



Designation: E1277 – 08

Standard Test Method for Chemical Analysis of Zinc-5 % Aluminum-Mischmetal Alloys by ICP Emission Spectrometry¹

This standard is issued under the fixed designation E1277; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 This test method covers the chemical analysis of zinc alloys having chemical compositions within the following limits:

Element	Concentration Range, %
Aluminum	3.0–8.0
Antimony	0.002 max
Cadmium	0.025 max
Cerium	0.03–0.10
Copper	0.10 max
Iron	0.10 max
Lanthanum	0.03–0.10
Lead	0.026 max
Magnesium	0.05 max
Silicon	0.015 max
Tin	0.002 max
Titanium	0.02 max
Zirconium	0.02 max

1.2 Included are procedures for elements in the following concentration ranges:

Element	Concentration Range, %
Aluminum	3.0–8.0
Cadmium	0.0016–0.025
Cerium	0.005–0.10
Iron	0.0015–0.10
Lanthanum	0.009–0.10
Lead	0.002–0.026

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use. Specific safety hazards statements are given in Section 8, 12.2, and 14.1.*

2. Referenced Documents

2.1 *ASTM Standards:*²

[D1193 Specification for Reagent Water](#)

¹ This test method is under the jurisdiction of ASTM Committee E01 on Analytical Chemistry for Metals, Ores, and Related Materials and is the direct responsibility of Subcommittee E01.05 on Cu, Pb, Zn, Cd, Sn, Be, their Alloys, and Related Metals.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

[E29 Practice for Using Significant Digits in Test Data to Determine Conformance with Specifications](#)

[E50 Practices for Apparatus, Reagents, and Safety Considerations for Chemical Analysis of Metals, Ores, and Related Materials](#)

[E55 Practice for Sampling Wrought Nonferrous Metals and Alloys for Determination of Chemical Composition](#)

[E88 Practice for Sampling Nonferrous Metals and Alloys in Cast Form for Determination of Chemical Composition](#)

[E135 Terminology Relating to Analytical Chemistry for Metals, Ores, and Related Materials](#)

[E173 Practice for Conducting Interlaboratory Studies of Methods for Chemical Analysis of Metals \(Withdrawn 1998\)](#)³

[E876 Practice for Use of Statistics in the Evaluation of Spectrometric Data \(Withdrawn 2003\)](#)³

[E1601 Practice for Conducting an Interlaboratory Study to Evaluate the Performance of an Analytical Method](#)

2.2 *NIST Standard Reference Materials:*⁴

[SRM 728 Zinc, Intermediate Purity](#)

3. Terminology

3.1 For definitions of terms used in this test method, refer to Terminology [E135](#).

4. Summary of Test Method

4.1 The sample is dissolved in mixed acids. The sample solution is introduced into the plasma source of an ICP spectrometer and the intensities at selected wavelengths from the plasma emission spectrum are compared to the intensities at the same wavelengths measured with calibration solutions.

5. Significance and Use

5.1 This test method for the chemical analysis of metals and alloys is primarily intended to test such materials for compliance with compositional specifications. It is assumed that all those who use this test method will be trained analysts capable

³ The last approved version of this historical standard is referenced on www.astm.org.

⁴ Available from National Institute of Standards and Technology (NIST), 100 Bureau Dr., Stop 1070, Gaithersburg, MD 20899-1070, http://www.nist.gov.

of performing common laboratory procedures skillfully and safely. It is expected that work will be performed in a properly equipped laboratory.

6. Apparatus

6.1 Inductively-Coupled Argon Plasma (ICP) Atomic Emission Spectrometer—The instrument is equipped with an argon-plasma source and a sample transport system for introducing the test sample and calibration solutions into the plasma. The monochromator or polychromator must be capable of isolating the required wavelengths shown in **Table 1** for measurement of their intensities by a linear photometer. Multielement programmed analysis including automatic data acquisition and computer-controlled calibration and determinations may be used if available, provided that, in addition to calculated results, the instrument records intensity readings each time a test sample or calibration solution is presented to the instrument.

NOTE 1—All elements (including aluminum) are calibrated as linear functions of intensity. If the instrument cannot be set to measure aluminum and ignore other elements in calibration solutions No. 1 and No. 4, then a separate determination of aluminum must be made using calibration solutions No. 1, No. 2, and No. 4. The other elements can then be determined together in another run using only calibration solutions No. 2 and No. 3. Use the calibration solutions prepared in **11.1** in determining the instrument settings for the elements in this matrix. Follow the manufacturer's instructions to set the wavelengths and parameters to provide as large a difference between the intensity readings for the high and low calibration concentrations as is consistent with stable instrument readings. If there is a question of linearity of the instrument's response over the range of solution concentrations given, a third standard, equidistant between the two listed standards, must be measured to verify linearity.

7. Reagents

7.1 Purity of Reagents—Unless otherwise indicated, all reagents used in this test method shall conform to the "Reagent Grade" Specifications of the American Chemical Society.⁵ Other chemicals may be used provided that it is first ascertained that the reagent used is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

⁵ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

TABLE 1 Wavelengths and Instrument Conditions^A

Element	Wavelength, nm	Time, s	No. Integ.	BCor1	BCor2
Aluminum	309.27	1.0	3
Cadmium	226.502	.5	3	226.446	226.558
Cerium	418.66	.5	2
Iron	259.94	.5	2
Lanthanum	398.85	.5	2	398.754	398.906
Lead	283.297	1.0	3	...	283.336

^A The tabulated conditions were those found satisfactory on one instrument. Wavelengths are expressed in nanometres (nm). Time = seconds for each integration, No. Integ. = number of integrations averaged for each reading, and BCor1 and BCor2 are off-peak background correction wavelengths.

7.2 Purity of Water—Unless otherwise indicated, references to water shall be understood to mean reagent water as defined by Type II of Specification **D1193**.

7.3 Aluminum, Standard Solution (1 mL = 20.0 mg Al)—Transfer 2.0000 g of aluminum (purity: 99.999 % min) to a 250-mL beaker. Cover, add 50 mL of HCl (1 + 1) and a small crystal of mercuric nitrate. Heat gently to accelerate the reaction, but avoid temperatures high enough to cause a noticeable volume loss. If the reaction slows, add more mercuric salt as needed. A number of hours may be required to complete the dissolution (only a small droplet of mercury will remain undissolved). Transfer the solution to a 100-mL volumetric flask, dilute to volume, and mix. Store in a polyethylene bottle.

7.4 Cadmium, Standard Solution (1 mL = 1.00 mg Cd)—Transfer 1.000 g of cadmium (purity: 99.95 % min) to a 250-mL beaker. Cover and add 40 mL of HNO₃ (1 + 1) and 10 mL of HCl. After dissolution is complete, heat to boiling to remove oxides of nitrogen. Cool, transfer to a 1-L volumetric flask, add 240 mL of HCl, dilute to volume, and mix. Store in a polyethylene bottle.

7.5 Cerium, Standard Solution A (1 mL = 1.00 mg Ce)—Dry ceric ammonium nitrate ((NH₄)₂Ce(NO₃)₆, also known as ammonium hexanitrate cerate) (purity: 99.95 % min) for 4 h at 85 °C and cool to room temperature in a desiccator. Dissolve 3.913 g of dry ceric ammonium nitrate in 100 mL of HCl (1 + 9). Transfer to a 1-L volumetric flask, add 240 mL of HCl and 20 mL of HNO₃, dilute to volume, and mix. Store in a polyethylene bottle.

7.6 Cerium, Standard Solution B (1 mL = 0.010 mg Ce)—Using a pipet, transfer 1.00 mL of Cerium Standard Solution A to a 100-mL volumetric flask. Dilute to volume with dilution solution and mix.

7.7 Dilution Solution—Half fill a 2-L volumetric flask with water. Add 500 mL of HCl and 40 mL of HNO₃, swirl to mix, dilute to the mark, and mix.

7.8 Iron, Standard Solution A (1 mL = 1.00 mg Fe)—Transfer 1.000 g of iron (purity: 99.95 % min) to a 250-mL beaker, cover, and add 100 mL of HCl (1 + 1). Boil gently to complete dissolution. Cool and transfer to a 1-L volumetric flask, add 200 mL of HCl and 20 mL of HNO₃, dilute to volume, and mix. Store in the polyethylene bottle.

7.9 Iron, Standard Solution B (1 mL = 0.010 mg Fe)—Using a pipet, transfer 1.00 mL of Iron Standard Solution A to a 100-mL volumetric flask. Dilute to volume with dilution solution and mix.

7.10 Lanthanum, Standard Solution A (1 mL = 0.010 mg La)—Ignite lanthanum oxide (La₂O₃) (purity: 99.9 % min) for 1 h at 1000 °C and cool to room temperature in a desiccator. Dissolve 1.173 g of dry lanthanum oxide in 100 mL of HCl (1 + 9) and transfer to a 1-L volumetric flask. Add 240 mL of HCl and 20 mL of HNO₃, dilute to volume, and mix. Store in a polyethylene bottle.

7.11 Lanthanum, Standard Solution B (1 mL = 0.010 mg La)—Using a pipet, transfer 1.00 mL of Lanthanum Standard