



Designation: ~~D5291 – 10~~ **D5291 – 10 (Reapproved 2015)**

Standard Test Methods for Instrumental Determination of Carbon, Hydrogen, and Nitrogen in Petroleum Products and Lubricants¹

This standard is issued under the fixed designation D5291; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope*Scope

1.1 These test methods cover the instrumental determination of carbon, hydrogen, and nitrogen in laboratory samples of petroleum products and lubricants. Values obtained represent the total carbon, the total hydrogen, and the total nitrogen.

1.2 These test methods are applicable to samples such as crude oils, fuel oils, additives, and residues for carbon and hydrogen and nitrogen analysis. These test methods were tested in the concentration range of at least 75 mass % to 87 mass % for carbon, at least 9 mass % to 16 mass % for hydrogen, and <0.1 mass % to 2 mass % for nitrogen.

1.3 The nitrogen test method is not applicable to light materials or those containing <0.75 mass % nitrogen, or both, such as gasoline, jet fuel, naphtha, diesel fuel, or chemical solvents.

1.3.1 However, using Test Method D levels of 0.1 ~~mass %~~ mass % nitrogen in lubricants could be determined.

1.4 These test methods are not recommended for the analysis of volatile materials such as gasoline, gasoline-oxygenate blends, or gasoline type aviation turbine fuels.

1.5 The results of these tests can be expressed as mass % carbon, hydrogen or nitrogen.

1.6 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.7 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*²

[D4057 Practice for Manual Sampling of Petroleum and Petroleum Products](#)

[D4177 Practice for Automatic Sampling of Petroleum and Petroleum Products](#)

[D6299 Practice for Applying Statistical Quality Assurance and Control Charting Techniques to Evaluate Analytical Measurement System Performance](#)

3. Summary of Test Methods

3.1 In these test methods, carbon, hydrogen, and nitrogen are determined concurrently in a single instrumental procedure. With some systems, the procedure consists of simply weighing a portion of the sample, placing the portion in the instrument, and initiating the (subsequently automatic) analytical process. In other systems, the analytical process, to some degree, is manually controlled.

3.2 The actual process can vary substantially from instrument to instrument, since a variety of means can be utilized to effect the primary requirements of the test methods. All satisfactory processes provide for the following:

3.2.1 The conversion of the subject materials (in their entirety) to carbon dioxide, water vapor, and elemental nitrogen, respectively, and

3.2.2 The subsequent, quantitative determination of these gases in an appropriate gas stream.

¹ These test methods are under the jurisdiction of ASTM Committee D02 on Petroleum Products-Products, Liquid Fuels, and Lubricants and are the direct responsibility of Subcommittee D02.03 on Elemental Analysis.

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² For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

*A Summary of Changes section appears at the end of this standard

3.3 The conversion of the subject materials to their corresponding gases takes place largely during combustion of the sample at an elevated temperature in an atmosphere of purified oxygen. Here, a variety of gaseous materials are produced, including the following:

- 3.3.1 Carbon dioxide from the oxidation of organic and elemental carbon,
- 3.3.2 Hydrogen halides from organic halides (and organic hydrogen, as required),
- 3.3.3 Water vapor from the oxidation of (the remaining) organic hydrogen and the liberation of moisture,
- 3.3.4 Nitrogen and nitrogen oxides from the oxidation of organic nitrogen, and
- 3.3.5 Sulfur oxides from the oxidation of organic sulfur. In some systems, sulfurous and sulfuric acids can also be obtained from a combination of the sulfur oxides and the water vapor.

3.4 There are several accepted ways of isolating the desired gaseous products and quantitatively determining them. These are as follows:

3.4.1 *Test Method A*^{3,4}—From the combustion product gas stream, oxides of sulfur are removed with calcium oxide in the secondary combustion zone. A portion of the remaining mixed gases is carried by helium gas over a hot copper train to remove oxygen, and reduce NO_x to N₂, over NaOH to remove CO₂, and over magnesium perchlorate to remove H₂O. The remaining elemental nitrogen is measured by the thermal conductivity cell. Simultaneously, but separately from the nitrogen measurement, the carbon and hydrogen selective infrared cells measure the CO₂ and H₂O levels.

3.4.2 *Test Method B*^{4,5}—From the combustion product gas stream (which is cleaned from sulfur oxides, excess oxygen, etc. as in 3.4.1), the remaining CO₂, water vapor, and N₂ are flushed into a mixing chamber and are thoroughly homogenized at a precise volume, temperature, and pressure. After homogenization, the chamber is depressurized to allow the gases to pass through a heated column, where the gases separate as a function of selective retention times. The separation occurs in a stepwise steady-state manner for nitrogen, carbon dioxide, and water.

3.4.3 *Test Method C*^{4,6}—The combustion product gas stream, after full oxidation of component gases, is passed over heated copper to remove excess oxygen and reduce NO_x to N₂ gas. The gases are then passed through a heated chromatographic column to separate and elute N₂, CO₂, and H₂O in that order. The individual eluted gases are measured by a thermal conductivity detector.

3.4.4 *Test Method D*^{4,7}—The organic samples are packed into lightweight containers of oxidizable metal and dropped at preset times into a vertical quartz, inconel, or stainless steel reactor, heated at about ~~1050°C, 1050 °C~~, through which a constant flow of helium is maintained. When the samples are introduced, the helium stream is temporarily enriched with pure oxygen. Flash combustion takes place primed by the oxidation of the container. Quantitative combustion is then achieved by passing the gases over chromium trioxide and cupric oxide. The mixture of the combustion gases is transferred over copper at about ~~640°C~~ (840°C/640 °C (840 °C in a steel reactor) to eliminate the excess of oxygen; then without stopping, it is introduced into the chromatographic column heated to about ~~120°C~~ (50°C/120 °C (50 °C for Flash EA 1112 units). The individual components are then separated by elution in the order nitrogen, carbon dioxide, and water by a dedicated Poropak column (active carbon column for Flash EA 1112 units for nitrogen determination) and measured by a thermal conductivity detector. With dedicated software the percentage of elements present in the sample are calculated. The instrument is calibrated with standard pure organic compounds. K-factors or linear regression can be used for instrument calibration. The typical operator analysis time for a single sample is about 4 min, and the total elapsed time is ~~8 min~~ 8 min.

3.5 In all cases, the concentrations of carbon, hydrogen and nitrogen are calculated as functions of the following:

- 3.5.1 The measured instrumental responses,
- 3.5.2 The values for response per unit mass for the elements (established via instrument calibration), and
- 3.5.3 The mass of the sample.

3.6 A capability for performing these computations automatically can be included in the instrumentation utilized for these test methods.

4. Significance and Use

4.1 This is the first ASTM standard covering the simultaneous determination of carbon, hydrogen, and nitrogen in petroleum products and lubricants.

4.2 Carbon, hydrogen, and particularly nitrogen analyses are useful in determining the complex nature of sample types covered by this test method. The CHN results can be used to estimate the processing and refining potentials and yields in the petrochemical industry.

³ The sole source of supply of the Leco CHN-600 instrument known to the committee at this time is Leco Corporation, 3000 Lakeview Ave., St. Joseph, MI 49085.

⁴ If you are aware of alternative suppliers, please provide this information to ASTM International Headquarters. Your comments will receive careful consideration at a meeting of the responsible technical committee¹, which you may attend.

⁵ The sole source of supply of the Perkin Elmer 240C, 2400 series and CEC 240XA and 440 instruments known to the committee at this time is Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

⁶ The sole source of supply of the Carlo Erba 1106, 1108, and 1500 instruments known to the committee at this time is Carlo Erba Strumentazione, Strada Rivoltana, 20090 Rodano, Milan, Italy.

⁷ The sole source of supply of the Flash EA instruments known to the committee at this time is Thermo Fisher Scientific, Strada Rivoltana, 20090 Milano, Italy.

4.3 The concentration of nitrogen is a measure of the presence of nitrogen containing additives. Knowledge of its concentration can be used to predict performance. Some petroleum products also contain naturally occurring nitrogen. Knowledge of hydrogen content in samples is helpful in addressing their performance characteristics. Hydrogen to carbon ratio is useful to assess the performance of upgrading processes.

5. Apparatus

5.1 Since a variety of instrumental components and configurations can be satisfactorily utilized for these test methods, no specifications are given here with respect to overall system design.

5.2 Functionally, however, the following are specified for all instruments:

5.2.1 The conditions for combustion of the sample must be such that (for the full range of applicable samples) the subject components are completely converted to carbon dioxide, water vapor (except for hydrogen associated with volatile halides and sulfur oxides), and nitrogen or nitrogen oxides. Generally, instrumental conditions that affect complete combustion include availability of the oxidant, temperature, and time.

5.2.2 Representative aliquots of the combustion gases must then be treated:

5.2.2.1 To liberate (as water vapor) hydrogen present as hydrogen halides and sulfur oxyacids, and

5.2.2.2 To reduce (to the element) nitrogen present as nitrogen oxides.

5.2.3 The water vapor and nitrogen so obtained must be included with the materials originally present in these aliquots.

5.2.4 Additional treatment of the aliquots (prior to detection) depends on the detection scheme utilized for the instrument (see **Note 1**).

NOTE 1—These additional treatments can be provided by the instrumental components utilized to satisfy 5.2.2.

5.2.5 The detection system (in its full scope) must determine the analytical gases individually and without interference. Additionally, for each analyte, either:

5.2.5.1 The detectors must provide linear responses with respect to concentration over the full range of possible concentrations from the applicable samples, or

5.2.5.2 The system must include provisions for appropriately evaluating nonlinear responses so that they can be accurately correlated with these concentrations.

5.2.6 Such provisions can be integral to the instrumentation, or they can be provided by (auxiliary) computation schemes.

5.2.7 Lastly, except for those systems where the concentration data are output directly, the instrument must include an appropriate readout device for the detector responses.

5.3 Additionally consumables needed for the analyses include:

5.3.1 *Tin Capsules*, large and small,

5.3.2 *Ceramic Crucibles*,

5.3.3 *Copper Capsules*,

5.3.4 *Tin Plugs*,

5.3.5 *Tin Boats*,

5.3.6 *Copper Plugs*,

5.3.7 *Aluminum Capsules*,

5.3.8 *Combustion Tubes*,

5.3.9 *Adsorption Tubes*,

5.3.10 *Nickel Capsules*, and

5.3.11 *Reduction Tubes*.

5.4 *Analytical Balance*, capable of weighing to the nearest 0.00001 g– 0.00001 g.

5.5 *Syringes or Pipettes*, to transfer the test specimens to capsules.

6. Reagents

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁸ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

6.2 *Calibration Standards*—**Table 1** lists the pure organic compounds most commonly used to calibrate the instruments operated according to 3.4.1 – 3.4.3; other suitable pure compounds can also be used.

6.3 *Carrier and Combustion Gases*:

⁸ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

TABLE 1 Calibration Standards for CHN Instrumental Analysis^{A,B}

Compound	Molecular Formula	Carbon, Mass%	Hydrogen Mass %	Nitrogen Mass %
Acetanilide	C ₈ H ₉ NO	71.09	6.71	10.36
Atropine	C ₁₇ H ₂₃ NO ₃	70.56	8.01	4.84
Benzoic acid	C ₇ H ₆ O ₂	68.84	4.95	...
Cyclohexanone-2,4-dinitrophenylhydrazone	C ₁₂ H ₁₄ N ₄ O ₄	51.79	5.07	20.14
Cystine	C ₆ H ₁₂ N ₂ O ₄ S ₂	29.99	5.03	11.66
Diphenyl	C ₁₂ H ₁₀	93.46	6.54	..
EDTA	C ₁₀ H ₁₆ N ₂ O ₈	41.10	5.52	9.59
Imidazol	C ₃ H ₄ N ₂	52.92	5.92	41.15
Nicotinic acid	C ₆ H ₅ NO ₂	58.53	4.09	11.38
Stearic acid	C ₁₈ H ₃₆ O ₂	75.99	12.76	..
Succinamide	C ₄ H ₈ N ₂ O ₂	41.37	6.94	24.13
Sucrose	C ₁₂ H ₂₂ O ₁₁	42.10	6.48	..
Sulphanilamide	C ₆ H ₈ N ₂ O ₂ S	41.84	4.68	16.27
Triethanol amine	C ₆ H ₁₅ NO ₃	48.30	10.13	9.39

^A The Merck Index, 10th Edition, Merck and Company, Inc., Rahway, New Jersey, 1983.

^B Many of these compounds can be obtained from the National Institute of Standards and Technology as well as other commercial chemical manufacturers, such as Aldrich, Alfa, Kodak and others. See 6.1 for the purity of these reagents.

6.3.1 Oxygen, high purity (99.998 %),

6.3.2 Helium, high purity (99.995 %),

6.3.3 Compressed Air, Nitrogen, or Argon, for operating pneumatic valves, if needed, and

6.3.4 Carbon Dioxide.

6.4 Additional Reagents (as Specified by the Instrument Manufacturer)—This specification covers the reagents utilized to provide for the functional requirements cited in 5.2.2 and 5.2.3. These reagents can vary substantially for different instruments. Consequently, these reagents shall be those recommended by the manufacturer. Specifically, these reagents will be for:

6.4.1 Test Method A^{3,4}:

6.4.1.1 Sodium Hydroxide Coated Silica,

6.4.1.2 Quartz Wool,

6.4.1.3 Magnesium Perchlorate,

6.4.1.4 Copper Turnings,

6.4.1.5 Coated Calcium Oxide (Furnace Reagent),

6.4.1.6 Nitrogen Catalyst, and

6.4.1.7 Magnesium Oxide,^{4,9} for liquids.

6.4.2 Test Method B^{4,5}:

6.4.2.1 EA 1000 Reagent,^{4,10}

6.4.2.2 Silver Tungstate on MgO,

6.4.2.3 Silver Vanadate,

6.4.2.4 Quartz Wool,

6.4.2.5 Silver Gauze,

6.4.2.6 Copper Oxide,¹¹

6.4.2.7 Tungstic Oxide,

6.4.2.8 Cobalt Oxide,

6.4.2.9 Copper Powder,

6.4.2.10 Sodium Hydroxide Coated Silica,

6.4.2.11 Alumina,

6.4.2.12 Magnesium Perchlorate, and

6.4.2.13 Platinum Gauze.

6.4.3 Test Method C and D^{4,6,7}:

6.4.3.1 Quartz Wool,

6.4.3.2 Chromic Oxide (oxidation catalyst),

⁹ The sole source of supply of Com-aid, a registered trademark of Leco, known to the committee at this time is Leco Corporation, 3000 Lakeview Ave., St Joseph, MI 49085.

¹⁰ The sole source of supply of the EA 1000 Reagent, a registered trademark of Perkin Elmer, known to the committee at this time is Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

¹¹ The sole source of supply of Cuprox, a registered trademark of Perkin Elmer, known to the committee at this time is Perkin Elmer Corporation, Main Ave., Norwalk, CT 06856.

- 6.4.3.3 *Silver Coated Cobalt Oxide*,
- 6.4.3.4 *Reduced Copper* (reduction catalyst),
- 6.4.3.5 *Magnesium Perchlorate*,
- 6.4.3.6 *Molecular Sieve*, 3A $\frac{1}{16}$ in. (1.6 mm),
- 6.4.3.7 *Sodium Hydroxide Coated Silica*,
- 6.4.3.8 Chromosorb, (Absorber,¹² for liquid samples; calcined silica), and
- 6.4.3.9 *Copper Grains*.

6.5 *Quality Control (QC) Samples*, preferably are portions of one or more liquid petroleum materials that are stable and representative of the samples of interest. These QC samples can be used to check the validity of the testing process as described in Section 12.

7. Sampling, Test Specimens, and Test Units

- 7.1 *Laboratory Sample*—Take a representative sample as specified in Practices D4057 or D4177.
- 7.2 *Test Specimen*—Take an aliquot from the laboratory sample for analysis as follows:
 - 7.2.1 *Preparation*—Warm viscous samples until they are fluid, and shake for ~~5 s~~ 5 s.
 - 7.2.2 *Transfer*—Use any convenient, clean syringe or pipet to transfer test specimens to the capsules as described in Section 9.

8. Preparation of Apparatus

- 8.1 Prepare the instrumental system (in its entirety) in strict accordance to the manufacturer's instructions.
- 8.2 Calibrate the system using acetanilide or other suitable calibration standard mentioned in Table 1, using the standardization procedure specified by the manufacturer. This procedure, with respect to the actual analytical process involved, must not differ from that specified for the samples.
- 8.3 *Tubes/Columns Preparation* (see Note 2)—Clean all quartz and glass parts prior to use with soap and water followed by acetone, and dry fingerprints must be removed with a grease solvent such as acetone, prior to insertion of the tubes into the furnace. Handle the tubes using appropriate gloves, such as lint-free cotton gloves, that will not leave fingerprints.

NOTE 2—All combustion tubes and absorption tubes need to be periodically replaced after 50 to 300 sample runs. The exact intervals of change should be determined as recommended by the manufacturer.

- 8.3.1 *Test Method A*^{3,4}:
 - 8.3.1.1 *Combustion Tube*—Pack 5 cm of quartz wool in the bottom of the tube on the primary (inlet) side. Set a ceramic crucible on top of the quartz wool on the primary side. Pack 3.1 cm of quartz wool in the bottom of the tube on the secondary (outlet) side. Fill ~~6.3 cm~~ 6.3 cm of furnace reagent on top of the quartz wool. Pack ~~4.9 cm~~ 11.9 cm of quartz wool on top of the furnace reagent.
 - 8.3.1.2 *Reduction Tube*—Insert a small copper plug in the bottom. Fill up to ~~8.8 cm~~ 8.8 cm of *N*-catalyst on top of the copper plug. Add ~~4.8 cm~~ 13.8 cm of copper turnings on top of the *N*-catalyst.

- 8.3.2 *Test Method B*^{4,5}:
 - 8.3.2.1 *Combustion Tube*—Precondition EA 1000^{4,10} and silver tungstate on magnesium oxide by preheating it at ~~900°C~~ 900 °C for ~~10 min~~ 30 min to ~~30 min~~ 30 min. Roll a strip of silver gauze to fit into the combustion tube. Clean the rolled gauze by either ultrasonically washing it in detergent and water, and drying it, or by passing it rapidly over a flame several times until the smoke disappears. Slide ~~2.5 cm~~ 2.5 cm of quartz wool into the tube from the outlet so that the end of the wool meets with the indentations in the tube. Add ~~5 cm~~ 5 cm of EA 1000^{4,10} reagent. Top with a small wad of quartz wool. Add ~~5 cm~~ 5 cm of silver tungstate on magnesium oxide. Top with a small wad of quartz wool. Add ~~2.5 cm~~ 2.5 cm of silver vanadate. Top with a small wad of quartz wool. Slide a roll of silver gauze into the combustion tube, so that there is about ~~1.2 cm~~ 1.2 cm empty space left in the tube. Insert ~~1.2 cm~~ 1.2 cm of quartz wool into the inlet end of the combustion tube, and another into the vial receptacle, and place the receptacle into the combustion tube.

- 8.3.2.2 *Reduction Tube*—Insert a small wad of quartz wool into the inlet end of the tube to where the tube widens. Then insert a conditioned roll of silver gauze. Fill the tube from the outlet end. Fill with copper powder, occasionally tapping till the copper column is 23.8 cm deep. Insert a small wad of quartz wool, then add ~~1.2 cm~~ 1.2 cm of cuprox (copper oxide), and add another small wad of quartz wool. Insert the copper plug into the outlet end of the reduction tube.

- 8.3.3 *Test Method C and D*^{4,6,7}:
 - 8.3.3.1 *Combustion Reactor*—First place a ~~2 mm~~ 2 mm plug of quartz wool in the bottom of the tube. Fill the tube with ~~50 mm~~ 50 mm of silver coated cobalt oxide. Insert a 10 mm thick plug of quartz wool. Add ~~120 mm~~ 120 mm of chromic oxide. Insert a ~~10 mm~~ 10 mm thick plug of quartz wool. Slide the combustion tube into the combustion furnace and secure with the O-ring fittings.

¹² The sole source of supply of Chromosorb, a registered trademark of Carlo Erba Strumentazione, known to the committee at this time is Carlo Erba Strumentazione, Strada Rivoltana, 20090 Rodano, Milan, Italy.

8.3.3.2 *Reduction Reactor*—First place a ~~5 mm~~ 5 mm plug of quartz wool in the bottom of the tube. Fill the tube with reduced copper. Insert a ~~10 mm~~ 10 mm thick plug of quartz wool. Slide the reduction tube into the reduction furnace and secure with the O-ring fittings.

8.3.3.3 *Water Trap*—First place a ~~10 mm~~ 10 mm plug of quartz wool in the bottom of the tube. Fill the tube with 3A ~~1/16 in.~~ (1.6 mm) in. (1.6 mm) molecular sieve or magnesium perchlorate.^{4,10} Insert a ~~10 mm~~ 10 mm thick plug of quartz wool. Secure the tube in the water trap bracket.

8.3.3.4 *Carbon Dioxide Trap*—First place a ~~10 mm~~ 10 mm plug of quartz wool in the bottom of the tube. Fill the tube with ~~50 mm~~ 50 mm of magnesium perchlorate. Add ~~130 mm~~ 130 mm of sodium hydroxide coated silica.^{4,9} Add 50 mm of magnesium perchlorate. Insert a ~~10 mm~~ 10 mm thick plug of quartz wool. Secure the tube in the carbon dioxide trap bracket.

8.3.3.5 *Scrubber Tube*—Place a small amount of cotton or quartz wool on the bottom and then fill it with about ~~1 inch~~ 1 in. of silica gel with indicator (magnesium perchlorate), about ¼ of the way fill with molecular sieve, and soda lime to the fill to the top. Ascarite can also be used for CO₂ removal. Magnesium perchlorate can be used for water removal. If using soda lime, one drop of water should be added to the soda lime to activate it.

NOTE 3—If the gases of stated purity in 9.4.3 are used, it may not be necessary to use the scrubber.

9. Procedure

9.1 Test Method A^{3,4}:

9.1.1 Accurately to within ± 1 ~~mg~~ ± 1 mg weigh out about ~~5050 mg~~ 200 mg to ~~200 mg~~ 200 mg of homogenous test specimen taken according to Practices **D4057** or **D4177**, in a capsule. Crimp the capsule with a forceps.

9.1.1.1 Weigh the solid samples in tin capsules.

9.1.1.2 Weigh the liquid samples in copper capsules and saturate with an absorbent (magnesium oxide¹¹) within the copper capsules. Seal the copper capsules with tin plugs. These precautions will induce a slower combustion to prevent any back flashes with the light samples, or incomplete sample combustion.

9.1.2 The following typical settings may be used.

9.1.2.1 Temperatures:

Combustion furnace; primary zone	950°C
Combustion furnace; secondary zone	950°C
Catalyst heater	750°C
Oven chamber	53°C
<hr/>	
Combustion furnace; primary zone	950 °C
Combustion furnace; secondary zone	950 °C
Catalyst heater	750 °C
Oven chamber	53 °C

9.1.2.2 Oxygen and helium carrier gas pressures should be ~~40 psi~~ 40 psi each.

9.1.2.3 Typical settings for gas flows are:

Helium	400 cm ³ /min (normal flow)
	70 cm ³ /min (conservation flow)
Oxygen	7 dm ³ /min
Air	6 dm ³ /min

9.1.3 To start the sequence of analysis, run two to four blank capsules, followed by five calibration standards. The results must agree within ± 10 % for blanks, and ± 1 % of the theoretical value for calibration standards.

9.1.4 Hold all the results in the instrument memory, and recall the best values for the calibration. Enter these results and the known calibrant concentrations into the microprocessor to generate a one point calibration curve.

9.1.5 Analyze the calibration standard again to check the new curve. The results should differ by less than 1 %, or repeat the calibration.

9.1.6 Combust the encapsulated sample in a manner similar to the calibrant. Depending on the sample matrices, set varying oxygen flow rates and combustion times to overcome incomplete combustion.

9.2 Test Method B^{4,5}:

9.2.1 Accurately to within ± 0.02 ~~mg~~ ± 0.02 mg weigh out ~~22 mg~~ 4 mg to ~~4 mg~~ 4 mg of a homogeneous test specimen taken according to Practices **D4057** or **D4177**, into the vial. Pinch the center of the vial with a forceps and fold it in half. Flatten the vial with the tweezers, and fold the vial in thirds.

9.2.2 The following typical settings may be used.

Combustion temperature	975°C
Combustion temperature	975 °C
Reduction temperature	640°C
Reduction temperature	640 °C
Detector oven temperature	80 to 84°C
Detector oven temperature	80 °C to 84 °C
Helium	137.9 kPa
Oxygen	110.3 kPa
Air, nitrogen, or argon	413.7 kPa