Designation: C761 – 11

Standard Test Methods for Chemical, Mass Spectrometric, Spectrochemical, Nuclear, and Radiochemical Analysis of Uranium Hexafluoride[1]

This standard is issued under the fixed designation C761; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 These test methods cover or give reference to procedures for subsampling and for chemical, mass spectrometric, spectrochemical, nuclear, and radiochemical analysis of uranium hexafluoride (UF_6). Most of these test methods are in routine use to determine conformance to UF_6 specifications in the Enrichment and Conversion Facilities.

1.2 The analytical procedures in this document appear in the following order:

NOTE 1—Subcommittee C26.05 will confer with C26.02 concerning the renumbered section in Test Methods C761 to determine how concerns with renumbering these sections are best addressed in subsequent publications as analytical methods are replaced with stand-alone analytical methods.

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1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* (For specific safeguard and safety consideration statements, see Section 6.)

2. Referenced Documents

2.1 The following documents of the issue in effect on date of material procurement form a part of this specification to the extent referenced herein:

2.2 *ASTM Standards:*[2]

C787 Specification for Uranium Hexafluoride for Enrichment

C799 Test Methods for Chemical, Mass Spectrometric, Spectrochemical, Nuclear, and Radiochemical Analysis of Nuclear-Grade Uranyl Nitrate Solutions

C859 Terminology Relating to Nuclear Materials

C996 Specification for Uranium Hexafluoride Enriched to Less Than 5 % ^{235}U

C1128 Guide for Preparation of Working Reference Materials for Use in Analysis of Nuclear Fuel Cycle Materials

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[1] These test methods are under the jurisdiction of ASTM Committee C26 on Nuclear Fuel Cycle and are the direct responsibility of Subcommittee C26.05 on Methods of Test.

Current edition approved May 15, 2011. Published July 2011. Originally approved in 1973. Last previous edition approved in 2004 as C761 – $04^{\varepsilon 1}$. DOI: 10.1520/C0761-11.

[2] For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

C1219 Test Methods for Arsenic in Uranium Hexafluoride (Withdrawn 2015)[3]

C1233 Practice for Determining Equivalent Boron Contents of Nuclear Materials

C1267 Test Method for Uranium by Iron (II) Reduction in Phosphoric Acid Followed by Chromium (VI) Titration in the Presence of Vanadium

C1287 Test Method for Determination of Impurities in Nuclear Grade Uranium Compounds by Inductively Coupled Plasma Mass Spectrometry

C1295 Test Method for Gamma Energy Emission from Fission and Decay Products in Uranium Hexafluoride and Uranyl Nitrate Solution

C1344 Test Method for Isotopic Analysis of Uranium Hexafluoride by Single-Standard Gas Source Mass Spectrometer Method

C1346 Practice for Dissolution of UF_6 from P-10 Tubes

C1380 Test Method for the Determination of Uranium Content and Isotopic Composition by Isotope Dilution Mass Spectrometry

C1413 Test Method for Isotopic Analysis of Hydrolyzed Uranium Hexafluoride and Uranyl Nitrate Solutions by Thermal Ionization Mass Spectrometry

C1428 Test Method for Isotopic Analysis of Uranium Hexafluoride by Single–Standard Gas Source Multiple Collector Mass Spectrometer Method

C1429 Test Method for Isotopic Analysis of Uranium Hexafluoride by Double-Standard Multi-Collector Gas Mass Spectrometer

C1441 Test Method for The Analysis of Refrigerant 114, Plus Other Carbon-Containing and Fluorine-Containing Compounds in Uranium Hexafluoride via Fourier-Transform Infrared (FTIR) Spectroscopy

C1474 Test Method for Analysis of Isotopic Composition of Uranium in Nuclear-Grade Fuel Material by Quadrupole Inductively Coupled Plasma-Mass Spectrometry

C1477 Test Method for Isotopic Abundance Analysis of Uranium Hexafluoride and Uranyl Nitrate Solutions by Multi-Collector, Inductively Coupled Plasma-Mass Spectrometry

C1508 Test Method for Determination of Bromine and Chlorine in UF_6 and Uranyl Nitrate by X-Ray Fluorescence (XRF) Spectroscopy

C1539 Test Method for Determination of Technetium-99 in Uranium Hexafluoride by Liquid Scintillation Counting

C1561 Guide for Determination of Plutonium and Neptunium in Uranium Hexafluoride and U-Rich Matrix by Alpha Spectrometry

C1636 Guide for the Determination of Uranium-232 in Uranium Hexafluoride

C1689 Practice for Subsampling of Uranium Hexafluoride

C1742 Test Method for Isotopic Analysis of Uranium Hexafluoride by Double Standard Single-Collector Gas Mass Spectrometer Method

D1193 Specification for Reagent Water

D3084 Practice for Alpha-Particle Spectrometry of Water

E60 Practice for Analysis of Metals, Ores, and Related Materials by Spectrophotometry

2.3 *American Chemical Society Specification:*

Reagent Chemicals [4]

2.4 *Other Specifications:*

Uranium Hexafluoride : Base Charges, Use Charges, Special Charges, Table of Enriching Services, Specifications, and Packaging[5]

USEC 651 Good Handling and Practices for UF_6

2.5 *ANSI Standards:*[6]

ANSI N 14.1 Nuclear Material-Uranium Hexafluoride-Packaging for Transport

2.6 *ISO Standards:*

ISO 7195 Nuclear Energy-Packaging of Uranium Hexafluoride (UF_6) for Transport

3. Significance and Use

3.1 Uranium hexafluoride is a basic material used to prepare nuclear reactor fuel. To be suitable for this purpose the material must meet criteria for uranium content, isotopic composition, metallic impurities, hydrocarbon and halohydrocarbon content. These test methods are designed to determine whether the material meets the requirements described in Specifications C787 and C996.

4. Reagents

4.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all procedures. Unless otherwise indicated, all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.[4] Other grades may be used, provided that it is first established that the reagent to be used is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

4.2 *Purity of Water*—Unless otherwise indicated, references to water shall mean reagent water conforming to Specification D1193.[7]

5. Rejection

5.1 Rejection or acceptance criteria are described in Specifications C787 and C996.

6. Safety Considerations

6.1 Since UF_6 is radioactive, toxic, and highly reactive, especially with reducing substances and moisture (see Uranium Hexafluoride: Handling Procedures and Container Criteria, sections 2.4 through 2.6), appropriate facilities and practices for sampling and analysis must be provided.

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[3] The last approved version of this historical standard is referenced on www.astm.org.

[4] *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K. and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

[5] United States Department of Energy, Oak Ridge, TN 37830.

[6] Available from American National Standards Institute (ANSI), 25 W. 43rd St., 4th Floor, New York, NY 10036.

[7] Type 1 and 2 water have been found to be suitable.

6.2 Hydrofluoric acid is a highly corrosive acid that can severely burn skin, eyes, and mucous membranes. Hydrofluoric acid is similar to other acids in that the initial extent of the burn depends on the concentration, the temperature, and the duration of the contact with the acid. Hydrofluoric acid differs from other acids because the fluoride ion readily penetrates the skin, causing destruction of deep issue layers. Unlike other acids that are rapidly neutralized, hydrofluoric acid reactions with tissue may continue for days if left untreated. Due to the serious consequences of hydrofluoric acid burns, prevention of exposure or injury of personnel is the primary goal. Utilization of appropriate laboratory controls (hoods) and wearing adequate personal protective equipment to protect from skin and eye contact is essential.

6.3 *Committee C-26 Safeguards Statement:*

6.3.1 The material (uranium hexafluoride) to which these test methods apply, is subject to nuclear safeguards regulations governing its possession and use. The following analytical procedures in these test methods have been designated as technically acceptable for generating safeguards accountability measurement data: Gravimetric Determination of Uranium; Titrimetric Determination of Uranium; All Isotopic Analyses.

6.3.2 When used in conjunction with appropriate certified Reference Materials (CRMs), these procedures can demonstrate traceability to the national measurement base. However, adherence to these procedures does not automatically guarantee regulatory acceptance of the resulting safeguards measurements. It remains the sole responsibility of the user of these test methods to assure that its application to safeguards has the approval of the proper regulatory authorities.

SUBSAMPLING OF URANIUM HEXAFLUORIDE

7. Scope

7.1 This test method has been discontinued (see C 761–04$^{\varepsilon 1}$). The subsampling of UF_6 from bulk sample containers into smaller containers suitable for laboratory analyses has been published as a separate Practice, C1689.

GRAVIMETRIC DETERMINATION OF URANIUM

8. Scope

8.1 Practice C1346 is applicable to the hydrolysis of uranium hexafluoride in polychlorotrifluoroethylene (P10) tubes. The following test method is then applicable to the direct gravimetric determination of uranium.

9. Summary of Test Method

9.1 A sample of uranium hexafluoride is weighed, cooled in liquid nitrogen, and hydrolyzed with water. The uranyl fluoride solution produced is evaporated to dryness and converted to uranic oxide by pyrohydrolysis. The uranium content is determined from the weight of the uranium oxide after correcting for stoichiometry based on isotopic content, ignition conditions, and nonvolatile impurities. Ref. (**1-4**).[8]

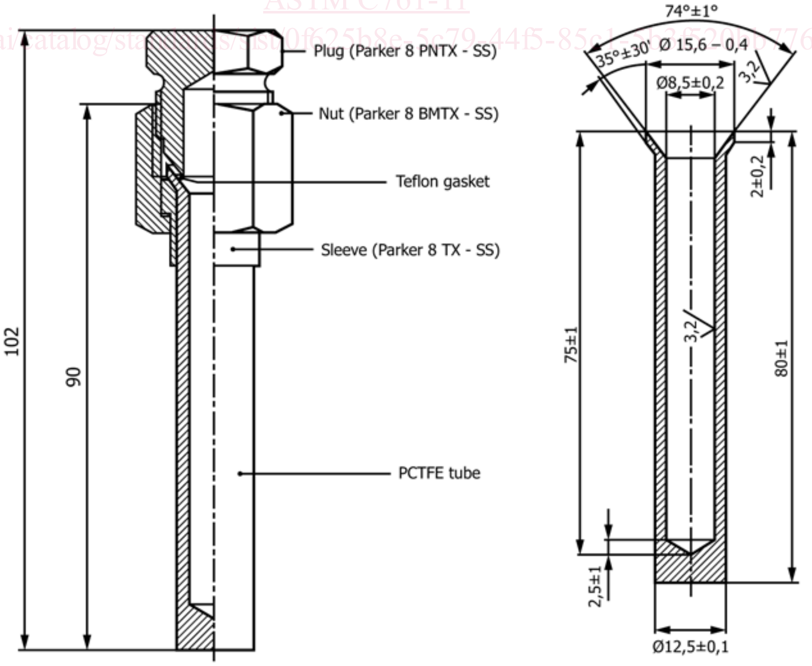

FIG. 1 Example of a Polychlorotrifluoroethylene P-10 Tube

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[8] The boldface numbers in parentheses refer to a list of references at the end of these test methods.

10. Interferences

10.1 Nonvolatile impurities affect the accuracy of the method and must be measured by spectrographic analysis with corrections applied.

11. Apparatus

11.1 *Polytrifluorochloroethylene (PTFCE) Sample Tube, TFCE Gasket, Flare Nut, and Plug,* see Fig. 1.

11.2 *Platinum Boat and Cover*—The cover should be platinum gauze (52 mesh) and shaped to cover the boat (Fig. 2).

11.3 *Muffle Furnace,* must be capable of operating continuously at 875°C and maintain this temperature within ±25°C. The furnace shall be equipped with a steam supply that is passed through a tube furnace to preheat the steam to 875°C.

11.4 *Tube Furnace,* must be capable of operating continuously at 875°C and maintain this temperature within 25°C.

11.5 *Infrared Heat Lamps,* 250 watts.

11.6 *Analytical Balance.*

11.7 *Vacuum Oven.*

11.8 *Dewar Flask,* stainless steel.

11.9 *Spatula,* platinum.

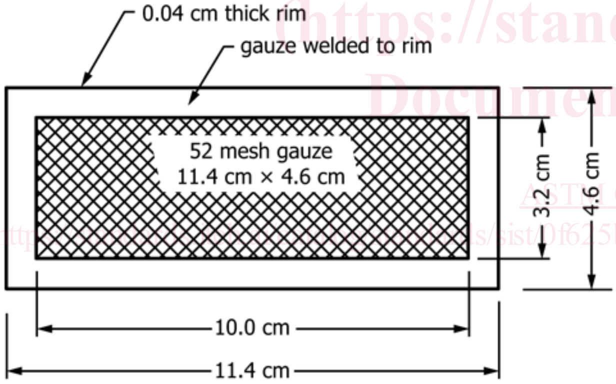
11.10 *PTFCE Rod,* 120 mm long and 1.6 mm in diameter.

11.11 *Forceps,* platinum tipped.

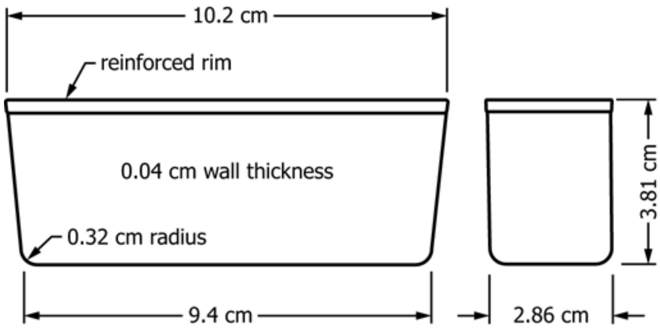
11.12 *Jig,* suitable for holding the TFCE sample tube so that it can be opened with a wrench.

11.13 *Box Wrench,* to fit sample tube plug.

11.14 *Beaker,* stainless steel, 125 mL capacity.



Cover



Boat

FIG. 2 Platinum Boat and Cover

12. Reagents

12.1 *Liquid Nitrogen.*

12.2 *Nitric Acid (sp gr 1.42)*—concentrated nitric acid (HNO_3).

12.3 *Nitric Acid* (4*M*)—Mix 500 mL of concentrated HNO_3 with 1500 mL of distilled water.

12.4 *Detergent.*

13. Sampling

13.1 A UF_6 sample is taken as described in Practice C1689.

14. Procedure

14.1 Inspect the PTFCE sample tube for leaks.

NOTE 2—An indication of a leak is a yellow-green residue on the flare nut and cap or a yellow discoloration in the tube. Discard the sample if a leak is indicated.

14.2 Allow the sample tube to stand overnight in the laboratory.

14.3 Wipe the sample tube with a lint-free tissue to remove any moisture or foreign material that might be adhering.

14.4 Weigh the sample tube to the nearest 0.1 mg.

14.5 Heat the platinum boat and screen in the pyrohydrolysis furnace at 875°C for 20 min.

14.6 Cool the platinum boat and store in a desiccator for 40 min. Weigh the boat and screen to the nearest 0.1 mg.

14.7 Freeze the sample by immersing the sample tube in liquid nitrogen for 10 min.

14.8 Add enough chilled water to the tared platinum boat to immerse the sample tube (about 50 mL).

14.9 Place the sample tube in the jig and loosen the plug with the box wrench.

14.10 Remove the sample tube from the jig and unscrew the plug while holding the sample tube in an upright position.

14.11 Remove the flare nut from the sample tube and immerse the tube and gasket in the chilled water in the tared platinum boat.

14.12 Let the gasket remain in the chilled water about 30 min.

14.13 Remove the gasket with the forceps and rinse well with deionized water into the boat.

14.14 Place the plug-nut assembly and gasket into a stainless steel beaker for drying.

14.15 Allow the tube to remain in the water until the UF_6 has been hydrolyzed (2 to 4 h).

14.16 Remove the tube from the sample solution by inserting the TFCE rod or platinum spatula into the tube and lifting directly above the boat.

14.17 Rinse the sample tube with deionized water into the boat using extreme care to prevent splashing.

14.18 Cover the sample boat containing the UO_2F_2 solution with the matching cover shown in Fig. 2. Place under the infrared head lamps and evaporate to dryness for 16 h.

14.19 Shake the excess water from the sample tube and place in the stainless beaker containing the plug-nut assembly and gasket.

14.20 Dry the sample tube parts in the vacuum oven at 80°C.

14.21 Allow the unassembled parts to sit in the room overnight.

14.22 Assemble the empty sample tube and weigh to the nearest 0.1 mg.

14.23 Disassemble the sample tube and soak the tube and gasket in 4*M* HNO_3 at 75° to 80°C for 1 h.

14.24 Rinse with deionized water and place in the stainless steel beaker.

14.25 Clean the metal parts with detergent and rinse with deionized water and acetone.

14.26 Place the metal parts to the stainless steel beaker and dry all parts in the vacuum oven at 80°C overnight.

14.27 Reassemble the sample tube for the next sample.

14.28 Set the temperatures of the furnace and tube furnace at 875°C.

14.29 Establish a steam flow to the furnace equal to 1 L of water per hour.

14.30 Place the boat into the furnace with the platinum cover on the boat and pyrohydrolyze the sample for 1 h.

14.31 Remove the boat from the furnace, cool, and place in a desiccator while still warm.

14.32 Desiccate the sample for 1 h and weigh quickly to the nearest 0.1 mg.

14.33 Transfer a portion of the U_3O_8 residue to a vial and submit for spectrographic analysis to determine the weight of nonvolatile impurities.

14.34 Place the platinum boat in hot 4*M* HNO_3 for 3 to 4 h and rinse with deionized water acetone.

15. Calculation

15.1 Calculate the weight fraction of uranium in the sample as follows:

$$\text{gU/g UF}_6 = (A - (AB))(\text{Gravimetric Factor})/W \tag{1}$$

where:

Gravimetric Factor = gU/g U_3O_8 which varies with isotopic composition. Theoretical stoichiomtry for U_3O_8 cannot be assumed and the actual gU/g U_3O_8 must be established by potentiometric titration **(1–4)**. (Tri-diffusion plant committee with DOE approval has established 0.8479 g U/g U_3O_8 by titration as the factor for natural uranium,

A = grams of U_3O_8 from the pyrohydrolysis of UO_2F_2,

B = grams of impurity metal oxides per gram of U_3O_8,

W = corrected sample weight in grams.

The correction is for the combined effects of cover gas trapped over the UF_6 in the sample tube and the air buoyancy correction **(5)**. The following equation has been determined for the sample tube in Fig. 1 and the subsampling conditions described in Practice C1689. The correction equation is applicable for sample weights in the range of 7 to 13 g.

$$W = (1.00047)x - 0.0058 \tag{2}$$

where:

x = observed UF_6 sample weight, g.

16. Precision and Bias

16.1 *Precision*—The precision within a laboratory and between laboratories was established by analyzing 15 samples at each laboratory. The sampling scheme is shown in Table 1. Within a laboratory, based on 15 measurements made on separate days the relative standard deviation is 0.021 %. The results from all the laboratories are shown in Table 2.

16.2 *Bias*—To establish an estimate of bias for the gravimetric method, a series of comparative analyses of UF_6 control batches were made using the gravimetric and potentiometric titration methods. The potentiometric titration was used as the reference method because the uranium was measured directly using NIST potassium dichromate.[9] The results are shown in Table 3.

TITRIMETRIC DETERMINATION OF URANIUM

17. Scope

17.1 A sample of the U_3O_8 produced by the hydrolysis of the UF_6 and ignition of the resulting UO_2F_2 is analyzed according to Test Method C1267.

PREPARATION OF HIGH-PURITY U_3O_8

18. Scope

18.1 High purity U_3O_8 can be prepared according to Preparation C1128. High purity uranium is needed for a blank matrix for analyses using ICP-MS, ICP-AES, AA, XRF, and MS equipment.

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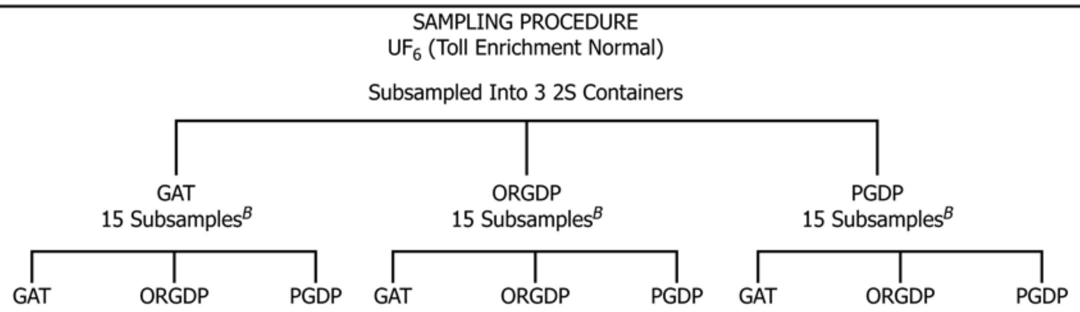
[9] Standard reference material, now available as NIST SRM 136e.

TABLE 1 Interlaboratory Study[A]—Determination of Uranium in Uranium Hexafluoride



[A] Date of study—October 1983; Participating laboratories—Goodyear Atomic Corporation, Portsmouth, Ohio (GAT); Martin Marietta Energy Systems, Inc., Oak Ridge, Tennessee (ORGDP); Martin Marietta Energy Systems, Inc., Paducah, Kentucky (PGDP).

[B] Five tubes analyzed at each laboratory.

TABLE 2 Results of Interlaboratory Study—U in UF_6

Analysis Site %U in UF_6		
GAT	ORGDP	PGDP
Subsampled at GAT:		
67.600	67.619	67.589
67.601	67.574	67.575
67.583	67.607	67.612
67.611	67.600	67.612
67.618	67.606	Sample Lost
Subsampled at ORGDP:		
67.614	67.580	67.611
67.611	67.621	67.598
67.587	67.600	67.501
67.599	67.606	67.610
67.617	67.596	67.624
Subsampled at PGDP:		
67.616	67.588	67.591
67.586	67.602	67.620
67.573	67.612	67.612
67.614	67.606	67.612
67.607		67.586
Mean and Standard Deviation:		
67.602 ± 0.014	67.601 ± 0.013	67.603 ± 0.014

ISOTOPIC ANALYSIS

19. Scope

19.1 The isotopic composition can be determined on either gaseous UF_6 or on hydrolyzed UF_6.

19.2 For gaseous UF_6, using single collector mass spectrometer instruments, Test Methods C1344 and C1742 have been developed and can be used for single or double standard method respectively. For multi-collector instruments, Test Methods C1428 and C1429, using single or double standard can be used.

19.3 For hydrolyzed UF_6, methods using Thermal Ionization Mass Spectrometry (TIMS) have been developed and can be used: Test Methods C1413 and C1380. Methods using ICP-MS can also be used: Test Methods C1474 and C1477.

DETERMINATION OF HYDROCARBONS, CHLOROCARBONS, AND HALOHYDROCARBONS

20. Scope

20.1 The determination of some forms of hydrocarbons, chlorocarbons, and halohydrocarbons in UF_6 vapor can be performed using Test Method C1441. As an alternative, a mass spectrometry technique may be used and is detailed below. Although this test method is only semiquantitative, it is adequate for certifying that the subject impurities do not exceed 0.01 mol % of the UF_6.

21. Summary of Test Method

21.1 UF_6 is admitted to a mass spectrometer through a gas sample leak, and magnetic scanning is employed to record a spectrum of peaks. A representative group of recorded peaks is compared to the same peaks in a pure UF_6 standard scan to determine whether appreciable ion fragments from subject impurities are present.

22. Interferences

22.1 If detectable impurities are present, a complete mass scan of the range from 12 to 400 is performed. All impurities are then identified from their cracking patterns, and calculations are performed using ionization efficiency factors for the compounds present. Since cracking patterns vary with ionization potential and ionization efficiencies vary with focus conditions, this measurement can only be performed by one proficient in analytical mass spectrometery.

23. Apparatus

23.1 A mass spectrometer with resolution adequate to distinguish between adjacent peaks at $m/e = 400$ is required. For example, a 152-mm radius, 60-deg, Nier-type spectrometer modified for spectrum recording **(6)** is suitable. The sample inlet system should be of nickel or Monel, equipped with an

TABLE 3 Determination of Uranium in Uranium Hexafluoride—Comparison of Gravimetric and Potentiometric Titration Methods

Control UF_6 Date[A]	Method	Number of Measurements	%Uranium Mean	%Uranium SD	Bias Estimate[B]
9/78	Gravimetric	24	67.610	0.009	−0.001
	Potentiometric Titration	9	67.611	0.015	
5/82–8/82	Gravimetric	30	67.596	0.010	−0.009
	Potentiometric Titration	8	67.605	0.011	
7/83–9/83	Gravimetric	25	67.610	0.006	+ 0.005
	Potentiometric Titration	8	67.605	0.010	

[A] Control UF_6 used in 9/78 was a different batch of material from that used in 1982 and 1983.

[B] Potentiometric titration results are used as the reference values for the bias estimates.

adjustable viscous-flow or molecular leak for delivering the sample to the ion source.

23.2 The ion source must be fabricated from nonmagnetic material such as Nichrome V, and must be designed so it can be disassembled for cleaning. The magnetic field of the analyzer magnet must be continuously variable from about 200 to 6500 gauss. A single ion collector electrode is suitable, and a vibrating-reed electrometer and 304-mm strip chart recorder are optimum for amplifying and recording ion signals.

23.3 It is quite possible that quadrupole or time-of-flight instruments could be adapted to this measurement.

24. Procedure

24.1 *UF_6 Standard Measurements:*

24.1.1 Select a standard material that has been given repetitive flash purifications to rid it of all volatile impurities. Isotopic UF_6 standards usually fall in this category.

24.1.2 With the electrometer sensitivity set at 1⁄100 of the most sensitive usable operating range, adjust the gas flow to the ion source to record a mass (Note 3) 333 peak (UF_5^+) approximately 80 % of full scale (80 divisions).

NOTE 3—The term “mass” in this procedure alludes to *m/e*, the mass-to-charge ratio (see also Terminology C859).

24.1.3 Measure the ratio of mass 333 (UF_5^+) to 147.5 (UF_3^{++}). Mass 333 is measured on a sensitivity range one-hundredth that of mass 147.5. Depending on focus conditions, a ratio of the order of 10^2 is obtained. Measure this ratio only once per day and use for calculating results of all samples analyzed that day.

24.1.4 With the electrometer sensitivity set at 1⁄100 of the most sensitive usable operating range, increase the gas flow to provide an output signal of approximately 80 divisions at the 147.5 mass position. This gives a detection limit of the order of 2 ppm per chart division: 50 to 100 due to ratio between UF_5^+ and UF_3^{++}, 100 due to sensitivity shunts, and approximately 80 on the recorder chart.

24.1.5 Record a scan of mass range from 12 to 150 using the most sensitive usable operating shunt, and use this scan as a background for all samples analyzed that day.

24.2 *UF_6 Sample Measurement:*

24.2.1 Introduce the sample to the spectrometer source such that an output intensity of approximately 80 chart divisions is obtained at the 147.5 mass number (UF_3^{++}), using 1⁄100 the most sensitive usable operating range.

24.2.2 Record a scan of mass range from 12 to 150 using the most sensitive usable operating range.

24.2.3 Repeat 24.2.1 and 24.2.2 for each sample to be analyzed that day.

25. Calculation

25.1 Due to mass spectrometer cracking patterns, low-mass ion fragments are produced from all compounds, even the high-mass ones.

25.2 It is practical to look for a representative group of such ion fragments at specific masses. Thus, initially monitor the following masses for purposes of this procedure:

Mass Number	Positively Charged Ion Fragment
15	CH_3
26	C_2H_2
27	C_2H_3
31	CF
43	C_3H_7
47	CCl^{35}
49	CCl^{37}
69	CF_3

25.3 Read sample intensities for the representative ion fragments from the recorder chart.

25.4 Subtract the background intensities observed on the pure standard from respective sample intensities.

25.5 Examine the net intensity at each of the eight mass numbers. (It will be recalled that one recorder chart division of net intensity is equivalent to about 2 ppm on a UF_6 basis; however, ionization efficiencies of compounds differ, and a specific ion fragment may result from many different compounds. Thus, the net intensity at a specific mass number is only qualitative and not a quantitative measurement of impurity.)

25.5.1 If the net ion intensity does not exceed 2 ppm at any of the mass positions, report the sample as containing less than 0.01 mol % of the subject impurities.

25.5.2 Where detectable impurities are apparent, perform a complete mass scan of the range from 12 to 400, identify impurities, and perform calculations using ionization efficiency factors for the compounds present.

26. Reliability

26.1 This simplified procedure was designed specifically to certify that a UF_6 sample contains less than 0.01 mol %

hydrocarbons, chlorocarbons, and partially substituted halohydrocarbons. Thus, the procedure is qualitative rather than quantitative in cases where the impurity level is below 100 ppm.

26.2 The detectability limit for any ion fragment is about 2 ppm. The detectability limit for the parent compound could be greater or less than 2 ppm depending on ionization efficiency and cracking pattern. If one of the subject compounds were present to 10 ppm or more, it would be evident from monitoring the eight masses. Thus, a quoted result of less than 0.01 mol % is conservative.

26.3 The results are considered quantitative when the impurity being determined is present to a level greater than 100 ppm. In such instances the impurity is identified and measured. The 95 % symmetrical confidence interval for such a measurement is ±50 % of the quoted impurity.

DETERMINATION OF ANTIMONY

27. Scope

27.1 The Atomic Absorption test method has been discontinued (see C761–96). Antimony can be determined by ICP-MS. Test Method C1287 can be used.

DETERMINATION OF BROMINE

28. Scope

28.1 The Spectrophotometric test method has been discontinued (see C761–96). Bromine can be determined by X-Ray spectroscopy. Test Method C1508 can be used.

DETERMINATION OF CHLORINE

29. Scope

29.1 Chlorine can be determined by X-Ray Spectroscopy. Test Method C1508 can be used. Chlorine can also be determined by titrimetry. This test method is described below. It is applicable over a range from 10 to 100 ppm chlorine; however, higher concentrations can be measured by appropriate sample dilution.

30. Summary of Test Method

30.1 The test method consists of treating a hydrolyzed sample of UF_6 with ferrous sulfate in sulfuric acid solution to reduce chlorates, and then with potassium permanganate to liberate free chlorine. The chlorine gas is carried by a nitrogen stream into a potassium iodide solution, and the liberated iodine is titrated with sodium thiosulfate. Bromine, if present, is determined separately, and a correction is applied to the chlorine result.

30.2 It is recommended that the potassium iodide-sodium acetate solution be made up fresh once each week. Any color change signals the need for a new solution. As the solution ages, the blank result increases; therefore, the same potassium iodide solution is used for both sample and blank.

30.3 If the sample solution is allowed to boil too vigorously when chlorine gas is being released, liquid droplets may be carried by the nitrogen stream into the potassium iodide solution, resulting in sample bias.

31. Apparatus

31.1 *Distillation Apparatus,* shown in Fig. 3.

32. Reagents

32.1 *Boric Acid* (H_3BO_3), reagent grade, crystal or powder.

32.2 *Ferrous Sulfate Solution*—5 g $FeSO_4 \times 7H_2O$ dissolved in 500 mL of 3.6 *M* sulfuric acid.

32.3 *Potassium Iodide* (KI), reagent grade.

32.4 *Potassium Iodide-Sodium Acetate Solution*—Dissolve 100 g of KI and 100 g of $NaC_2H_3O_2 \times 3H_2O$ in distilled water and dilute to 2 L.

32.5 *Potassium Permanganate Solution (1 %)*—Prepare a 1 % solution of $KMnO_4$ in water.

32.6 *Sodium Acetate* ($NaC_2H_3O_2$), reagent grade.

32.7 *Sodium Thiosulfate Solution* (0.025 *N*)—Prepare a 0.025 *N* solution of $Na_2S_2O_3$ in water.

32.8 *Starch Indicator Solution,* pH 7.

32.9 *Sulfuric Acid (sp gr 1.84)*—Concentrated sulfuric acid (H_2SO_4).

33. Procedure

33.1 *Sample Preparation:*

33.1.1 Hydrolyze the sample of UF_6 with distilled water. Approximately 250 g of UF_6 from a nickel knockout container may be hydrolyzed to provide a sample for various chemical measurements, or about 20 g from two polychlorotrifluoroethylene sample tubes may be hydrolyzed to provide the chlorine sample. The procedure for hydrolyzing the contents of the polychlorotrifluoroethylene tubes is described here.

33.1.1.1 Immerse the tubes in liquid nitrogen and cool for 10 min.

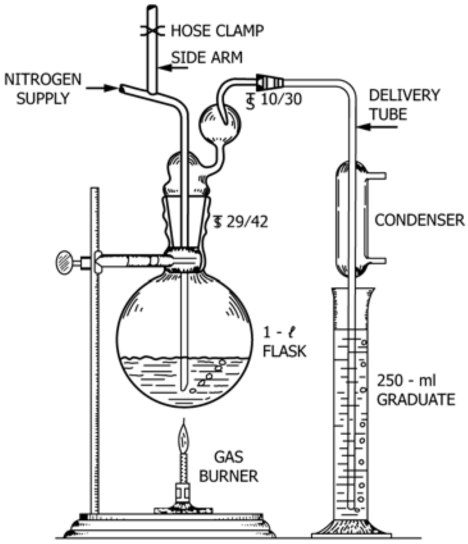


FIG. 3 Apparatus for Distillation of Chlorine

33.1.1.2 Remove the top flare plugs and collars, and place the tubes into a platinum dish or a polychlorotrifluoroethylene beaker containing 100 mL of chilled distilled water.

33.1.1.3 After hydrolysis of the UF_6, remove the polychlorotrifluoroethylene tubes and rinse with distilled water. Add the rinse solution to the UO_2F_2 solution.

33.2 *Analysis:*

33.2.1 Fill the graduated cylinder in Fig. 2 to the 150-mL level with KI-$NaC_2H_3O_2$ solution. Then connect the delivery tube so its tip is near the bottom of the solution in the receiving graduate.

33.2.2 Dispense 20 g of H_3BO_3 into the 1-L round-bottom flask.

33.2.3 Transfer the sample solution containing UO_2F_2 from approximately 20 g of UF_6 in 100 mL of solution to the flask.

33.2.4 Add 10 mL of concentrated H_2SO_4 (sp gr 1.84) to the flask, and swirl the contents for mixing.

33.2.5 Add 10 mL of the $FeSO_4$ solution, rinse the mouth of the flask, and connect the flask immediately to the apparatus as in Fig. 3.

33.2.6 Initiate nitrogen flow through the solution at a rate of 2 to 3 bubbles per second and start the water flow through the condenser.

33.2.7 Heat the contents of the flask until boiling and allow to boil for 30 s.

33.2.8 Remove the heat, add 10 mL of 1 % $KMnO_4$ solution through the sidearm, and close the sidearm by clamping the rubber tube that is attached to the end of the sidearm.

33.2.9 Reapply heat and allow the contents of the flask to simmer for 5 min.

33.2.10 Remove the heat, but continue the nitrogen purge for an additional 5 min.

33.2.11 Rinse the delivery tube into the receiving graduate and transfer the contents of the graduate to a 300-mL Erlenmeyer flask. Add 1 mL of starch solution and titrate the iodine with 0.025 N $Na_2S_2O_3$ solution to the starch end point. (The iodine may be measured spectrophotometrically rather than titrimetrically.)

33.2.12 Perform a blank analysis by carrying 100 mL of distilled water through procedural steps, 33.2.1 through 33.2.11, and subtracting from the sample titration.

34. Calculation

34.1 Calculate the concentration of chlorine in ppm chlorine on a uranium basis as follows:

$$Cl, ppm = (V_1 - V_2)\,(N)\,(0.03545 \times 10^4)/S$$

where:

V_1 = millilitres of thiosulfate for sample,
V_2 = millilitres of thiosulfate for blank,
N = normality of $Na_2S_2O_3$ solution, and
0.03545 = grams of chlorine per milliequivalents, and
S = grams of uranium.

35. Reliability

35.1 The precision at the 95 % confidence level is ±10 % at the 100-ppm level.

DETERMINATION OF SILICON AND PHOSPHORUS

36. Scope

36.1 Phosphorus and Silicon can be determined by ICP-MS (see Test Method C1287).

36.2 Silicon can be analyzed by Atomic Absorption without matrix separation.

36.3 Phosphorus and Silicon can be analyzed by spectrophotometry (as described below). With these procedures about 0.5 µg silicon or phosphorus per gram of uranium can be detected.

37. Summary of Test Method

37.1 The test methods are based on the development of the color known as molybdenum blue obtained by the reduction of silico- or phosphomolybdate ions.

37.2 Reduction of the silico- or phosphomolybdate ions with a combination 1-amino-2-naphthol-4-sulfonic acid, sodium sulfite, sodium pyrosulfite solution produces the same molybdenum blue complex that is measured spectrophotometrically, directly in the uranium solution at 710 nm. One gram of uranium absorbs slightly at this wavelength, but its absorbance is easily corrected by the use of an additional aliquot to which no reducing agent is added as the blank.

38. Interferences

38.1 Phosphate interference in the silicon determination is eliminated by the addition of oxalic acid to decompose any phospho-molybdate formed. Silicon in small amounts does not interfere in the phosphorus analysis since silicomolybdate does not form at the acid concentration at which the phosphomolybdic acid is formed.

38.2 Fluoride, which would be a serious interference in the silicon analysis, is complexed with boric acid. A high concentration of silicon which could interfere in the phosphorus analysis is prevented by handling hydrolyzed UF_6 samples in platinum or plastic. Other potential interferences rarely present in significant amounts are arsenic and tungsten.

39. Apparatus

39.1 *Polyethylene Bottles,* 100 and 500 mL.

39.2 *Polyethylene Beakers,* 100 mL.

39.3 *Polyethylene Pipets;* 1, 2, 5, and 10 mL.

39.4 *Spectrophotometer,* equipped with 1 and 5 cm cells as described in Practice E60.

40. Reagents

40.1 *Ammonium Hydroxide Silicon-Free*—Distill 500 mL of saturated NH_4OH through plastic tubing into 300 mL of distilled water.

40.2 *Ammonium Molybdate Solution (10 %)*—Dissolve 100 g of reagent grade $(NH_4)_6Mo_7O_{24}$ in water, and dilute the solution to 1 L with distilled water.

NOTE 4—Not all commercially available $(NH_4)_6Mo_7\,O_{24}$ is suitable. Material supplied by J. T. Baker Chemical Co. or Baker and Adamson

Products, however, has been found to be satisfactory consistently.

40.3 *Boric Acid Solution (5 %)*—Dissolve 25 g of reagent grade H_3BO_3 in water, and dilute the solution to 500 mL.

40.4 *Oxalic Acid Solution (5 %)*—Dissolve 25 g of reagent grade $H_2C_2\ O_4$ in water, and dilute the solution to 500 mL. (This solution is not required for determination of phosphorus.)

40.5 *Phosphorus Standard Solution (25 µg P/mL)*—Dissolve 4.6422 g of ammonium dihydrogen phosphate $[(NH_4)H_2PO_4]$ in distilled water and dilute the solution to 1 L with distilled water. Transfer 20 mL of this solution to a 1-L volumetric flask and dilute to 1 volume with distilled water to obtain a solution containing 2.0 µg P/mL.

40.6 *Reducing Mix*—Dissolve 0.1 g of 1-amino-2-naphthol-4-sulfonic acid, 1.0 g of sodium sulfite (Na_2SO_3), and 10.0 g of sodium pyrosulfite ($Na_2S_2O_5$) in water; then dilute the solution to 100 mL.

40.7 *Silicon Standard Solution (2.5 µg Si/mL)*—Dissolve 10.6 mg of precipitated silica (SiO_2) and 0.5 g of sodium hydroxide (NaOH) in a platinum dish. Transfer the solution to a 2-L plastic bottle and dilute to volume.

NOTE 5—Silicon in solution as sodium silicate is not stable when stored in polyethylene bottles. New standard solutions should be prepared monthly.

40.8 *Sulfuric Acid—Boric Acid Solution (10 % H_2SO_4—4 %* H_3BO_3)—Dissolve 20 g of reagent grade H_3BO_3 in water. Add 50 mL of concentrated H_2SO_4 (sp gr 1.84) and dilute the solution to 500 mL.

40.9 *Uranium Oxide* (U_3O_8, UO_2, or UO_3), phosphorus- and silicon-free.

NOTE 6—All standard solutions should be made and stored in plastic containers to prevent silicon contamination from glassware.

41. Procedure

41.1 *Sample Preparation:*

41.1.1 Hydrolyze a weighed portion of 6 to 10 g of UF_6 in a platinum boat in 80 mL of distilled water as described in 14.7 through 14.17.

41.1.2 Transfer the solution to a 100-mL plastic bottle and dilute to 100 mL.

41.1.3 Transfer an aliquot equivalent to 1 g of UF_6 to a 100-mL TFE-fluorocarbon beaker, and add 1 mL of 6 *N* H_2SO_4. (Phosphorus aliquots should contain 1 to 50 µg of phosphorus.)

41.1.4 Add 20 mL of 5 % H_3BO_3, and heat the solution for 20 min to complex the fluoride.

41.1.5 For phosphorus analysis only, transfer the solution to a 100-mL borosilicate beaker and evaporate to 20 mL (see 41.3).

41.2 *Determination of Silicon:*

41.2.1 *Preparation of Calibration Curve:*

41.2.1.1 Pipet standard aliquots containing 0, 2.5, 5.0, 7.5, 10.0, and 12.5 µg of silicon into plastic beakers.

41.2.1.2 Add silicon-free uranium, 0.1 g as uranyl nitrate solution, to each beaker. Prepare the uranyl nitrate solution by dissolving silicon-free uranium oxide in nitric acid in a TFE-fluorocarbon beaker.

41.2.1.3 Add from 1.0 to 1.5 mL of 18 *N* H_2SO_4 to each beaker and dilute to 25 mL. Then proceed with the analysis starting with 41.2.2.4. Plot the absorbances corrected for the blank against the known quantities of silicon taken to obtain a calibration curve. In a typical case, 10 µg of silicon gave a corrected absorbance of about 0.285 in a 5-cm cell. Up to about 150 µg can be handled using a 1-cm absorbance cell and an appropriate calibration curve.

41.2.2 *Analysis:*

41.2.2.1 A blank containing all the reagents in the amounts used in the sample aliquot must be analyzed with the samples. Normally 10 mL of 1 *N* NaOH solution gives an absorbance of 0.030 to 0.050 in this procedure. Most other reagents were found to be nearly silicon-free.

41.2.2.2 Dilute the aliquot of the sample in a 100-mL plastic beaker to 25 mL with water.

41.2.2.3 Add from 1 to 1.5 mL of 18 *N* H_2SO_4.

41.2.2.4 Place the TFE-fluorocarbon beaker containing the sample in a water bath or an oven and heat to 90 to 95°C.

41.2.2.5 Remove the beaker from the water bath or the oven, and add 5 mL of 10 % $(NH_4)_6Mo_7O_{24}$ solution immediately.

41.2.2.6 Adjust the acidity to a pH of 1.2 to 1.3 while the solution is still warm by adding silicon-free NH_4OH or HCl.

41.2.2.7 Allow the sample to stand 10 min to permit the formation of the silico-molybdate complex.

41.2.2.8 Add 10 ml of 5 % $H_2C_2O_4$ solution to the beaker and swirl. Allow the solution to stand for 2 min to decompose any phosphomolybdate.

41.2.2.9 Add 2 mL of reducing mix to the beaker and swirl.

41.2.2.10 Add sufficient 6 *N* HCl immediately to the sample to obtain a 1 *N* acid solution.

41.2.2.11 After all the precipitate is dissolved, transfer the solution to a 50 or 100-mL volumetric flask and dilute to volume with 1 *N* HCl.

41.2.2.12 Determine the absorbance of the solutions in a 5-cm cell at 710 nm.

NOTE 7—The uranyl ion shows a slight absorbance at 710 nm, and samples must be corrected for this absorbance. This is best determined by taking an additional aliquot from the sample solution and treating it as indicated in the procedure up to the point the pH is adjusted with NH_4OH. Any precipitate is dissolved with a minimum of H_2SO_4 and the solution diluted to 50 mL. The absorbance of this solution is used as an additional blank correction.

41.2.2.13 Determine the quantity of silicon in the aliquot from a previously prepared calibration curve.

41.3 *Determination of Phosphorus:*

41.3.1 *Preparation of Calibration Curve*—Using a TFE-fluorocarbon beaker, dissolve sufficient uranium oxide (phosphorus-free) containing 20 g of uranium in HNO_3. Dilute it to 200 mL in a plastic bottle. To separate 10-mL aliquots, add 0, 5, 10, 25, and 50 µg of phosphorus. Analyze by the procedure described below. Plot the absorbances corrected for the blanks against the known quantities of phosphorus to obtain a calibration curve.

41.3.2 *Analysis:*