



Designation: **D3352—08a D3352 – 15**

Standard Test Method for Strontium Ion in Brackish Water, Seawater, and Brines¹

This standard is issued under the fixed designation D3352; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope—Scope*

1.1 This test method covers the determination of soluble strontium ion in brackish water, seawater, and brines by atomic absorption spectrophotometry.

1.2 Samples containing from 5 to 2100 mg/L of strontium may be analyzed by this test method.

1.3 The values stated in SI units are to be regarded as standard. ~~No other units of measurement are included in this~~ The values given in parentheses are mathematical conversions to inch-pound units that are provided for information only and are not considered standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*²

[D1129 Terminology Relating to Water](#)

[D1193 Specification for Reagent Water](#)

[D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water](#)

[D3370 Practices for Sampling Water from Closed Conduits](#)

[D5810 Guide for Spiking into Aqueous Samples](#)

[D5847 Practice for Writing Quality Control Document Specifications for Standard Test Methods for Water Analysis](#)

3. Terminology

3.1 *Definitions*—For definitions of terms used in this test method, refer to Terminology [D1129](#).

4. Summary of Test Method

4.1 This test method is dependent on the fact that metallic elements, in the ground state, will absorb light of the same wavelength they emit when excited. When radiation from a given excited element is passed through a flame containing ground state atoms of that element, the intensity of the transmitted radiation will decrease in proportion to the amount of the ground state element in the flame. A hollow cathode lamp whose cathode is made of the element to be determined provides the radiation. The metal atoms³ to be measured are placed in the beam of radiation by aspirating the specimen into an oxidant-fuel flame. A monochromator isolates the characteristic radiation from the hollow cathode lamp and a photosensitive device measures the attenuated transmitted radiation.

4.2 Since the variable and sometimes high concentrations of matrix materials in the waters and brines affect absorption differently, it is difficult to prepare standards sufficiently similar to the waters and brines. To overcome this difficulty, the method

¹ This test method is under the jurisdiction of ASTM Committee [D19](#) on Water and is the direct responsibility of Subcommittee [D19.05](#) on Inorganic Constituents in Water. Current edition approved Nov. 15, 2008/Feb. 1, 2015. Published November 2008/April 2015. Originally approved in 1974. Last previous edition approved in 2008 as [D3352—08](#). [D3352 – 08a](#). DOI: [10.1520/D3352-08A-10.1520/D3352-15](#).

² For referenced ASTM standards, visit the ASTM website, [www.astm.org](#), or contact ASTM Customer Service at [service@astm.org](#). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

³ For additional information on atomic absorption, see the following references: Angino, E. E., and Billings, G. K., *Atomic Absorption Spectrophotometry in Geology*, Elsevier Publishing Co., New York, N.Y., 1967; Dean, J. A., and Rains, T. C., Editors, *Flame Emission and Atomic Absorption Spectrometry Vol 1 – Theory*, Marcel Dekker, New York, NY, 1969; For additional information on atomic absorption, see the following references: Angino, E. E., and Billings, G. K., *Atomic Absorption Spectrophotometry in Geology*, Elsevier Publishing Co., New York, N.Y., 1967; Dean, J. A., and Rains, T. C., Editors, *Flame Emission and Atomic Absorption Spectrometry Vol 1 – Theory*, Marcel Dekker, New York, NY, 1969.

*A Summary of Changes section appears at the end of this standard

of additions is used in which three identical samples are prepared and varying amounts of a standard added to two of them. The three samples are then aspirated, the concentration readings recorded, and the original sample concentration calculated.

5. Significance and Use

5.1 This test method⁴ can be used to determine strontium ions in brackish water, seawater, and brines.

6. Interferences

6.1 The chemical suppression caused by silicon, aluminum, and phosphate is controlled by adding lanthanum. The lanthanum also controls ionization interference.

7. Apparatus

7.1 *Atomic Absorption Spectrophotometer—Spectrophotometer*—The instrument shall consist of atomizer and burner, suitable pressure-regulating devices capable of maintaining constant oxidant and fuel pressure for the duration of the test, a hollow cathode lamp for each metal to be tested, an optical system capable of isolating the desired line of radiation, an adjustable slit, a photomultiplier tube or other photosensitive device as a light measuring and amplifying device, and a read-out mechanism for indicating the amount of absorbed radiation.

7.1.1 *Multi-Element Hollow Cathode Lamps* are available and have been found satisfactory.

7.2 *Pressure-Reducing Valves*—The supplies of fuel and oxidant shall be maintained at pressures somewhat higher than the controlled operating pressure of the instrument by suitable valves.

8. Reagents and Materials

8.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society,⁵ where such specifications are available. Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

8.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification **D1193**, Type I. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the precision and bias of the test method. Type III water was specified at the time of round robin testing of this test method.

8.3 *Filter Paper*—Purchase suitable filter paper. Typically the filter papers have a pore size of 0.45- μm membrane. Material such as fine-textured, acid-washed, ashless paper, or glass fiber paper are acceptable. The user must first ascertain that the filter paper is of sufficient purity to use without adversely affecting the bias and precision of the test method.

8.4 *Lanthanum Solution (5 % La)*—Wet 58.65 g of lanthanum oxide (La_2O_3) with water. Add 250 mL of concentrated hydrochloric acid (sp gr 1.19) very slowly until the material is dissolved. Dilute solution to 1 litre with water.

⁴ Additional information is contained in the following references: Fletcher, G. F., and Collins, A. G., "Atomic Absorption Methods of Analysis of Oilfield Brines: Barium, Calcium, Copper, Iron, Lead, Lithium, Magnesium, Manganese, Potassium, Sodium, Strontium, and Zinc," U.S. Bureau of Mines, Report of Investigations 7861, 1974, 14 pp.; Collins, A. G., *Geochemistry of Oilfield Waters*, Elsevier Publishing Co., Amsterdam, The Netherlands, 1975. Additional information is contained in the following references: Fletcher, G. F., and Collins, A. G., "Atomic Absorption Methods of Analysis of Oilfield Brines: Barium, Calcium, Copper, Iron, Lead, Lithium, Magnesium, Manganese, Potassium, Sodium, Strontium, and Zinc," U.S. Bureau of Mines, Report of Investigations 7861, 1974, 14 pp.; Collins, A. G., *Geochemistry of Oilfield Waters*, Elsevier Publishing Co., Amsterdam, Netherlands, 1975.

⁵ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopoeia and National Formulary*, U.S. Pharmacopoeial Convention, Inc. (USPC), Rockville, MD.

TABLE 1 Compositions of Artificial Brine Samples

| Sample No. | g/L | | | |
|-------------------|-------|-------|-------|-------|
| | 1 | 2 | 3 | 4 |
| Sr | 0.060 | 0.100 | 1.600 | 2.100 |
| NaCl | 24.0 | 170.0 | 80.0 | 200.0 |
| KCl | 0.5 | 2.0 | 1.5 | 3.0 |
| KBr | 1.0 | 2.0 | 2.0 | 2.0 |
| KI | 0.1 | 0.5 | 0.5 | 1.0 |
| CaCl ₂ | 1.5 | 3.0 | 2.0 | 5.0 |
| MgCl ₂ | 4.5 | 5.0 | 2.0 | 1.0 |
| BaCl ₂ | 0.05 | 1.0 | 0.5 | 0.5 |

8.5 *Strontium Solution, Standard* (1 mL = 1 mg Sr)—Dissolve 2.415 g of strontium nitrate [Sr(NO₃)₂] in 10 mL of concentrated hydrochloric acid (sp gr 1.19) and about 700 mL of water. Dilute solution to 1 L with water. One millilitre of this solution contains 1 mg of strontium. ~~Alternatively, certified strontium stock solutions are commercially available through chemical supply vendors and may be used.~~ A purchased strontium stock solution of appropriate known purity is also acceptable.

8.6 *Oxidant, for Atomic Absorption Spectrophotometer:*

8.6.1 *Air*, which has been cleaned and dried through a suitable filter to remove oil, water, and other foreign substances, is the usual oxidant.

8.6.2 *Nitrous Oxide* may be required as an oxidant for refractory-type metals.

8.7 *Fuel, for Atomic Absorption Spectrophotometer:*

8.7.1 *Acetylene*—Standard, commercially available acetylene is the usual fuel. Acetone, always present in acetylene cylinders, can be prevented from entering and damaging the burner head by replacing a cylinder which only has ~~100 psig~~ 689 kPa (100 psi) of acetylene remaining.

9. Sampling

9.1 Collect the sample in accordance with Practices **D3370**.

10. Procedure

10.1 Strontium is determined at the 460.7-nm wavelength with an air-acetylene flame.

10.2 *Preliminary Calibration*—Using micropipets prepare standard strontium solutions containing 1 to 10 mg/L of strontium using the standard strontium solution (8.5) and 50-mL volumetric flasks. Before making up to volume, add to each of these and to a blank, 5 mL of the lanthanum ~~solution~~ solution (8.4). Analyze at least three working standards containing concentrations of strontium that bracket the expected sample concentration, prior to analysis of samples, to calibrate the instrument. Aspirate these standards and the blank (for background setting) and adjust the curvature controls, if necessary, to obtain a linear relationship between absorbance and the actual concentration of the standards.

10.3 Transfer an aliquot of water or brine (previously filtered (8.3) through a 0.45-µm filter) to a 50-mL volumetric flask. The specific gravity of the water or brine can be used to estimate the strontium content of the sample and, thereby, serve as a basis for selecting the aliquot size that will contain about 0.1 mg of strontium. **Fig. 1** shows the relationship between strontium concentration and specific gravity for some oilfield brines from the Smackover formation. The concentrations of strontium in the Smackover brines will not necessarily correlate with the concentrations found in other formations. Therefore, the user of this test method may find it necessary to draw a similar curve for brine samples taken from other formations. Add 5 mL of the lanthanum stock ~~solution~~ solution (8.4), dilute to volume, and aspirate. Calculate the approximate sample concentration from the preliminary calibration readings, and determine the aliquot size that will contain about 0.1 mg of strontium.

10.4 Transfer equal aliquots containing about 0.1 mg of strontium to three 50-mL volumetric flasks. Add no strontium standard to the first flask. With a micropipet add 0.1 mg to the second and 0.2 mg to the third.

10.5 Add 5 mL of the lanthanum solution (8.4) to each of the three flasks and dilute to volume. Aspirate and record the direct concentration of the sample, if this capability is provided with the instrument or record the absorbance readings for each sample.

11. Calculation

11.1 Calculate the concentration of strontium ion in the original sample in milligrams per litre as follows:

$$\text{Strontium concentration, mg/L} = \frac{V_1(A_s \times C_{\text{std}})}{V_2(A_{\text{std}} - A_s)}$$

where:

V_1 = volume of the diluted samples, mL,

V_2 = volume of the original sample, mL,

A_s = absorbance of dilute sample,

TABLE 2 Determination of Precision and Bias of Strontium Ion

| Amount Added, mg/L | Amount Found, mg/L | S_o | S_T | ± Bias | Statistically Significant (95 % Confidence Level) |
|--------------------|--------------------|-------|-------|--------|---|
| 60 | 63.48 | 2.96 | 8.49 | + 5.8 | yes |
| 100 | 99.5 | 4.12 | 11.84 | -0.5 | no |
| 1600 | 1665.6 | 54.87 | 157.3 | + 4.1 | no |
| 2100 | 2167.2 | 71.12 | 203.9 | + 3.2 | no |