



Designation: ~~D1126~~—~~12~~ **D1126** – 17

## Standard Test Method for Hardness in Water<sup>1</sup>

This standard is issued under the fixed designation D1126; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reappraisal. A superscript epsilon ( $\epsilon$ ) indicates an editorial change since the last revision or reappraisal.

*This standard has been approved for use by agencies of the U.S. Department of Defense.*

### 1. Scope\*

1.1 This test method covers the determination of hardness in water by titration. This test method is applicable to waters that are clear in appearance and free of chemicals that will complex calcium or magnesium. The lower detection limit of this test method is approximately 2 to 5 mg/L as CaCO<sub>3</sub>; the upper limit can be extended to all concentrations by sample dilution. It is possible to differentiate between hardness due to calcium ions and that due to magnesium ions by this test method.

1.2 This test method was tested on reagent water only. It is the user's responsibility to ensure the validity of the test method for waters of untested matrices.

1.3 The values stated in SI units are to be regarded as standard. No other units of measurement are included in this standard.

1.4 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety, health, and environmental practices and determine the applicability of regulatory limitations prior to use.*

1.5 *This international standard was developed in accordance with internationally recognized principles on standardization established in the Decision on Principles for the Development of International Standards, Guides and Recommendations issued by the World Trade Organization Technical Barriers to Trade (TBT) Committee.*

### 2. Referenced Documents

2.1 *ASTM Standards:*<sup>2</sup>

[D1066 Practice for Sampling Steam](#)

[D1129 Terminology Relating to Water](#)

[D1193 Specification for Reagent Water](#)

[D2777 Practice for Determination of Precision and Bias of Applicable Test Methods of Committee D19 on Water](#)

[D3370 Practices for Sampling Water from Closed Conduits](#)

[D5847 Practice for Writing Quality Control Specifications for Standard Test Methods for Water Analysis](#)

### 3. Terminology

~~3.1 Definitions—For definitions of terms used in this test method, refer to Terminology [D1129](#).~~

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~~3.1.1 For definitions of terms used in this standard, refer to Terminology [D1129](#).~~

3.2 *Definitions of Terms Specific to This Standard:*

3.2.1 *equivalent per million (epm), n*—a unit chemical equivalent weight of solute per million unit weights of solution.

3.2.2 *laboratory control sample (LCS), n*—a solution with certified hardness.

### 4. Summary of Test Method

4.1 Calcium and magnesium ions in water are sequestered by the addition of disodium ethylenediamine tetraacetate. The end point of the reaction is detected by means of Chrome Black T<sup>3</sup>, which has a red color in the presence of calcium and magnesium and a blue color when they are sequestered.

<sup>1</sup> This test method is under the jurisdiction of ASTM Committee [D19](#) on Water and is the direct responsibility of Subcommittee [D19.05](#) on Inorganic Constituents in Water. Current edition approved March 1, 2012; Dec. 1, 2017. Published March 2012; December 2017. Originally approved in 1950. Last previous edition approved in 2007; 2012 as ~~D1126 – 02 (2007)~~ [D1126 – 12](#).<sup>1,2</sup> DOI: [10.1520/D1126-12.10.1520/D1126-17](#).

<sup>2</sup> For referenced ASTM standards, visit the ASTM website, [www.astm.org](#), or contact ASTM Customer Service at [service@astm.org](#). For *Annual Book of ASTM Standards* volume information, refer to the standard's Document Summary page on the ASTM website.

<sup>3</sup> 3-Hydroxy-4-(1-hydroxy-2-naphthyl) azo-7-nitro-1 naphthalenesulfonic acid, sodium salt, Color Index 14645.

\*A Summary of Changes section appears at the end of this standard

**TABLE 1 Freedom of Reaction from Interferences**

Substance	Maximum Concentration Without Interference in the Total Hardness Test, mg/L	Maximum Concentration Without Interference in the Calcium Hardness Test, mg/L
Aluminum, Al <sup>+++</sup>	20	5
Ammonium, NH <sub>4</sub> <sup>+</sup>	<sup>A</sup>	2 000
Bicarbonate, HCO <sub>3</sub> <sup>-</sup>	...	500
Bromine, Br	...	2
Cadmium, Cd <sup>++</sup>	20	...
Carbonate, CO <sub>3</sub> <sup>--</sup>	1 000	50
Chloride, Cl <sup>-</sup>	10 000	...
Chlorine, Cl	...	2
Chromate, CrO <sub>4</sub> <sup>--</sup>	500	500
Cobalt, Co <sup>++</sup>	0.3	...
Copper, Cu <sup>++</sup>	20	2
Iron, ferric, Fe <sup>+++</sup>	10 <sup>B</sup>	20
Iron, ferrous, Fe <sup>++</sup>	10 <sup>B</sup>	20
Lead, Pb <sup>++</sup>	20	5
Manganese, Mn <sup>++</sup>	1 <sup>C</sup>	10 <sup>C</sup>
Nickel, Ni <sup>++</sup>	0.5 <sup>D</sup>	...
Nitrate, NO <sub>3</sub> <sup>-</sup>	500	500
Nitrite, NO <sub>2</sub> <sup>-</sup>	500	500
Phosphate, PO <sub>4</sub> <sup>---</sup>	100	...
Silicate, SiO <sub>3</sub> <sup>--</sup>	200	100
Strontium, Sr <sup>++</sup>	<sup>E</sup>	<sup>E</sup>
Sulfate, SO <sub>4</sub> <sup>--</sup>	10 000	10 000
Sulfite, SO <sub>3</sub> <sup>--</sup>	500	500
Tannin, Quebracho	200	50
Tin, stannic, Sn <sup>++++</sup>	10	5
Tin, stannous, Sn <sup>++</sup>	10	5
Zinc, Zn <sup>++</sup>	20	5

<sup>A</sup> No data are available.

<sup>B</sup> Iron will not interfere in concentrations up to 200 mg/L. However, the red color of the end point may return in about 30 s.

<sup>C</sup> Manganese will not interfere in concentrations up to 10 mg/L if a few crystals of K<sub>4</sub>Fe(CN)<sub>6</sub>·3H<sub>2</sub>O are added to the buffer immediately before use.

<sup>D</sup> Accurate results can be obtained in the presence of 1 mg/L nickel, but the end point is slow under these conditions.

<sup>E</sup> If strontium is present, it will be titrated with calcium and magnesium.

<https://standards.iteh.ai/catalog/standards/sist/fb47cb6a-5096-4bad-90c8-6fa272f9e792/astm-d1126-17>

## 5. Significance and Use

5.1 Hardness salts in water, notably calcium and magnesium, are the primary cause of tube and pipe scaling, which frequently causes failures and loss of process efficiency due to clogging or loss of heat transfer, or both.

5.2 Hardness is caused by any polyvalent cations, but those other than Ca and Mg are seldom present in more than trace amounts. The term hardness was originally applied to water in which it was hard to wash; it referred to the soap-wasting properties of water. With most normal alkaline water, these soap-wasting properties are directly related to the calcium and magnesium content.

## 6. Interferences

6.1 The substances shown in **Table 1** represent the highest concentrations that have been found not to interfere with this determination.

6.2 The test method is not suitable for highly colored waters, which obscure the color change of the indicator.

## 7. Reagents

7.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society.<sup>4</sup> Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

<sup>4</sup> *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For Suggestions on the testing of reagents not listed by the American Chemical Society, see *Annual Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD.

7.2 *Purity of Water*—Unless otherwise indicated, reference to water shall be understood to mean reagent water conforming to Specification **D1193**, Type I. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the precision and bias of the test method. Type II water was specified at the time of round robin testing of this test method.

7.3 *Ammonium Hydroxide Solution (1 + 4)*—Mix 1 volume of  $\text{NH}_4\text{OH}$  (sp gr 0.90) with 4 volumes of water.

7.4 *Buffer Solution*—Prepare the buffer solution in three steps as follows:

7.4.1 Dissolve 40 g of sodium tetraborate ( $\text{Na}_2\text{B}_4\text{O}_7 \cdot 10\text{H}_2\text{O}$ ) in 800 mL of water.

7.4.2 Dissolve 10 g of sodium hydroxide ( $\text{NaOH}$ ), 10 g of sodium sulfide ( $\text{Na}_2\text{S} \cdot 9\text{H}_2\text{O}$ ), and 10 g of potassium sodium tartrate ( $\text{KNaC}_4\text{O}_6 \cdot 4\text{H}_2\text{O}$ ) in 100 mL of water.

7.4.3 When cool mix the two solutions and add 1 g of magnesium disodium ethylenediamine tetraacetate, having a magnesium-to-EDTA mole ratio of 1 to 1. Make up to 1 L with water. Keep the solution bottle stoppered when not in use. The reagent will be effective for at least 1 month.

7.5 *Calcium Solution, Standard (1 mL = 0.20 mg  $\text{CaCO}_3$ )*—Dissolve 0.2000 g of  $\text{CaCO}_3$  in 3 to 5 mL of  $\text{HCl}$  (1 + 4). Dilute to 1 L with water.

7.6 *Calcium Indicator*—Use powdered hydroxynaphthol blue,<sup>5</sup> or grind solid hydroxynaphthol blue to 40 to 50 mesh size.

7.7 *Hardness Indicator*—The hardness indicator can be prepared, stored, and used in liquid or powder form.

7.7.1 *Hardness Indicator Solution*—Dissolve 0.5 g of Chrome Black  $\text{T}^3$  in 50 mL of diethanolamine or triethanolamine. Store the solution in a dark-colored bottle. This solution has a storage life of several months.

7.7.2 *Hardness Indicator Powder*—Grind 0.5 g of Chrome Black  $\text{T}^3$  with 100 g of powdered sodium chloride. Use a dark-colored bottle for storage. The powder has a storage life of at least 1 year.

7.8 *Hydrochloric Acid (1 + 4)*—Mix 1 volume of concentrated hydrochloric acid (sp gr 1.19) with 4 volumes of water.

7.9 *Disodium Ethylenediamine Tetraacetate ( $\text{Na}_2\text{H}_2\text{EDTA}$ ) Solution, Standard (1 mL = 1.0 mg  $\text{CaCO}_3$ )*—Dissolve 3.8 g of disodium ethylenediamine tetraacetate dihydrate in approximately 800 mL of water. Adjust the pH of the solution to 10.5 with  $\text{NaOH}$  solution (50 g/L). Determine the concentration of this solution using the standard calcium (7.5) solution, and that procedure in Section 9 that will be used for the sample analysis (9.1, 9.2, or 9.3). Adjust the concentration of the EDTA so that 1 mL will be equivalent to 1.0 mg of  $\text{CaCO}_3$ . Store the standard EDTA in polyethylene, plastic, or hard rubber bottles and restandardize monthly.

7.10 *Sodium Hydroxide Solution (50 g/L)*—Dissolve 50 g of sodium hydroxide in water and dilute to 1 L.

## 8. Sampling

8.1 Collect the sample in accordance with Practice **D1066** or Practices **D3370** as applicable.

## 9. Procedure

9.1 *Hardness*—Measure 50 mL of clear sample into an opaque white container or a clear colorless container utilizing a white background. Adjust the pH of the sample to 7 to 10 by adding  $\text{NH}_4\text{OH}$  (7.3) solution or  $\text{HCl}$  (7.8) solution. Add 0.5 mL of buffer (7.4) solution, and approximately 0.2 g of hardness (7.7) indicator powder or 2 drops of liquid and stir. Add standard  $\text{Na}_2\text{H}_2\text{EDTA}$  (7.9) solution slowly from a burette with continuous stirring until the color changes from red to blue. Complete the titration within 5 min after the buffer addition. If the titration requires more than 20 mL of the titrating solution, dilute the sample and repeat the test.

9.2 *Low Hardness*—Determine low-hardness values (0.5 to 5.0 ppm as  $\text{CaCO}_3$ ) in accordance with 9.1, but use a 100 mL sample and titrate by means of micro-burette. When employing a 100-mL sample, add twice the quantity of the reagents as indicated in 9.1.

9.3 *Calcium Hardness*—Measure 50 mL of the sample into an opaque white container, or a clear colorless container utilizing a white background. Add 2 mL of  $\text{NaOH}$  (7.10) solution and stir. Add approximately 0.2 g of calcium (7.6) indicator and stir. Add standard  $\text{Na}_2\text{H}_2\text{EDTA}$  (7.9) solution slowly from a burette with continuous stirring until the color changes from red to royal blue. Complete the titration within 5 min after the  $\text{NaOH}$  addition. If the titration requires more than 15 mL of the titrating solution, dilute the sample and repeat the test.

## 10. Calculations

10.1 Calculate the hardness, epm, of the sample as follows:

$$\text{Hardness, epm} = 20 C/S \quad (1)$$

<sup>5</sup> 3-Hydroxy-4 (2-hydroxy-4 sulfo-1 naphthyl) azo-2, 7-naphthalenedisulfonic acid, trisodium salt.