Designation: ~~D3920 – 02 (Reapproved 2007)$^{\varepsilon1}$~~ D3920 – 12

Standard Test Method for Strontium in Water[1]

This standard is issued under the fixed designation D3920; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

~~ε^1 NOTE—Editorial changes were made throughout in September 2007.~~

1. ~~Scope~~ Scope*

1.1 This test method covers the determination of dissolved and total recoverable strontium in water and wastewater by atomic absorption spectroscopy.

1.2 The test method is applicable in the range from 0.1 to 1 mg/L of strontium. The range may be extended by dilution of the original sample.

1.3 Round-robin data were obtained in natural and reagent water matrices. It is the user's responsibility to ensure the validity of the test method for waters of untested matrices.

1.4 The values stated in either SI units or inch-pound units are to be regarded separately as standard. The values stated in each system are mathematical conversions and may not be exact equivalents; therefore, each system shall be used independently of the other.

1.5 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* Specific precautionary statements are given in 8.4 and 8.9.

2. Referenced Documents

2.1 *ASTM Standards:*[2]

D1129 Terminology Relating to Water
D1193 Specification for Reagent Water
D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water
D3370 Practices for Sampling Water from Closed Conduits
D4691 Practice for Measuring Elements in Water by Flame Atomic Absorption Spectrophotometry
D4841 Practice for Estimation of Holding Time for Water Samples Containing Organic and Inorganic Constituents
D5810 Guide for Spiking into Aqueous Samples
D5847 Practice for Writing Quality Control Specifications for Standard Test Methods for Water Analysis

3. Terminology

3.1 *Definitions:* For definitions of terms used in this test method, refer to Terminology D1129.

3.2 *Definitions of Terms Specific to This Standard:*

3.2.1 *total recoverable strontium, n*—an arbitrary analytical term relating to the forms of strontium that are determinable by the digestion procedure described in this test method.

3.2.2 *laboratory control sample, n*—a solution with a certified concentration of the strontium.

4. Summary of Test Method

4.1 Strontium is determined by atomic absorption spectrophotometry. The sample is aspirated into an air-acetylene flame following the addition of lanthanum chloride/potassium chloride solution. Samples containing particulate matter that may clog the aspirator capillary or burner, thus producing inaccurate results, are filtered through a 0.45-µm membrane filter prior to testing.

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[1] This test method is under the jurisdiction of ASTM Committee D19 on Water and is the direct responsibility of Subcommittee D19.05 on Inorganic Constituents in Water.

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[2] For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website. DOI: 10.1520/D3920-02R07E01.

***A Summary of Changes section appears at the end of this standard**

5. Significance and Use

5.1 Although most potable supplies contain little strontium, some well waters in the midwestern part of the United States have levels as high as 39 mg/L.[3]

5.2 This test method affords a reliable means of accurately determining strontium and correcting calcium results obtained by the methods cited in 6.2.

6. Interferences

6.1 Chemical interference caused by silicon, aluminum, and phosphate is controlled by adding lanthanum chloride. Potassium chloride is added to suppress the ionization of strontium.

NOTE 1—A nitrous oxide-acetylene flame has been used successfully by some to remove chemical interferences.

6.2 Strontium chemically resembles calcium and causes a positive error in gravimetric and titrimetric methods for calcium determination.

7. Apparatus

7.1 *Atomic Absorption Spectrophotometer* for use at 460.7 nm. A general guide for the use of flame atomic absorption applications is given in Practice D4691.

NOTE 2—The manufacturer's instructions should be followed for setting instrumental parameters.

7.2 *Strontium Hollow-Cathode Lamp.*

8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available. [4] Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification D1193, Type I, II, and III water. Type I is preferred and more commonly used. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the precision and bias of the test method. Type II water was specified at the time of round robin testing of this test method.

NOTE 3—The user must ensure the type of reagent water chosen is sufficiently free of interferences. The water should be analyzed using the test method.

8.3 *Hydrochloric Acid* (sp gr 1.19)—Concentrated hydrochloric acid (HCl).

8.4 *Lanthanum Chloride/Potassium Chloride Solution*—Dissolve 11.73 g of lanthanum oxide (La_2O_3) in a minimum amount of concentrated hydrochloric acid (approximately 50 mL.). Add 1.91 g of potassium chloride (KCl). Allow solution to cool to room temperature and dilute to 100 mL with water. (**Warning**—Add acid slowly and in small portion to control the reaction rate upon mixing.)

8.5 *Nitric Acid* (sp gr 1.42)—Concentrated HNO_3.

8.6 *Strontium Solution, Stock* (1.0 mL = 1.0 mg Sr)—Dissolve 2.415 g of strontium nitrate ($Sr(NO_3)_2$) in a 1-L flask containing 10 mL of concentrated HCl and 700 mL of water. Dilute to 1000 mL with water. A purchased strontium stock solution of appropriate known purity is also acceptable.

8.7 *Strontium Solution, Standard* (1 mL = 0.010 mg Sr)—Dilute 5.0 mL of strontium solution standard to 500 mL with water.

8.8 *Oxidant*—Air that has been cleaned and dried through a suitable filter to remove oil, water, and other foreign substances, is the usual oxidant.

8.9 *Fuel*—Standard, commercially available acetylene is the usual fuel. Acetone, always present in acetylene cylinders, can be prevented from entering the burner system by replacing the cylinder when the pressure reaches 490 kPa (70 psig). (**Warning**—Purified grade acetylene containing a special proprietary solvent other than acetone should not be used with poly (vinyl chloride) tubing as weakening of the walls may result and cause a potentially hazardous situation.)

9. Sampling

9.1 Collect the samples in accordance with instructions in Practices D3370. The holding time for the samples may be calculated in accordance with Practice D4841.

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[3] *Standard Method for the Examination of Water and Wastewater,* 14th Ed. American Public Health Assn., Washington, DC 20005.

[4] *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.