
International Standard



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Water quality — Determination of calcium and magnesium — Atomic absorption spectrometric method

Qualité de l'eau — Dosage du calcium et du magnésium — Méthode par spectrométrie d'absorption atomique

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Foreword

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International Standard ISO 7980 was prepared by Technical Committee ISO/TC 147, *Water quality*.

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Water quality — Determination of calcium and magnesium — Atomic absorption spectrometric method

1 Scope and field of application

This International Standard specifies a method for the determination of dissolved calcium and magnesium by flame atomic absorption spectrometry. It is intended for the analysis of raw and drinking waters and can be used for waters having a calcium content of up to 50 mg/l and a magnesium content of up to 5 mg/l. For samples containing higher concentrations of calcium or magnesium a smaller volume of the sample must be taken for the analysis.

When using the air/acetylene flame and the dilution factor 1 in 10, as described in 6.1, the optimum range is 3 to 50 mg/l for calcium and 0,9 to 5 mg/l for magnesium.

2 Principle

Measurement by flame atomic absorption spectrometry after adding lanthanum chloride (if an air/acetylene flame is used) or caesium chloride (if a nitrous oxide/acetylene flame is used) to reduce interferences. For calcium the absorbance is measured at 422,7 nm and for magnesium at 285,2 nm.

3 Reagents and materials

During the analysis, use only reagents of recognized analytical grade and only distilled water or water of equivalent purity. (Commercially available, ready-made solutions may be used.)

3.1 Hydrochloric acid (HCl), $\rho = 1,18$ g/ml.

3.2 Hydrochloric acid (HCl), 0,1 mol/l.

Dilute 8 ml of hydrochloric acid (3.1) to 1 litre.

3.3 Lanthanum chloride (LaCl₃) solution, containing 20 g of La per litre.

To a 1 litre one-mark volumetric flask add 24 g of lanthanum oxide (La₂O₃) (atomic absorption spectrometry grade). Slowly and cautiously add 50 ml of hydrochloric acid (3.1) while stirring to dissolve the lanthanum oxide. Make up to the mark with water.

3.4 Caesium chloride (CsCl) solution, containing 20 g of Cs per litre.

Dissolve 25 g of caesium chloride in 1 litre of hydrochloric acid (3.2).

3.5 Calcium, stock solution, 1 000 mg/l.

Dry a portion of calcium carbonate (CaCO₃) at 180 °C for 1 h and allow it to cool in a desiccator. Weigh $2,50 \pm 0,01$ g of the dried material and suspend this in 100 ml of water. Add slowly the minimum amount of hydrochloric acid (3.2) necessary to dissolve the calcium carbonate (approximately 250 ml). Boil briefly to expel dissolved carbon dioxide, then cool. Transfer the solution quantitatively to a 1 000 ml one-mark volumetric flask and make up to the mark with hydrochloric acid (3.2).

Store the solution in a polyethylene or polypropylene bottle.

3.6 Magnesium, stock solution, 1 000 mg/l.

Dry a portion of magnesium oxide (MgO) at 180 °C for 1 h. Weigh $1,66 \pm 0,01$ g and dissolve in hydrochloric acid (3.2). Dilute with the same acid to 1 000 ml in a one-mark volumetric flask.

Store the solution in a polyethylene bottle.

3.7 Calcium-magnesium, standard solution corresponding to 20 mg of Ca and 2 mg of Mg per litre.

With pipettes, transfer 20,0 ml of the calcium stock solution (3.5) and 2,0 ml of the magnesium stock solution (3.6) to a 1 000 ml one-mark volumetric flask. Make up to the mark with hydrochloric acid (3.2).