
**Water quality — Determination of
dissolved perchlorate — Method using
ion chromatography (IC)**

*Qualité de l'eau — Détermination du perchlorate dissous — Méthode
par chromatographie ionique (IC)*

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Foreword

ISO (the International Organization for Standardization) is a worldwide federation of national standards bodies (ISO member bodies). The work of preparing International Standards is normally carried out through ISO technical committees. Each member body interested in a subject for which a technical committee has been established has the right to be represented on that committee. International organizations, governmental and non-governmental, in liaison with ISO, also take part in the work. ISO collaborates closely with the International Electrotechnical Commission (IEC) on all matters of electrotechnical standardization.

The procedures used to develop this document and those intended for its further maintenance are described in the ISO/IEC Directives, Part 1. In particular, the different approval criteria needed for the different types of ISO documents should be noted. This document was drafted in accordance with the editorial rules of the ISO/IEC Directives, Part 2 (see www.iso.org/directives).

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For an explanation on the voluntary nature of standards, the meaning of ISO specific terms and expressions related to conformity assessment, as well as information about ISO's adherence to the World Trade Organization (WTO) principles in the Technical Barriers to Trade (TBT) see the following URL: www.iso.org/iso/foreword.html.

This document was prepared by Technical Committee ISO/TC 147, *Water quality*, Subcommittee SC 2, *Physical, chemical and biochemical methods*.

Water quality — Determination of dissolved perchlorate — Method using ion chromatography (IC)

WARNING — Persons using this document should be familiar with normal laboratory practice. This document does not purport to address all of the safety problems, if any, associated with its use. It is the responsibility of the user to establish appropriate safety and health practices and to determine the applicability of any other restrictions.

IMPORTANT — It is absolutely essential that tests conducted in accordance with this document be carried out by suitably qualified staff.

1 Scope

This document specifies a method for the determination of dissolved perchlorate in water (e.g. drinking water, mineral water, raw water, surface water, partially treated water or swimming pool water, waste water from drinking/swimming pool water treatment plants).

Appropriate pre-treatment of the sample (e.g. matrix elimination) allows a direct determination of perchlorate $\geq 1 \mu\text{g/l}$.

The working range is restricted by the ion-exchange capacity of the separator column. Dilution of the sample to the working range can be necessary.

2 Normative references

The following documents are referred to in the text in such a way that some or all of their content constitutes requirements of this document. For dated references, only the edition cited applies. For undated references, the latest edition of the referenced document (including any amendments) applies.

ISO 3696, *Water for analytical laboratory use — Specification and test methods*

ISO 8466-1, *Water quality — Calibration and evaluation of analytical methods and estimation of performance characteristics — Part 1: Statistical evaluation of the linear calibration function*

ISO 8466-2, *Water quality — Calibration and evaluation of analytical methods and estimation of performance characteristics — Part 2: Calibration strategy for non-linear second-order calibration functions*

3 Terms and definitions

No terms and definitions are listed in this document.

ISO and IEC maintain terminological databases for use in standardization at the following addresses:

- IEC Electropedia: available at <http://www.electropedia.org/>
- ISO Online browsing platform: available at <https://www.iso.org/obp>

4 Interferences

Perchlorate is known to be susceptible to microbiological degradation in the absence of nitrate and by anaerobic bacteria (References [5] and [6]).

Any substance that has a retention time coinciding with perchlorate and producing a detector response can interfere. Co-elution can be solved by changing separator columns, eluent strength (e.g. gradient elution), modifying the eluent with organic solvents or by selective removal of the interference by sample pre-treatment.

Higher chloride, sulfate, nitrate, orthophosphate, hydrogen carbonate and carbonate concentrations (e.g. 1 g/l) can interfere with the determination of perchlorate (co-elution) or can have an impact on the retention time or the peak shape (distortion) or recovery of perchlorate. This effect can be checked for every matrix by standard addition and the recovery of perchlorate should be within 85 % to 115 %. Interferences can be reduced by sample dilution or with the aid of special cation exchangers (see [7.2](#) and [Clause 9](#)) or be resolved by the application of advanced inline cutting or re-injection techniques (see [Annexes A, B and C](#)).

Metals like iron or aluminium present in samples and eluent will bind to the resin material of the separator column or suppressor resulting in a loss of performance. Metal ions can be eliminated with the aid of special cation exchangers (see [Annex A](#) and [Clause 9](#)).

Solid particles and organic compounds (such as mineral oils, detergents and humic acids) shorten the lifetime of the pre-column and the separator column (see [Clause 9](#)).

5 Principle

The method requires the application of high-capacity separator columns which allow the injection of sample volumes, e.g. up to 4 ml.

If necessary, the sample is pre-treated in order to remove anions, metals, organics and solids (see [Clause 9](#)).

Measurement of perchlorate is performed with or without matrix elimination and with or without pre-concentration (see [10.3](#) and [Annexes A, B and C](#)).

Perchlorate is separated by ion chromatography (IC) with suppressed conductivity detection (CD).

As stationary phase, an anion exchange resin is used. Aqueous solutions of salts of monobasic acids and dibasic acids are used as eluents for isocratic or gradient elution e.g. carbonate-, hydrogen carbonate-, hydroxide-eluent and an organic modifier like acetone or acetonitrile ([6.3](#)).

Resolution, R , shall be checked to ensure that it complies with the required separation conditions ([8.1](#)).

The concentration of perchlorate is determined after a calibration according to ISO 8466-1 or ISO 8466-2.

Control experiments are necessary to check the validity of the calibration function. Replicate determinations can be necessary. The method of standard addition can be applied if matrix interferences are expected.

6 Reagents

Use only reagents of pro analysis grade free of compounds containing perchlorate. Weigh the reagents with an accuracy of ± 1 % of the nominal mass, unless stated otherwise. Prepare alternative concentrations or volumes of solutions as described in [6.3](#) to [6.5](#), if necessary. Alternatively, use commercially available solutions of the required concentration.

6.1 Water, ISO 3696, grade 1 and with a resistivity of $\geq 18,2$ M Ω cm (25 °C).

6.2 Sodium perchlorate, NaClO₄, > 99 %.

6.3 Eluents.

Degas all eluents used. Take steps to avoid any renewed air pick-up during operation (e.g. by helium sparging, inline degasser).

The choice of eluent (e.g. based on sodium carbonate or sodium hydroxide solutions, potassium hydroxide, mixed with an organic modifier, if needed) depends on the choice of column and detector; seek advice from the column supplier. Apply eluents that were prepared manually, by inline dilution or electrochemically *in situ*. The chosen combination of separator column and eluent shall conform to the resolution requirements stated in 8.1. Use eluents as long as the requirements in Clause 8 and in 10.3 are met.

An example for an appropriate eluent manually prepared is given in 6.3.2.

6.3.1 Sodium hydroxide solution, $w(\text{NaOH}) = 50 \%$.

6.3.2 Sodium hydroxide eluent, $\rho(\text{NaOH}) = 0,05 \text{ mol/l}$.

Dissolve 4 g of sodium hydroxide solution (6.3.1) in approximately 900 ml of water (6.1) in a 1 000 ml volumetric flask and dilute to volume with water (6.1). Mix and degas, e.g. by sparging with helium for approximately 10 min. Store the eluent in a glass or polyethene eluent reservoir under an inert helium atmosphere to minimize carbonate contamination. Prepare the eluent on the day of use.

NOTE Solutions of manually prepared sodium hydroxide can be susceptible to carbonate contamination resulting from adsorption of carbon dioxide from the atmosphere. This contamination can lead to irreproducible perchlorate retention times, elevated instrument background conductivity and increased baseline noise/drift.

6.4 Perchlorate standard solution.

Depending on the concentrations expected, prepare the following standard solutions of different perchlorate concentrations from the stock solution (6.4.1). Note the possible risk of changes in concentration caused by interaction with the vessel material which will increase with decreasing perchlorate concentration. Store the standard solutions in polyethene or glass bottles.

6.4.1 Perchlorate stock solution, $\rho(\text{ClO}_4^-) = 1\,000 \text{ mg/l}$.

Dry approximately 1,5 g sodium perchlorate (6.2) at $(110 \pm 5)^\circ\text{C}$ for 2 h. Dissolve $(1,231 \pm 0,001)$ g of the dried sodium perchlorate in approximately 800 ml of water (6.1) in a 1 000 ml volumetric flask and dilute to volume with water (6.1).

Store the solution at 2°C to 8°C in polyethene or glass bottles. The solution is stable for 12 months.

6.4.2 Perchlorate standard solution I, $\rho(\text{ClO}_4^-) = 100 \text{ mg/l}$.

Pipette 10,0 ml of sodium perchlorate stock solution (6.4.1) into a 100 ml volumetric flask and dilute to volume with water (6.1).

Store the solution at 2°C to 8°C in polyethene or glass bottles. The solution is stable for six months.

6.4.3 Perchlorate standard solution II, $\rho(\text{ClO}_4^-) = 1 \text{ mg/l}$.

Pipette 1,0 ml of standard solution I (6.4.2) into a 100 ml volumetric flask, dilute to volume with water (6.1).

Store the solution at 2°C to 8°C in polyethene or glass bottles. The solution is stable for three months.

6.5 Perchlorate calibration solutions

Depending on the perchlorate concentration expected in the sample, use the perchlorate standard solutions I or II (6.4.2 or 6.4.3) to prepare 5 to 10 calibration solutions distributed over the expected working range as evenly as possible.

For example, proceed as follows for the range 1 µg/l to 5,5 µg/l perchlorate. Pipette into a series of 100 ml volumetric flasks the volumes described in Table 1.

Table 1 — Example for the preparation of calibration solutions

Volume standard solution II (6.4.3)	100 µl	150 µl	200 µl	250 µl	300 µl	350 µl	400 µl	450 µl	500 µl	550 µl
$\rho(\text{ClO}_4^-)$ (µg/l)	1	1,5	2	2,5	3	3,5	4	4,5	5	5,5

Prepare the calibration solutions on the day of use.

6.6 Blank solution

Fill a volumetric flask (e.g. 100 ml) with water (6.1).

7 Apparatus

The usual laboratory apparatus and, in particular, the following.

7.1 Ion chromatographic system, complying with the quality requirements given in Clause 8, i.e. resolution. In general, it shall consist of the following components (see Figure 1).

7.1.1 Separator column, with the specified separating performance (8.1).

7.1.2 Precolumn, if necessary (see Clause 4).

NOTE The use of a precolumn is advantageous not only for the analyses of waters highly loaded with organics (see Clause 4). It serves as well to protect the separator column. In general, two different types of precolumns are available: those containing the same or similar resin material as the separator column, and those packed with a non-functionalised resin.

7.1.3 Conductivity detector (CD), with a suppressor device.

7.1.4 Eluent reservoir and a degassing unit.

7.1.5 IC-pump, suitable for isocratic or gradient technique, respectively.

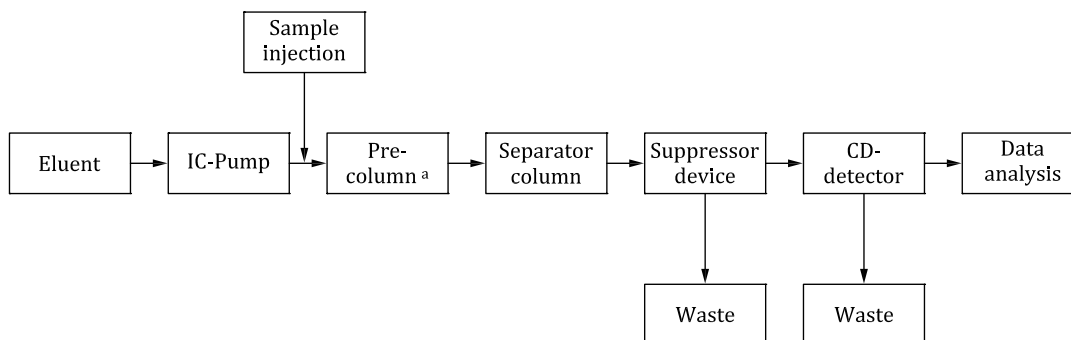
7.1.6 Sample delivery device, e.g. sample pump, including a sample injection system incorporating a sample loop of appropriate volume (e.g. 1 ml to 4 ml) or an autosampler device.

7.1.7 Recording device, e.g. PC with software for chromatography data acquisition and evaluation.

7.2 SPE-cartridges.

7.2.1 Cation exchanger in the Na-form, optional: metal clean up column for inline use.

7.2.2 Cartridges, with non-polar phases to be used for sample preparation (e.g. polyvinylpyrrolidone, RP C18).



^a Optional.

Figure 1 — Schematic representation of an ion chromatographic system

8 Quality requirements

8.1 Performance of the separator column

In chromatograms of samples and standard solutions, the peak resolution, R , between the anion of interest and its nearest peak shall not fall below 1,3 (see [Formula \(1\)](#) and [Figure 2](#)). Separation conditions shall be such that possible interfering anions will not interfere with the anion of interest.

If R fails the criteria $\geq 1,3$ or perchlorate elutes as a rider peak (see [Figure A.1](#)) due to higher concentrations of matrix ions (e.g. chloride, sulfate, nitrate and carbonate), then if necessary,

- dilute the sample,
- change separator columns,
- change the eluent strength (e.g. gradient elution),
- modify the eluent with organic solvents,
- remove the interference selectively by sample pre-treatment ([Annex A](#)), and/or
- apply column cut or re-injection techniques according to [Annex B](#) or [Annex C](#), and
- determine recovery or apply standard addition.