Designation: E50 – 11 (Reapproved 2016)

Standard Practices for Apparatus, Reagents, and Safety Considerations for Chemical Analysis of Metals, Ores, and Related Materials[1]

This standard is issued under the fixed designation E50; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ε) indicates an editorial change since the last revision or reapproval.

This standard has been approved for use by agencies of the U.S. Department of Defense.

1. Scope

1.1 These practices cover laboratory apparatus and reagents that are required for the chemical analysis of metals, ores and related materials by standard methods of ASTM. Detailed descriptions of recommended apparatus and detailed instructions for the preparation of standard solutions and certain nonstandardized reagents will be found listed or specified in the individual methods of analysis. Included here are general recommendations on the purity of reagents and protective measures for the use of hazardous reagents.

1.2 These recommendations are intended to apply to the ASTM methods of chemical analysis of metals when definite reference is made to these practices, as covered in Section 4.

1.3 The values stated in inch-pound units are to be regarded as standard. The values given in parentheses are mathematical conversions to SI units that are provided for information only and are not considered standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of whoever uses this standard to consult and establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* Specific hazards are given in Section 8.

NOTE 1—The use of the verb "shall" (with its obligatory third person meaning) in this standard has been confined to those aspects of laboratory safety where regulatory requirements are known to exist. Such regulations, however, are beyond the scope of these practices.

2. Referenced Documents

2.1 *ASTM Standards:*[2]

D1193 Specification for Reagent Water

E1 Specification for ASTM Liquid-in-Glass Thermometers

E77 Test Method for Inspection and Verification of Thermometers

E100 Specification for ASTM Hydrometers

E126 Test Method for Inspection, Calibration, and Verification of ASTM Hydrometers

E287 Specification for Laboratory Glass Graduated Burets

E288 Specification for Laboratory Glass Volumetric Flasks

E438 Specification for Glasses in Laboratory Apparatus

E542 Practice for Calibration of Laboratory Volumetric Apparatus

E694 Specification for Laboratory Glass Volumetric Apparatus

E969 Specification for Glass Volumetric (Transfer) Pipets

E1044 Specification for Glass Serological Pipets (General Purpose and Kahn)

E1621 Guide for Elemental Analysis by Wavelength Dispersive X-Ray Fluorescence Spectrometry

3. Terminology

3.1 For definitions of terms used in these practices, refer to Terminology E135.

4. Significance and Use

4.1 The inclusion of the following paragraph, or a suitable equivalent, in any standard (preferably after the section on Scope) is due notification that the apparatus and reagents required in that standard are subject to the recommendations set forth in these practices.

"*Apparatus and Reagents*—Apparatus and reagents required for each determination are listed in separate sections preceding the procedure. Apparatus, standard solutions, and certain other reagents shall conform to the requirements prescribed in ASTM Practices E50, for Apparatus, Reagents, and Safety Considerations for Chemical Analysis of Metals, Ores, and Related Materials."

4.2 It is assumed that the users of these practices will be trained analysts capable of performing common laboratory procedures skillfully and safely. It is expected that work will be performed in a properly-equipped laboratory.

5. Purity of Water and Reagents

5.1 *Purity of Water*—Unless otherwise indicated, references to water shall be understood to mean reagent water conforming

[1] These practices are under the jurisdiction of ASTM Committee E01 on Analytical Chemistry for Metals, Ores, and Related Materials and are the direct responsibility of Subcommittee E01.20 on Fundamental Practices.

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[2] For referenced ASTM standards, visit the ASTM website, www.astm.org, or contact ASTM Customer Service at service@astm.org. For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

TABLE 1 Chemical Reagents Specified in ASTM Methods for Chemical Analysis of Metals

Name	Formula
* Acetic acid	CH_3COOH
Acetone	CH_3COCH_3
Acetylacetone (2,4-pentanedione)	$CH_3COCH_2COCH_3$
Alizarin-Red-S	$C_6H_4COC_6H$-1,2-$(OH)_2$-3-SO_3NaCO
Aluminon (aurintricarboxylic acid-ammonium salt)	$(4\text{-}HOC_6H_3\text{-}3\text{-}COONH_4)_2C{:}C_6H\text{-}3\text{-}(COONH_4){:}O$
Aluminum metal (99.9 % min)	Al
* Aluminum metal (sheet or rolled foil)	Al
Aluminum ammonium sulfate	$Al_2(NH_4)_2(SO_4)_4{\cdot}24H_2O$
Aluminum nitrate	$Al(NO_3)_3{\cdot}9H_2O$
Aluminum sulfate	$Al_2(SO_4)_3{\cdot}18H_2O$
Aluminum oxide, fused (Alundum)	
1-Amino-2-naphthol-4-sulfonic acid	$NH_2C_{10}H_5(OH)SO_3H$
Ammonium acetate	CH_3COONH_4
Ammonium benzoate	$C_6H_5COONH_4$
Ammonium bifluoride	NH_4FHF
Ammonium bisulfate	NH_4HSO_4
Ammonium bisulfite	NH_4HSO_3
Ammonium carbonate	$(NH_4)_2CO_3$
* Ammonium chloride	NH_4Cl
* Ammonium citrate	$CH_2(COONH_4)C(OH)(COOH)CH_2COONH_4$
Ammonium fluoride	NH_4F
* Ammonium hydroxide[A]	NH_4OH
Ammonium iodide	NH_4I
Ammonium molybdate	$(NH_4)_2MoO_4$
* Ammonium heptamolybdate tetrahydrate	$(NH_4)_6Mo_7O_{24}{\cdot}4H_2O$
Ammonium nitrate	NH_4NO_3
* Ammonium oxalate	$NH_4OCOCOONH_4{\cdot}H_2O$
* Ammonium phosphate, dibasic (diammonium acid phosphate)	$(NH_4)_2HPO_4$
* Ammonium persulfate (ammonium peroxydisulfate)	$(NH_4)_2S_2O_8$
* Ammonium sulfate	$(NH_4)_2SO_4$
* Ammonium tartrate	$NH_4OCO(CHOH)_2COONH_4$
Ammonium thiocyanate	NH_4SCN
Ammonium vanadate	NH_4VO_3
Antimony metal (powder)	Sb
Antimony trichloride	$SbCl_3$
* Arsenic trioxide	As_2O_3
Asbestos (for use with Gooch crucible)	
Barium Chloride	$BaCl_2{\cdot}2H_2O$
Barium diphenylamine sulfonate	$(C_6H_5NHC_6H_4\text{-}4\text{-}SO_3)_2Ba$
* Benzoic acid	C_6H_5COOH
α-Benzoin oxime (benzoin anti-oxime)	$C_6H_5CHOHC{:}NOHC_6H_5$
Beryllium sulfate	$BeSO_4{\cdot}4H_2O$
Bismuth metal (99.9 % min)	Bi
Boric acid	H_3BO_3
Bromocresol green (3',3″,5',5″-tetrabromo-*m*-cresolsulfonephthalein)	$C_6H_4SO_2OC(C_6H\text{-}3,5\text{-}Br_2\text{-}2\text{-}CH_3\text{-}4\text{-}OH)_2$
Bromocresol purple (5',5″-Dibromo-*o*-cresolsulfonephthalein)	$C_6H_4SO_2OC(C_6H_2\text{-}3\text{-}CH_3\text{-}5\text{-}Br\text{-}4\text{-}OH)_2$
Bromine (liquid)	Br_2
Bromophenol blue (3',3″,5',5″-tetrabromophenolsulfonephthalein)	$C_6H_4SO_2OC(C_6H_2\text{-}3,5\text{-}Br_2\text{-}4\text{-}OH)_2$
1-Butanol	$CH_3CH_2CH_2CH_2OH$
Butyl acetate (normal)	$CH_3COOCH_2CH_2CH_2CH_3$
* Cadmium chloride	$CdCl_2{\cdot}2½\ H_2O$
Cadmium chloride, anhydrous	$CdCl_2$
* † Calcium carbonate (low-boron)	$CaCO_3$
Carbon dioxide (gas)	CO_2
Carbon dioxide (solid)	CO_2
Carbon tetrachloride	CCl_4
Carminic acid	$1,3,4\text{-}(HO)_3\text{-}2\text{-}C_6H_{11}O_6C_6COC_6H\text{-}5\text{-}COOH\text{-}6\text{-}OH\text{-}8\text{-}CH_3CO$
* Chloroform	$CHCl_3$
Cinchonine	$C_{19}H_{22}N_2O$
Citric acid	$HOC(COOH)(CH_2COOH)_2$
Cobalt metal	Co
Cobalt sulfate	$CoSO_4$
Coke	
Congo red test paper	
Copper metal (99.9 % min)	Cu
* Copper metal (powder or turnings)	Cu

TABLE 1 *Continued*

Name	Formula
Copper metal (P-free)	Cu
Copper metal (Mn, Ni, and Co-free, less than 0.001 % of each)	Cu
Copper-rare earth oxide mixture	
m-Cresol purple (*m*-cresolsulfonephthalein)	$C_6H_4SO_2OC(C_6H_3\text{-}2\text{-}CH_3\text{-}4\text{-}OH)_2$
Cupferron	$C_6H_5N(NO)ONH_4$
Cupric chloride	$CuCl_2{\cdot}2H_2O$
* Cupric nitrate	$Cu(NO_3)_2{\cdot}3H_2O$
* Cupric oxide (powder)	CuO
Cupric potassium chloride	$CuCl_2{\cdot}2KCl{\cdot}2H_2O$
* Cupric sulfate	$CuSO_4{\cdot}5H_2O$
Curcumin	$(2\text{-}CH_3OC_6H_3\text{-}1\text{-}OH\text{-}4\text{-}CH{:}CHCO)_2CH_2$
Devarda's alloy	50Cu-45Al-5Zn
Diethylenetriamine pentaacetic acid ([[(carboxymethyl)imino]bis(ethylenenenitrilo)] tetraacetic acid)	$((HOCOCH_2)_2NCH_2CH_2)_2NCH_2COOH$
* Dimethylglyoxime	$CH_3C{:}NOHC{:}NOHCH_3$
N,N' Diphenylbenzidine	$C_6H_5NHC_6H_4C_6H_4NHC_6H_5$
Diphenylcarbazide (1,5-diphenylcarbohydrazide)	$C_6H_5NHNHCONHNHC_6H_5$
* Disodium (ethylenedinitrilo) tetraacetate dihydrate	See (ethylenedinitrilo) tetraacetic acid disodium salt
Dithiol (toluene-3,4-dithiol)	$CH_3C_6H_3(SH)_2$
Dithizone (diphenylthiocarbazone)	$C_6H_5NHNHCSN{:}NC_6H_5$
Eriochrome black-T (1(1-hydroxy-2-naphthylazo)-6-nitro-2-naphthol-4-sulfonic acid sodium salt)	$1\text{-}HOC_{10}H_6\text{-}2\text{-}N{:}N\text{-}1\text{-}C_{10}H_4\text{-}2\text{-}OH\text{-}4\text{-}SO_3Na\text{-}6\text{-}NO_2$
* EDTA (Disodium salt)	See (ethylenedinitrilo) tetraacetic acid disodium salt
* Ethanol	C_2H_5OH
* Ethyl ether (diethyl ether)	$C_2H_5OC_2H_5$
* (Ethylenedinitrilo) tetraacetic acid disodium salt	$HOCOCH_2(NaOCOCH_2)NCH_2N(CH_2COONa)CH_2COOH{\cdot}2H_2O$
Ethylene glycol monomethyl ether (2-methoxy-ethanol)	$CH_3OCH_2CH_2OH$
* Ferric chloride	$FeCl_3{\cdot}6H_2O$
* Ferric nitrate	$Fe(NO_3)_3{\cdot}9H_2O$
Ferric sulfate	$Fe_2(SO_4)_3{\cdot}nH_2O$
* Ferrous ammonium sulfate	$Fe(NH_4)_2(SO_4)_2{\cdot}6H_2O$
* Ferrous sulfate	$FeSO_4{\cdot}7H_2O$
Fluoroboric acid	HBF_4
Fluorescein, sodium salt	$2NaOCOC_6H_4C{:}C_6H_3\text{-}3(:O)OC_6H_3\text{-}6\text{-}ONa$
Formaldehyde	HCHO
* Formic acid[A]	HCOOH
Gelatin	
Graphite	C
Glass wool	
Glycerol	$CH_2OHCHOHCH_2OH$
Hydrazine sulfate	$NH_2NH_2{\cdot}H_2SO_4$
* Hydrobromic acid[A]	HBr
* Hydrochloric acid[A]	HCl
* Hydrofluoric acid[A]	HF
Hydrogen chloride gas	HCl
* Hydrogen peroxide	H_2O_2
Hydrogen sulfide gas	H_2S
Hydroquinone	$1,4\text{-}(OH)_2C_6H_4$
* Hydroxylamine hydrochloride	$NH_2OH{\cdot}HCl$
* Hypophosphorous acid[B]	H_3PO_2
Invert sugar	
* Iodine	I_2
Iron metal or wire (99.8 % min)	Fe
Isopropyl ether	$(CH_3)_2CHOCH(CH_3)_2$
Lead metal	Pb
* Lead acetate	$Pb(CH_3COO)_2$
Lead chloride	$PbCl_2$
* Lead nitrate	$Pb(NO_3)_2$
Litmus	
Lithium fluoride	LiF
Magnesium metal (Sn-free)	Mg
Magnesium perchlorate, anhydrous	$Mg(ClO_4)_2$

TABLE 1 *Continued*

Name	Formula
* Magnesium sulfate	$MgSO_4 \cdot 7H_2O$
Manganese metal (99.8 % min)	Mn
Manganous nitrate	$Mn(NO_3)_2$
Manganous sulfate	$MnSO_4 \cdot H_2O$
Mannitol	$CH_2OH(CHOH)_4CH_2OH$
Marble chips	
* Mercuric chloride	$HgCl_2$
* Mercury	Hg
* Methanol	CH_3OH
Methyl isobutyl ketone (4-methyl-2-pentanone)	$CH_3COCH_2CH(CH_3)_2$
* Methyl orange (*p*[[*p*-dimethylamino)phenyl]azo]benzenesulfonic acid sodium salt)	$4\text{-}NaOSO_2C_6H_4N{:}NC_6H_4\text{-}4\text{-}N(CH_3)_2$
Methyl purple	formula unknown, patented
* Methyl red (*o* -[[(*p*-dimethylamino)phenyl]azo]benzoic acid)	$4\text{-}(CH_3)_2NC_6H_4N{:}NC_6H_4\text{-}2\text{-}COOH$
Molybdenum metal (99.8 % min)	Mo
Molybdic acid, anhydride (molybdenum trioxide)	MoO_3
Molybdic acid (ammonium paramolybdate)	Assay: as MoO_3—85 %
Morin, anhydrous (2',3,4',7-penta hydroxyflavone)	$5,7\text{-}(HO)_2C_6H_2\ OC(C_6H_3\text{-}2,4\text{-}(OH)_2){:}C(OH)CO$
β-Naphthoquinoline (5,6-benzoquinoline)	$C_{10}H_6CH{:}CHCH{:}N$
Neocuproine (2,9-dimethyl-1,10-phenanthroline)	$(CH_3)_2C_{12}H_6N_2 \cdot {}_{12}H_2O$
Nickel metal (99.8 % min)	Ni
Nickel metal (sheet)	Ni
Nickelous nitrate	$Ni(NO_3)_2 \cdot 6H_2O$
Nickelous sulfate	$NiSO_4 \cdot 6H_2O$
* Nitric acid[A]	HNO_3
Nitrogen gas (oxygen-free)	N_2
Nitrogen, liquid	N_2
m-Nitrophenol	$NO_2C_6H_4OH$
1-Nitroso-2-naphthol(α-nitroso-β-naphthol)	$NOC_{10}H_6OH$
Nitroso-R-salt (1-nitroso-2-naphthol-3,6-disulfonic acid disodium salt)	$1\text{-}NOC_{10}H_4\text{-}2\text{-}(OH)\text{-}3,6\text{-}(SO_3Na)_2$
Osmium tetraoxide	OsO_4
Oxalic acid	$(COOH)_2$
Oxygen gas	O_2
* Perchloric acid[A]	$HClO_4$
1,10-Phenanthroline (*o* -phenanthroline)	$CH{:}CHCH{:}NC{:}CCH{:}CHC{:}CN{:}CHCH{:}CH \cdot H_2O$
* Phenolphthalein	$C_6H_4COOC(C_6H_4\text{-}4\text{-}OH)_2$
* Phosphoric acid	H_3PO_4
Piperidine	$NH(CH_2)_4CH_2$
Platinized quartz	
Platinized silica gel	
Platinum gauze	Pt
* Potassium biphthalate	$1\text{-}KOCOC_6H_4\text{-}2\text{-}COOH$
Potassium bisulfate	$KHSO_4$
* Potassium bromate	$KBrO_3$
* Potassium bromide	KBr
* Potassium chlorate	$KClO_3$
* Potassium chloride	KCl
* Potassium chromate	K_2CrO_4
Potassium columbate	$4K_2O \cdot 3Cb_2O_5 \cdot 16H_2O$
* Potassium cyanide	KCN
* Potassium dichromate	$K_2Cr_2O_7$
* Potassium ferricyanide	$K_3Fe(CN)_6$
Potassium ferrocyanide	$K_4Fe(CN)_6 \cdot 3H_2O$
* Potassium fluoride	$KF \cdot 2H_2O$
* Potassium hydroxide	KOH
* Potassium iodate	KIO_3
* Potassium iodide	KI
Potassium iodide starch paper	
* Potassium nitrate	KNO_3
* Potassium *m*-periodate	KIO_4
* Potassium permanganate	$KMnO_4$
Potassium persulfate	$K_2S_2O_8$
Potassium phosphate, monobasic	KH_2PO_4
* Potassium pyrosulfate	$K_2S_2O_7$
* Potassium sulfate	K_2SO_4
Potassium tantalum fluoride	K_2TaF
Potassium thiocarbonate	K_2CS_3
* Potassium thiocyanate	KSCN

TABLE 1 *Continued*

Name	Formula
Pyrogallic acid (pyrogallol)	C_6H_3-1,3-$(OH)_3$
Quinine sulfate	$(C_{20}H_{24}N_2O_2)_2{\cdot}H_2SO_4{\cdot}2H_2O$
8-Quinolinol (8-hydroxyquinoline)	HOC_6H_3N:CHCH:CH
Sebacic acid	$HOCO(CH_2)_8COOH$
Selenium (powder)	Se
Silicon dioxide (silica)	SiO_2
* Silver nitrate	$AgNO_3$
Soda-lime	
Soda-mica mineral (CO_2 absorbent)	
Sodium acetate	CH_3COONa
Sodium arsenite	$NaAsO_2$
Sodium azide	NaN_3
* Sodium bicarbonate	$NaHCO_3$
* Sodium bismuthate	$NaBiO_3$
Sodium bisulfate	see sodium hydrogen sulfate
* Sodium bisulfate, fused	see sodium hydrogen sulfate, fused
Sodium bisulfite	$NaHSO_3$
* Sodium borate	$Na_2B_4O_7{\cdot}10H_2O$
* Sodium carbonate, anhydrous	Na_2CO_3
Sodium chlorate	$NaClO_3$
Sodium chloride	NaCl
Sodium citrate	$HOC(COONa)(CH_2COONa)_2{\cdot}2H_2O$
Sodium cyanide	NaCN
Sodium diethyldithiocarbamate	$(C_2H_5)_2NCSSNa{\cdot}3H_2O$
Sodium dimethylglyoximate	$CH_3C(:NONa)C(:NONa)CH_3{\cdot}8H_2O$
Sodium diphenylamine sulfonate	$C_6H_5NHC_6H_4$-4-SO_3Na
Sodium dithionite (hydrosulfite)	$Na_2S_2O_4$
* Sodium fluoride	NaF
Sodium hydrogen sulfate	$NaHSO_4$
Sodium hydrogen sulfate, fused	A mixture of $Na_2S_2O_7$ and $NaHSO_4$
* Sodium hydroxide	NaOH
Sodium hypophosphite	$NaH_2PO_2{\cdot}H_2O$
Sodium molybdate	$Na_2MoO_4{\cdot}2H_2O$
Sodium nitrate	$NaNO_3$
Sodium nitrite	$NaNO_2$
Sodium oxalate	NaOCOCOONa
Sodium perchlorate	$NaClO_4$
Sodium peroxide	Na_2O_2
Sodium phosphate, dibasic, anhydrous	Na_2HPO_4
Sodium pyrophosphate	$Na_4P_2O_7{\cdot}10H_2O$
Sodium pyrosulfate	$Na_2S_2O_7$
Sodium sulfate, anhydrous	Na_2SO_4
Sodium sulfide	$Na_2S{\cdot}9H_2O$
Sodium sulfite	$Na_2SO_3{\cdot}7H_2O$
Sodium sulfite, anhydrous	Na_2SO_3
Sodium thiocyanate	NaSCN
* Sodium thiosulfate	$Na_2S_2O_3{\cdot}5H_2O$
* Sodium tungstate	$Na_2WO_4{\cdot}2H_2O$
* Stannous chloride	$SnCl_2{\cdot}2H_2O$
* Starch	$(C_6H_{10}O_5)_x$
Succinic acid	$HOCOCH_2CH_2COOH$
Sulfamic acid	NH_2SO_3H
Sulfatoceric acid (ceric sulfate)	$H_4Ce(SO_4)_4$
5-Sulfosalicylic acid	2-HOC_6H_3-1-COOH-5-$SO_3H{\cdot}2H_2O$
Sulfur dioxide gas	SO_2
* Sulfuric acid[A]	H_2SO_4
* Sulfurous acid[A]	H_2SO_3
Talc	
* Tartaric acid	$HOCO(CHOH)_2COOH$
Test lead	Pb
Tetrapropylammonium hydroxide	$(CH_3CH_2CH_2)_4NOH$
Thioglycollic acid (mercaptoacetic acid)	$CH_2SHCOOH$
Thiourea	NH_2CSNH_2
Tin metal (99.9 %min)	Sn
Titanium dioxide	TiO_2
Titanium metal (low Sn)	Ti
Triethanolamine (2,2′,2″-nitrilotriethanol)	$(CH_2OHCH_2)_3N$
Uranium oxide	U_3O_8
* Uranyl nitrate	$UO_2(NO_3)_2{\cdot}6H_2O$
Urea	NH_2CONH_2

TABLE 1 *Continued*

Name	Formula
Zinc (99.9 % min)	Zn
Zinc metal (S-free)	Zn
Zinc oxide	ZnO
Zinc sulfate	$ZnSO_4 \cdot 7H_2O$
Zirconium oxide	ZrO_2
Zirconium metal	Zr
Zirconyl chloride	$ZrOCl_2 \cdot 8H_2O$

[A] * Reagent on which ACS specifications exist.
† ACS specification exists but does not cover all requirements.
For concentration of laboratory reagent, see Table 2.
[B] Contains at least 50 % H_3PO_2.

to Type I or II of Specification D1193. Type III or IV may be used if they effect no measurable change in the blank or sample.

5.2 *Reagents*—Unless otherwise indicated, it is intended that all reagents conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society when such specifications are available.[3] Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination. In addition to this, it is desirable in many cases for the analyst to ensure the accuracy of his results by running blanks or checking against a comparable sample of known composition.

6. Reagents

6.1 *Concentrated Acids, Ammonium Hydroxide, and Hydrogen Peroxide*—When acids, ammonium hydroxide, and hydrogen peroxide are specified by name or chemical formula only, it is understood that concentrated reagents of the specific gravities or concentrations shown in Table 2 are intended. The specific gravities or concentrations of all other concentrated acids are stated wherever they are specified.

6.2 *Diluted Acids and Ammonium Hydroxide*—Concentrations of diluted acids and ammonium hydroxide, except when standardized, are specified as a ratio stating the number of volumes of the concentrated reagent to be diluted with a given number of volumes of water, as in the following example: HCl (5 + 95) means 5 volumes of concentrated HCl (sp gr 1.19) diluted with 95 volumes of water.

6.3 *Standard Solutions*—Concentrations of standard solutions are stated as molarities or normalities, expressed decimally; or the equivalent of 1 mL of solution in terms of grams, milligrams, or micrograms of a given element expressed as "1 mL = x.xx—g, mg, or µg of..."

6.4 *Nonstandard Solutions*—Composition of nonstandard solutions prepared by dissolving a given mass of the solid reagent in a solvent are specified in grams of the salt as weighed per litre of solution, and it is understood that water is the solvent unless otherwise specified. For example, to prepare barium chloride solution (100 g/L) dissolve 100 g of barium chloride ($BaCl_2 \cdot 2H_2O$) in water and dilute to 1 L. In the case of certain reagents, the composition may be specified as a mass fraction percent. For example, H_2O_2 (3 %) means a solution containing 3 g of H_2O_2 per 100 g of solution. Other nonstandard solutions may be specified by name only and the designation of the composition of such solutions will be governed by the instructions for their preparation.

7. Laboratory Ware (1,2)[4,5]

7.1 *Glassware*—Unless otherwise stated all analytical methods are conducted in borosilicate glassware.

[3] *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC, www.chemistry.org . For suggestions on the testing of reagents not listed by the American Chemical Society, see the *United States Pharmacopeia* and *4.2 National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD, www.usp.org.

[4] The boldface numbers in parentheses refer to the list of references at the end of these practices.

TABLE 2 Composition of Acids, Ammonium Hydroxide, and Hydrogen Peroxide

Name	Formula	Specific Gravity, Approximate	Reagent, Mass Fraction, %		
			Nominal	Min	Max
Acetic acid	CH_3COOH	1.05	...	99.5	...
Formic acid	HCOOH	1.20	...	88.0	...
Hydrobromic acid	HBr	1.49	48	47.0	49.0
Hydrochloric acid	HCl	1.19	...	35.0	38.0
Hydrofluoric acid	HF	1.15	...	48.0	51.0
Nitric acid	HNO_3	1.42	...	69.0	71.0
Perchloric acid	$HClO_4$	1.67	...	70.0	72.0
Phosphoric acid	H_3PO_4	1.69	...	85.0	...
Sulfuric acid	H_2SO_4	1.84	...	95.0	98.0
Sulfurous acid	H_2SO_3	1.03	...	6.0(SO_2)	...
Ammonium hydroxide	NH_4OH	0.90	...	27.0(NH_3)	30.0 (NH_3)
Hydrogen peroxide	H_2O_2	1.10	30	28.0	...