decimal.

NOTE 2— SP_{Ca} can be obtained from the supplier of the reverse osmosis or nanofiltration system. For most reverse osmosis and nanofiltration devices, SP_{Ca} can be considered to be zero, in which case the equation simplifies to:

$$Ca_c = Ca_f \times \frac{1}{1 - Y}$$

This assumption will introduce only a small error.

7.2 Calculate the SO_4^{--} concentration in the concentrate stream from the SO_4^{--} concentration in the feed solution, from the recovery of the reverse osmosis or nanofiltration system, and from the sulfate ion passage by using the appropriate substitutions in the equation given in 7.1. The simplified equation can be used.

7.3 Calculate the concentration of the major ions in the concentrate stream using the appropriate substitutions in the equation given in 7.1. The simplified equation can be used.

7.4 Calculate the ionic strength of the concentrate stream as follows:

$$I_c = \frac{1}{2} \sum \bar{m}_i Z_i^2$$

where:

- I_c = ionic strength of concentrate stream,
- \bar{m}_i = molal concentration of ion, i (moles/1000 g of water) in the concentrate stream, and
- Z_i = ionic charge of ion, i .

NOTE 3—The molal concentration is calculated as follows:

$$m_i = \frac{C_i}{1000 MW_i \left[\frac{10^6 - TDS}{10^6} \right]} = \frac{1000 C_i}{MW_i (10^6 - TDS)}$$

where:

- C_i = concentration of ion, i , in concentrate stream, mg/L,
- MW_i = molecular weight of ion, i , and
- TDS = total dissolved solids in concentrate stream, mg/L.

7.5 Calculate the ion product (IP_c) for CaSO_4 in the concentrate stream as follows:

$$IP_c = ({}^m\text{Ca}^{++})_c ({}^m\text{SO}_4^{--})_c$$

where:

- $({}^m\text{Ca}^{++})_c$ = $M \text{Ca}^{++}$ in concentrate, mol/L and
- $({}^m\text{SO}_4^{--})_c$ = $M \text{SO}_4^{--}$ in concentrate, mol/L.

7.6 Compare IP_c for CaSO_4 with the solubility product (K_{sp}) of CaSO_4 at the ionic strength of the concentrate stream (Fig. 1).⁴ If $IP_c > K_{sp}$, CaSO_4 scaling will occur and adjustment is required.

⁴ Marshall, W. L. and Slusher, R., "Solubility to 200°C of Sulfate and its Hydrates in Sea Water and Saline Water Concentrates and Temperature, Concentration Limits," *Journal of Chemical and Engineering Data*, Vol 13, No. 1, 1968, p. 83.