



Designation: **D715 – 86 (Reapproved 2008) D715 – 86 (Reapproved 2014)**

Standard Test Methods for Analysis of Barium Sulfate Pigment¹

This standard is issued under the fixed designation D715; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

- 1.1 These test methods cover the analysis of barium sulfate pigment.
- 1.2 The values stated in SI units are to be regarded as the standard. The values given in parentheses are for information only.
- 1.3 *This standard does not purport to address the safety concerns associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

- 2.1 *ASTM Standards:*²
 - D185 Test Methods for Coarse Particles in Pigments
 - D280 Test Methods for Hygroscopic Moisture (and Other Matter Volatile Under the Test Conditions) in Pigments
 - D1193 Specification for Reagent Water
 - D1208 Test Methods for Common Properties of Certain Pigments
 - E832 Specification for Laboratory Filter Papers

3. Significance and Use

3.1 These test methods are used to determine the purity of barium sulfate and to determine the concentration of known impurities. This information is useful to producers and users as an aid in the manufacture of coatings.

4. Purity of Reagents

4.1 Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.³ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

4.2 Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Type II of Specification **D1193**.

BARIUM SULFATE

5. Reagents

- 5.1 *Ammonium Hydroxide (sp gr 0.90)*—Concentrated ammonium hydroxide (NH₄OH).
- 5.2 *Ammonium Sulfate ((NH₄)₂SO₄)*.
- 5.3 *Hydrochloric Acid (sp gr 1.19)*—Concentrated hydrochloric acid (HCl).
- 5.4 *Hydrochloric Acid (1+1)*—Mix equal volumes of concentrated HCl (sp gr 1.19) and water.

¹ These test methods are under the jurisdiction of ASTM Committee **D01** on Paint and Related Coatings, Materials, and Applications and are the direct responsibility of Committee **D01.31** on Pigment Specifications.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the *standard's* Document Summary page on the ASTM website.

³ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

5.5 *Methyl Red Indicator Solution*—Dissolve 0.2 g of methyl red in 100 mL of methanol, ethanol, or isopropanol.

5.6 *Sodium Carbonate Solution (30 g/L)*—Dissolve 30 g of Na_2CO_3 in water and dilute to 1 L.

6. Procedure

6.1 Weigh to 0.1 mg approximately 0.5 g of the sample into a platinum crucible, add 3 g of Na_2CO_3 , mix thoroughly, and fuse until the melt is clear. Allow the melt to cool and then leach in a platinum dish with hot water until it is entirely disintegrated. Filter on a close-texture paper and wash thoroughly with hot Na_2CO_3 solution (30 g/L).

6.2 Transfer the filter paper containing the insoluble carbonates to a 250-mL beaker and acidify with concentrated HCl (sp gr 1.19). Wash the fusion crucible with HCl (sp gr 1.19) so that no barium is lost. Boil the solution, filter into a 600-mL beaker, and wash well with water. Add methyl red indicator solution, and add NH_4OH (sp gr 0.90) until the solution is slightly alkaline. Add 6 mL of HCl (1+1), and dilute to 300 mL.

6.3 Heat the solution to boiling, and add 5 g of $(\text{NH}_4)_2\text{SO}_4$ dissolved in 40 mL of water. If low-grade material is being analyzed, the $(\text{NH}_4)_2\text{SO}_4$ solution should be added drop by drop from a buret to minimize inclusion. If the barytes is a rather pure product (95 to 99 % BaSO_4), this is not necessary, since the only nonvolatile constituents of the solution will be barium salts. Allow the precipitate of BaSO_4 to digest for 4 h or overnight, and filter through a close-texture paper. Wash the precipitate with as little cold water as possible (consistent with the purity of the precipitate), ignite in an oxidizing atmosphere, and weigh as BaSO_4 .

7. Calculation

7.1 Calculate the percent of BaSO_4 as follows:

$$\text{BaSO}_4, \% = (P/S_1) \times 100 \quad (1)$$

where:

P = BaSO_4 , g, and

S_1 = sample used, g.

iTeh Standards

FERRIC OXIDE

(<https://standards.iteh.ai>)

8. Apparatus

8.1 *Colorimetric Apparatus*—Nessler type or other similar 100-mL colorimetric tubes.

9. Reagents

9.1 *Ammonium Thiocyanate Solution (76.1 g/L)*—Dissolve 76 g of ammonium thiocyanate (NH_4CNS) in water and dilute to 1 L. <https://standards.iteh.ai/catalog/standards/sist/96c46fc1-d426-46a3-ad9b-2d475cc71808/astm-d715-862014>

9.2 *Iron, Standard Solution (100 mL = 0.00002 g Fe)*—Dilute and divide a ferric solution of known iron content so as to obtain 0.4 mg of iron. Dilute the solution to 2 L with water containing 200 mL of iron-free H_2SO_4 .

9.3 *Potassium Permanganate Solution (0.1 g/L)*—Dissolve 0.1 g of potassium permanganate (KMnO_4) in water and dilute to 1 L.

9.4 *Standard Color Solution*—Mix thoroughly 10 parts by volume of NH_4CNS and 90 parts by volume of standard iron solution (100 mL = 0.02 mg Fe). One hundred millilitres of the solution will thus contain 0.000018 mg of Fe.

9.5 *Sulfuric Acid (sp gr 1.84)*—Concentrated sulfuric acid (H_2SO_4).

9.6 *Sulfuric Acid (1+1)*—Carefully mix 1 volume of concentrated H_2SO_4 (sp gr 1.84) with 1 volume of water.

10. Procedure

10.1 Dissolve the soluble portion of a 1-g specimen in H_2SO_4 (1+1), filter, and wash, keeping the volume of the solution under 100 mL. Oxidize any iron present in the filtrate by adding potassium permanganate (KMnO_4) until a faint pink color is obtained. Dilute the solution to 100 mL and pour into a buret graduated in 0.1-mL divisions.

10.2 Pour 100 mL of the standard color solution into a 100-mL colorimetric tube. Into a second colorimetric tube pour 10 mL of concentrated H_2SO_4 (sp gr 1.84) and 10 mL of NH_4CNS , dilute to 60 or 70 mL and then add the test solution from the buret until the depth of color thus produced on dilution to 100 mL exactly matches that of the standard. Record the number of millilitres required.

11. Calculation

11.1 Calculate the percent of Fe_2O_3 as follows:

$$\text{Fe}_2\text{O}_3, \% = [(I \times 1.4298)/S_2] \times 100 \quad (2)$$