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**Iron ores — Determination of  
chromium content — Flame atomic  
absorption spectrometric method**

*Minerais de fer — Dosage du chrome — Méthode par spectrométrie  
d'absorption atomique dans la flamme*

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## Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

The procedures used to develop this document and those intended for its further maintenance are described in the ISO/IEC Directives, Part 1. In particular the different approval criteria needed for the different types of ISO documents should be noted. This document was drafted in accordance with the editorial rules of the ISO/IEC Directives, Part 2 (see [www.iso.org/directives](http://www.iso.org/directives)).

Attention is drawn to the possibility that some of the elements of this document may be the subject of patent rights. ISO shall not be held responsible for identifying any or all such patent rights. Details of any patent rights identified during the development of the document will be in the Introduction and/or on the ISO list of patent declarations received (see [www.iso.org/patents](http://www.iso.org/patents)).

Any trade name used in this document is information given for the convenience of users and does not constitute an endorsement.

For an explanation on the meaning of ISO specific terms and expressions related to conformity assessment, as well as information about ISO's adherence to the WTO principles in the Technical Barriers to Trade (TBT) see the following URL: [Foreword - Supplementary Information](#)

The committee responsible for this document is ISO/TC 102, *Iron ore and direct reduced iron*, Subcommittee SC 2, *Chemical analysis*.

This second edition cancels and replaces the first edition (ISO 15634:2005), which has been technically revised.

# Iron ores — Determination of chromium content — Flame atomic absorption spectrometric method

**WARNING** — This International Standard may involve hazardous materials, operations, and equipment. This International Standard does not purport to address all of the safety problems associated with its use. It is the responsibility of the user of this International Standard to establish appropriate health and safety practices and determine the applicability of regulatory limitations prior to use.

## 1 Scope

This International Standard specifies a flame atomic absorption spectrometric method for the determination of the chromium content of iron ores.

This method is applicable to chromium contents between 0,001 6 % and 0,1 % (mass fractions) in natural iron ores, iron ore concentrates, and agglomerates including sinter products.

## 2 Normative references

The following documents, in whole or in part, are normatively referenced in this document and are indispensable for its application. For dated references, only the edition cited applies. For undated references, the latest edition of the referenced document (including any amendments) applies.

ISO 648, *Laboratory glassware — Single-volume pipettes*

ISO 1042, *Laboratory glassware — One-mark volumetric flasks*

ISO 3082, *Iron ores — Sampling and sample preparation procedures*

ISO 3696, *Water for analytical laboratory use — Specification and test methods*

ISO 7764, *Iron ores — Preparation of predried test samples for chemical analysis*

## 3 Principle

The test portion is decomposed by treatment with hydrochloric and nitric acids.

The major portion of iron in the filtrate is removed by extraction with 4-methylpentan-2-one.

The insoluble residue is ignited and the silicon dioxide is removed by evaporation with hydrofluoric and sulfuric acids. The residue is fused with a mixture of sodium carbonate and sodium tetraborate, and then dissolved with hydrochloric acid and combined with the main solution.

The solution is aspirated into the flame of an atomic absorption spectrometer using a nitrous oxide-acetylene burner.

The absorbance values obtained are compared with those obtained from the calibration solutions.

## 4 Reagents

During the analysis, use only reagents of recognized analytical grade.

**4.1 Water**, complying with grade 2 as defined in ISO 3696.

- 4.2 **Sodium carbonate**, ( $\text{Na}_2\text{CO}_3$ ), anhydrous powder.
- 4.3 **Sodium tetraborate**, ( $\text{Na}_2\text{B}_4\text{O}_7$ ), anhydrous powder.
- 4.4 **Lithium tetraborate**, ( $\text{Li}_2\text{B}_4\text{O}_7$ ), anhydrous powder.
- 4.5 **Hydrochloric acid**,  $\rho$  1,16 g/ml to 1,19 g/ml.
- 4.6 **Hydrochloric acid**,  $\rho$  1,16 g/ml to 1,19 g/ml, diluted 2 + 1.
- 4.7 **Hydrochloric acid**,  $\rho$  1,16 g/ml to 1,19 g/ml, diluted 1 + 1.
- 4.8 **Hydrochloric acid**,  $\rho$  1,16 g/ml to 1,19 g/ml, diluted 2 + 100.
- 4.9 **Nitric acid**,  $\rho$  1,4 g/ml.
- 4.10 **Nitric acid**,  $\rho$  1,4 g/ml, diluted 1 + 1.
- 4.11 **Hydrofluoric acid**,  $\rho$  1,13 g/ml, 40 % (mass fraction) or  $\rho$  1,19 g/ml, 48 % (mass fraction).
- 4.12 **Sulfuric acid**,  $\rho$  1,84 g/ml.
- 4.13 **Sulfuric acid**,  $\rho$  1,84 g/ml, diluted 1 + 1.
- 4.14 **Methylpentan-2-one**, (methyl isobutyl ketone, MIBK).
- 4.15 **Chromium standard solution A**, 100  $\mu\text{g Cr/ml}$ .

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Dissolve 0,100 0 g of chromium metal [purity >99,9 % (mass fraction) (see Note of 4.16)] in 20 ml of hydrochloric acid (4.7). After cooling, transfer to a 1 000 ml one-mark volumetric flask quantitatively, dilute to volume with water (4.1), and mix.

4.16 **Chromium standard solution B**, 10  $\mu\text{g Cr/ml}$ .

Transfer 100,0 ml of chromium standard solution A (4.15) to a 1 000 ml one-mark volumetric flask. Dilute to volume with water and mix.

NOTE The purity of the metals stated on the certificates does not generally take into account the presence of adsorbed gases like oxygen, carbon monoxide, etc.

## 5 Apparatus

Ordinary laboratory apparatus including one-mark pipettes and one-mark volumetric flasks complying with the specifications of ISO 648 and ISO 1042, and the following.

- 5.1 **Platinum crucible**, of minimum capacity 25 ml.
- 5.2 **Muffle furnace**.
- 5.3 **Atomic absorption spectrometer**, equipped with a dinitrogen oxide-acetylene burner.

**WARNING** — Follow the manufacturer's instructions for igniting and extinguishing the dinitrogen oxide-acetylene flame to avoid possible explosion hazards. Wear tinted safety glasses whenever the burner is in operation.

The atomic absorption spectrometer used in this method shall meet the following criteria:

- a) *Minimum sensitivity* — The absorbance of the most concentrated calibration solution (see 7.4.4) is at least 0,3;
- b) *Graph linearity* — The slope of the calibration graph covering the top 20 % of the concentration range (expressed as a change in absorbance) is not less than 0,7 of the value of the slope for the bottom 20 % of the concentration range determined in the same way;
- c) *Minimum stability* — The standard deviation of the absorbance of the most concentrated calibration solution and that of the zero calibration solution each being calculated from a sufficient number of repetitive measurements are less than 1,5 % and 0,5 %, respectively of the mean value of the absorbance of the most concentrated solution.

The use of a strip-chart recorder and/or digital readout device is recommended to evaluate criteria a), b), and c) and for all subsequent measurements.

NOTE Instrument parameters will vary with each instrument. The following parameters were successfully used in several laboratories and they can be used as guidelines:

- hollow-cathode lamp            10 mA;
- wavelength                        357,9 nm;
- dinitrogen oxide flow rate     $\approx 10$  l/min;
- acetylene flow rate               5 l/min.

A dinitrogen oxide-acetylene flame was used.

In systems where the values shown above for gas flow rates do not apply, the ratio of the gas flow rates can still be a useful guideline.

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## 6 Sampling and samples

### 6.1 Laboratory sample

For analysis, use a laboratory sample of particle size of less than 100  $\mu\text{m}$  which has been taken and prepared in accordance with ISO 3082. In the case of ores containing significant contents of combined water or oxidizable compounds, use a particle size of less than 160  $\mu\text{m}$ .

NOTE A guideline on significant contents of combined water and oxidizable compounds is incorporated in ISO 7764.

Ensure that the sample has not been pulverized in a nickel/chromium pot.

### 6.2 Preparation of predried test samples

Thoroughly mix the laboratory sample and taking multiple increments, extract a test sample in such a way that it is representative of the whole contents of the container. Dry the test sample at  $(105 \pm 2)$  °C as specified in ISO 7764 (this is the predried test sample).

## 7 Procedure

### 7.1 Number of determinations

Carry out the analysis at least in duplicate in accordance with [Annex A](#), independently on one predried test sample.

NOTE The expression “independently” means that the second and any subsequent result is not affected by the previous result(s). For this particular analytical method, this condition implies that the repetition of the procedure be carried out either by the same operator at a different time or by a different operator including, in either case, appropriate recalibration.

### 7.2 Test portion

Taking several increments, weigh to the nearest 0,000 2 g, approximately 1 g of the predried test sample obtained in accordance with [6.2](#).

The test portion should be taken and weighed quickly in order to avoid reabsorption of moisture.

### 7.3 Blank test and check test

In each run, one blank test and one analysis of a certified reference material of the same type of ore shall be carried out in parallel with the analysis of the ore sample(s) under the same conditions. A predried test sample of the certified reference material shall be prepared as specified in [6.2](#).

The certified reference material should be of the same type as the sample to be analysed and the properties of the two materials should be sufficiently similar to ensure that in either case no significant changes in the analytical procedure will become necessary. Where a certificate reference material is not available, a reference material can be used (see [8.2.4](#)).

Where the analysis is carried out on several samples at the same time, the blank value can be represented by one test provided that the procedure is the same and the reagents used are from the same reagent bottles.

Where the analysis is carried out on several samples of the same type of ore at the same time, the analytical value of one certified reference material can be used.

### 7.4 Determination

#### 7.4.1 Decomposition of the test portion

Transfer the test portion ([7.2](#)) to a 250 ml tall-form beaker. Moisten with a few millilitres of water ([4.1](#)). Add 25 ml of hydrochloric acid ([4.5](#)), cover with a watch-glass, and heat for about 1 h on a hotplate adjusted to give a temperature of 100 °C in a test beaker containing a similar volume and depth of sulfuric acid ([4.12](#)).

If the amount of insoluble residue is high, continue heating on a higher temperature zone of the hotplate without boiling the solution.

After heating, the volume of the solution should have decreased to about 10 ml.

Add 5 ml of nitric acid ([4.9](#)) and 0,2 ml of sulfuric acid ([4.13](#)), digest for about 15 min, and evaporate the solution almost to dryness.

If the test sample contains a significant content of barium, the addition of sulfuric acid ([4.13](#)) should be omitted.

Be sure to cover with a watch-glass in order to prevent sublimation of chlorides. During evaporation, shift the glass slightly.



Add 20 ml of hydrochloric acid (4.7) and heat to dissolve the salts. Cool, wash the watch-glass and the walls of the beaker, and filter the solution through a close-texture filter paper containing filter pulp (0,3 g to 0,4 g dry mass) into a 200 ml beaker. Carefully remove all adhering particles with a rubber-tipped glass rod or piece of moistened filter paper and transfer to the filter washing the paper with warm hydrochloric acid (4.8) until it is visibly free from iron and again, wash the paper with three or four portions of warm water (4.1). Reserve the filtrate and washings in the 200 ml beaker as the main solution. Transfer the filter paper and residue to a platinum crucible (5.1).

#### 7.4.2 Removal of iron

Heat the main solution and evaporate almost to dryness. Dissolve the salts in 15 ml of hydrochloric acid (4.6) and transfer to a 200 ml separating funnel. Rinse the beaker with 20 ml of hydrochloric acid (4.6) and transfer the washings to the funnel.

Add 50 ml of 4-methylpentan-2-one (4.14) to the funnel and shake it for 1 min. Allow the two layers to separate and drain the lower aqueous layer into the original 200 ml beaker. Wash the organic layer by adding 10 ml of hydrochloric acid (4.6) to the funnel and shaking it for 30 s. Allow the two layers to separate and drain the lower aqueous layer into the 200 ml beaker to combine with the first aqueous layer. Discard the organic layer.

Heat the solution gently and expel almost all of the 4-methylpentan-2-one in the solution. Then, add 5 ml of nitric acid (4.9) and evaporate to dryness. Dissolve the salts in 20 ml of hydrochloric acid (4.7).

#### 7.4.3 Treatment of the residue

Dry and burn off the filter paper in the platinum crucible (5.1) at a low temperature (500 °C to 800 °C) and ignite the residue. Cool, moisten with one to two drops of water (4.1), and add three drops of sulfuric acid (4.12) and 5 ml of hydrofluoric acid (4.11). Evaporate slowly to expel silicon dioxide and continue heating to remove the sulfuric acid. Ignite at 800 °C for several minutes and then cool. Add 1,2 g of a mixture of 0,8 g of sodium carbonate (4.2) and 0,4 g of sodium tetraborate (4.3) to the residue and mix. Heat gently for several minutes, then heat at 1 000 °C for 15 min in a muffle furnace (5.2) or over a pressurized air burner for a time sufficient to produce a clear melt.

Lithium tetraborate (4.4) can be used instead of sodium tetraborate (4.3). In this case, sodium tetraborate should be replaced with lithium tetraborate in the calibration solutions (7.4.4) and the flame setting solutions (7.4.5).

Allow the crucible to cool, add 10 ml of hydrochloric acid (4.7), and heat gently to dissolve the melt. Combine this solution with the main solution from 7.4.2, heat to remove carbon dioxide, and cool. Transfer to a 100 ml one-mark volumetric flask quantitatively (for samples containing less than a mass fraction of 0,01 % chromium, use a 50 ml one-mark volumetric flask), dilute to volume with water (4.1), and mix (this is the test solution).

#### 7.4.4 Preparation of the calibration solutions

Introduce into a series of six 200 ml beakers, 30 ml of hydrochloric acid (4.7), 0,2 ml of sulfuric acid (4.13) and a mixture of 0,8 g of sodium carbonate (4.2) and 0,4 g of sodium tetraborate (4.3). Heat to remove carbon dioxide and cool.

Using pipettes or burettes, add 0 ml, 1,0 ml, 3,0 ml, 5,0 ml, 7,0 ml, and 10,0 ml of chromium standard solution A (4.15) for samples containing 0,01 % chromium to 0,1 % chromium or chromium standard solution B (4.16) for samples containing 0,001 % chromium to 0,01 % chromium.

Transfer to six 100 ml one-mark volumetric flasks respectively (50 ml one-mark volumetric flasks for standard solution B), dilute to volume with water (4.1), and mix. These calibration solutions cover the concentration ranges 0 µg/ml to 10 µg/ml in the case of standard solution A or 0 µg/ml to 2 µg/ml in the case of standard solution B.