



Designation: E449 – 08 (Reapproved 2013)

## Standard Test Methods for Analysis of Calcium Chloride<sup>1</sup>

This standard is issued under the fixed designation E449; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reapproval.

*This standard has been approved for use by agencies of the U.S. Department of Defense.*

### 1. Scope

1.1 These test methods cover the analysis of calcium chloride and solutions.

1.2 Procedures are given for the determination of calcium chloride, magnesium chloride, potassium chloride, sodium chloride, and calcium hydroxide. The test methods appear in the following order:

	Sections
Calcium Chloride	8 to 16
Magnesium Chloride, Potassium Chloride, and Sodium Chloride	17 to 26
Calcium Hydroxide	27 to 33

1.3 The values stated in SI units are to be regarded as standard. The values given in parentheses are for information only.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* Specific hazard statements are given in Section 5.

1.5 Review the current Material Safety Data Sheet (MSDS) for detailed information concerning toxicity, first aid procedures, handling, and safety precautions.

### 2. Referenced Documents

#### 2.1 ASTM Standards:<sup>2</sup>

[D345 Test Method for Sampling and Testing Calcium Chloride for Roads and Structural Applications](#)

[D1193 Specification for Reagent Water](#)

[E180 Practice for Determining the Precision of ASTM Methods for Analysis and Testing of Industrial and Spe-](#)

[cialty Chemicals \(Withdrawn 2009\)<sup>3</sup>](#)

[E200 Practice for Preparation, Standardization, and Storage of Standard and Reagent Solutions for Chemical Analysis](#)

[E663 Practice for Flame Atomic Absorption Analysis \(Withdrawn 1997\)<sup>3</sup>](#)

### 3. Significance and Use

3.1 Calcium chloride is available in various forms and purities. A major use is the de-icing and dust control of roads. It is also used in the coal industry for dustproofing and freezeproofing, in foods, in electrolytic cells, and in refrigeration brines. The test methods listed in 1.2 provide procedures for analyzing calcium chloride to determine if it is suitable for its intended use.

### 4. Reagents

4.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, these shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.<sup>4</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

4.2 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean Type II or Type III reagent water conforming to Specification [D1193](#).

### 5. Hazards

5.1 While calcium chloride is a relatively harmless material, some of the reagents used in these methods present possible safety hazards. Potassium cyanide is extremely hazardous and must be handled with great care. In addition to being poisonous, solutions containing cyanide should never be mixed

<sup>1</sup> These test methods are under the jurisdiction of ASTM Committee [D16](#) on Aromatic Hydrocarbons and Related Chemicals and are the direct responsibility of Subcommittee [D16.16](#) on Industrial and Specialty Product Standards.

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<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](http://www.astm.org), or contact ASTM Customer Service at [service@astm.org](mailto:service@astm.org). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

<sup>3</sup> The last approved version of this historical standard is referenced on [www.astm.org](http://www.astm.org).

<sup>4</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

with acids to preclude the release of poisonous hydrogen cyanide gas. Concentrated hydrochloric acid and sodium hydroxide also are hazardous chemicals which may produce serious burns on contact.

## 6. Atomic Absorption Spectrophotometers

6.1 Photometers and photometric practice used in these methods shall conform to Practice E663.

## 7. Sampling

7.1 Sampling of calcium chloride is not within the scope of these test methods. See the appropriate sections of Test Method D345.

7.2 The sample to be analyzed shall be considered to be that sample in a single bottle submitted to the analytical laboratory.

7.3 The size of the sample shall be sufficient to perform all analyses without the reuse of any portion of the sample.

## CALCIUM CHLORIDE

## 8. Scope

8.1 This test method covers the determination of calcium chloride in the range from 0 to 100 %.

## 9. Summary of Test Method

9.1 Calcium in an alkaline solution is titrated with standard ethylenediaminetetraacetate solution, using modified calcein II as an indicator. The color change is from green to purple.  $\alpha$ -hydroxynaphthol blue is also suitable as an indicator, in which case the color change is from red to blue.

## 10. Interferences

10.1 Strontium and other cations not complexed with cyanide at pH of at least 10 will consume ethylenediaminetetraacetate solution and will affect the accuracy of this test method.

## 11. Apparatus

11.1 *Buret*, 50-mL, Class A.

11.2 *Weighing Bottle*, glass-stoppered, 100-mL.

## 12. Reagents

12.1 *Calcium Chloride, Standard Solution* (1 mL = 0.00832 g CaCl<sub>2</sub>)—Weigh 7.500 g of primary standard calcium carbonate (CaCO<sub>3</sub>). Transfer to a 600-mL beaker and add 300 mL of water. Cover with a watch glass and slowly add to the beaker, while stirring, 15 mL of concentrated hydrochloric acid (HCl) delivered from a pipet inserted between the lip of the beaker and the edge of the watch glass. When dissolution of the CaCO<sub>3</sub> is complete, boil gently to expel CO<sub>2</sub>. Cool, and transfer to a 1-L volumetric flask. Dilute to volume with water and mix.

12.2 *Ethylenediaminetetraacetate, Standard Solution* (0.1 mol/L (M))—Dissolve 37.22 g of disodium dihydrogen ethylenediaminetetraacetate dihydrate (EDTA) in water. Transfer to a 1-L volumetric flask, dilute to volume with water, and mix. Standardize as follows: Transfer a 50-mL aliquot of CaCl<sub>2</sub> standard solution (1 mL = 0.00832 g CaCl<sub>2</sub>) to a 500-mL

Erlenmeyer flask and dilute to 200 mL with water. Proceed as directed in 13.2. Calculate the CaCl<sub>2</sub> equivalent of the EDTA solution as follows:

$$\text{Calcium chloride equivalent, g/mL} = 0.416/A \quad (1)$$

where:

A = millilitres of EDTA solution required for the titration of the CaCl<sub>2</sub> solution.

12.3 *Hydrochloric Acid* (1 + 3)—Mix 1 volume of concentrated hydrochloric acid (HCl, sp gr 1.19) and 3 volumes of water.

12.4 *Hydroxylamine Hydrochloride Solution*—(10 %)—Dissolve 10 g of hydroxylamine hydrochloride (NH<sub>2</sub>OH·HCl) in 90 mL of water.

12.5  *$\alpha$ -Hydroxynaphthol Blue*.<sup>5</sup>

12.6 *Modified Calcein Indicator*.<sup>6</sup>

12.7 *Potassium Cyanide* (KCN).

12.8 *Sodium Hydroxide Solution* (80 g/L)—Add slowly 80 g of sodium hydroxide (NaOH) in 300 mL of water stirring constantly. Cool, transfer to a 1-L volumetric flask, dilute to volume with water, and mix.

12.9 *Sugar*, granulated.

## 13. Procedure

13.1 *Solid Samples*—Weigh 100.0 g of sample and wash into a 1000-mL volumetric flask with water. Add 10 mL of HCl (1 + 3) and swirl to dissolve the sample. Cool to room temperature, make to volume with water, and mix. Pipet a 20-mL aliquot into a 500-mL volumetric flask, dilute to volume, and mix. Proceed as in 13.3.

13.2 *Liquid Samples*—Weigh 100.0 g of sample and wash into a 1000-mL volumetric flask with water. Add 10 mL of HCl (1 + 3) and mix. Cool to room temperature, make up to volume with water, and mix. Pipet an aliquot containing about 2 g of CaCl<sub>2</sub> into a 500-mL volumetric flask, dilute to volume, and mix. Appropriate aliquot volumes are indicated in the table below. Interpolate if necessary.

Expected CaCl <sub>2</sub> Concentration, %	Aliquot Size, mL
10	200
20	100
30	75
40	50
50	40

13.3 Pipet a 100-mL aliquot of the solution prepared in 13.1 or 13.2 into a 500-mL Erlenmeyer flask and dilute to about 200 mL with water. Add in order 10 mL of hydroxylamine hydrochloride solution and 3 g of sugar. Swirl to dissolve. Add 40 mL of NaOH solution (12.8) and swirl to mix. Add 0.1 g of KCN, and swirl to dissolve and mix. Add about 0.2 g of calcium indicator.

13.4 Titrate with 0.1 M EDTA solution until the indicator changes from green to purple.

<sup>5</sup> Available from Mallinckrodt Inc., Paris, KY.

<sup>6</sup> Available from the G. Frederick Smith Co., Columbus, OH.