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Standard Terminology Relating to Hazard Potential of Chemicals¹

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1. Scope

1.1 This standard is a compilation of terminology used in the area of hazard potential of chemicals. Terms that are generally understood or adequately defined in other readily available sources are not included.

1.2 Although some of these definitions are general in nature, many must be used in the context of the standards in which they appear. The pertinent standard number is given in parentheses after the definition.

1.3 In the interest of common understanding and standardization, consistent word usage is encouraged to help eliminate the major barrier to effective technical communication.

2. Referenced Documents

2.1 ASTM Standards:²

- E476 Test Method for Thermal Instability of Confined Condensed Phase Systems (Confinement Test) (Withdrawn 2008)³
- E487 Test Method for Constant-Temperature Stability of Chemical Materials
- E537 Test Method for The Thermal Stability of Chemicals by Differential Scanning Calorimetry
- E582 Test Method for Minimum Ignition Energy and Quenching Distance in Gaseous Mixtures
- E659 Test Method for Autoignition Temperature of Liquid Chemicals
- E680 Test Method for Drop Weight Impact Sensitivity of Solid-Phase Hazardous Materials
- E681 Test Method for Concentration Limits of Flammability of Chemicals (Vapors and Gases)

¹ This terminology is under the jurisdiction of ASTM Committee E27 on Hazard Potential of Chemicals and is the direct responsibility of Subcommittee E27.01 on Editorial and Nomenclature.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

³ The last approved version of this historical standard is referenced on www.astm.org.

- E698 Test Method for Arrhenius Kinetic Constants for Thermally Unstable Materials Using Differential Scanning Calorimetry and the Flynn/Wall/Ozawa Method
- E771 Test Method for Spontaneous Heating Tendency of Materials (Withdrawn 2001)³
- E918 Practice for Determining Limits of Flammability of Chemicals at Elevated Temperature and Pressure
- E1226 Test Method for Explosibility of Dust Clouds
- E1231 Practice for Calculation of Hazard Potential Figures-of-Merit for Thermally Unstable Materials
- E1232 Test Method for Temperature Limit of Flammability of Chemicals
- E1491 Test Method for Minimum Autoignition Temperature of Dust Clouds
- E1515 Test Method for Minimum Explosible Concentration of Combustible Dusts
- E1981 Guide for Assessing Thermal Stability of Materials by Methods of Accelerating Rate Calorimetry
- E2012 Guide for the Preparation of a Binary Chemical Compatibility Chart
- E2019 Test Method for Minimum Ignition Energy of a Dust Cloud in Air
- E2021 Test Method for Hot-Surface Ignition Temperature of Dust Layers
- E2046 Test Method for Reaction Induction Time by Thermal Analysis

3. Terminology

3.1 Definitions:

adiabatic calorimeter, n —an instrument capable of making calorimetric measurements while maintaining a minimal heat loss or gain between the sample and its environment, which is verifiable by the capability to continuously measure the temperature differential between the sample and its surroundings. **E1981**

adiabatic decomposition temperature rise, $(T)_d, n$ —an estimation of the computed temperature which a specimen would attain if all of the enthalpy (heat) of decomposition reaction were to be absorbed by the sample itself. High values represent high hazard potential. **E1231**

anvil, n —the smooth, hardened surface upon which the test sample or cup containing the sample rests. **E680**

Arrhenius equation— $k = Ze^{-E/RT}$ where k is the specific reaction rate constant in reciprocal minutes for first order, Z is the pre-exponential factor in reciprocal minutes, E is the Arrhenius activation energy in J/mol, R is the gas constant, 8.32 J/mol K, and T is the temperature in kelvin. **E698**

autoignition, n —the ignition of a material commonly in air as the result of heat liberation due to an exothermic oxidation reaction in the absence of an external ignition source such as a spark or flame. **E659**

autoignition temperature, n —the minimum temperature at which autoignition occurs under the specified conditions of test. **E659**

DISCUSSION—Autoignition temperature is also referred to as spontaneous ignition temperature, self-ignition temperature, autogenous ignition temperature, and by the acronyms AIT and SIT. AIT is the lowest temperature at which the substance will produce hot-flame ignition in air at atmospheric pressure without the aid of an external energy source such as spark or flame. It is the lowest temperature to which a combustible mixture must be raised, so that the rate of heat evolved by the exothermic oxidation reaction will over-balance the rate at which heat is lost to the surroundings and cause ignition.

compatibility, adj —the ability of materials to exist in contact without specified (usually hazardous) consequences under a defined scenario. **E2012**

constant-temperature stability (CTS) value, n —the maximum temperature at which a chemical compound or mixture may be held for a 2-h period under the conditions of the test without exhibiting a measurable exothermic reaction. **E487**

cool-flame, n —a faint, pale blue luminescence or flame occurring below the autoignition temperature (AIT). **E659**

DISCUSSION—Cool-flames occur in rich vapor-air mixtures of most hydrocarbons and oxygenated hydrocarbons. They are the first part of the multistage ignition process.

critical half thickness, (a), n —an estimation of the half thickness of a sample in an *unstirred container*, in which the heat losses to the environment are less than the retained heat. This buildup of internal temperature leads to a thermal-runaway reaction. **E1231**

critical temperature, (T_c), n —an estimation of the lowest temperature of an *unstirred container* at which the heat losses to the environment are less than the retained heat leading to a buildup of internal temperature. This temperature buildup leads to a thermal-runaway reaction. **E1231**

DISCUSSION—This description assumes perfect heat removal at the reaction boundary. This condition is not met if the reaction takes place in an insulated container such as when several containers are stacked together or when a container is boxed for shipment. These figures-of-merit underestimate the hazard as a result of this underestimation of thermal conductivity.

deflagration index, (K_{St}), n —maximum dP/dt normalized to a 1.0 m³ volume. It is measured at the optimum dust concentration. K_{St} is defined according to the following cubic relationship:

$$K_{St} = (dP/dt)_{max} V^{1/3}$$

where:

P = pressure, (bar),

t = time, (s),

V = volume, (m³), and

K_{St} = (bar m/s).

E1226

differential scanning calorimetry (DSC), n —a technique in which the difference in energy inputs into a substance and a reference material is measured as a function of temperature, while the substance and the reference material are subjected to a controlled temperature program. **E698**

DISCUSSION—Two modes, power compensation differential scanning calorimetry (power compensation DSC) and heatflux differential scanning calorimetry (heatflux DSC), can be distinguished depending on the method of measurement used.

differential thermal analysis (DTA), n —a technique in which the temperature difference between a substance and reference material is measured as a function of temperature while the substance and the reference material are subjected to a controlled temperature program. **E698**

(dP/dt)_{ex}, n —the maximum rate of pressure rise during the course of a single deflagration. **E1226**

(dP/dt)_{max}, n —maximum value for the rate of pressure increase per unit time reached during the course of a deflagration for the optimum concentration of the dust tested. It is determined by a series of tests over a large range of concentrations. It is reported in bar/s. **E1226**

drop weight, n —that weight which is raised to a selected height and released. This weight does not impact the sample directly; rather it strikes another stationary weight that is in contact with the sample. **E680**

DTA (DSC) curve, n —a record of a thermal analysis where the temperature difference (ΔT) or the energy change (Δq) is plotted on the ordinate and temperature or time is plotted on the abscissa (see Figs. 3 and 4). **E537**

dust concentration, n —the mass of dust divided by the internal volume of the test chamber. **E1491**

extrapolated onset temperature, n —empirically, the temperature found by extrapolating the baseline (prior to the peak) and the leading side of the peak to their intersection (see Fig. 3). **E537**

final temperature (T_{final}), n —the lowest temperature, corrected to a pressure of 101.3 kPa (760 mm Hg, 1013 mbar), at which application of an ignition source causes the vapors of the specimen to ignite under specified conditions of test. **E1232**

flash point, n —the observed system temperature at the end of an exotherm, generally at the temperature where the self-heat rate of the reaction has decreased below the operator-defined slope sensitivity threshold. (1981) **E1232**

general rate law— $dC/dt = k(1 - C)^n$ where C is fractional conversion, t is the time in minutes, and n is the reaction order. **E698**

guide bushing, n —the steel bushing that surrounds, aligns, and holds the stationary intermediate weight in place. **E680**