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**Fine ceramics (advanced ceramics,  
advanced technical ceramics) —  
Methods for chemical analysis of  
boron nitride powders**

*Céramiques techniques — Méthodes pour l'analyse chimique de  
poudres de nitrure de bore*

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## Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

The procedures used to develop this document and those intended for its further maintenance are described in the ISO/IEC Directives, Part 1. In particular the different approval criteria needed for the different types of ISO documents should be noted. This document was drafted in accordance with the editorial rules of the ISO/IEC Directives, Part 2 (see [www.iso.org/directives](http://www.iso.org/directives)).

Attention is drawn to the possibility that some of the elements of this document may be the subject of patent rights. ISO shall not be held responsible for identifying any or all such patent rights. Details of any patent rights identified during the development of the document will be in the Introduction and/or on the ISO list of patent declarations received (see [www.iso.org/patents](http://www.iso.org/patents)).

Any trade name used in this document is information given for the convenience of users and does not constitute an endorsement.

For an explanation on the meaning of ISO specific terms and expressions related to conformity assessment, as well as information about ISO's adherence to the WTO principles in the Technical Barriers to Trade (TBT) see the following URL: Foreword - Supplementary information.

The committee responsible for this document is ISO/TC 206, *Fine ceramics*.

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## Introduction

This International Standard describes the chemical analysis of boron nitride powders for use in fine ceramics applications. Determination methods of both main and trace constituents of interest are fully covered.

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# Fine ceramics (advanced ceramics, advanced technical ceramics) — Methods for chemical analysis of boron nitride powders

## 1 Scope

This International Standard specifies the methods for the chemical analysis of fine boron nitride powders used as the raw material for fine ceramics.

This International Standard stipulates the analysis methods of total boron, free boron, total nitrogen, aluminium, cadmium, calcium, chromium, cobalt, copper, iron, lead, magnesium, manganese, nickel, potassium, silicon, sodium, titanium, vanadium, zinc, carbon, and oxygen in boron nitride powders for fine ceramics. Total boron is determined by using either a fusion–titration method or a fusion–inductively coupled plasma–optical emission spectrometry (ICP–OES). Free boron is determined by using either an acid digestion–ICP–OES or a methanol extraction–ICP–OES. If necessary, the boron amount which arises from the hydrolysis of boron nitride during sample treatment is corrected using spectrophotometry. Total nitrogen is determined by using either an acid pressure decomposition–distillation separation–titration method or an inert gas fusion–thermal conductivity method. Silicon is determined by using a fusion–ICP–OES. Aluminium, cadmium, calcium, chromium, cobalt, copper, iron, lead, magnesium, manganese, nickel, titanium, vanadium, and zinc are determined by using an acid pressure decomposition–ICP–OES or a fusion–ICP–OES. Sodium and potassium are determined by using atomic absorption spectrometry (AAS), flame emission spectrometry (FES), or ICP–OES following acid pressure decomposition. Carbon is determined by using a combustion–IR absorption spectrometry or a combustion–thermal conductivity method. Oxygen is determined by using an inert gas fusion–IR absorption spectrometry.

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## 2 Normative references

The following documents, in whole or in part, are normatively referenced in this document and are indispensable for its application. For dated references, only the edition cited applies. For undated references, the latest edition of the referenced document (including any amendments) applies.

EN 12698-1:2007, *Chemical analysis of nitride bonded silicon carbide refractories - Part 1: Chemical methods*

ISO 3696, *Water for analytical laboratory use — Specification and test methods*

ISO 6353-3, *Reagents for chemical analysis — Part 3: Specifications — Second series*

ISO 9138, *Abrasive grains — Sampling and splitting*

ISO 21068-2:2008, *Chemical analysis of silicon-carbide-containing raw materials and refractory products — Part 2: Determination of loss on ignition, total carbon, free carbon and silicon carbide, total and free silica and total and free silicon*

## 3 General rules in chemical analysis

Chemical analysis shall be performed in accordance with general rules regulated as standards, if available.

## 4 Analysis items

Analysis items specified in this International Standard shall be as follows:

- a) Total boron (T.B)

- b) Free boron (F.B)
- c) Total nitrogen (T.N)
- d) Silicon (Si)
- e) Aluminium (Al)
- f) Cadmium (Cd)
- g) Calcium (Ca)
- h) Chromium (Cr)
- i) Cobalt (Co)
- j) Copper (Cu)
- k) Iron (Fe)
- l) Lead (Pb)
- m) Magnesium (Mg)
- n) Manganese (Mn)
- o) Nickel (Ni)
- p) Titanium (Ti)
- q) Vanadium (V)
- r) Zinc (Zn)
- s) Sodium (Na)
- t) Potassium (K)
- u) Carbon (C)
- v) Oxygen (O)

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## **5 Sampling, weighing, and drying of sample**

The method of preparing samples shall be in accordance with ISO 9138 unless otherwise mutually agreed upon between the analyser and customer.

### **5.1 Sampling**

Take the sample in accordance with ISO 9138.

### **5.2 Weighing**

Weigh the sample of the required quantity to the nearest 0,1 mg using an electric balance.

### **5.3 Drying**

Take about 10 g of the sample into a flat-type weighing bottle (60 mm × 30 mm) and spread it uniformly over the bottom of the bottle. Place the bottle in an air bath at (120 ± 5) °C for 2 h without a lid, and then cool in a desiccator (desiccant: magnesium perchlorate for drying) with a lid for 1 h.



## 6 Reporting analytical values

### 6.1 Number of analyses

Analyse the sample twice on different days.

### 6.2 Blank test

Upon analysis, perform a blank test to correct the measured values.

### 6.3 Evaluation of analytical values

- a) When the difference between two analytical values does not exceed the tolerance shown in [Table 1](#), the average value shall be reported.
- b) When the difference between two analytical values exceeds the tolerance shown in [Table 1](#), perform additional two analyses. When the difference does not exceed the tolerance, the average value thereof shall be reported. If the difference also exceeds the tolerance, the median of four analytical values shall be reported.

### 6.4 Expression of analytical values

The analytical values shall be given in % (mass fraction) in dryness.

- a) **Total boron and total nitrogen** — express the results with two to four figures according to the method used.
- b) **Free boron** — express the results with two to four figures according to the method used.
- c) **Others** — express the results to the third decimal place.

**Table 1 — Tolerance on analytical values**

Unit: % (mass fraction)

Component	total-B	free-B	total-N	Si, Al, Cd, Ca, Cr, Co, Cu, Fe, Pb, Mg, Mn, Ni, Ti, V, Zn, K, Na	C	O
Tolerance	0,30		0,30 <sup>a</sup> 0,40 <sup>b</sup>	0,001 <sup>c</sup> 0,005 <sup>d</sup>	0.005	0,01 <sup>e</sup> 0,05 <sup>f</sup>
<sup>a</sup> Acid pressure decomposition–distillation separation–acidimetric titration method. <sup>b</sup> Inert gas fusion–thermal conductivity method. <sup>c</sup> Applicable to the content of less than 0,01 % (mass fraction). <sup>d</sup> Applicable to the content of not less than 0,01 % (mass fraction). <sup>e</sup> Applicable to the content of less than 1,0 % (mass fraction). <sup>f</sup> Applicable to the content of not less than 1,0 % (mass fraction).						

## 7 Determination methods of total boron

### 7.1 Classification of determination methods

Total boron shall be determined by either of the following methods. If analytical results with three or four figures are required, use method A; whereas if two figures are required, method B can be used.

- Method A, fusion–acidimetric titration method

— Method B, fusion-ICP-OES

## 7.2 Fusion-acidimetric titration method

### 7.2.1 Principle

Fuse the sample with sodium carbonate and dissolve the melt in hydrochloric acid. Adjust the solution to pH 6,4 after removal of carbon dioxide. Add mannitol, and then titrate with sodium hydroxide solution to pH 6,4.

This method shall be used to obtain the results of three or four figures.

### 7.2.2 Reagents

Reagents of analytical grade shall be used. Reagent solutions shall be preserved in plastic bottles (products made of polyethylene, polypropylene, and tetrafluoroethylene resin are available.). The standard solutions being SI traceable shall be used.

**7.2.2.1 Water**, Grade 1 or superior as specified in ISO 3696.

**7.2.2.2 Sodium carbonate, anhydrous**, as specified in ISO 6353-3 or that of higher grade.

**7.2.2.3 Hydrochloric acid (1+1), (1+50)**, prepare from hydrochloric acid, respectively.

**7.2.2.4 Sodium hydroxide**, more than 97,0 % (mass fraction) of purity.

**7.2.2.5 Sodium hydroxide solution (200 g/l), (40 g/l), (20 g/l)**, prepare from sodium hydroxide, respectively.

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**7.2.2.6 Ethanol (95)**.

**7.2.2.7 Ethanol-Sodium hydroxide (1+1) mixed solution**, prepare by mixing the same volume of ethanol (95) and a sodium hydroxide solution (200 g/l).

**7.2.2.8 Mannitol**.

**7.2.2.9 Boric acid**, more than 99,5 % (mass fraction) of purity.

**7.2.2.10 Standard boron solution**, weigh 1 g of boric acid (7.2.2.9) in a platinum crucible (for example, 30 ml) and heat it initially at a low temperature and finally at about 1 100 °C. After cooling in a desiccator, weigh the mass of boric oxide as B<sub>2</sub>O<sub>3</sub>. Take the product into a 200 ml beaker together with the crucible and add 100 ml of hot water. Heat and dissolve it. After cooling, transfer the solution into a 200 ml volumetric flask and dilute with water to the mark.

**7.2.2.11 Methyl-red indicator solution**, prepare by dissolving 0,10 g of methyl-red with 50 ml of ethanol (95) (7.2.2.6) to make up to 100 ml by adding water.

**7.2.2.12 0.2 mol/ l Sodium hydroxide solution**, weigh 50 g of sodium hydroxide in a polyethylene bottle and dissolve in 50 ml of water under cooling. Stand for a few days with a cap to precipitate sodium carbonate. Take 20 ml of the supernatant and dilute with water to 2 000 ml. Preserve it into a polyethylene bottle equipped with an absorption tube of carbon dioxide. The volume of equivalent to the boron shall be obtained according to the following. Take exactly 50 ml of the standard boron solution (7.2.2.10) into a 200 ml beaker and add 12 ml of hydrochloric acid (1+1) (7.2.2.3) and a few drops of methyl-red indicator

solution (7.2.2.11). After neutralization with the sodium hydroxide solution (40 g/l) (7.2.2.5), operate in accordance with 7.2.5 g). Calculate the factor of this solution according to the following formula:

$$F = \left( G \times \frac{0,1748}{V_0} \right) \times \frac{50}{200} \quad (1)$$

where

$F$  is the factor of boron content of 0,2 mol/l sodium hydroxide solution per 1 ml (g/ml);

$G$  is the weighed amount of boric acid in 7.2.2.9 (g);

$V_0$  is the consumed volume of 0,2 mol/l sodium hydroxide solution after addition of mannitol in 7.2.2.12 (ml).

### 7.2.3 Apparatus

Apparatus shall be as follows:

7.2.3.1 **Platinum crucible** (30 ml), with a platinum lid.

7.2.3.2 **Electric furnace**, adjustable to the temperature of  $(1\ 100 \pm 25)$  °C.

7.2.3.3 **pH meter**, readable to the smallest value of 0,1 equipped with a glass electrode.

### 7.2.4 Amount of sample

The weighed amount of the sample shall be 0,10 g.  
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### 7.2.5 Operation

The operation shall be as follows.

- a) Take the sample aliquot into a platinum crucible (30 ml) and 2 ml of the ethanol–sodium hydroxide mixed solution (7.2.2.7).
- b) Heat on a hot plate, and then evaporate until dry.
- c) Add 2,0 g of sodium carbonate and heat initially at a low temperature. Elevate the temperature gradually to start fusing (heat carefully to prevent the contents from scattering) by using an electric furnace or a Bunsen burner and keep the temperature for 15 min ~ 30 min during fusion. Fuse the sample at 950 °C ~ 1 000 °C. Keep the temperature as low as possible.
- d) Stand the crucible for cooling with a lid and place the crucible into a 300 ml beaker. Add 50 ml of water and 10 ml of hydrochloric acid (1+1) (7.2.2.3) and heat it on a hot plate with a watch glass to dissolve the melt.
- e) After dissolving, wash the crucible with small amount of water and add a few drops of the methyl-red indicator solution (7.2.2.11). Neutralize the solution with a sodium hydroxide solution until its colour turns yellow, and then add hydrochloric acid (1+50) (7.2.2.3) to return to red again.
- f) Remove carbon dioxide in the solution by boiling or by purging nitrogen gas for 5 min with the watch glass. Cool down to room temperature and wash the watch glass with small amount of water before removing. Keep the total water volume not more than 150 ml while washing.
- g) Stirring with a magnetic stirrer, titrate with 0,2 mol/l sodium hydroxide solution (7.2.2.12), depict a titration curve around pH 6,4 using a pH meter, and determine the inflection point of the curve. After removing the electrode, add 25 g of mannitol (7.2.2.8). Titrate again with the sodium hydroxide

solution (7.2.2.12), depict a titration curve around pH 8,5, and determine the inflection point of the curve. Calculate the total volume of titrant between two inflection points.

### 7.2.6 Blank test

Perform operations of 7.2.5 a) to 7.2.5 g) without taking a sample to obtain the blank test value of boron.

### 7.2.7 Calculation

Calculate the content of total boron in the sample according to the following formula:

$$T.B = \frac{(V_1 - V_0) \times F}{m} \times 100 \quad (2)$$

where

$T.B$  is the content of total boron in the sample [% (mass fraction)];

$V_1$  is the total volume of the 0,2 mol/l sodium hydroxide solution after the first addition of mannitol in 7.2.5 g) (ml);

$V_0$  is the total volume of the 0,2 mol/l sodium hydroxide solution of the blank test after the first addition of mannitol in 7.2.6 (ml);

$F$  is the factor of the 0,2 mol/l sodium hydroxide solution in 7.2.2.12 (g);

$m$  is the weighed amount of the sample in 7.2.5 a) (g).

## 7.3 Fusion-ICP-OES

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### 7.3.1 Principle

To prepare a test solution, decompose the sample using base, spray into the argon plasma flame of ICP-OES, and measure the emission intensity at a selected wavelength.

This method can be used to obtain the results of two figures.

### 7.3.2 Reagents

Use only reagents of analytical grade unless stated to the contrary and prepare all solutions using distilled or deionized water and store in polyethylene bottles.

**WARNING — Hydrochloric acid causes burns and has an irritating vapour. Prevent inhalation and contact with skin and eyes. Proceed as for sulfuric acid, carrying out the dilution in a fume cupboard.**

**7.3.2.1 Hydrochloric acid**, concentrated, density 1,095 g/ml.

**7.3.2.2 Hydrochloric acid (1+1 solution)**, carefully add 50 ml of acid to 50 ml of water to a beaker cooled in a water bath. Stir the mixture continuously. Allow to cool and store.

**7.3.2.3 Nitric acid**,  $\rho$  1,42 g/ml.

**7.3.2.4 Sodium peroxide** ( $\text{Na}_2\text{O}_2$ ).

**7.3.2.5 Sodium carbonate**, anhydrous.

### 7.3.2.6 Ethanol.

**7.3.2.7 Boron standard solution**, containing 1,000 g of boron (*B*) per litre. Do not dry. Dissolve 5,716 g of anhydrous  $H_3BO_3$  (B fraction = 0,174 9), accurately weighed to at least four significant figures, in reagent water and dilute in a 1 l volumetric flask with reagent water. Transfer immediately after mixing in a clean polytetrafluoroethylene (PTFE) bottle to minimize any leaching of boron from the glass container. The use of a non-glass volumetric flask is recommended to avoid boron contamination from glassware. A 1 ml of this standard solution contains 1,0 mg of boron.

### 7.3.3 Apparatus

Apparatus are ordinary laboratory apparatus and the following:

**7.3.3.1 Glass beaker**, with a range of appropriate volumes (100 ml, 250 ml).

**7.3.3.2 Volumetric flask**, with a range of appropriate volumes (100 ml, 250 ml, 500 ml, and 1 000 ml).

**7.3.3.3 Burette**, with a 0,1 ml scale and a maximum volume of 50 ml.

**7.3.3.4 One-mark pipette**, suitable for the transfer of each sample solution or standard solution.

**7.3.3.5 Nickel crucibles.**

**7.3.3.6 Burner**, such as a Mecker burner or a Bunsen burner.

**7.3.3.7 Desiccator**, containing dried silica gel as the drying agent.

**7.3.3.8 Inductively coupled plasma emission spectrometer.**

NOTE Refer to the manufacturer's instructions for operation.

### 7.3.4 Procedure

#### 7.3.4.1 Preparation of test solution

- a) Weigh, to the nearest 0,001 g, approximately 1,0 g of the test sample, and transfer the test sample in a nickel crucible.
- b) Add 2,0 g of anhydrous sodium carbonate and 2,0 g of sodium peroxide and mix intimately. Heat over a burner slowly at first and gradually raise the temperature to the full heat until complete fusion.

**CAUTION — When heating the sample in the burner, carbon dioxide is produced. So be careful when heating the sample.**

- c) Cool to ambient temperature and place in a 250 ml beaker. Extract with approximately 150 ml of water containing 2 ml to 3 ml ethanol, heat to boiling until the melt from the walls of the crucible has been separated. Remove the crucible and wash it with a small amount of water.
- d) Add 10 ml of hydrochloric acid ([7.3.2.1](#)) and heat to 100 °C until the melt is completely dissolved. Pour the dissolved sample into a 250 ml volumetric flask, make up to the mark, and mix.
- e) Pipette 25 ml of the solution from the 250 ml volumetric flask into a new 250 ml volumetric flask, make up to the mark, and mix. This solution is used as the test solution.

### 7.3.4.2 Blank test

Carry out the procedure given in [7.3.4.1](#) without the sample, but omit the fusion of the flux. Designate the resulting solution as blank solution.

### 7.3.4.3 Preparation of the calibration curve

- To a series of six 100 ml volumetric flasks, add respectively 0 ml, 1 ml, 5 ml, 10 ml, 20 ml, and 50 ml boron standard solution ([7.3.2.7](#)). Add to each flask 5 ml hydrochloric acid 1+1, dilute to mark, and mix.
- Spray a portion of each solution into the argon plasma flame of the inductively coupled plasma atomic emission spectrometer, and measure the emission intensity for boron at 249,677 nm, 249,772 nm, or 208,809 nm. Interferences may be encountered. Carefully choose the optimum wavelength free from concomitants.

### 7.3.5 Calculation

Determine the concentration of boron in the test solution and in the blank from the calibration curve. Calculate the boron content,  $B$ , as total boron, expressed in percent mass fraction, from Formula (3):

$$T.B = \frac{(C_1 - C_2) \times V}{m} \times 10^4 \quad (3)$$

where

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$T.B$  is the total boron content in mass fraction (%);

$C_1$  is the concentration, expressed in milligrams per litre, of boron in test solution;

$C_2$  is the concentration, expressed in milligrams per litre, of boron in blank test;

$V$  is the volume, in millilitres, of the test solution described in [7.3.4.1 d](#));

$m$  is the mass, in grams, of the test portion.

## 8 Determination methods of free boron

### 8.1 Classification of determination methods

Boron nitride may somewhat tend to receive hydrolysis and yield equivalent of boron oxide and ammonia; only the latter evaporates. The determination of boron oxide in boron nitride may be critical when the equivalence between total boron and total nitrogen is evaluated. There are two approaches to extract boron oxide in boron nitride with acid or methanol, but either of them is accompanied by the slight dissolution of boron nitride. Extraction of only boron oxide without dissolving boron nitride is difficult. However, the contribution of boron generated from boron nitride can be evaluated by the determination of ammonia whose amount is equivalent to the boron.

Boron oxide shall be determined by any of the following methods. Both method A and method B include the correction procedure for the boron amount which arises from the hydrolysis of boron nitride during sample treatment stages. If analytical results with three or four figures are required for the total boron, method A or C shall be used; whereas if two figures, either method A or C, without the correction procedure or method B, can be used.

- Method A, Sulfuric acid digestion–ICP-OES and spectrophotometry
- Method B, Nitric acid digestion–ICP-OES

— Method C, Methanol extraction–borate separation–ICP-OES and spectrophotometry

## 8.2 Sulfuric acid digestion-ICP-OES and spectrophotometry

### 8.2.1 Principle

Dissolve the sample in sulfuric acid and filtrate the insoluble. Determine boron by using ICP-OES. If analytical results of three or four figures for the total boron are determined, the correction procedure including ammonia determination of [8.2.5 a\)](#) and [8.2.5 c\)](#) to [8.2.5 f\)](#) shall be necessary.

### 8.2.2 Reagents

Reagents shall be as follows.

**8.2.2.1 Water**, as specified in [7.2.2.1](#).

**8.2.2.2 Sulfuric acid**, 0,5 mol/l, prepare with sulfuric acid.

**8.2.2.3 EDTA solution (5 g/100 ml)**, prepare with EDTA.

**8.2.2.4 Sodium hydroxide solution (20 g/100 ml)**, prepare with sodium hydroxide.

**8.2.2.5 Phenol.** iTeh STANDARD PREVIEW

**8.2.2.6 Acetone.** (standards.iteh.ai)

**8.2.2.7 Sodium phenoxide solution**, take 55 ml of a sodium hydroxide solution (20 g/100 ml) ([8.2.2.4](#)) in a beaker and gradually add 25 g of phenol ([8.2.2.5](#)) to the solution cooling in cold water. After cooling, add 6 ml of acetone ([8.2.2.6](#)) and dilute with water to 200 ml. Store in a dark room below 10 °C.

**8.2.2.8 Sodium hypochlorite solution**, [1 % (mass fraction) of available chlorine]. Prepare by diluting a commercial sodium hypochlorite solution, 5 % ~ 12 % (mass fraction) of available chlorine content, with water. Prepare fresh immediately before use.

**8.2.2.9 Ammonium ion standard solution** (NH<sub>4</sub><sup>+</sup> 1 mg/ml).

NOTE A commercial standard solution being SI traceable is available.

**8.2.2.10 Ammonium ion standard solution** (NH<sub>4</sub><sup>+</sup> 0.1 mg/ml), take 10 ml of an ammonium ion standard solution (NH<sub>4</sub><sup>+</sup> 1 mg/ml) ([8.2.2.9](#)) into a 100 ml plastic volumetric flask and dilute with water to the mark.

NOTE Products made of polyethylene, polypropylene, and tetrafluoroethylene resin are available for the plastic volumetric flask.

**8.2.2.11 Boron standard solution (B 1 mg/ml).**

NOTE A commercial standard solution being SI traceable is available.

**8.2.2.12 Boron standard solution (B 0,1 mg/ml)**, take 10 ml of a boron standard solution (1 mg/ml) ([8.2.2.11](#)) into a 100 ml plastic volumetric flask, and dilute with water to the mark. Prepare fresh immediately before use.

NOTE Products made of polyethylene, polypropylene, and tetrafluoroethylene resin are available for the plastic volumetric flask.