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Standard Test Method for Reserve Alkalinity of Engine Coolants and Antirusts¹

This standard is issued under the fixed designation D1121; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

This standard has been approved for use by agencies of the Department of Defense.

1. Scope

1.1 This test method covers the determination of the reserve alkalinity of new, unused engine coolants, and liquid antirusts as received, of used or unused aqueous dilutions of the concentrated materials, and of aqueous dilutions of solid antirusts.

1.2 The values stated in SI units are to be regarded as the standard.

1.3 *This standard does not purport to address all of the safety problems, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*²

D1123 Test Methods for Water in Engine Coolant Concentrate by the Karl Fischer Reagent Method

D1176 Practice for Sampling and Preparing Aqueous Solutions of Engine Coolants or Antirusts for Testing Purposes

D1287 Test Method for pH of Engine Coolants and Antirusts

3. Terminology

3.1 *Definition:*

3.1.1 *reserve alkalinity, n*—a term applied to engine coolant concentrates and antirusts to indicate the amount of alkaline inhibitors present in the product.

4. Summary of Test Method

4.1 A 10-mL sample of concentrated coolant, antirust, coolant additive, or an aqueous engine coolant containing these products, is diluted to about 100 mL with water and titrated potentiometrically with 0.100 *N* hydrochloric acid to a pH of 5.5. The volume of acid used is reported to the nearest 0.1 mL. The concentration of coolant or additive in the coolant solution shall be recorded if required, for calculation of depletion of the reserve alkalinity during performance tests.

5. Significance and Use

5.1 Reserve alkalinity is the number of millilitres, to the nearest 0.1 mL of 0.100 *N* hydrochloric acid (HCl) required for the titration to a pH of 5.5 of a 10-mL sample of an undiluted, unused coolant, antirust, or coolant additive, and unused or used solutions thereof.

5.2 Reserve alkalinity is a term applied to engine coolants and antirusts to indicate the amount of alkaline components present in the product. It is frequently used for quality control during production, and values are often listed in specifications. When applied to used solutions, reserve alkalinity gives an approximate indication of the amount of remaining alkaline components. Unfortunately, the term is sometimes misused in that its numerical value is said to be directly related to coolant quality, the higher the number, the better the coolant. ASTM Committee ~~D-15~~D15 on Engine Coolants believes there is a need to correct some misconceptions and place the term in its proper perspective.

5.3 Reserve alkalinity (R.A.) as defined in Section 3 of this method is the number of millilitres of 0.1-*N* hydrochloric acid required to titrate 10 mL of concentrated coolant to a pH of 5.5. The term is not a completely accurate description of the property being measured because “alkalinity” usually refers to the pH range above 7.0.

¹ This test method is under the jurisdiction of ASTM Committee D15 on Engine Coolants and Related Fluids and is the direct responsibility of Subcommittee D15.04 on Chemical Properties.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

5.4 A review of the origin of the term may be helpful in attempting to understand its proper use. When ethylene glycol was first used as an engine coolant, it was uninhibited. The need for inhibition soon became apparent and triethanolamine was incorporated. When solutions of this inhibited glycol were titrated with dilute hydrochloric acid, it was found that the steepest part of the neutralization curve occurred at a pH of about 5.0. Following the introduction of triethanolamine, other buffers, such as borates and phosphates, have been used. Titration to 5.5 was employed with these buffers which have an end point close to 5.5.

5.5 Generally, most metals in an automotive cooling system corrode less in solutions which are mildly alkaline. The commonly used alkaline buffers, borates and phosphates, help to maintain this desirable alkalinity and a stable pH for substantial additions of acid. A well inhibited coolant contains smaller amounts of other inhibitors (in addition to the buffers), to give broad range corrosion protection for all of the metals found in the cooling system. These additional inhibitors may contribute very little to the titration, but they can provide excellent corrosion protection.

5.6 The alkaline inhibitors provide buffering action and neutralize acids introduced into the coolant by exhaust gas leakage, by residual acid cleaner, or by the oxidation of ethylene glycol and propylene glycol. Some inhibitors which contribute little or no reserve alkalinity may give excellent corrosion protection to certain metals but have little capacity to combat acid contamination. In view of this, the magnitude of the R. A. in a coolant is not always a good criterion in determining its potential protective properties.

5.7 In summary, the precaution against the misuse of reserve alkalinity is that the reserve alkalinity of an engine coolant solution is not a dependable measure of its ability to prevent corrosion, nor can it satisfactorily indicate the additional life of the solution.

6. Apparatus

6.1 The pH meter, glass electrode, and calomel electrode, shall be as specified in the section on Apparatus (Section 5) of Test Method D1287.

6.2 It is recognized that many laboratories use a combination electrode system for measuring the reserve alkalinity of engine coolants. It should be noted that results obtained from using a combination electrode system have been found to be statistically different from results obtained when using the calomel and reference electrode system.³

7. Reagents

7.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁴ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

7.2 *Purity of Water*—References to water shall be understood to mean distilled water or water of equal purity. Distilled water shall be boiled thoroughly, or purged with carbon dioxide-free air, to remove carbon dioxide, and shall be protected with a soda-lime tube or the equivalent while cooling and in storage. The pH of the water shall be between 6.2 and 7.2 at 25°C. Take precautions to prevent contamination of the distilled water with traces of the material used for protection against carbon dioxide.

7.3 *Standard Buffer Solution*⁵—The buffer solutions for checking pH assemblies and glass electrodes shall be stored in bottles of chemically resistant glass or polyethylene and shall be prepared from salts sold specifically for use, either singly or in combination, as pH standards. The salts shall be dried for 1 h at 110°C before use.

7.4 *Hydrochloric Acid* (0.100 N)—Prepare and standardize 0.100 N hydrochloric acid (HCl).

7.5 *Phthalate Buffer Solution* (0.05 M, pH = 4.01 at 25°C)—Dissolve 10.21 g of potassium hydrogen phthalate in distilled water and dilute to 1 L.

7.6 *Phosphate Buffer Solution* (0.025 M with respect to each phosphate salt, pH = 6.86 at 25°C)—Dissolve 3.40 g of potassium dihydrogen phosphate (KH₂PO₄) and 3.55 g of anhydrous disodium hydrogen phosphate (Na₂HPO₄) in distilled water and dilute to 1 L.

7.7 *Potassium Chloride Electrolyte*—Prepare a saturated solution of potassium chloride (KCl) in water.

8. Sampling

8.1 Sample the material in accordance with Standard Practice D1176.

9. Preparation of Electrode System

9.1 *Maintenance of Electrodes*—Clean the glass electrodes or combination electrode at frequent intervals (not less than once every week during continual use) in accordance with the manufacturer's instructions. Drain the calomel electrodes at least once each week and fill with KCl electrolyte. Keep the electrolyte level in the calomel electrode above that of the liquid in the titration

³ Refer to ASTM Report RR:D15-1004.

⁴ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

⁵ Commercially available buffer tablets or prepared buffer solutions may be used.